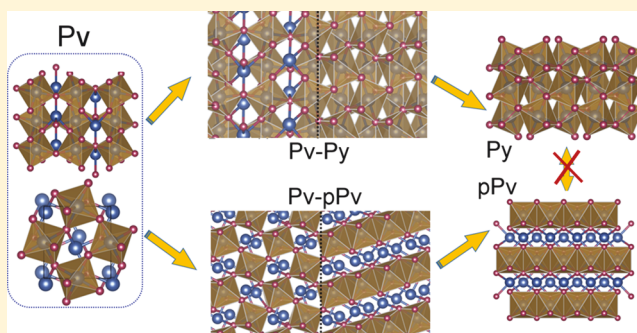


Structure-Controlled Oxygen Concentration in Fe_2O_3 and FeO_2 Sheng-cai Zhu,^{†,‡,§,||} Jin Liu,^{§,||} Qingyang Hu,^{*,‡,||} Wendy L. Mao,[§] Yue Meng,[⊥] Dongzhou Zhang,^{||} Ho-kwang Mao,^{‡,#} and Qiang Zhu^{*,†}[†]Department of Physics and Astronomy, High Pressure Science and Engineering Center, University of Nevada, Las Vegas, Nevada 89154, United States[‡]Center for High Pressure Science and Technology Advanced Research (HPSTAR), Shanghai 201203, P. R. China[§]Department of Geological Sciences, Stanford University, Stanford, California 94305, United States^{||}Hawai'i Institute of Geophysics and Planetology, School of Ocean and Earth Science and Technology, University of Hawai'i at Manoa, Honolulu, Hawaii 96822, United States[⊥]High Pressure Collaborative Access Team, X-ray Science Division, Argonne National Laboratory, Argonne, Illinois 60439, United States[#]Geophysical Laboratory, Carnegie Institution of Washington, Washington, D.C. 20015, United States

Supporting Information

ABSTRACT: Solid–solid reaction, particularly in the Fe–O binary system, has been extensively studied in the past decades because of its various applications in chemistry and materials and earth sciences. The recently synthesized pyrite- FeO_2 at high pressure suggested a novel oxygen-rich stoichiometry that extends the achievable O–Fe ratio in iron oxides by 33%. Although FeO_2 was synthesized from Fe_2O_3 and O_2 , the underlying solid reaction mechanism remains unclear. Herein, combining in situ X-ray diffraction experiments and first-principles calculations, we identified that two competing phase transitions starting from Fe_2O_3 : (1) without O_2 , perovskite- Fe_2O_3 transits to the post-perovskite structure above 50 GPa; (2) if free oxygen is present, O diffuses into the perovskite-type lattice of Fe_2O_3 leading to the pyrite-type FeO_2 phase. We found the O–O bonds in FeO_2 are formed by the insertion of oxygen into the Pv lattice via the external stress and such O–O bonding is only kinetically stable under high pressure. This may provide a general mechanism of adding extra oxygen to previous known O saturated oxides to produce unconventional stoichiometries. Our results also shed light on how O is enriched in mantle minerals under pressure.



1. INTRODUCTION

Oxygen-rich (peroxides and superoxides) materials are often found in alkali metal oxides or alkaline-earth metal oxides because of the small cation to anion radius ratio at ambient pressure.¹ Because of the fragile O–O bonds, the peroxides $[\text{O}_2]^{2-}$ and superoxides $[\text{O}_2]^-$ species are often metastable and prone to release O and form stable oxides at ambient conditions, making them widely used as decolorizers and aerophores in our daily life. External pressure can alter the ionic radius,² stiffen bonds, and thus give rise to thermal stability of oxygen-rich materials, such as LiO_2 ,³ MgO_2 ,^{4,5} and FeO_2 .⁶ These phenomena may have great implications in materials sciences and planetary sciences. For example, MgO_2 peroxides are predicted to exist as planet forming materials under strong oxidizing conditions.⁵

