

Challenges in Lithium Metal Anodes for Solid-State Batteries

Kelsey B. Hatzell,* Xi Chelsea Chen, Corie L. Cobb, Neil P. Dasgupta, Marm B. Dixit, Lauren E. Marbella, Matthew T. McDowell, Partha P. Mukherjee, Ankit Verma, Venkatasubramanian Viswanathan, Andrew S. Westover, and Wolfgang G. Zeier



Cite This: *ACS Energy Lett.* 2020, 5, 922–934



Read Online

ACCESS |

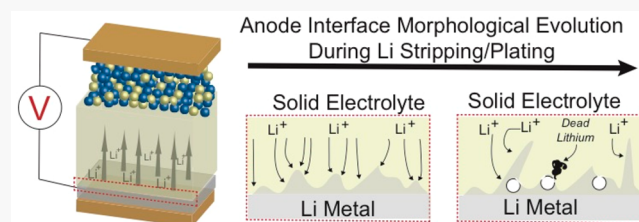


Metrics & More



Article Recommendations

ABSTRACT: In this Perspective, we highlight recent progress and challenges related to the integration of lithium metal anodes in solid-state batteries. While prior reports have suggested that solid electrolytes may be impermeable to lithium metal, this hypothesis has been disproven under a variety of electrolyte compositions and cycling conditions. Herein, we describe the mechanistic origins and importance of lithium filament growth and interphase formation in inorganic and organic solid electrolytes. Multimodal techniques that combine real and reciprocal space imaging and modeling will be necessary to fully understand nonequilibrium dynamics at these buried interfaces. Currently, most studies on lithium electrode kinetics at solid electrolyte interfaces are completed in symmetric Li–Li configurations. To fully understand the challenges and opportunities afforded by Li-metal anodes, full-cell experiments are necessary. Finally, the impacts of operating conditions on solid-state batteries are largely unknown with respect to pressure, geometry, and break-in protocols. Given the rapid growth of this community and the diverse portfolio of solid electrolytes, we highlight the need for detailed reporting of experimental conditions and standardization of protocols across the community.



There is considerable interest in replacing graphite anodes in rechargeable batteries with Li metal because of its high theoretical capacity (3860 mAh/g) and large negative potential (−3.06 vs NHE).¹ However, Li metal is prone to unstable side reactions, leading to solid electrolyte interphase (SEI) and dendrite formation when in contact with most liquid electrolytes. These side effects can lead to considerable safety concerns and a loss of efficiency. Solid electrolytes that are chemically stable against Li metal may be able to suppress these negative effects. However, there remain several unknown aspects to the nature of ion transport and kinetics at Li-metal/solid-electrolyte (Li-metal/SE) interfaces. For future adoption, it is critical to understand Li-metal/SE interfaces to mitigate Li filament formation, achieve reliable power densities, and maintain a high Coulombic efficiency over the cycling lifetime.

Controlling the morphology of Li metal during electrochemical stripping and plating processes at Li-metal/SE interfaces is necessary for the realization of energy-dense Li-metal solid-state batteries (SSBs).^{2–4} Li-metal SSBs could potentially increase energy density and reduce space and weight requirements (Figure 1a).⁵ However, achieving >10 mAh cm^{−2} cumulative capacity plated at failure, >3 mA cm^{−2} plating current density, >3 mAh cm^{−2} per cycle areal capacity,

and >80% fraction of Li passed per cycle are all necessary at the device level to displace the current state-of-the-art graphite anodes.⁴ Few reports come close to these targets because Li metal is prone to interfacial instabilities and significant volume changes. This Perspective summarizes the recent progress in understanding Li-metal/SE interfaces for SSBs, with the goal of ultimately controlling factors such as Li filament growth and interfacial reaction/delamination during electrochemical cycling. The mechanistic origins of lithium filament growth in SEs is discussed with respect to interphase formation and lithium mechanics. Next, the Perspective highlights state-of-the-art characterization tools equipped to observe Li/SE interfaces and discusses future needs, including techniques that provide complementary chemical and structural information. Finally, the Perspective concludes with a discussion of

