

ZnO nanoparticles modulate the ionic transport and voltage regulation of lysenin nanochannels

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## Abstract

**Background:** The insufficient understanding of unintended biological impacts from nanomaterials (NMs) represents a serious impediment to their use for scientific, technological, and medical applications. While previous studies have focused on understanding nanotoxicity effects mostly resulting from cellular internalization, recent work indicates that NMs may interfere with transmembrane transport mechanisms, hence enabling contributions to nanotoxicity by affecting key biological activities dependent on transmembrane transport. In this line of inquiry, we investigated the effects of charged nanoparticles (NPs) on the transport properties of lysenin, a pore-forming toxin that shares fundamental features with ion channels such as regulation and high transport rate.

**Results:** The macroscopic conductance of lysenin channels greatly diminished in the presence of cationic ZnO NPs. The inhibitory effects were asymmetrical relative to the direction of the electric field and addition site, suggesting electrostatic interactions between ZnO NPs and a binding site. Similar changes in the macroscopic conductance were observed when lysenin channels were reconstituted in neutral lipid membranes, implicating protein-NP interactions as the major contributor to the reduced transport capabilities. In contrast, no inhibitory effects were observed in the presence of anionic SnO<sub>2</sub> NPs. Additionally, we demonstrate that inhibition of ion transport is not due to the dissolution of ZnO NPs and subsequent interactions of zinc ions with lysenin channels.

**Conclusion:** We conclude that electrostatic interactions between positively charged ZnO NPs and negative charges within the lysenin channels are responsible for the inhibitory effects on

the transport of ions. These interactions point to a potential mechanism of cytotoxicity, which may not require NP internalization.

**Keywords:** ZnO, nanoparticles, lysenin, ion transport, electrophysiology, SnO<sub>2</sub>, toxicity, voltage gated channels, ligand gated channels

## Background

The rapid development of certain nanomaterials (NMs) has led to their extensive use in many commercial applications including cosmetics, sporting goods, automotive parts, and electronics [1-4], while many others are under intense investigation for scientific, technological, and biomedical applications [5-9]. The large surface area to volume ratio of these materials yields novel physical and chemical properties that enable applications that are unachievable using micron-sized bulk material of identical composition. The scientific community has spent decades developing an understanding of NMs in order to control their fundamental physical and chemical properties. However, early investigations demonstrated that some of the same properties that make NMs attractive for multiple applications may cause unintended hazardous interactions with biological systems. Therefore, environmental and human exposure poses potentially significant risks [10], and this paradigm has led to intense investigations on the potential biological impact of NMs [11, 12]. While we have thus far attained a tremendous body of knowledge on end-point effects such as cytotoxicity, neurotoxicity, genotoxicity and oxidative stress [13-16], we lack a thorough understanding of the principles by which modulation of size, charge, composition, dissolution levels and surface chemistry affect the interaction of NMs with living cells.

ZnO nanoparticles (NPs) are considered to be one of the more toxic of the metal oxide NMs [17, 18]. Studies on ZnO NPs have demonstrated toxicity towards a large number of cell lines and model organisms, however, the mechanism of cytotoxicity is still under debate. Certain physicochemical properties, such as surface chemistry, dissolution potential, and their intrinsic ability to produce reactive oxygen species (ROS) have a strong impact on their cytotoxic effects [19-21]. Several studies have demonstrated that cytotoxicity stems from high dissolution rates, causing elevated levels of Zn<sup>2+</sup> ions in cellular media that eventually disrupt homeostasis and leads to cell death [22, 23]. Other groups have suggested that their intrinsic ability to produce ROS (which may arise from surface defects, such as oxygen vacancies) is responsible for the high cytotoxic potential of ZnO NPs [24, 25]. In the same line, SnO<sub>2</sub> NPs have been shown to inhibit kinetic growth and cytotoxicity towards certain cell lines and organisms [18, 26, 27], while other publications have demonstrated modest to no cytotoxic effects [28, 29]. Similar to other NPs, the crystal and hydrodynamic size of SnO<sub>2</sub> NPs play an important role on their toxic effects, and smaller sizes have been shown to correlate with increased toxicity [27].

Our inability to correctly predict how physical and chemical properties relate to toxicity stems from the fact that biological systems are elaborate and structurally and functionally interconnected, making it very difficult to isolate distinct interactions responsible for cytotoxicity. Therefore, investigations utilizing a simplified model system that mimics the structure and function of a biological assembly can be an important step towards a more complete understanding of mechanisms of nanotoxicity. In these regards, we address how the directional flow of ions across lipid membranes containing specialized transmembrane ion transporters are affected by NPs. This work is motivated by the tremendous biological relevance of ionic transport for any living cell, and by the evidence that malfunctions of the mechanisms

that control the transmembrane transport may have catastrophic consequences for cell functionality [30].

Among transmembrane transporters, voltage-regulated channels play key roles in fundamental cellular processes such as creating and maintaining electrochemical gradients, transmission of information, ion transport, signaling, and metabolism [31]. A salient feature of such transporters is the regulation of their activity by transmembrane electric fields interacting with voltage-sensing domains present in the channel's structure [32]. The presence of charged domains in different regions of protein channels presents opportunities for electrostatic interactions with charged NPs, which may affect the transmembrane transport and functionality of the host cells.

Given the large variety of ion transporters in the cell membrane, isolating a particular one in a specific cell for relevant studies on transport modulation induced by NPs is not an easy task. Moreover, reconstitution of a particular ionic transporter in an artificial membrane system, although feasible, may require multiple, extensive and costly preparation steps. A simplified system featuring fundamental characteristics of ion channels may constitute an excellent model for investigating potential nanotoxicity effects originating from the disruption of transmembrane transport of ions. Therefore, we propose a simplified model that explores the effects of charged NPs on the transport of ions through lysenin channels inserted into an artificial bilayer lipid membrane (BLM).

Lysenin is a pore-forming protein extracted from the coelom of the earthworm *E. foetida*, which self-assemble as a large conductance nonameric pore (~3 nm) in artificial and natural lipid membranes containing sphingomyelin (SM) [33-35]. The recently deciphered crystal structure indicates large charged domains present within the channel [36, 37], thus presenting a strong

potential for electrostatic interactions with charged NPs. The physiological role of lysenin is still obscure but the cytolytic and hemolytic activity is indicative of a pore-forming toxin [38]. Nonetheless, its relevance for nanotoxicity studies stem from several remarkable biophysical properties it shares with ion channels. Unlike many other pore-forming toxins and similar to voltage-gated ion channels, lysenin channels present asymmetrical voltage-induced gating [33, 39]. They adopt an open state at negative voltages, while positive voltages larger than  $\sim +20$  mV induce gating and closing [39, 40]. This salient feature is complimented by reversible ligand-induced gating, manifested as conformational changes in the presence of low concentrations of multivalent metal cations leading to channel closure [41, 42]. Once the multivalent cations bind and induce conformational changes, the channel adopts a sub-conducting or closed state [41, 42]. Another advantageous property of lysenin channels is that voltage and ligand-gating properties can be easily discriminated. This is achieved by reconstituting the channels in neutral lipid membranes which maintains the ligand-induced gating mechanism but renders lysenin unresponsive to the applied voltage [41, 42]. The high transport rate of lysenin channels yield large ionic currents which facilitate data recording and analysis. Lastly, lysenin channels are easily reconstituted in artificial membrane systems containing SM, are stable for extended time periods, and the monomer form of the protein is commercially available.

## Methods

### Chemicals and nanoparticles

Asolectine (Aso), cholesterol (Chol), SM (from Sigma-Aldrich) and diphyanoyl phosphatidylcholine (DiPhytPC, from Avanti Polar Lipids) were purchased as powders and