Iron oxides are major naturally occurring compounds in the earth and have a variety of industrial applications.^{7,8} Seeking new iron oxides with different valence and spin states of iron has been a long-term interest for both materials scientists and

geoscientists. To date, besides the common stoichiometry of FeO , Fe_3O_4 , and Fe_2O_3 , several novel stable iron oxides such as Fe_4O_5 ,⁹ Fe_5O_6 ,¹⁰ and Fe_5O_7 ¹¹ have been reported under various pressure–temperature conditions. All these newly discovered iron oxides are regarded as mixed-valence oxides, falling into the stoichiometry range between FeO and Fe_2O_3 . Most of them can be actually considered as the homologous series of $n\text{FeO} \cdot m\text{Fe}_2\text{O}_3$.¹¹ Given the abundance of both Fe and O, their chemistries play a critical role in redox equilibria of the solid earth, which thus in turn affect the oxygen fugacity, chemical differentiation, and physical properties.

Along this track, it was reported that iron could be oxidized further beyond Fe_2O_3 , similar to what happened on alkali and alkali-earth metals. Indeed, the pyrite-type FeO_2 (denoted as Py-FeO_2) was synthesized from the mixture of Fe_2O_3 and O_2 at 74 GPa and 1700 K, which has the highest O concentration in

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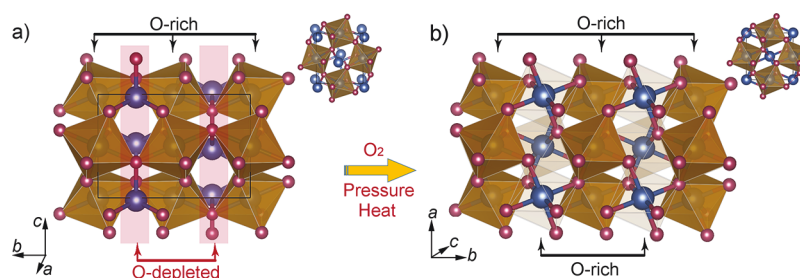


Figure 1. Atomic structure of (a) Pv-Fe₂O₃ and (b) Py-FeO₂, the inset is top view from Pv(010) and Py(010). The similar arrangement of Fe and O in Py-FeO₂ and Pv-Fe₂O₃ indicates that they may form a low-energy interface. Red: O, blue: Fe.

all known iron oxides.⁶ The Py-FeO₂ bears O–O bonding^{6,12,13} and is only stable at high pressure. Similar to many peroxides, it becomes unstable upon decompression and releases O₂.⁶ For example, below 40 GPa, Py-FeO₂ starts to degas oxygen and reverts to Fe₂O₃. While the thermodynamic phase stability fields for Fe₂O₃^{14–17} and FeO₂^{6,18,19} have been extensively investigated, the general structure modification mechanism involving stoichiometric variation remains elusive. The kinetics in the oxygenation reaction is fundamentally important to understand how they are stabilized at high pressure. This motivated us to perform a detailed investigation on the atomistic mechanism in the oxygenation reaction from Fe₂O₃ to Py-FeO₂ and study how this reaction competes against polymorphic phase transitions to other high-pressure forms of Fe₂O₃.

2. RESULTS

2.1. Density Functional Theory Calculation Results.

The key to understand the solid–solid reaction is to establish the structural relation between Fe₂O₃ and FeO₂ and construct the reaction interface for the O diffusion. Among the many high-pressure polymorphs of Fe₂O₃, including hematite phase (α -Fe₂O₃), Rh₂O₃-II-type, perovskite-type (Pv-Fe₂O₃), post-perovskite-type (pPv-Fe₂O₃) and Aba2-type, we found that the Pv-Fe₂O₃ shares fairly strong structural similarity with Py-FeO₂ (see Supporting Information Figure S1 for details). On the basis of the available data, the Pv-type phase is only thermodynamically stable within a narrow pressure window of 45–50 GPa and transforms to the pPv phase above 50 GPa upon heating.^{15,16} It was also reported that Pv-Fe₂O₃ is slightly triclinic distorted²⁰ but this would not be considered in this study.