Received: December 8, 2019

Accepted: February 18, 2020

Published: February 18, 2020



ACS Publications

© 2020 American Chemical Society

922

<https://dx.doi.org/10.1021/acsenenergylett.9b02668>
ACS Energy Lett. 2020, 5, 922–934

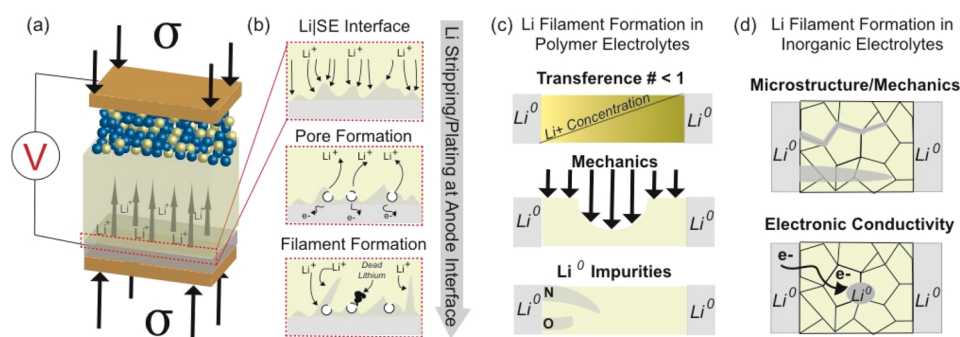


Figure 1. Li-metal solid-state battery architecture. To achieve adequate rate performance, pressure (σ) is applied to the cell (a). The morphology of the interface governs local ionic flux, pore formation in the Li-metal upon stripping, and Li filament growth (b). Li filament growth in a polymer solid electrolyte is governed by electrolyte mechanics, transference number, and the presence of interfacial impurities (c). Li filaments in inorganic SEs are attributed to electronic conductivity, microstructural properties, and interfacial contact (d).

frontiers in modeling and remaining open questions in the quest for Li-metal SSBs (Figure 1a).

Lithium Penetration through Solid Electrolytes. Graphite has been the standard anode material for decades because reactive Li metal can lead to dendrite formation, short-circuits, and fires. Concentration gradients across a cell can create excess charge at an interface (local space charge) and induce instabilities (e.g., dendrites). The term “Li dendrite” broadly describes the formation of a wide range of nonplanar/branched metal morphologies and may not accurately describe the complex mechanisms involved in SEs (Figure 1b). Instead, we suggest the use of the term Li filaments, and use this terminology here. Li filaments through SEs plague SSBs, leading to a loss of efficiency and cell failure. Understanding the nature of this growth is imperative for next-generation SSBs.

The term “Li dendrite” broadly describes the formation of a wide range of nonplanar/branched metal morphologies and may not accurately describe the complex mechanisms involved in SEs. Li filaments through SEs plague SSBs, leading to a loss of efficiency and cell failure. Understanding the nature of this growth is imperative for next-generation SSBs.

The current density at which a Li filament propagates across a SE is known as the critical current density (CCD). The magnitude of the CCD is important because it dictates the power density of a cell. Low CCDs are generally ascribed to iR losses that emerge as a result of interfacial impedances between the SE and the electrodes. These interfacial impedances are largely governed by chemical, electrochemical, and mechanical stability issues at interfaces. Deconvoluting these loss mechanisms and interpretation of CCD measurements can be challenging because SEs may be susceptible to “soft shorts” that can go undetected in symmetric cell experiments that utilize thick lithium metal ($>30\ \mu\text{m}$).⁴ Furthermore, the CCD is dependent on cell stack pressure and plating capacity, and thus, it is necessary to report detailed experimental conditions (e.g., Li-metal thickness and stack pressure) to accurately