dissolved in n-decane at a final concentration of 50 mg/mL. For the support electrolyte, NaCl (Fisher Scientific) was dissolved in nanopure water at a final concentration of 130 mM (if not otherwise indicated) and buffered with 20 mM 4-(2-hydroxyethyl)-1-piperazineethanesulfonic acid (HEPES) at pH = 7.2. ZnO and SnO<sub>2</sub> NPs were synthesized using wet chemical methods as previously described [43, 44]. Briefly, for ZnO NP samples, the precursor zinc acetate dihydrate (Zn[CH<sub>3</sub>CO<sub>2</sub>]<sub>2</sub>·2H<sub>2</sub>O) was suspended in diethylene glycol. The solution was heated and nanopure water was added when the solution reached 80<sup>0</sup> C. The temperature was then brought to and held at 150<sup>0</sup> C for 90 minutes. The NPs were collected by centrifugation and subsequently washed with ethanol. For SnO<sub>2</sub> NPs, sodium stannate (Na<sub>2</sub>[Sn(OH)<sub>6</sub>]) and urea were used as precursors with nanopure water as the solvent. The solution was heated to 90 <sup>0</sup>C and held for 90 minutes. The NPs were collected via centrifugation and subsequently washed with nanopure water. Characterizations were performed using X-ray diffraction (XRD) (Figure S1, Additional file 1), transmission electron microscopy (TEM) (Figure S2 and S3, Additional file 1), zeta potential (ZP) measurements, dynamic light scattering (DLS) (Figure S4, Additional file 1), X-ray photoelectron spectroscopy (XPS) (Figure S5, Additional file 1) and Fourier-transform infrared spectroscopy (FTIR) (Figure S6, Additional file 1). XPS confirmed sample purity and atomic concentrations for stoichiometric ratios. XRD was employed to ensure crystal phase purity and to obtain average crystalline size for both samples. XRD confirmed the expected hexagonal wurzite crystal structure for ZnO and cassiterite for SnO<sub>2</sub>. The average crystal size for ZnO and SnO<sub>2</sub> NPs was analyzed with Rietveld refinement using Materials Analysis Using Diffraction (MAUD) software and estimated at 8.3 +/- 2 nm and 4.3 +/- 0.04 nm respectively. A JEOL JEM-2100 HR analytical TEM was used to confirm spherical morphology and average crystal sizes. FTIR spectra was collected using a Bruker Tensor 247 spectrometer and FTIR

pellets were produced by first grinding 1.6 mg of each NP sample with 0.200 g of spectroscopic grade KBr. The ground powder mixture was then pressed with 8 tons of pressure for 3 minutes and pellets were analyzed after removing the KBr background. Zeta potential and DLS measurements were performed, after dispersing the powders in nanopure water at a concentration of 1 mg/mL, using a Malvern Zetasizer NanoZS. ZnO NP clusters had an average hydrodynamic size (HDS) of 276 nm and average ZP of +32mV, whereas SnO<sub>2</sub> NP clusters average HDS was 176 nm with an average ZP of -42.0 mV.

### **Bilayer lipid membrane setup**

The experimental setup employed the use of a planar BLM chamber consisting of two polytetrafluoroethylene (PTFE) reservoirs separated by a thin (~120 $\mu$ m) PTFE film that had been pierced with an electric spark to create a circular hole of ~ 70  $\mu$ m diameter [45, 46]. The reservoirs were filled with 1mL buffered electrolyte and connected via two Ag/AgCl electrodes inserted in the solution to an Axopatch 200B amplifier (Molecular Devices). The amplified analog signal fed the DigiData 1440A digitizer (Molecular Devices) which provided the digital signal for visualization, recording, and further analysis. Continuous stirring of the solutions in the BLM chamber was assured by a low-noise magnetic stirrer (Warner Instruments). All the experiments were performed in voltage-clamp mode upon manual or automatic voltage stimulation. The signal recorded during various voltage stimulations was further analyzed with ClampFit 10.6.2.2 (Molecular Devices) and Origin 8.5.1 (Origin Lab) software packages.

### **Experimental procedure**

Lipid membrane preparation was performed by “painting” the hole in the PTFE film with small amounts of lipid mixtures composed of 4mg Aso or DiPhytPC, 2mg Chol, and 2mg SM dissolved in ~400 $\mu$ L n-decane [46, 47]. The successful creation of the BLM was indicated by measuring the capacitance in response to an applied triangular voltage stimulation, while achievement of a seal resistance larger than 1000 G $\Omega$  was assessed by measuring the leakage current in response to a DC voltage stimulation (100 mV). Channel insertion was performed by adding the lysenin monomer (from Sigma-Aldrich, 0.3 nM final concentration) to the ground (*cis*) reservoir under continuous stirring and at -60 mV bias potential applied to the *trans* (headstage) reservoir. The application of a negative voltage was required to prevent the voltage-induced gating which manifests at positive transmembrane potentials [33, 39, 40]. After the insertion process was completed, as indicated by a steady state value of the open current, an extensive flushing of the *cis* reservoir with lysenin-free electrolyte was performed to remove the bulk monomer and prevent additional insertions. To avoid potential changes in the lysenin functionality originating in congestion effects [48], the total number of channels inserted into the membranes was limited to ~1000. To facilitate quantitative comparison of the influence of NPs on the transport properties of lysenin channels in parallel experiments comprising different numbers of inserted channels, we used the relative changes in the macroscopic conductance ( $G_r = G/G_0$ ) for data plotting, where  $G$  is the conductance after addition of NPs and  $G_0$  is the conductance before addition. In order to avoid premature dissolution and/or aggregation, the NPs (powder form) were dispersed by sonication for 5 minutes in the support electrolyte solution in a sonication bath before each addition to the reservoirs.

## Results and Discussion

Once a steady state current through the population of lysenin channels was achieved, the NPs were introduced into either side of the chamber with both negative and positive voltages applied across the membrane to assess their effect on the macroscopic conductance (see Fig. 1 for a schematic of the setup). The addition of ZnO NPs (20 µg/mL final concentration) to either side of the membrane containing lysenin channels, when biased by -60 mV, yielded only a modest decrease of the macroscopic conductance, i.e. a few percent, irrespective of the side of addition (Fig. 2). This slight decrease in the conductance suggests a minimal influence of ZnO NPs on the lysenin channels' ability to transport ions in these particular experimental conditions.

To explain the small reduction in conductance, one may hypothesize several different mechanisms such as ligand gating induced by small amounts of Zn<sup>2+</sup> ions provided from low NP dissolution, ligand gating induced by NP binding to a specific binding site, or physical occlusion by transient NP attachment to the opening of the nanopore. Past investigations show a dramatic yet reversible decrease of the macroscopic conductance of lysenin channels in the presence of low concentrations of multivalent cations [41, 42], indicative of strong interactions with lysenin channels. Those interactions have been elucidated in single-channel experiments, which provide evidence of gating, i.e. transition from the open state to a sub-conducting or closed state [41, 42]. To explain lysenin's reversible gating in the presence of multivalent cations, it is assumed that the channel's structure contains at least one negatively-charged binding site with high affinity for cations, which triggers gating upon binding. A potential leakage of Zn<sup>2+</sup> ions from NPs may affect the macroscopic conductance of lysenin channels, as observed. In addition, if exposed this binding site could electrostatically interact with cationic NPs and yield a significant decrease in conductance either by induced gating or physical occlusion of the conducting pathway. However,

such strong effects were not observed in the above experimental conditions, which prompted us to look closer to the lysenin's structure for alternative explanations. The assembled lysenin channel shows the presence of multiple anionic domains [36, 37, 49], hence presenting opportunities for physical occlusions of the channels through electrostatic interactions even in the absence of gating. We may account for the weak conductance inhibition by considering the position of a binding site and the orientation of the external electric field. A deep-buried binding site would be inaccessible from either side to NPs larger than the channel's diameter (~3 nm), which is mostly the case in our investigations. Nanoparticle interaction with a binding site present at the *trans* opening of the channel would be prevented at -60 mV by the electric field orientation. Although the electric field in the bulk is very low, its amplitude increases substantially when approaching the channel opening (fringe effects), therefore keeping the NPs far from a binding site located at the *cis* opening. The same electric field will drive the NPs added to the *cis* side towards the membrane but the lack of changes in macroscopic conductance suggest the absence of a binding site at this location. The hypothesis of an exposed binding site at the *trans* opening was further sustained in similar experiment comprising *trans* NP addition and no transmembrane voltage; in such experimental conditions, a marked decrease of the macroscopic conductance was observed at -60 mV after two hours of NP incubation in the absence of a bias potential (data not shown). However, this result could be an artifact originating from dissolution during the prolonged NP exposure to the electrolyte solution.

To identify if the elusive binding site is located either deep within the channel or at the *trans* side, we performed the experiments under positive bias potentials (Figure 3). After the channel insertion process, the influence of ZnO NPs was assessed in experiments comprising of *cis* or *trans* addition and opposite orientations of the electric field. Lysenin channels are voltage-

gated at positive voltages greater than  $\sim +20$  mV but are stable in the open state for extended time periods as long as the applied voltage is less than this critical value [33, 39]. Interestingly, addition of ZnO NPs to the *trans* side under positive biasing (+15 mV to prevent voltage gating) induced a rapid and sustained decrease of the macroscopic conductance (Fig. 3), while *cis* addition elicited only a weak response in otherwise similar conditions. Consequently, we concluded that the electric field plays a major role in preventing ZnO NPs accumulation near the membrane when biased by -60 mV, however, in the absence of an electric field or when positive voltages are applied, ZnO NPs may interact with a binding site situated at the *trans* opening of the channel.

