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ARTICLE

Direct Triplet Sensitization of Oligothiophene by Quantum Dots

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Effective sensitization of triplet states is essential to many applications, including triplet-triplet annihilation based photon upconversion schemes. This work demonstrates successful triplet sensitization of a CdSe quantum dot (QD)-bound oligothiophene carboxylic acid (T6). Transient absorption spectroscopy provides direct evidence of Dexter-type triplet energy transfer from the QD to the acceptor without populating the singlet excited state or charge transfer intermediates. Analysis of T6 concentration dependent triplet formation kinetics shows that the intrinsic triplet energy transfer rate in 1:1 QD-T6 complexes is 0.077 ns⁻¹ and the apparent transfer rate and efficiency can be improved by increasing the acceptor binding strength. This work demonstrates a new class of triplet acceptor molecules for QD-based upconversion systems that are more stable and tunable than the extensively studied polyacenes.

Introduction

Triplet excitons in organic materials exhibit long lifetime, extended diffusion length and low-lying energy levels,¹⁻³ and have found promising applications in photon upconversion.⁴ For example, an optical upconversion layer on a solar cell can capture sub-bandgap photons and emit above bandgap photons,⁵ increasing the efficiency of the conventional single-junction devices beyond the Shockley-Queisser limit.⁶ Various strategies for improving the overall efficiency of upconversion systems have been developed.⁷⁻¹⁰ In recent years, quantum dot (QD) sensitized upconversion systems¹¹⁻¹⁴ have attracted intense interest as a versatile and promising approach because of their large absorption coefficient,¹⁵ small singlet-to-triplet energy gap and fast intersystem crossing (ISC) rate,¹⁶ and tunable band gap and band alignment.¹⁷

In a typical photon-upconversion system, triplet excitons are generated through a sensitizer that undergoes intersystem crossing from an excited singlet state to a triplet state. This process is followed by a sequential triplet energy transfer first to the transmitter then to the emitter. The latter, can undergo triplet-triplet-annihilation (TTA) and emit a higher energy photon.^{4, 11, 13, 18-22} Thus, the efficiency of a typical sensitizer-emitter-based upconversion system can be represented by the following equation:⁴

$$\Phi_{UC} = \Phi_{ISC}\Phi_{TET}\Phi_{TTA}\Phi_{FL} \text{ (Eq. 1)}$$

The overall upconversion efficiency (Φ_{UC}) is the product of efficiencies in each step involved, namely the ISC of the

sensitizer (Φ_{ISC}), the triplet energy transfer (TET)²³ from the sensitizer to transmitter to emitter (Φ_{TET}), TTA of the emitter (Φ_{TTA}), and the emitter's fluorescence (Φ_{FL}). For a specific sensitizer and emitter, Φ_{ISC} , Φ_{TTA} and Φ_{FL} are determined by the material's properties. A promising area for performance improvement is the design of QD/transmitter complexes to enable efficient TET.²⁴

In most reported QD sensitized upconversion systems, the transmitter/emitter molecules are solely limited to acenes and their derivatives, which have limited structural and energetic tunability and poor stability.^{11, 13, 18-22, 25} Oligothiophenes have wide range of tunability in energetics and molecular structure,²⁶ making them a very desirable class of triplet acceptors/transmitters in QD-organic hybrid TTA upconversion systems. Although an example of oligothiophene phosphonic acid grafted to cadmium selenide (CdSe) QDs was previously reported,²⁷ the photoluminescence quenching was attributed to charge transfer and TET was not observed. We hypothesize that with appropriate design of the energetics of QDs and oligothiophene acceptors, efficient TET transfer can be facilitated while competing single energy and charge transfer pathways may be suppressed (Fig. 1).

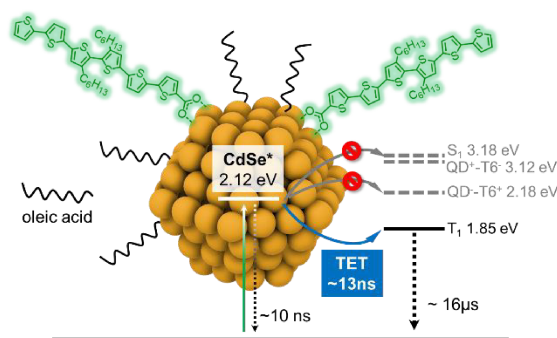


Figure 1. Cartoon of QD-T6 complexes showing the energetics of the relevant states and associated excited state decay processes.