Pv-Fe₂O₃ has an orthorhombic unit cell (*Pnma*, no. 62) which contains two different Fe cation sites. Using the prototypical ABO₃ structure for the Pv-structure, Fe cations at A sites fill the interstices of 8 Fe–O-octahedral and hold 4 short Fe–O bonds of ~ 1.87 Å, 2 long Fe–O bonds of ~ 2.22 Å, and 2 even longer bonds of ~ 2.45 Å at 75 GPa. In light of the first shell of the Fe–O, the A site layer corresponds to a FeO formula,²¹ thus labeled as the O-depleted layer. Meanwhile, Fe cations in the B site fill the O octahedral interstices (all the Fe–O bonds are ~ 1.86 Å) that correspond to a FeO₂ formula, denoted as the O-rich layer. In the Pv-Fe₂O₃ structure, the shortest O...O distance is 2.52 Å. The O-rich and the O-depleted layers are alternating along the [010] direction, leading to the stoichiometry of Fe₂O₃.

In FeO₂ (*Pa* $\bar{3}$, no. 205), all Fe atoms fill the octahedral sites, thus making them fully O-rich layers. Each O is shared by 3 [FeO₆] octahedral and connected with a neighboring O in the O–O bond. The O–O bond length is 1.9 Å from the

experiment at 76 GPa and is estimated as ~ 1.8 Å according to theoretical calculations with Hubbard $U = 5$ eV (~ 1.9 Å when $U = 4$ eV) at 75 GPa.^{12,13} The structure of FeO₂ is analogous to a certain group of peroxides (e.g., pyrite-MgO₂) and dioxides (e.g. PdF₂-type SiO₂). Although the O–O bond length in FeO₂ is larger than values of 1.29–1.53 Å in typical peroxides,¹ both experiment and theory suggest an unusual O–O bonding in FeO₂ (see the electron localization function map in Figure S2 in Supporting Information). Therefore, we regard FeO₂ as the pyrite-type phase.²²

From the top view, Py-FeO₂ and the O-rich layers of Pv-Fe₂O₃ have the same atomic arrangement (insets in Figure 1). However, Fe atoms of Fe₂O₃ in the O-depleted layer are positioned at different sites, thus distinguishing two structures. Once all the O-depleted layers in Pv-Fe₂O₃ are filled by O, they will turn to O-rich layers and transform to Py-FeO₂.

On the basis of this structural relation, we constructed an interface model between the (010) plane of Py-FeO₂ and the (010) plane of Pv-Fe₂O₃, namely, (010)_{Py}//(010)_{Pv} + [100]_{Py}//[100]_{Pv}. Its stability can be assessed by the interfacial energy γ via first-principles calculation. We defined interface energy γ as $\gamma = (E_{a/b} - E_a - E_b)/2S$, where S is the interfacial area, E_a and E_b are the energies of the pure phases, and $E_{a/b}$ is the energy of the mixed phase in a supercell containing two equivalent interfaces under periodic boundary conditions.²³ To describe localized 3d electrons of Fe, we used the DFT + U scheme²⁴ with a U value of 5.0 eV following a previous work.²⁵ The on-site U parameter may change the total energy for each individual structure but will have limited effects on the energy barrier and interface energy. Our calculation enabled spin polarization, and the structures were relaxed for atomic forces. In our calculation, both Pv-Fe₂O₃ and pPv-Fe₂O₃ have mixed high-spin and low-spin states. Different spin states are alternately aligned along the [010] direction for Pv and the [001] direction for pPv as were described in the ref 26. The Fe atoms in Py-FeO₂ are in the low spin-state. For the magnetic structure, both Pv-Fe₂O₃ and pPv-Fe₂O₃ structures are treated as antiferromagnetic configuration, while FeO₂ is nonmagnetic. The magnetization at the interface is extremely complex. After geometric relaxation, we managed to optimize the interface supercell to a local minimum. For this relaxed structure, Fe atoms at the reaction interface form a pseudo antiferromagnetic configuration (see Figure S3 in Supporting Information), with Fe atoms of the Py-phase further away from the interface staying nonmagnetic. After optimization, the interface energy value is calculated as 0.53 J/m² at 75 GPa. This is considered as a low-energy interface compared with many other well-defined systems. For example, the rutile (101)/anatase (112) has an interface energy of 1.49 J/m² because of the pronounced lattice mismatch,²⁷ and the interface energy of