Dissolution of ZnO NPs can result in high extracellular  $Zn^{2+}$  concentrations which have been proposed as one of the main mechanisms of ZnO NPs cytotoxic effects [22, 23, 50]. Zinc ions inhibit the macroscopic conductance of lysenin channels by a ligand-induced gating mechanism [41, 42]. Due to the high sensitivity of lysenin channels to  $Zn^{2+}$ , dissolution may explain the observed inhibition of conductance upon exposure to ZnO NPs. To eliminate such potential experimental artifacts, we performed investigations in similar conditions but added  $Zn^{2+}$  ions ( $ZnSO_4$ ; 2 mM final concentration) to the reservoirs instead of ZnO NPs. Addition of  $Zn^{2+}$  to the either side, biased by -60 mV, yielded a sudden decrease of the macroscopic conductance in agreement with previous reports (Fig. 4) [41, 42]. Addition of the same amount of  $Zn^{2+}$  to a similar BLM containing lysenin channels and biased by +15 mV (to prevent voltage-induced gating) yielded a similar relative decrease of the macroscopic open current (Fig. 4). If conductance inhibition elicited by ZnO NPs had been induced by the  $Zn^{2+}$  ions dissipating from the NPs, then addition to either side would have displayed a similar pattern of conductance inhibition. However, addition of  $Zn^{2+}$  ions yielded fundamentally different results compared with

the experiments involving ZnO NPs.  $Zn^{2+}$  ions affected the macroscopic conductance irrespective of the side of addition and direction of the electric field, while the inhibitory activity of ZnO NPs depended on both these experimental parameters. The total concentration of ZnO NPs was only 20  $\mu$ g/mL (corresponding to ~0.25 mM  $Zn^{2+}$  ions) and resulted in a 70% decrease in the macroscopic conductance. In order to obtain an approximate decrease of only 45% in conductance measurements with  $Zn^{2+}$  ions, the experiment employed a final concentration of 2.0 mM. Assuming complete dissolution of ZnO NPs, this would correlate to approximately eight times the amount of  $Zn^{2+}$  ions from  $ZnSO_4$  in the solution. To further eliminate the possibility that the  $Zn^{2+}$  ions contributed to the observed conductance inhibition, experiments with ZnO NPs were carried out in the presence of the strong  $Zn^{2+}$  chelator EDTA. EDTA (10 mM) was added to the solutions prior to nanoparticle addition, thus effectively preventing any interactions of the free zinc ions from the NPs with lysenin channels. These experiments yielded almost identical decreases in the macroscopic conductance when compared with ZnO NPs with no EDTA (Figure S7, Additional file 1). Our results clearly indicate that the conductance inhibition elicited by ZnO NPs was not a consequence of  $Zn^{2+}$  ions from dissolution. These experiments revealed that the extent of the conductance inhibition depended on both the orientation of the lysenin channels and the electric field relative to the site of ZnO NP addition. The observed conductance inhibition may originate from local accumulation of NPs by electrophoretic effects, specific interactions with the membrane itself, or preferential interactions with binding sites of lysenin.

Next, we asked whether or not exposure to ZnO NPs changes the voltage-induced gating profile. To answer this question, the voltage-induced gating of lysenin channels was assessed from the I-V plot recorded in the range -60 to +60 mV (Fig. 5) at a voltage rate of 0.2 mV/s with and without the addition of ZnO NPs. The macroscopic current recorded in absence of NPs (Fig.

5) featured the well-known characteristics of voltage-induced gating, i.e. a linear behavior in the negative voltage range, indicative of the absence of gating, and a non-linear behavior at positive voltages higher than +20 mV, indicative of channel closure [33, 39, 40]. A typical feature of the macroscopic current recorded at positive voltages is the transition from high current to low current through a dynamic negative resistance region [39, 40]. The macroscopic currents recorded in the same voltage range after addition of ZnO NPs (20  $\mu$ g/mL final concentration) to the *trans* side of the membrane yielded a fundamentally different I-V plot (Fig. 5). The addition of ZnO NPs elicited a slight decrease in the open current recorded in the negative voltage range, however, the I-V characteristic preserved quasi-linearity between -60 mV and -20 mV. Once the voltage approached neutral values, the macroscopic current greatly deviated from the control I-V plot and the ionic transport capabilities of lysenin channels were strongly diminished in the presence of ZnO NPs. Interestingly, the macroscopic conductance started to decrease at small negative voltages, as indicated by the diminished slope of the I-V plot. This is consistent with the hypothesis that the fringe effect of the electric field prevents the NPs from interacting with the binding site. The magnitude of the electric field decreases with decreasing applied voltage and the weak electrophoretic force, although opposed, is not sufficient to prevent interactions with the binding site and channel conductance modulation. The consistently lower macroscopic currents indicated that addition of ZnO NPs induced severe channel conductance inhibition as demonstrated by the large decrease of the macroscopic current at any positive voltage. At positive voltages, the currents recorded in the presence of ZnO NPs were consistently lower than the currents recorded in the absence of ZnO NPs up to  $\sim$ +40 mV, after which the recorded currents were similar to the control when the channels are in a closed state.

This experiment demonstrated that ZnO NPs affect the macroscopic conductance of lysenin channels in a voltage-dependent manner but it does not offer a complete mechanistic description. The significant changes in the I-V curve after addition of ZnO NPs potentially stem from multiple mechanisms. Experimental evidence and theoretical modeling have demonstrated that electrostatic interactions between membrane components and NPs are key factors that contribute to toxicity and the ability of NPs to internalize into cells [51-53]. Our experiments comprised a simple system consisting of lysenin channels inserted into an artificial BLM composed of charged lipids. We assumed that the conductance of lysenin channels was affected by interactions between the proteins channels and NPs but we could not exclude interactions between the charged lipids and ZnO NPs as a source of conductance modulation. The Aso lipid mixture used for BLM preparation contains several anionic components that may interact electrostatically with voluminous cationic NPs unable to penetrate the lumen, which would lead to channel conductance modulation. To elucidate the potential role played by the charged lipids, we performed experiments by replacing Aso with neutral DiPhytPC. The use of neutral lipids abolishes the voltage-induced gating at positive voltages while preserving the ligand-induced gating observed in the presence of multivalent cations [41, 42]. Addition of ZnO NPs to the *cis* side of a neutral membrane containing lysenin channels and biased by +60 mV elicited no change in the macroscopic conductance (Fig. 6). However, addition of ZnO NPs to the *trans* side of the same membrane, biased by an identical positive voltage, yielded a massive decrease in conductance similar to the results obtained using charged lipids (Fig. 6). The non-symmetrical response and preservation of the inhibitory capabilities of ZnO NPs recorded for the neutral BLM suggest that the inhibition mechanism excludes electrostatic interactions between NPs and

lipids. The interaction between lysenin channels and ZnO NPs is therefore likely responsible for the observed inhibitory activity.

We have shown that lysenin channels interact with positively charged ZnO NPs but have not yet demonstrated the electrostatic nature of those interactions. Therefore, we asked whether or not any NPs electrophoretically driven towards a specific or non-specific yet accessible binding site would interact with lysenin channels and inhibit their conductance. In this respect, we performed conductance experiments by replacing positively charged ZnO NPs with negatively charged SnO<sub>2</sub> NPs (-42 mV ZP). Irrespective of the applied voltage and the addition site, SnO<sub>2</sub> NPs did not affect the macroscopic conductance of lysenin channels (Fig. 7). In order to try to elicit interactions with SnO<sub>2</sub> NPs, 200 µg/mL (final concentration) of SnO<sub>2</sub> NPs were used, which is 10x the concentration of ZnO NPs that induced rapid decreases in the macroscopic conductance (Fig. 3). The crystal and hydrodynamic sizes of SnO<sub>2</sub> NPs used in this experiment were much smaller than ZnO NPs, suggesting that SnO<sub>2</sub> NPs would be better suited to inhibit conductance by physical occlusion. The absolute magnitude of the ZP for SnO<sub>2</sub> NPs was also larger than ZnO NPs, further strengthening the hypothesis of a mechanism that requires strong electrostatic interactions between cationic ZnO NPs and an anionic domain present at the *trans* side of the lysenin channel to induce conductance inhibition. Also, to further support the hypothesis that electrostatic interactions between the lysenin channels and ZnO NPs initiate a decrease in conductance, we investigated the effects of electrostatic screening induced by an increased ion concentration in the bulk electrolyte solutions. Addition of 20 µg/mL ZnO NPs to the *trans* side of the bilayer containing lysenin channels in 500 mM NaCl and under positive bias reduced the conductance by ~15% (Figure S8, Additional file 1), which is much smaller than what we observed at 130 mM NaCl concentration (~70%, Fig. 2). In addition, the time required

to reach equilibrium increased to more than 2500 s, indicating that ionic screening weakened the interactions between NPs and lysenin channels, and supporting the hypothesis that electrostatic interactions are at the origin of the observed changes in conductance. However, we may not eliminate potential artifacts arising from the effects of screening on the ZnO NPs. At high salt concentration, screening may accelerate NP aggregation, which is what we observed when attempting to further increase the ionic concentration of the bulk electrolyte solutions. The ZnO NPs rapidly aggregated into large clusters at the bottom of the vials in a matter of minutes, which prevented further experimentation in high ionic strength conditions.