interpret data.³ Recently, Sakamoto and co-workers described a new terminology, critical stack pressure, which highlights the pressure dependency in CCD measurements.³ This parameter is an important additional metric for evaluating Li electrode properties at SE interfaces. Pores in Li metal formed during electrochemical dissolution (stripping) and uneven Li⁺ transport at anode–electrolyte interfaces may be origins for dendrite nucleation.^{6–8} A nominal stack pressure can aid in mitigating pore formation during Li stripping. However, Li filament growth and propagation occur upon Li plating. When these occur simultaneously in Li–Li cells, the pressure will significantly affect the CCD. However, it is unclear exactly how these effects will translate to full cells, where plating and stripping are independent processes. Further work is necessary to resolve differences between whole cell and symmetric experiments.

Li filament growth at a solid polymer electrolyte (SPE) interface is dependent on the electrolyte’s transference number, the electrolyte’s mechanical properties, and Li-metal impurities (Figure 1c). Ionic transport in SPEs (for example, poly(ethylene oxide) (PEO)) mimics the transport characteristics of liquid electrolytes and can be accurately described by the concentrated solution theory for binary electrolytes based on the Onsager–Stefan–Maxwell formalism.⁹ Thus, current flux is a consequence of ion transport under both concentration and potential gradients in the polymer electrolyte, with a transference number of the cation being less than unity.⁸ Consequently, the CCD in SPEs is primarily determined by the limiting current density, wherein the cationic concentration goes to zero at Sand’s time, engendering unmitigated dendritic growth.¹⁰ Sand’s time is inversely proportional to the square of the anionic transference number, and thus, single-ion solid conducting SEs (inorganic or polymer) are theoretically stable against Li filament growth.^{11,12} However, a large number of single-ion conducting polymers have been developed without significant improvements to the CCD.^{12,13}

Monroe and Newman predicted that if the shear modulus of a SPE exceeded 6 GPa, Li filaments could not penetrate the SPE.⁶ Recent work by the Balsara group has shown that improving the modulus can increase the CCD, especially at high temperatures, but it cannot entirely eliminate Li filament formation.^{14–17} Thus, neither the mechanical properties nor the transference number of the SE can fully determine the growth of Li filaments through SPEs. In models such as the Monroe and Newman model, Li is assumed to be pure, which

may not be fully representative of the interfacial composition. X-ray imaging has shown that Li-metal surface impurities (Li_2O , Li_3N , or Li_2CO_3) can cause inhomogeneities in the local current density and promote the nucleation of Li protrusions (Figure 1c).^{18–20} Thus, while increasing the elastic modulus and transference number of SPEs is critical, the purity of the Li-metal source and surface may play a significant role in Li penetration at high current densities. Pretreatment strategies may be necessary for integration of Li metal.²¹

Single-ion conducting inorganic SE (glass and/or ceramic) have failure mechanisms that are distinctly different than that of SPEs (Figure 1d). Metal filament formation has been attributed to physical and/or microstructural properties of the SE, Li-metal/SE interfacial contact, and/or electronic properties of the electrolyte (Figure 1d).²² Pioneering work on solid-state Na β -alumina batteries revealed two modes for Na filament formation. Mode 1 describes filament formation and propagation with respect to SE microstructural features (grain boundaries, voids, scratches, etc.). This failure mode is based on Griffith's theories of fracture mechanics and was used to derive an empirical relationship for the critical current density of a Na β -alumina system²²

$$i_{\text{crit}} = \left[\frac{\pi(1 - \nu^2)F\gamma_{\text{eff}}^2}{8V_{\text{m}}E\eta} \right] \left(\frac{c^2}{l^3} \right)$$