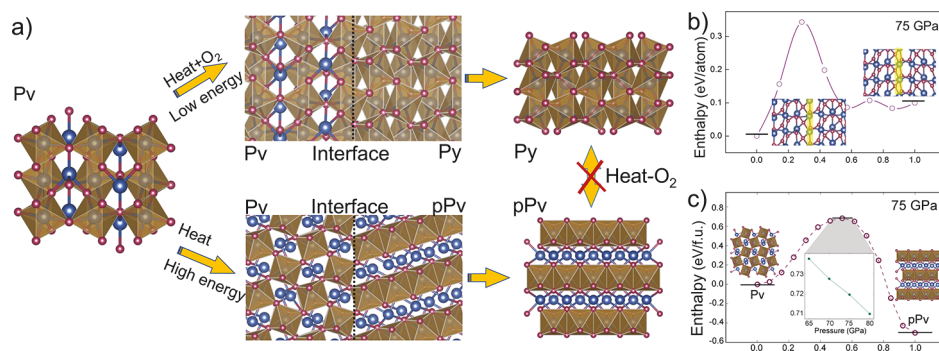


Figure 2. Kinetics of O diffusion and Pv-to-pPv phase transition in Fe_2O_3 . (a) Py- Fe_2O_3 and Pv- Fe_2O_3 interface with oxygen diffusion (upper panel) and the interface for the phase transition from Pv- Fe_2O_3 to pPv- Fe_2O_3 (bottom panel). (b) Enthalpy barrier for O diffusion at 75 GPa; (c) transition path from Pv- Fe_2O_3 to pPv- Fe_2O_3 from the ssNEB method; no direct path between pPv- Fe_2O_3 and Py- Fe_2O_3 can be found.