A few assumptions can be proposed about the mechanism responsible for the observed decrease in macroscopic conductance elicited by addition of ZnO NPs. Electrostatic interactions may bring ZnO NPs close enough to the channels such that the resulting physical blockage reduces the individual currents. In such case, an opposite electric field of appropriate magnitude may drive the NPs away from the binding site therefore unblocking the channels. Our attempts to apply higher voltages across the BLM and to force the unblocking were not successful. However, it is possible for the binding site to have a relatively strong affinity for charged ZnO NPs and consequently the force required to remove the NPs from the binding sight may require much higher electric fields. Unfortunately, such experiments are very difficult to achieve as the BLM is prone to disruption at high transmembrane voltages.

Another potential inhibition mechanism mimics ligand-induced gating. It has been established that lysenin channels interact with multivalent cations and undergo conformational transitions that force the channel into closed or sub-conducting states [41, 42]. This ligand-induced gating mechanism relies on electrostatic interactions between cations and one or more binding sites but ionic current blockage stems from the induced gating. It is possible that charged

ZnO NPs interact electrostatically with one or more binding sites, yet not necessarily the same one(s) involved in the ligand-induced gating and would force the channels to adopt a subconducting or a closed state. Lastly, defects on the surface of ZnO NPs such as oxygen vacancies have been shown to correlate with ROS production [24, 54]. Since the electrostatic interactions induce close contact of ZnO NPs with the channels, the highly reactive surface of ZnO NPs may interact with cysteine and methionine residues in their structure which may alter channel functionality and conduction similar to reports of oxidation of cysteine residues in  $\text{Ca}^{2+}/\text{K}^+$  channels [55, 56].

## Conclusions

Our work demonstrates that the transport properties of lysenin channels change significantly in the presence of cationic ZnO NPs. The modulation of the transport properties by NPs is strongly dependent on the net charge, and the orientation of the electric field and channel with respect to the NPs. There is little doubt that the primary interaction between NPs and lysenin channels is electrostatic. Nonetheless, the simplicity of the experimental system investigated here does not necessarily warrant biological interpolation to other protein channels interacting with NPs, not even ZnO. In complex biological environments, the binding of various functional groups on the NP surface may significantly alter their ability to interact with membrane components irrespective of the surface charge of the pristine nanomaterial. Given the aggregation tendency of the investigated NPs, we may not exclude aggregation at the membrane surface as being at the origin of conductance changes. Even the neutral lipids used for our investigations present a dipole moment that may initiate NP binding; further NP aggregation at

these binding sites may impede the ionic flow by physical occlusion or by introducing supplementary electrostatic energy barriers for ions. However, if an induced dipole moment that initiated binding of NPs to lipids occurred, then SnO<sub>2</sub> NPs should have yielded a similar response due to their higher net charge. Nonetheless, dipole-charge interactions have a much smaller magnitude than the charge-charge ones, and we did not observe such effects when using neutral lipids. In spite of these shortcomings, an important conclusion of this report pertains to the potential ability of NPs to interact with transmembrane transporters without the need of internalization. Many previous studies assume that cytotoxic effects of NPs are due to translocation of NPs into the cytosol by various transport mechanisms and/or dissolution of the NPs, disrupting homeostasis and interfering with vital cellular processes. Our work suggests that NPs may tamper with ionic transport mechanisms by basic electrostatic interactions. Given the physiological relevance of controlled transmembrane transport, such alterations may have catastrophic effects for cells. While this observation is generally valid for any cell, it may prove extremely helpful for understanding the potential neuro-toxic effects of NPs [57]. The physiology of the neural cell is based on the transport properties and regulation of voltage-gated ion channels, which are transmembrane structures with multiple charged domains that may interact electrostatically with NPs. Changes in the voltage-induced gating mechanism or blockage of ionic transport induced by NPs [57] may dramatically affect the correct functionality of the nerve cell. Such interactions may explain why certain NPs specifically alter the individual currents through specific channels while the transport properties of other channels are not affected by various NPs [58, 59]. The local distribution of charge within the structure of several ion channels is currently known so it may be possible to predict potential toxic effects based on

interactions with charged NPs, or to design NPs intended to alter the activity of transmembrane transporters.

This foray into deciphering the effects of NPs on the transmembrane transport of ions indicates alterations in the transporters' functionality as a potential mechanism of cytotoxicity. A previous study shows that ZnO NPs may induce neuronal cytotoxicity and genotoxicity in the absence of internalization or free Zn<sup>2+</sup> ions released from the NPs [60]. Future experiments will shed more light on intimate mechanistic details and the role that electrostatic interactions play in modulating the biological activity of protein channels.

### **List of abbreviations**

NMs, Nanomaterials; NPs, nanoparticles; ROS, reactive oxygen species; BLM, bilayer lipid membrane; SM, sphingomyelin; Aso, Asolectine; Chol, cholesterol; DiPhytPC, diphyanoyl phosphatidylcholine; HEPES, 4-(2-hydroxyethyl)-1-piperazineethanesulfonic acid; XRD, X-ray diffraction; TEM, transmission electron microscopy; ZP, zeta potential; DLS, dynamic light scattering; XPS, X-ray photoelectron spectroscopy; HDS, hydrodynamic size; PTFE, polytetrafluoroethylene.

### **Declarations**

#### **Ethics approval and consent to participate**

Not applicable.

#### **Consent for publication**

Not applicable.

#### **Availability of data and materials**

The datasets used and/or analyzed during the current study are available from the corresponding author on reasonable request.

### **Competing interests**

The authors declare that they have no competing interests.

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### **Author's contributions**

Conceived and designed experiments: DF. Wrote the paper: DF, JEE, SLB, NS. Performed nanoparticle/lysenin interaction experiments: SLB, DF, SR, HA, CH, NS, MM. Performed nanoparticle/lysenin interaction experiments, data analysis, and interpretation: DF, SLB, JEE, SR, HA, CH, NS, MM. Nanoparticles fabrication and characterization/analysis: JEE, AP. All authors read and approved the final manuscript.

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SLB and JEE contributed equally to this work.