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Results and discussion

In this work, we demonstrate for the first time a successful Dexter-type TET from CdSe QDs to carboxylic acid functionalized oligothiophene (3''',4''-dihexyl-[2,2':5',2''':5'',2''':5''',2''''':5''''',2''''':6''-sexithiophene]-5-carboxylic acid or **T6**, Fig 1). In this system, only the triplet energy transfer from the CdSe QD to T6 is energetically favored while the charge and singlet transfer are energetically uphill. We used transient absorption (TA) spectroscopy to first identify the lowest energy singlet (S_1) and triplet (T_1) state spectral features of free T6 molecules in solution. Then TA study of QD-T6 complexes provide direct evidence for triplet sensitization of T6 by CdSe QD. We also employed Density Function Theory (DFT) computations to characterize the low-lying excited states involved in the triplet-triplet transfer mechanism.

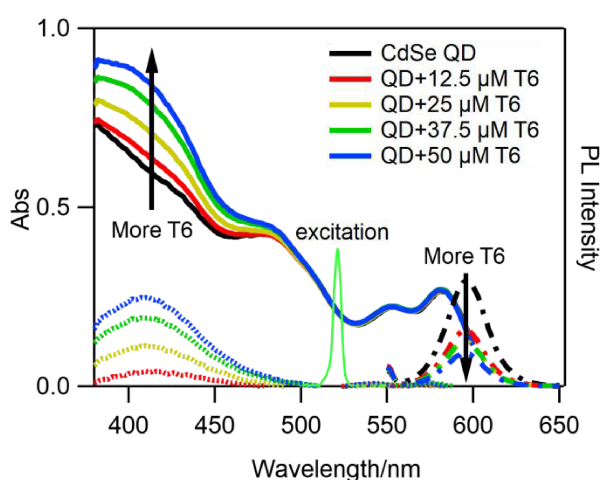


Figure 2. UV-Vis absorption (solid lines) and emission (dash-dot lines) spectra of QD-T6 samples with increasing T6 concentrations (0 to 50 μM) in toluene. The emission spectra were measured with 520 nm excitation of the QDs. Black lines are QD samples without T6. The adsorbed T6 UV-Vis absorption spectra are shown in dashed lines. The excitation pulse used in the transient spectroscopy study is also shown here. The 520 nm pulse has no overlap with T6 absorbance.

T6 (shown in Figure 1) was synthesized following a literature procedure^{28, 29} and characterized by ¹H-NMR, ¹³C-NMR and high resolution mass spectrometry (see SI and Fig. S1-3 for details). The absorption spectra of four QD-T6 complex samples with increasing T6 concentrations ranging from 12.5 to 50 μM are shown in Figure 2. These plots show a T6 band at 415 nm and the first excitonic peak of CdSe QDs at 584 nm. In contrast to the absorption spectrum of T6 in toluene (Fig 2 and Fig S4), the QD-bound T6 shows a red-shift of ~20 nm, which is attributed to a change in the dielectric environment and has been observed in a similar system.³⁰ The total amounts of T6 in the samples were obtained by subtracting QD only spectra from QD-T6 complex spectra. However, the total T6 amount contains both QD-adsorbed and free T6 molecules, which cannot be easily differentiated by their absorption spectra alone. The amount of QD-bound T6 is determined by fitting the transient kinetics of triplet formation in QD-T6 complexes (see below).

QD photoluminescence (PL) intensity decreased with increasing T6 concentration (Fig 2), suggesting possible energy or charge transfer from the QD to T6.

To understand the fluorescence quenching mechanism in QD-T6 complexes, we studied pristine T6 first to characterize its excited states spectral signature via TA spectroscopy. The TA spectra of T6 measured with 400 nm excitation (Figure 3A & B) show clear evolution from a singlet to triplet excited state on the sub-nanosecond time scale. The TA spectra at < 0.5 ns (Figure 3A) show : (i) a ground (S_0) state bleach (GSB) centered at 400 nm caused by the decrease of ground state T6 molecules; (ii) a stimulated emission from the T6 singlet excited state appearing as a negative signal at 500-600 nm; and (iii) a broad singlet (S_1) excited state absorption (ESA) from 600 nm to 900 nm. The assignment of GSB and stimulated emission of T6 is based on comparison with the steady state absorption and emission spectra shown in Figure S4. The (S_1) ESA peak is formed at an early delay time (2-5 ps) and its amplitude grows from 2-5 ps (red) to 10-50 ps (yellow) as the stimulated emission band shifts to longer wavelength (by 20 nm) at the same time, which is attributed to a fast relaxation from the initial excited state to the $v=0$ vibrational level of the S_1 excited state. Similar fast relaxation processes have also been observed in other oligothiophene molecules.^{26, 31-34} Both the S_1 ESA and stimulated emission decay on the sub-nanosecond time scale to form the first triplet excited state (T_1) with an absorption peak (T_1 to T_n transition) centered at 715 nm (Figure 3B). The lifetime of this species is longer than 10 μ s and can be shortened by more than 1000 times in the presence of oxygen (Figure S7), consistent with the assignment to a T_1 state.³⁵