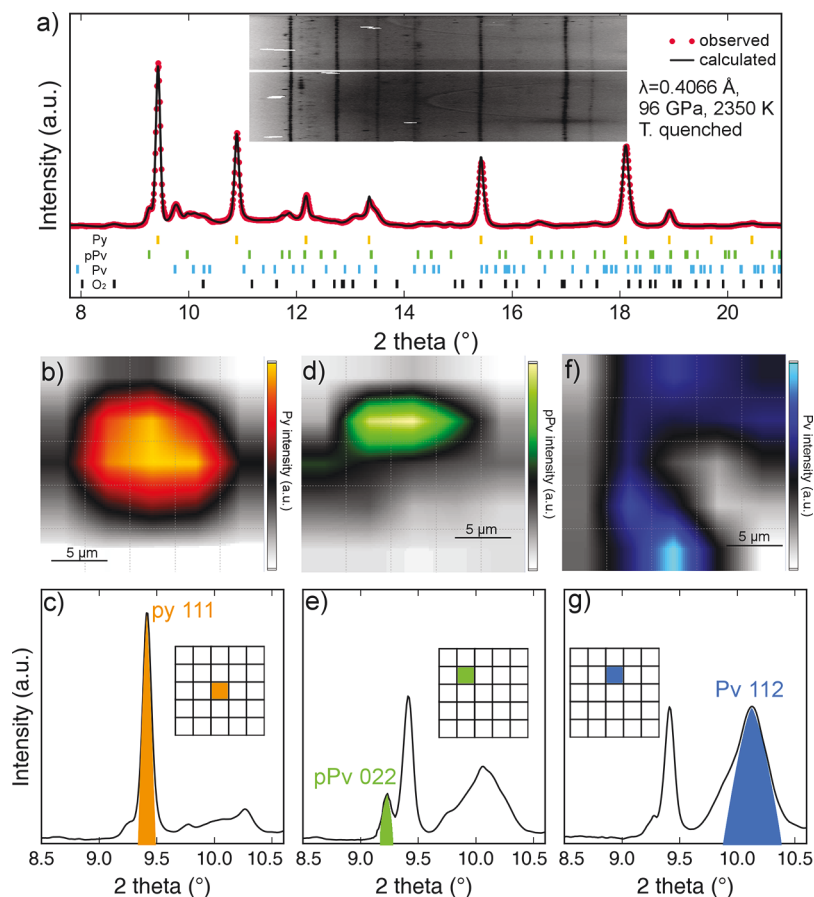


Figure 3. XRD mapping of synthesized FeO_2 at 96 GPa. (a) Fully indexed Py- FeO_2 at the heating center. The O–O distance is 1.87 Å. The inset is the caked diffraction pattern. (b–g). Distributions of Py- FeO_2 , pPv- Fe_2O_3 , and Pv- Fe_2O_3 are characterized by the intensities of their signature 111, 022, and 112 peaks, respectively. Double-sided heating laser spots generally cover all the sample region and thermal gradient is estimated up to 200 K. The dominant Py- FeO_2 at the heating center is fitted by the Rietveld method, while the rest spotty peaks from Fe_2O_3 and O_2 are analyzed by *LeBail* fitting. The final refinement factors are $R_1 = 0.069$, $wR_2 = 0.14$. The O–O distance is 1.87 Å for reference only. Lattice parameters are reported in Supporting Information, Table S1.

the symmetric $\Sigma 3(110)$ grain boundary in silicon is 0.76 J/m^2 .²⁸ Therefore, we conclude that such an interface is energetically favorable under high pressure conditions. It is worth mentioning that Rh_2O_3 -II-type phase is close to Pv- Fe_2O_3 in the atomic structure and also form reaction interface with the pyrite phase (see Figure S5). However, the Rh_2O_3 -II type phase transforms to Pv phase with a shallow energy barrier (e.g., 0.019 eV/f.u. at 65 GPa) when pressure is greater than 45 GPa.²⁹ Therefore, Rh_2O_3 -II-type Fe_2O_3 is kinetically

unstable even at 300 K and 75 GPa. Moreover, a direct interface of Rh_2O_3 -II/Py leads to a much higher interface energy (0.922 J/m^2). This is consistent with our high-pressure experiment ($>75 \text{ GPa}$) that Pv- Fe_2O_3 is the preferred phase to react with O_2 .

The satisfactory atomic match between $(010)_{\text{Py}}$ and $(010)_{\text{Pv}}$ provides the O diffusion channel for the oxygenation process. Mimicking O diffusion, moving the O atom from the O-rich to O-depleted layer only overcomes a tiny energy barrier of 0.34

eV per atom (equivalent to $0.032 \text{ eV}/\text{\AA}^2$) at 75 GPa by the nudged elastic band (NEB) method.³⁰ In this process, the O-rich layer transits to the O-depleted layer, while the rest of the structure remains unchanged. This ensures that the frame structure can be retained during O diffusion. Owing to the relatively small kinetic barrier, it is safe to say that Pv-Fe₂O₃ readily transforms to the Py-FeO₂ phase when free O atoms are present.

At the same pressure range (>50 GPa), Pv-Fe₂O₃ transforms to the polymorphic pPv-Fe₂O₃ when O₂ is depleted.^{11,14} It is generally believed that the archetypal Pv to pPv transition requires a rather high activation energy barrier.³¹ Here, we explore the Pv-to-pPv Fe₂O₃ phase transition pathways by an alternative sampling scheme, namely, stochastic surface walking (SSW) pathway method^{32,33} (details in the Method section). SSW produced the most kinetically favored phase transition pathway which was further optimized by the solid-state NEB method (ssNEB),³⁰ variable-cell NEB (VC-NEB),³⁴ and variable-cell double-ended surface walking (VC-DESW)³⁵ method. As shown in Figure 2a, by shearing along the [100] direction, Pv-Fe₂O₃ transforms to pPv-Fe₂O₃ with an enthalpy barrier 0.69 eV/f.u. from ssNEB (similarly 0.68 eV/f.u. for VC-NEB and 0.72 eV/f.u. for VC-DESW), equivalent to $0.046 \text{ eV}/\text{\AA}^2$, at 75 GPa. The high energy barrier is consistent with the experimental observation that Pv-to-pPv phase transition only takes place as high as 2000–2320 K as Ono and Ohishi reported.¹⁶ Pv–pPv transitions in bridgmanite MgSiO₃ was proposed to follow a stacking fault mechanism by Oganov et al.³¹ and it was confirmed by Tschauner through an experiment.³⁶ Here, we found that the phase transition between Pv- and pPv-Fe₂O₃ followed the same mechanism. Using the finite strain theory,³⁷ we derived that the phase transition orientation relation is $(001)_{\text{Pv}} // (110)_{\text{pPv}} + [010]_{\text{Pv}} // [001]_{\text{pPv}}$. Figure 2a shows a supercell interface model following this relation. Its interface energy is $0.38 \text{ J}/\text{m}^2$ after full optimization, which is smaller than the interface of $(010)_{\text{Py}} // (010)_{\text{Pv}} + [100]_{\text{Py}} // [100]_{\text{Pv}}$. Comparing with the O diffusion in Py-FeO₂, the structure phase transition to pPv-Fe₂O₃ forms a same stable reaction interface as Py but overcomes a much higher energy barrier than Pv to Py.