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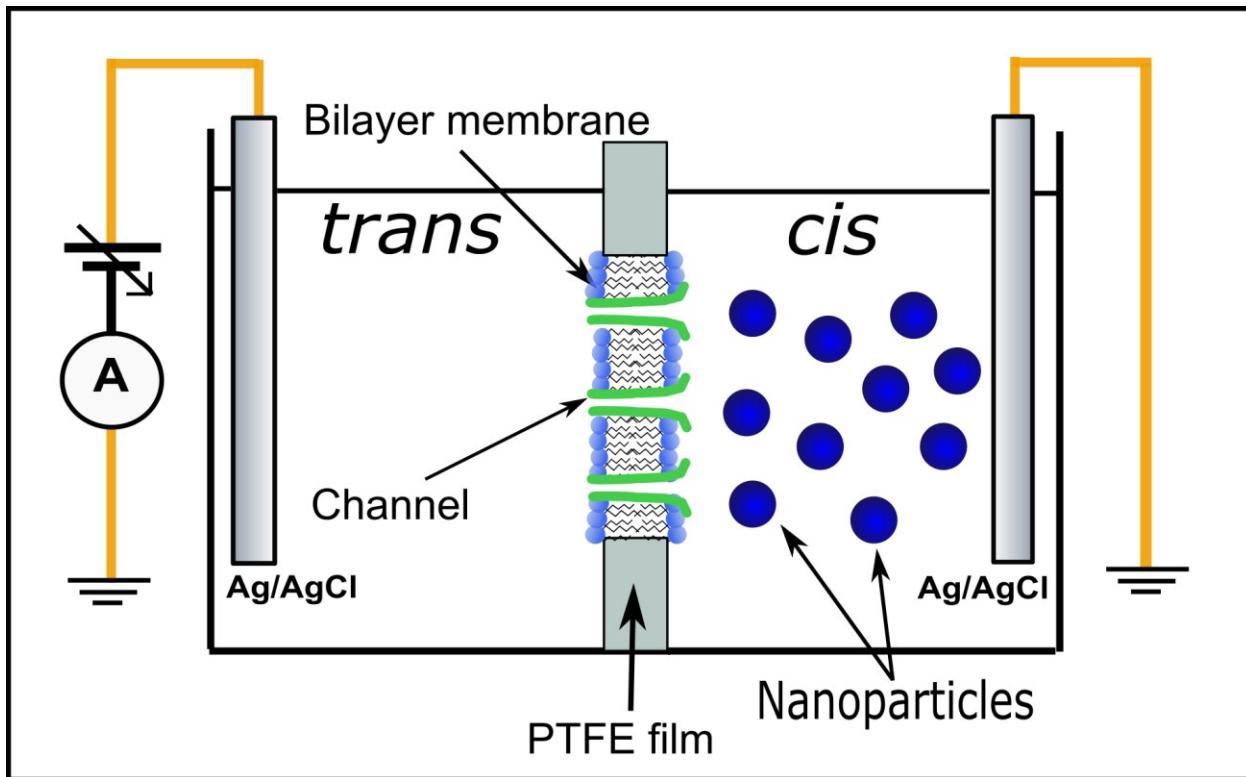
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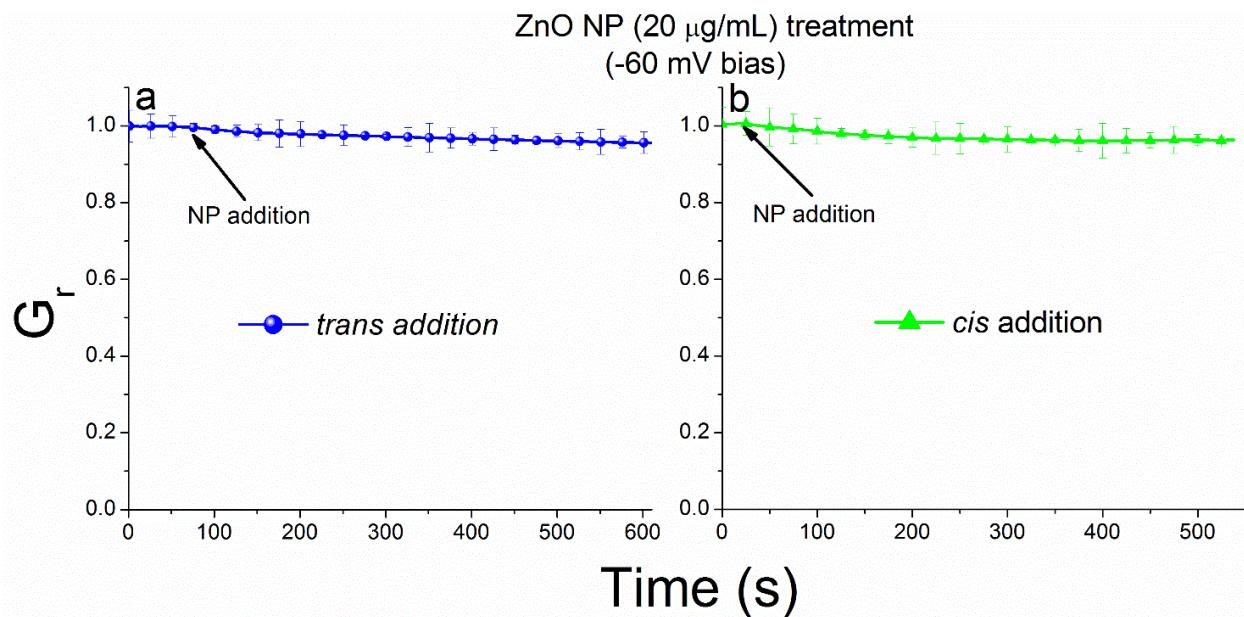
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**Additional file.** Additional file 1, pdf. Supplementary material that includes complementary figures of material characterization (Fig. S1-Fig. S6: XRD, TEM, DLS, XPS, and FTIR for ZnO and SnO<sub>2</sub> NPs), changes in relative macroscopic conductance of lysenin channels in the presence of ZnO NPs and EDTA (Fig. S7), and conductance inhibition at high ionic strength (Fig. S8).

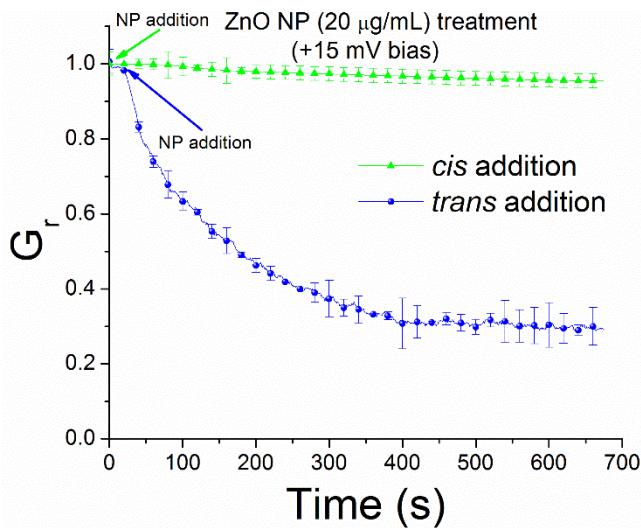
## Figures



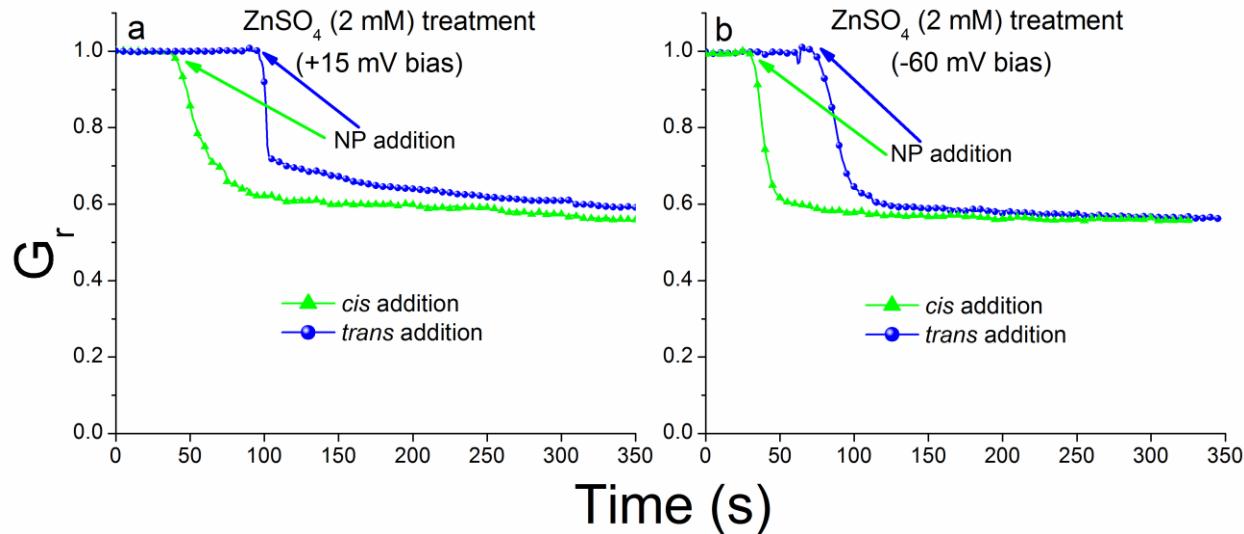
**Figure 1** The experimental setup comprises lysenin channels reconstituted into planar lipid bilayer membranes. The modulation of ionic transport and regulation by ZnO NPs is assessed in classic voltage-clamp experiments.



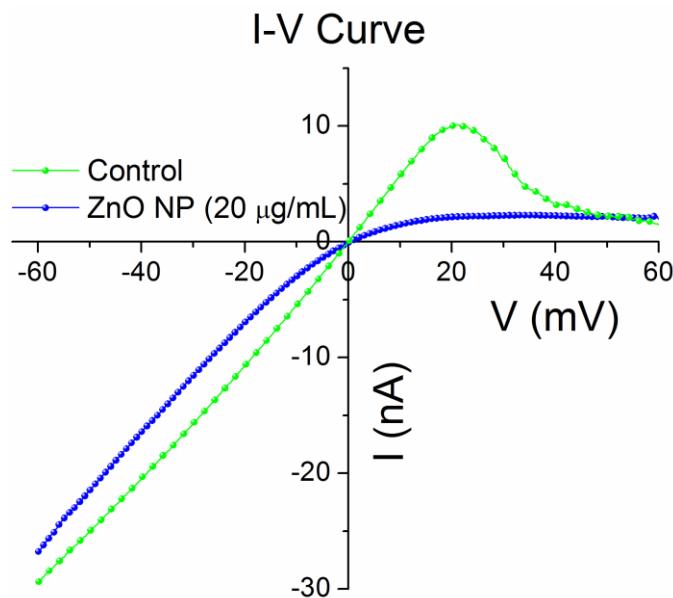
**Figure 2** ZnO NPs do not alter the ionic conductance of lysenin channels when biased by a -60 mV transmembrane voltage. Addition of ZnO NPs to either *trans* (a) or *cis* (b) reservoirs induces only negligible changes of the macroscopic conductance. The experimental values are reported as mean  $\pm$  S.D.,  $n = 3$ . All the data points represent experimental values but some symbols have been removed for improved visibility.