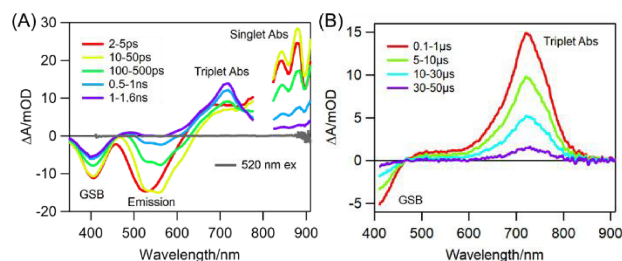


Figure 3. TA spectra and kinetics of T6 in toluene. TA spectra of T6 at indicated delay times after 500 $\mu\text{J}/\text{cm}^2$ 400 nm excitation: (A) 0–1.6 ns and (B) 0.1–50 μs . Data between 780–820 nm in (A) are cut out due to saturated probe light. 520 nm at 1 mJ/cm^2 could not directly excite free T6 as shown in panel A (grey lines, averaged from 2 ps to 1 ns).

We conducted DFT computations to determine the triplet excited state energetics on a model of the T6 molecule by replacing the C₆H₁₃ alkyl side chains with methyl groups. All DFT computations were performed with the ORCA³⁶ and Psi4³⁷ programs using the B3LYP³⁸ functional and the def2-SVP basis set.^{39, 40} Equilibrium geometries for the S₀ and T₁ states are shown in Fig S10 and S11. From these geometries, we computed vertical excitation energies to higher states by time-dependent density functional theory (TD-DFT) to identify low-lying electronic excited states involved in the triplet-triplet transfer mechanism. At the S₀ geometry, T₁ lies 1.85 eV above the

ground state and is characterized by a spin-flip HOMO $\rightarrow$ LUMO excitation. Geometric relaxation of the T_1 state stabilizes it by 0.47 eV, giving an adiabatic $S_0 \rightarrow T_1$ transition energy of 1.38 eV. The second triplet state corresponds to a HOMO $\rightarrow$ LUMO + 1 excitation, with vertical transition energy of 2.10 eV. The third and fourth triplet states are predicted to lie at higher energies (2.64 and 2.82 eV) and therefore should not be operative in the observed triplet-triplet energy transfer.

To compute oscillator strengths for the $T_1 \rightarrow T_N$ transitions, we performed a TD-DFT calculation at the optimized geometry of the T_1 state using the triplet state as a reference. The vertical excitation spectrum of T_1 is dominated by a transition at 1.88 eV (659 nm, $f_{osc} = 3.70$) with a less intense peak at 2.18 eV (568 nm, $f_{osc} = 0.24$). These results are in good agreement with the experimental observation of a transient peak at 715 nm in the transient absorption spectra for $t > 100$ ns.

The mechanism of exciton quenching in CdSe QD/T6 complexes were then studied by transient absorption spectroscopy, which probes directly not only the quenching kinetics, but also the identity of intermediates and products. The TA spectra of the QD-T6 complexes at early delay times (1–10 ps) show that the selective excitation of CdSe QDs at 520 nm (where T6 does not absorb) leads to a QD ground state bleach and excited state absorption^{41–44} at 420–620 nm (Figure 4A), similar to those observed in QD only samples (Figure S5) and indicative of excitons in the QD. The TA spectra at later delay times (Figure 4A) show the decay of the QD exciton and the formation of a new absorption peak at 640–800 nm (Figure S8d for all T6 concentration) and a S_0 ground state bleach (Figure 4B). These features are absent in QD only samples and can be assigned to T6 triplet formation by comparing to the T6 triplet TA spectrum (Figure 4B). Control experiment of T6 only samples at 520 nm excitation (the grey circle in Fig 4D, grey lines in Fig 3A) shows negligible T6 triplet signal in the absence of QDs. Here the TA spectra data directly proved a successful triplet sensitization of T6 by QD.