where ν is Poisson's ratio, F Faraday's constant, γ_{eff} the surface energy, V_{m} the molar volume of Na, E Young's modulus, η the Na viscosity, $2c$ the crack width, and l the crack length. In simpler terms, the CCD is determined by the aspect ratio of surface defects (c^2/l^3) such as voids, scratches, or cracks, the wetting of the alkali metal to the solid electrolytes (γ_{eff}), and the mechanical properties of the electrolyte. Temperature also plays a key role as this will directly affect the surface wetting (γ_{eff}) and viscosity of the alkali metal (η). Previously, it was shown that heating and constant pressure enable improved contact, lower interfacial resistances, and higher critical current densities.²³ In contrast, mode 2 describes how the electronic properties of the SE governs filament growth.^{24,25} During operation, Na β -alumina changes colors as the electronic conductivity of the material changes. When the SE is electrically conductive, Na can deposit within the SE and act as a nucleation site for filament growth. These two growth modes provide valuable insights into the driving mechanism for filament propagation in Li SEs. Recent research with Li metal strongly suggests that Li filament propagation is analogous to that of the Na metal filament propagation.

Parallels between Na and Li filament growth mechanisms can be drawn from a series of reports by Porz and Swamy.^{26,27} Porz et al.²⁶ studied Li filament formation and growth in a wide range of inorganic solid electrolytes ($\text{Li}_7\text{La}_4\text{Zr}_2\text{O}_{12}$, crystalline β - Li_3PS_4 , and lithium thiophosphate glasses). Their work demonstrates that lithium filaments prefer to grow through cracks and defects and that plating-induced lithium infiltration can take place in pre-existing microstructural flaws. Later work by Swamy et al. observed that lithium filaments preferentially grew at the electrode edge in grain-boundary free, single-crystal LLZO. Preferential Li growth at the electrode edge, rather than at engineered defects, reveals that filament formation can be driven by local electric-field hot spots. This current focusing leads to preferential Li deposition and results in enhanced mechanical stresses at these locations. Another recent paper by Westover

et al.²⁸ corroborates the role that current focusing and surface flaws (e.g., grain boundaries) play on filament propagation. Li filaments demonstrated preferential growth at artificial Lipon–Lipon interfaces and completely avoided the bulk electrolyte. Li filament penetration along the interface is attributed to current focusing, the presence of Li_2CO_3 defects, and the interfacial mechanical properties. All of these factors can contribute to the relatively wide spread in reported CCD values for Li-metal SSBs. There is currently a crucial need for unified, standard operating protocols (ramp rate, current hold duration, etc.) and standard cell architectures for CCD measurements for effective comparison across the community.

Although the most well-studied mechanism for Li filament penetration is mode 1 (microstructure, defect, etc.), a recent paper by Han et al.²⁹ highlighted mode 2 (electronic conductivity) mechanisms. *In situ* neutron depth profiling experiments revealed increases in Li-metal content within the SE (Lipon, LLZO, and β - Li_3PS_4) during electrochemical cycling. Increases in Li content within the SE suggests that local depositions of Li metal can occur within the bulk SE (Figure 1d). These deposits can act as nucleation sites for filament growth and can cause a short circuit. Electron microscopy³⁰ and X-ray tomography³¹ results also showed isolated Li deposits that support this finding. It has been hypothesized that the SE electronic conductivity is responsible for these localized Li deposits;³² however, recent theoretical studies suggest that local reduction of La or Zr at the grain boundaries³³ or electrons trapped at surface defects may drive deposition events.³⁴ Further research is needed to fully understand this mechanism, or combination of mechanisms, and to guide the development of materials and engineering solutions that enable high CCDs at room temperature.

Aside from SE properties, the mechanical properties of the lithium metal anode also play an important role in Li filament formation in SEs. Recent studies of the bulk mechanical response of Li metal have demonstrated that power-law creep is the dominant deformation mechanism over a wide range of strain rates and temperatures.^{35,36} These results help contextualize the evolution of mechanical stresses as Li is plated out at a solid-electrolyte surface from a viscous-flow perspective. The coupled relationships between current density and strain rate directly impact stress accumulation at the solid–solid interface. This has significant implications for the model of Porz et al.,²⁶ where extrusion of Li metal out of structural heterogeneities on the SE surface will contribute to the current-density dependence of mechanical failure of the SE.