2.2. In Situ High-Pressure Experiment. To confirm our hypothesis, we investigated the solid reaction pathways by in situ synchrotron X-ray diffraction (XRD) experiments. We compressed a piece of Fe₂O₃ in O₂ pressure medium to 90 GPa and heated it by laser (Supporting Information Figure S4). In Figure 3a, intensities of Py-FeO₂ diffraction gradually increased under around 2350 K laser heating and FeO₂ pattern eventually overwhelms all other Fe₂O₃ phase's after 30 min heating. Pressure equilibrated at 96 GPa after quenching to room temperature. At the center of heating area (Figure 3a), strong and smooth powder diffraction peaks are readily indexed for Py-FeO₂ using the Rietveld method while minor residual spotty peaks are associated with Fe₂O₃ and ϵ -O₂. The final refinement factors are $R_1 = 0.069$ and $wR_2 = 0.14$. The O–O distance is $1.87 \text{ \AA}$ that falls between traditional peroxide and PdF₂-dioxide. Figure 3b–g shows the 2 dimensional XRD mapping over the heated sample. In the spatial view, we can find that both Py-FeO₂ and pPv-Fe₂O₃ are contingent with the Pv-Fe₂O₃. Py-FeO₂ is the dominant phase among the coexistence of py-FeO₂ and Pv- and pPv-Fe₂O₃, indicating that the phase transition to the Py-FeO₂ is kinetically favored. In our previous experiment, we did not attempt to associate any diffraction peaks with Pv-Fe₂O₃ because it is considered as

a low-pressure phase.⁶ However, because of the kinetics of the two competing phase transitions, the polymorphic transition to pPv-Fe₂O₃ may be hindered by its higher energy barrier. As a result, the Pv-Fe₂O₃ appears as a metastable interfacial structure in the mixed products even if the pressure is 75 GPa.

We further studied the kinetics of Pv-Fe₂O₃ to Py-FeO₂ and polymorphic transition to pPv-Fe₂O₃ by the time-dependent XRD analysis; see Figure 4. XRD patterns were collected at 96