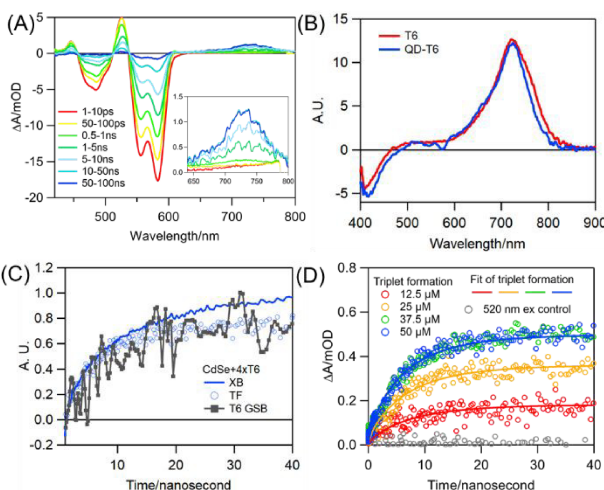
Figure 4. Transient spectra and kinetics of QD/T6 complexes measured at 520 nm excitation. (A) TA spectra of QD-T6 excited at indicated delay times. Inset: expanded view of the triplet spectra at 540–800 nm. (B) Comparison of the averaged TA spectra (from 100 ns to 5 μ s) of QD-T6 complexes (excited at 520 nm) and pristine T6 (excited at 400 nm). (C) Comparison of the kinetics of triplet formation of T6 (715 nm, blue circle), T6 GSB (415 nm, grey square, inverted and scaled) and QD exciton bleach (584 nm, blue line, inverted and scaled) in QD-T6 complexes. (D) Kinetics of T6 triplet formation (TF) with increasing T6 concentration (red, yellow, green, blue) in QD-T6 complexes. Solid lines: fits according to a model described in the main text. Grey dots: T6 signal of a control sample with 50 μ M T6 and no QD measured under same condition at 520 nm excitation, showing no triplet signal. Excitation pulse energy density is 35 μ J/cm² for panel A, C and D, 500 μ J/cm² for panel B.

To determine the mechanism of triplet formation, we compare the kinetics of T6 triplet state growth and exciton decay in QDs. The kinetics of triplet formation (Figure 4C) can be monitored by both the triplet ESA (TF, blue circle) and S_0 ground state bleach (GSB, black square). Here, triplet ESA kinetics was obtained by averaging TA signals from 650 nm to 780 nm to increase the signal/noise ratio. The T6 GSB is obtained by subtracting the overlapping QD contribution at 415 nm from the total signal. As reported previously, the CdSe XB signal is dominated by state filling of electrons in the conduction band edge.^{41–45} As shown in Figure 4C, the triplet formation kinetics agree well with the exciton bleach (XB, blue line) up to ~ 20 ns, consistent with a direct Dexter type TET from the QD to T6. These kinetics deviate after ~ 20 ns, which suggests the presence of other exciton decay pathways. As discussed below, we attribute this to the presence of QDs without adsorbed T6, in which excitons decay by radiative and nonradiative pathways within the QD. It has been previously reported that QD sensitized triplet formation can also occur via sequential charge transfer pathways: electron (hole) transfer followed by hole (electron) transfer.^{46–48} These pathways would produce $T6^+$ or $T6^-$ intermediates, which are not observed in our study. This observation is also supported by the energetics alignment that prohibits charge transfer or singlet energy transfer (Figure S12). Therefore, we attribute this sensitization process to be a direct Dexter-type triplet energy transfer from a triplet exciton state in the QD.

To examine whether TET occurs via static or dynamic quenching mechanisms, we measure TET kinetics as a function of total T6 concentration in the sample. As shown in Figure 4D, the final amplitude and apparent formation rate of the T6 triplet state increase with the total T6 concentration, until it reaches saturation. Similar acceptor concentration dependent kinetics has been observed in previous studies of QD/electron acceptor complexes.^{44, 49–52} This kinetics can be well described by a model that assumes that the number of adsorbed acceptor molecules in QD/acceptor complexes follows a Poisson distribution and the rate of transfer is proportional to the number of adsorbed acceptors. Assuming the same model for TET in QD/acceptor complexes, we can derive an expression of the kinetics of triplet formation (Eq. 2) and QD exciton bleach recovery (see SI).

$$[T6]_T^*(t) = [QD(0)] * \sum_i P(m,i) \sum_j a_{jikTET} \left(1 - \frac{e^{-(k_{1j} + ik_{TET})t}}{k_{1j} + ik_{TET}} \right)$$

(Eq. 2)