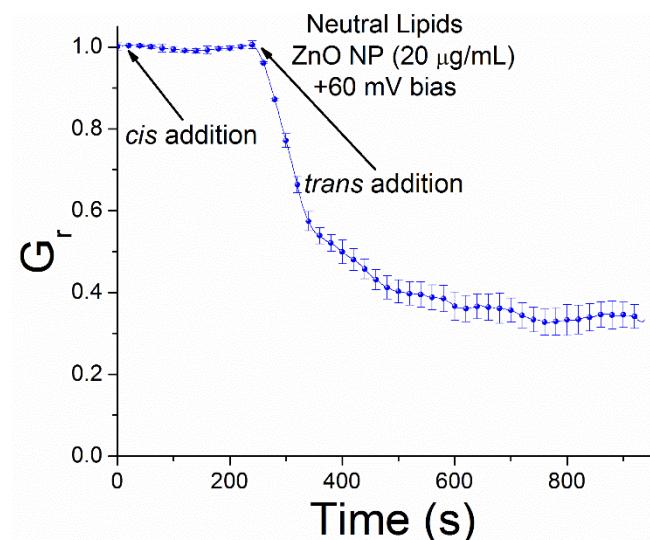
**Figure 3** Interactions between lysenin channels and ZnO NPs at +15 mV bias potential. *Cis* addition (green) of ZnO NPs yield minor changes in the macroscopic conductance. In contrast, *trans* addition (blue) elicits a significant decrease of the macroscopic conductance by ~70 %. The experimental values are reported as mean  $\pm$  S.D.,  $n = 3$ . All the data points represent experimental values but some symbols have been removed for improved visibility.

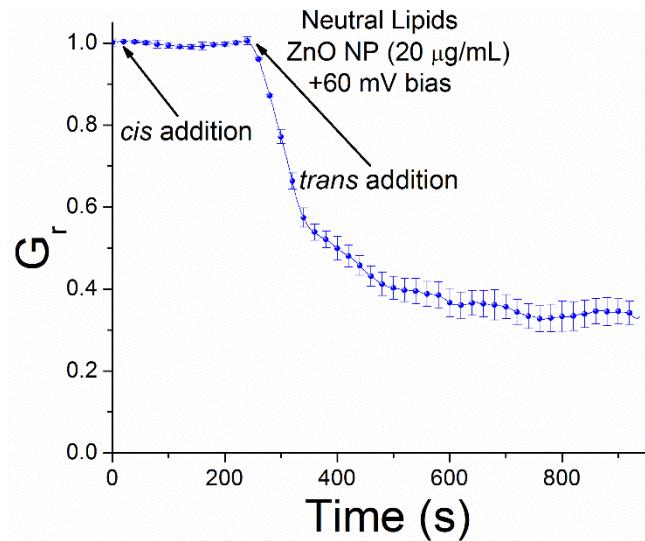


**Figure 4**  $Zn^{2+}$  ions decrease the macroscopic conductance of lysenin channels irrespective of bias potential and site of addition. At +15 mV transmembrane voltage (a),  $Zn^{2+}$  addition to either the *cis* or *trans* reservoir reduces the macroscopic conductance by ~40%. Similar decreases are recorded upon  $Zn^{2+}$  interactions with lysenin channels biased by -60 mV (b). The presented data represents a typical run for each experiment.

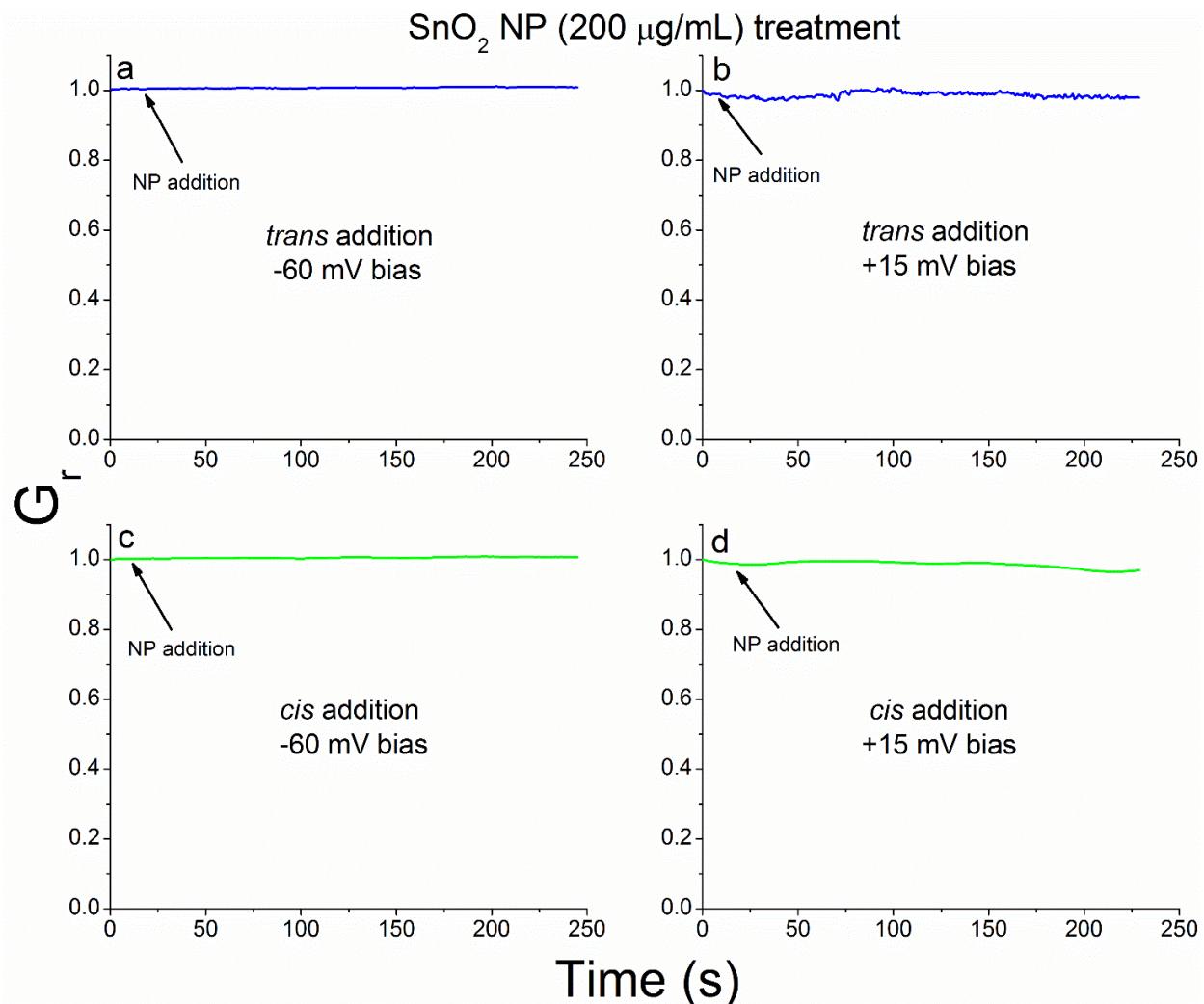


**Figure 5** Effects of ZnO NPs on lysenin voltage-induced gating. In the absence of NPs, lysenin channels begin to close at transmembrane potentials greater than 20 mV (green curve). ZnO NPs (20  $\mu$ g/mL final concentration) almost completely abrogate the conductance in the positive voltage range (blue curve) and indicate a strong interaction with the lysenin channels. All points on the curves are experimental data and symbols have been added for discrimination. The presented data represents a typical run for each experiment.





**Figure 6** Lysenin channels reconstituted in neutral lipid membranes interact with ZnO NPs at +60 mV transmembrane voltage. *Cis* addition of ZnO NPs elicits no changes in the macroscopic conductance. In contrast, ZnO NPs added to the *trans* reservoir interact with lysenin channels and significantly diminish their ionic transport capabilities. The experimental values are reported as mean  $\pm$  S.D.,  $n = 3$ . All the data points represent experimental values but some symbols have been removed for improved visibility.



**Figure 7** Interactions between anionic  $\text{SnO}_2$  NPs and lysenin channels reconstituted into a planar bilayer lipid membrane. Addition of  $\text{SnO}_2$  NPs to the *trans* reservoir at  $-60 \text{ mV}$  (a) and  $+15 \text{ mV}$  (b) indicates insignificant changes of the macroscopic conductance. Similarly,  $\text{SnO}_2$  NP addition to the *cis* reservoir at  $-60 \text{ mV}$  (c) and  $+15 \text{ mV}$  (d) yields negligible changes in the ionic transport capabilities. The presented data represents a typical run for each experiment.