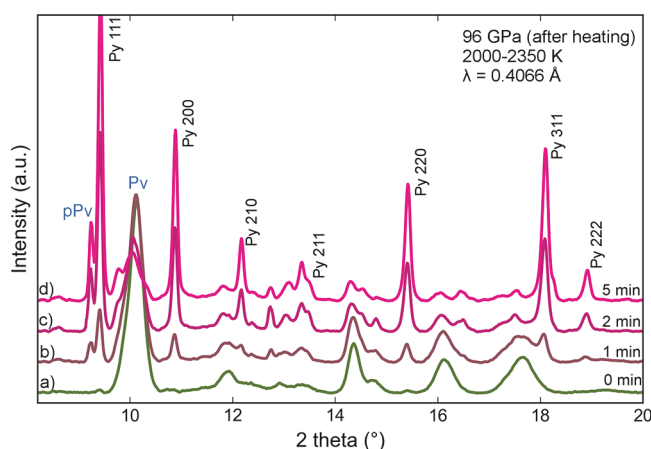


Figure 4. Time-resolved XRD patterns for a high-pressure sample during laser heating. All patterns were collected after heating for a certain amount of time and quenched to 300 K. (a) Before heating (0 min), (b) after 1 min heating, both Py-FeO₂ and pPv-Fe₂O₃ were found in the heating center; (c,d) after 2 and 5 min heating, Pv-Fe₂O₃ gradually transforms to Py-FeO₂ and pPv-Fe₂O₃. The reaction rate to Py-FeO₂ is faster than to pPv-Fe₂O₃. Each pattern was collected after 10 s X-ray exposure time. The pressure is measured after temperature quench with an uncertainty of ± 2 GPa; X-ray wavelength is $0.4066 \text{ \AA}$.

GPa and quenched room temperature after heating at 2000–2350 K for a certain amount of time. Comparing the peak intensity of Py- and pPv-phase, the transition of Pv to Py phase has clearly higher productivity. The low productivity of pPv-Fe₂O₃ is reproduced in another experimental run, where nearly pure Py-FeO₂ was found within the heating area (Supporting Information, Figure S5). According to the time-dependent XRD data, transition to Py-FeO₂ has a priority over Pv-Fe₂O₃ when O is present.

To trace the degassing process, the sample was decompressed from high-pressure at ambient temperature. We decrease the pressure slowly and trace the products by XRD. As shown in Figure 5, while the diffraction intensities from Py-FeO₂ attenuate by releasing pressure, signals from Pv-Fe₂O₃ stay strong. This is due to the reverse reaction by depleting O from Py-FeO₂ crystals. Because FeO₂ and Pv-Fe₂O₃ form a well-matched reaction interface, a small portion of FeO₂ will transform back to Pv-Fe₂O₃ without destroying the crystal framework. Below 46 GPa, the Pv-Fe₂O₃ slips out of its stability field and started to fade away. At 31 GPa, the sample completely decomposed to α -Fe₂O₃. Degassing of oxygen indicates that O–O bonds become fragile upon releasing pressure.

3. DISCUSSION

The chemical reaction of $2\text{Fe}_2\text{O}_3 + \text{O}_2 = 4\text{FeO}_2$ is only observed at high pressure. After filling the O-depleted layer with O atoms, the original Pv-Fe₂O₃ lattice suffers considerable compressive stresses in both [100] and [001] directions,

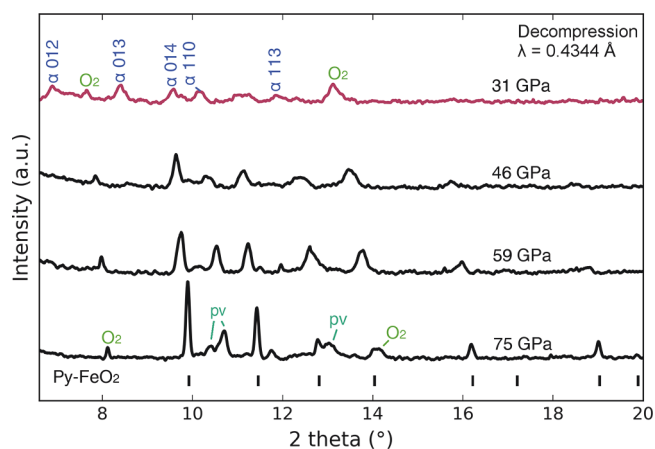


Figure 5. XRD pattern along the decompression process at 75, 59, 46, and 31 GPa. Signals from Py-FeO₂ become weak with decreasing pressure and eventually disappear at 31 GPa.

forcing the shortening of O–O length and eventually leading to the emergence of O–O pairs. As a result, O-depleted layers turn into O-rich layers and eventually form the Py-FeO₂ structure. On the other hand, the O–O bonds readily vanish by a reversible unpairing process. Decomposition of FeO₂ is ultimately triggered when the pyrite structure collapses upon releasing pressure. In a word, the O concentration changes in between Fe₂O₃ and FeO₂ are largely due to the pressure effect, controlled by the atomistic structure.