In Eq. 2, $P(m, i)$ is the Poisson distribution probability of finding QD/T6 complexes with i adsorbed T6 molecules in a sample with m average number of acceptors. k_{TET} is the intrinsic energy transfer rate in a 1:1 QD/T6 complex. The decay of QD exciton bleach through radiative and nonradiative pathways within the QD is described by multiple exponential decay with amplitude a_j and rate constant k_{1j} . $[QD(0)]^*$ is the initial concentration of excited QDs. Further details of the model and derivation can be found in the SI. Because the decay rate of T6 triplet excited state (k_3) is three orders of magnitude slower than triplet formation (Figure S7), it is not considered in the fitting of the triplet formation kinetics. The T6 triplet formation (Figure 4D) and QD exciton bleach recovery (Figure S8a) can be well fit by this model with an intrinsic TET of $0.077 \pm 0.002 \text{ ns}^{-1}$. The fit also yields the average number of adsorbed T6 molecules as a function of total T6 concentration, ranging from 0.16 to 0.54. This dependence can be fit to a Langmuir isotherm, as shown in Figure S8b, to obtain a maximum number of adsorption sites and binding constant of T6 on CdSe QDs of $\theta = 1.4 \pm 0.8$, and $K = 0.014 \pm 0.012 \mu\text{M}^{-1}$, respectively. It is important to note that in the range of m (from 0.16 to 0.54) studied in this work, over 90% of the QD population in the samples has zero or one adsorbed T6 (Table S2) and the triplet formation kinetics is mostly contributed by QD with one bounded T6. This is confirmed by the comparison of normalized triplet formation kinetics for different samples (Fig S8c), which are identical within the S/N ratio of this measurement. The increased T6 signal amplitude in Fig 4D reflects a larger percentage of T6 bounded QDs in these samples.

From the relative amplitude of the triplet state absorption and QD exciton bleach and their extinction coefficients, we can estimate a TET efficiency of $15.4 \pm 0.6\%$ in the sample with $50 \mu\text{M}$ T6 concentration (see SI for details). The low efficiency can be attributed to both the lower number of adsorbed T6 molecules and slow intrinsic TET rate. The average number of adsorbed T6 per QD is only around 0.5 at the highest coverage level, presumably limited by the competitive adsorption between T6 and the stabilization oleic acid ligand on the QD surface. In comparison, the Castellano group reported CdSe QD to 9-anthracene carboxylic acid (ACA) Dexter TET with an overall rate of $2 \times 10^9 \text{ s}^{-1}$ with an average number of ACA per QD of 12.¹² This yields a similar intrinsic TET rate constant of 0.17 ns^{-1} . To verify this point, we vigorously sonicated the sample with excess amount of T6 ($375 \mu\text{M}$) and measured the TA spectrum within an hour. This temporarily created a high loading which leads to more efficient TET at $31.8 \pm 1.2\%$ as shown in Figure S9. Thus, a promising approach to increase the overall TET rate and efficiency in the T6/QD complex is to improve the binding strength of T6, through perhaps a different anchoring group.

Conclusions

We have demonstrated the first example of triplet energy sensitization of an oligothiophene (T6) by CdSe QDs. By TA spectroscopy we confirmed this process occur via direct Dexter-

type triplet energy transfer. The efficiency of TET increases with the concentration of T6 in solution, reaching a value of $15.4 \pm 0.6\%$ in the sample with $50 \mu\text{M}$ T6 and $31.8 \pm 1.2\%$ in the sample with $375 \mu\text{M}$ of T6, although the latter exhibits poor long-term stability. The T6 concentration-dependent TET kinetics can be understood by a model that assumes a Poisson distribution of adsorbed T6 molecule on the QD. This analysis suggests that TET rates are similar in the 1:1 QD-T6 and previously reported QD-anthracene complexes, but T6 acceptors have relatively weaker binding constant on the surface, which leads to a smaller average number of acceptors and slower apparent TET rate. Our finding suggests that TET efficiency QD-T6 complexes can likely be further improved by increasing the acceptor binding strength (through modification of the acceptor and QD surface ligands) and fine-tuning energetic levels. Finally, this work identifies a new class of stable and highly tunable/functionalizable organic TET acceptors.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

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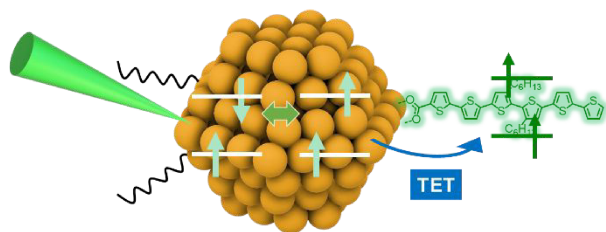
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