## Supplementary Material

### ZnO Nanoparticles Modulate the Ionic Transport and Voltage Regulation of Lysenin Nanochannels

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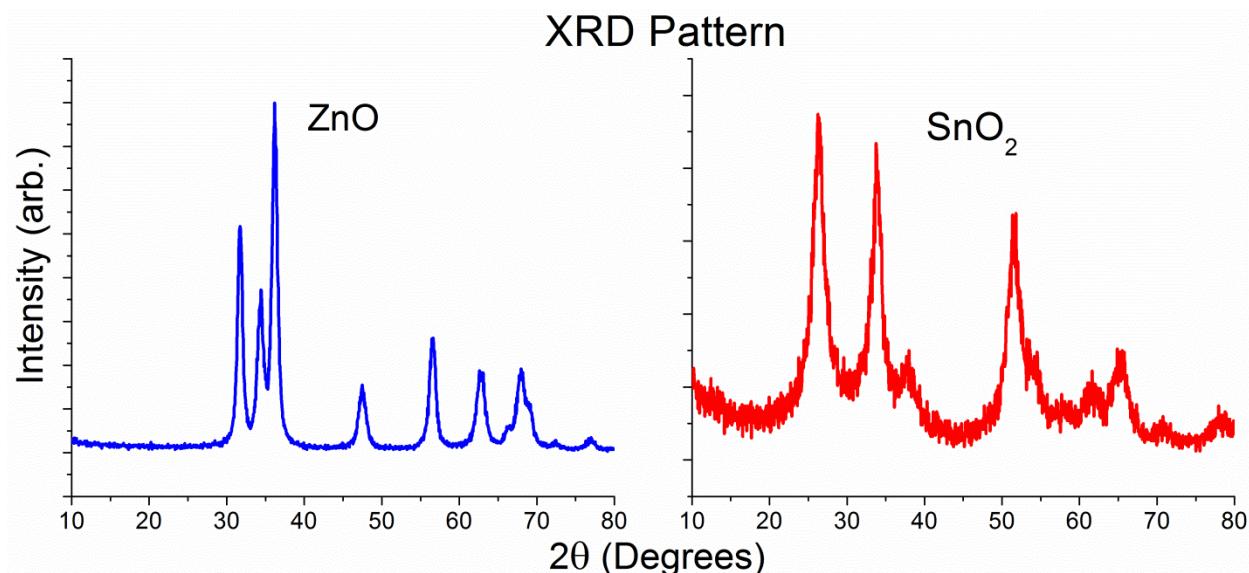
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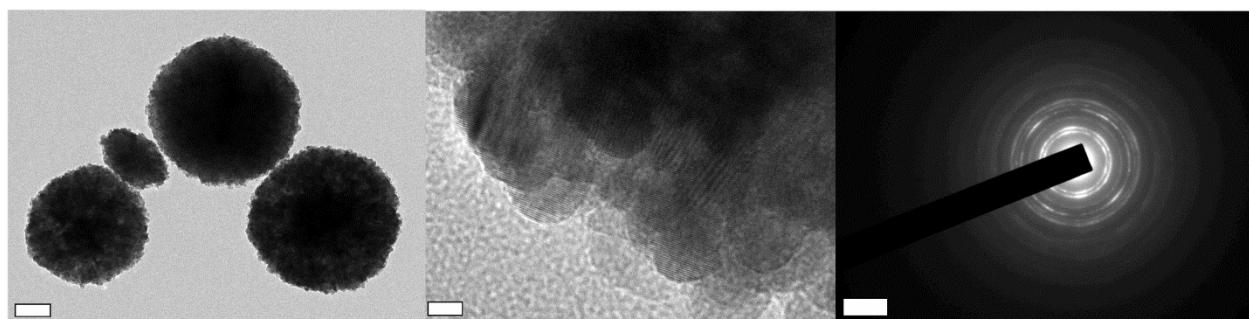
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Supplementary Data for Material Characterization



*Figure S1.* The XRD spectra of ZnO (left) and SnO<sub>2</sub> (right). For ZnO, XRD confirmed the hexagonal wurzite crystal structure with no alternate crystal phases detected. The average crystal size determined for ZnO was estimated at 8.3 nm. The XRD spectra of SnO<sub>2</sub> demonstrated the crystal phase obtained was cassiterite with no alternate phases present. The average crystal size for SnO<sub>2</sub> was determined to be 4.3 nm from the XRD spectra.



*Figure S2.* TEM images and electron diffraction pattern of ZnO nanoparticles. (Left; scale bar:

100 nm) The lower magnification image shows that the ZnO crystals aggregate to form larger spheres. (Middle; scale bar: 5 nm) Higher magnification images demonstrate that the large aggregates are comprised of small ZnO nanoparticles with an average size of  $\sim$ 10 nm. (Right; scale bar: 5 l/nm) The electron diffraction pattern obtained from the ZnO nanoparticle sample.

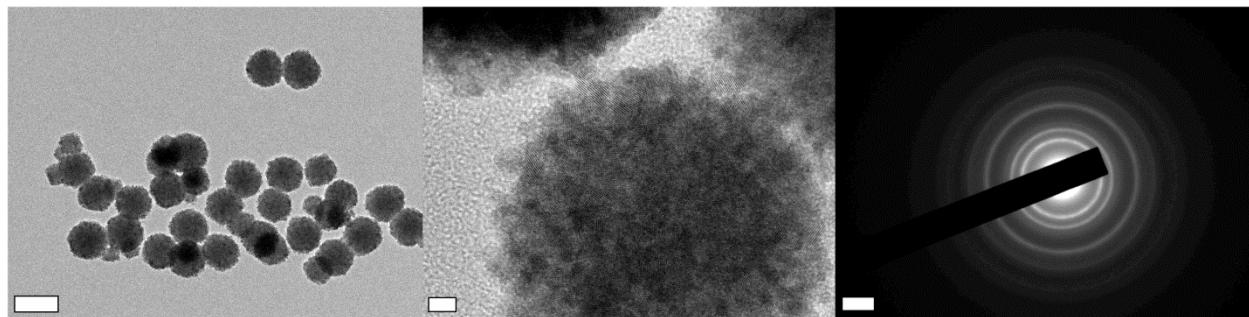


Figure S3. TEM images and electron diffraction pattern of  $\text{SnO}_2$  nanoparticles. (Left; scale bar: 100 nm) The lower magnification image shows that  $\text{SnO}_2$  nanoparticles also form larger aggregates which are generally smaller than the aggregates seen from the ZnO nanoparticles used. (Middle; scale bar: 5 nm) Higher magnification reveals that the aggregates formed are from small  $\text{SnO}_2$  nanoparticles with an average size of  $\sim$ 5 nm. (Right; scale bar: 2 l/nm) The electron diffraction pattern obtained from the  $\text{SnO}_2$  nanoparticle sample.

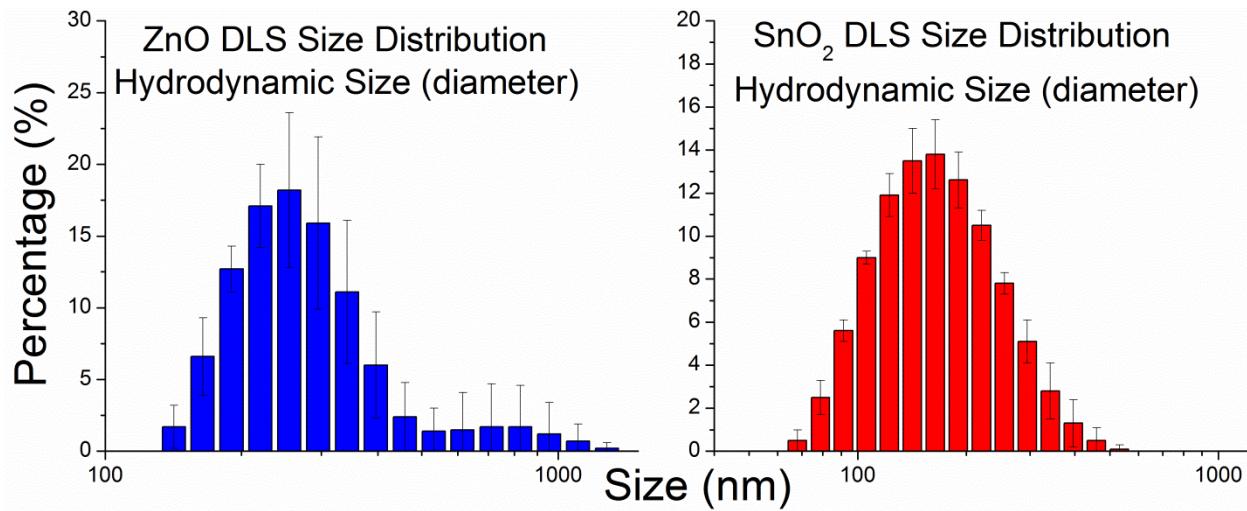


Figure S4. Hydrodynamic size distributions of ZnO (left) and SnO<sub>2</sub> (right) obtained from dynamic light scattering measurements. Hydrodynamic size distributions of SnO<sub>2</sub> demonstrate smaller average aggregates than ZnO, consistent with images obtained from TEM. The smaller particle size of SnO<sub>2</sub> suggests that these NPs would be better suited to inhibit conductance of lysenin channels by physical occlusion but this was not observed experimentally.