The mechanism finds its most important application in the chemistry of earth's interiors. Although Fe₂O₃ have many polymorphs, such as α -Fe₂O₃, t -Fe₂O₃, Pv-Fe₂O₃, pPv-Fe₂O₃, and θ -Fe₂O₃, our work found that only the phase transition between Pv and Py possesses such “switchable” characteristic. We did not observe similar reactions between any other Fe₂O₃ polymorphs and Py-FeO₂. This observation can be explained by the atomistic mechanism. Pv-Fe₂O₃ and Py-FeO₂ have a strong structural relation, which provided them well-matched interface and low energy barrier. As a result, at the same pressure and temperature condition, the formation of Py-FeO₂ will be more favorable than the competing phase transition to pPv-Fe₂O₃. In other words, if the kinetic of Pv to pPv is more favorable than to Py, the transition of pPv-Fe₂O₃ + O₂ to Py-FeO₂ would have to overcome a much higher energy barrier. Such transition kinetics may play a different role in the degassing or regassing of oxygen in earth's rocky materials.

It was reported that Py-FeO₂ can accumulate at earth's core–mantle boundary through iron–water reaction.^{38–40} The archetypal Pv structure in the simple form of ABO₃ allows numerous cation combinations on the A and B sites,⁴¹ creating an extraordinary Pv-type family. Replacing iron partially with other major earth's elements such as Mg or Al in Py-FeO₂ can generate solid solution-like compounds of (Mg,Fe)O₂ or (Al,Fe)O₂. Indeed, hydrogen bearing pyrite-(Al,Fe)O₂ has been observed in the experiment by Zhang et al.⁴² It provides important evidence that the pyrite phase can be a more viable material type in the earth. Therefore, the O diffusion mechanism between pyrite and perovskite structures in the current study can possibly apply to the materials in the general form of (Mg,Fe)(Al,Fe,Si)O₃ perovskite.⁴³ Although our current study only considers Fe and O, the alternating O-rich and O-depleted layers in the perovskite structure have already provided the basis to apply this mechanism. Finding

the appropriate pressure, temperature, and stoichiometry conditions to realize the Pv-to-Py phase transition in other minerals containing earth-abundant elements is the subject of future work. In addition to its geoscience implication, the Pv-to-Py reaction by adding extra oxygen to previous known O saturated oxides to produce unconventional stoichiometries may also apply to other materials, such as AlO₂ in the Al–O⁴⁴ system and GaO₂ in the Ga–O system. It would be a novel method to create a broader spectrum of O-rich materials that will have potential industrial applications.

4. CONCLUSIONS

In summary, we establish an oxygenation mechanism between perovskite Fe₂O₃ and pyrite FeO₂ in this work. Combining the first-principles calculation and in situ experiment, we identified that the most likely transition pathway is via O diffusion on the Pv-Fe₂O₃(010) and Py-FeO₂(010) interface under high-pressure and oxygen-rich conditions. Thereby, we propose two competing phase transition pathways on Fe₂O₃ controlled by its transforming structure: (1) polymorphic transition from Pv- to pPv-Fe₂O₃ when O is depleting and (2) oxygenation to Py-FeO₂ via O diffusion when excess O₂ is present. The mechanisms may have great impact in understanding the oxygen enrichment in the interior of the earth.

■ ASSOCIATED CONTENT

Supporting Information

The Supporting Information is available free of charge on the ACS Publications website at DOI: 10.1021/acs.inorgchem.8b02764.

High-pressure in situ XRD experiment method; the calculation detail of SSW sampling method, the interface energy and O diffusion; the atomic structure of Py-FeO₂ and Fe₂O₃ polymorphs; the electron localization function of MgO₂, pyrite-type FeO₂, and PdF₂-type SiO₂ at 75 GPa; XRD pattern of reaction products of Pv-Fe₂O₃ and oxygen at 75 GPa; and the Py-FeO₂ and Pv-Fe₂O₃ interface and Py-FeO₂ and Rh₂O₃-II-type Fe₂O₃ interface module (PDF)

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Notes

The authors declare no competing financial interest.

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