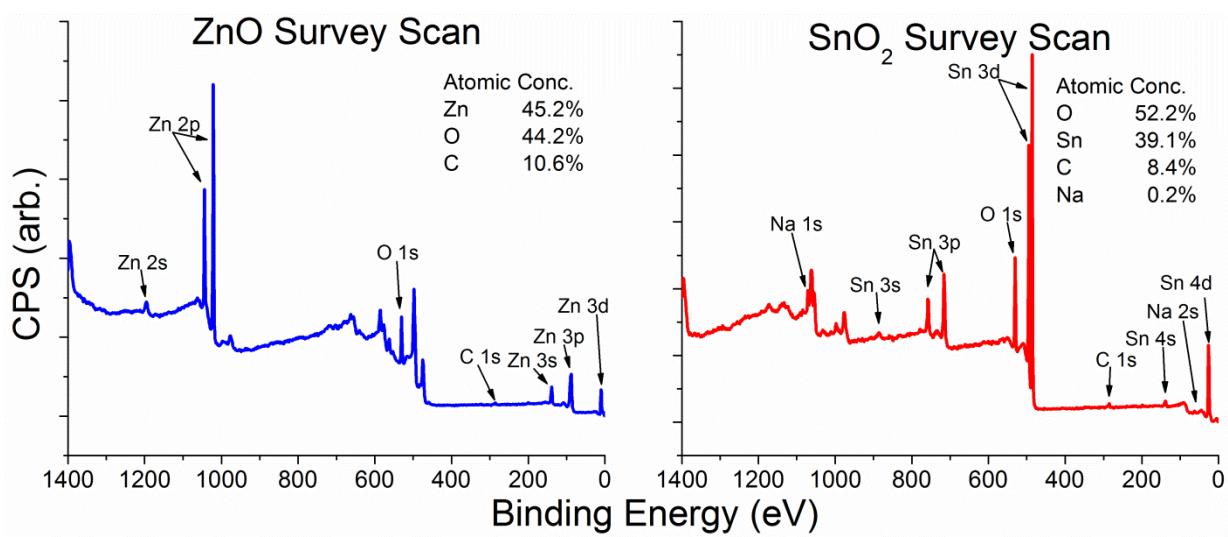


Figure S5. X-ray photoelectron spectroscopy survey spectra of  $ZnO$  (left) and  $SnO_2$  (right) nanoparticles. Spectra demonstrate high sample purity, small retention of carbon species and sodium in the case of  $SnO_2$ , likely from carbon dioxide from atmospheric exposure and species retained from the chemical precursors.

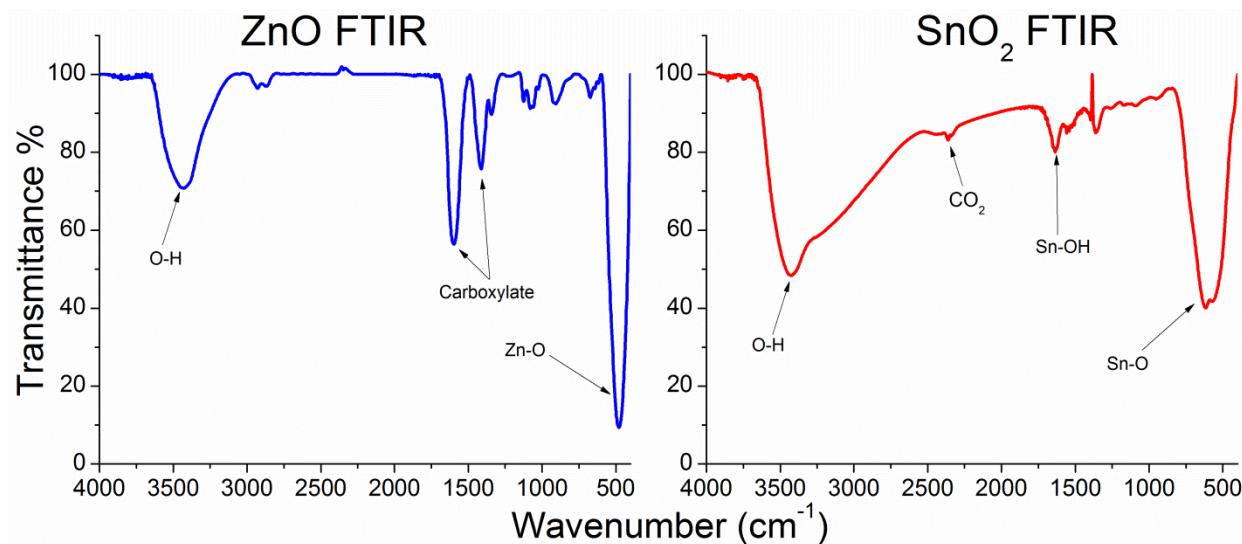
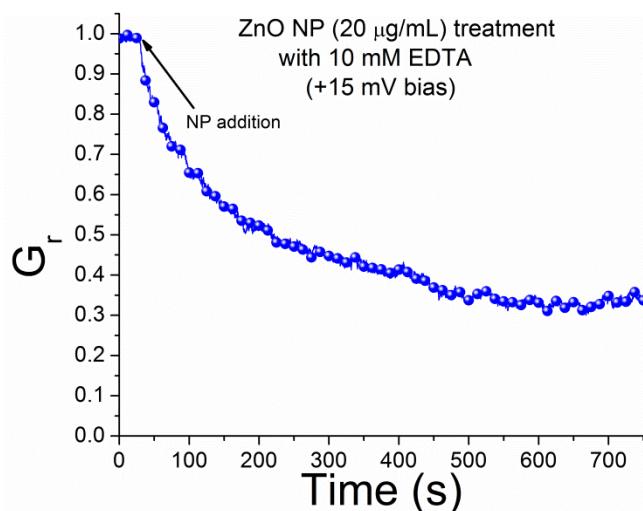
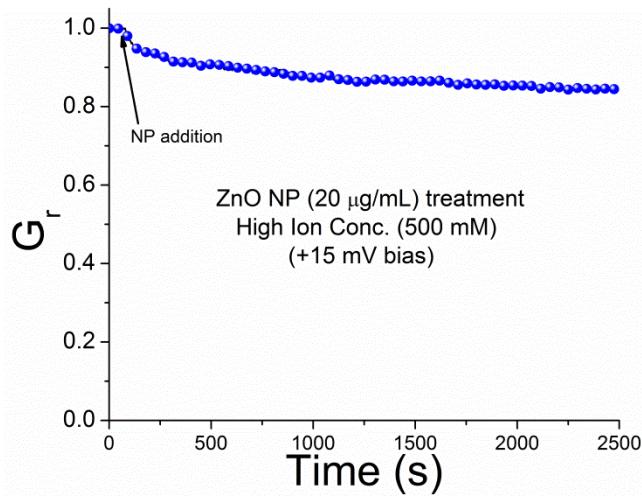


Figure S6. Fourier-transform infrared (FTIR) spectra of  $ZnO$  (left) and  $SnO_2$  (right) nanoparticles. Spectra confirm retained species from chemical precursors and carbon dioxide, as well as demonstrate hydroxide species are present in both nanoparticles synthesized.

Supplementary data for conductance inhibition of lysenin channels by NPs



*Figure S7. Interactions between ZnO nanoparticles and lysenin channels with EDTA (10 mM) present in the solution reservoirs. EDTA is a strong chelator of  $Zn^{2+}$  ions and thus prevents interactions between lysenin nanochannels and zinc ions. Nevertheless, ZnO nanoparticles elicited a strong conductance inhibition of lysenin channels, even with EDTA present, clearly demonstrating that inhibition is due to interactions between ZnO and lysenin, not free zinc ions.*



*Figure S8. Ionic conductance of lysenin channels in the presence of ZnO nanoparticles while using a significantly higher electrolyte concentration in solution (500 mM). When compared with experiments under low electrolyte solutions (130 mM), the ZnO nanoparticles elicited a much weaker response. The reduction in conductance inhibition may be due to: 1) Debye screening effects that reduce the electrostatic interactions between ZnO nanoparticles and lysenin channels, preventing the initial binding event required for conductance inhibition and 2) increased aggregation of ZnO nanoparticles, which reduces interactions between lysenin channels and ZnO nanoparticles.*