Furthermore, both nanoindentation^{37,38} and micropillar compression³⁹ experiments have reported that lithium can support significantly higher stresses when confined to small length scales. Using nanoindentation, Herbert et al. observed a change in the underlying deformation mechanism as a function of indentation depth.³⁸ At shallow indentation depths, self-diffusion dominated the flow, and at deeper depths, shear-driven dislocation motion was the primary deformation mechanism. Within the diffusive flow regime, the pressure that Li is capable of supporting was found to be significantly higher than the bulk yield stress and strongly dependent on the strain rate, which is directly impacted by the current density. Both of these results have implications for the propagation of Li filaments within solid-state electrolytes. More research is needed to understand the role that Li-metal mechanics plays on filament formation and growth. In particular, improved knowledge regarding Li purity (especially at the SE interface),

Li mechanical properties, and how the morphology and microstructure of Li metal change during electrochemical cycling is needed.

Interphase Formation and Engineered Interfaces. Solid electrolyte redox stability depends on the alignment of the solid electrolyte's valence and conduction bands relative to the Li chemical potential (μ_{Li}) and decomposition thermodynamics.^{40,41} Thus, most SEs will form an interphase at Li metal due to redox reactions. The formation of interphase layers between Li and SEs has widespread consequences for the operation of SSBs. Many SEs under consideration are unstable in contact with Li metal, and the evolution of the reacted interphase is expected to be dependent on its transport properties.^{42–44} An interphase that conducts both ions and electrons (a mixed conductor) will continue to grow with time because of direct electrochemical reaction occurring at the SE. An interphase that conducts only ions and not electrons could grow to a stable thickness, which would ideally result in a passivating layer that enables long-term stability. An interphase with insufficient ionic conductivity, however, will cause increased impedance. Finally, for stability in contact with Li metal, the interphase must prevent atomic Li diffusion from the Li anode to the SE, which is challenging given the relatively high diffusion rate of Li in many materials.

The chemo-mechanics of interphase formation and growth have important ramifications for battery operation. Li insertion and microstructural transformation during interphase growth causes the evolution of mechanical stress both within the interphase and within the SE.^{45,46} In SE materials where the interphase grows to be relatively thick, such stresses can be large enough to mechanically fracture the SE. Fracture of $\text{Li}_{1-x}\text{Al}_x\text{Ge}_{2-x}(\text{PO}_4)_3$ (LAGP) caused by interphase growth was recently monitored with *in situ* X-ray imaging, and it was found that this mechanical degradation was the primary cause of increased cell impedance, rather than the transport properties of the interphase itself.⁴⁷ The chemo-mechanics of interphase

The chemo-mechanics of interphase formation is also important in the context of electrochemical deposition and stripping of Li metal, but this topic has received less attention to date.

formation is also important in the context of electrochemical deposition and stripping of Li metal, but this topic has received less attention to date. For instance, stress within the interphase or in the SE near the Li-metal interface could alter diffusion barriers and pathways for Li ions. Furthermore, local variations in interphase morphology could lead to stress nonuniformities that may locally alter the potential, which could impact Li deposition/stripping and lead to the formation of Li filaments or dendrites.⁷ The interrelationship between interphase growth and Li deposition/stripping mechanics thus requires dedicated investigation by the community. To date, there have been two primary approaches to mitigating interphase effects: (1) engineering the lithium metal and/or (2) engineering an interfacial layer material (e.g., interlayer).

One promising approach toward engineering the anode is through controlled alloying. Li-rich alloys (e.g., Si, Sn, Al, Mg, In, etc.) may play an important future role in SSBs. Such materials have received significant attention as high-capacity

anodes for liquid-based Li-ion batteries and are now being integrated in small quantities into commercial graphite anodes to boost capacity.⁴⁸ In SSBs, they could be useful either as interfacial layers or as anode active materials. In principle, alloy materials can experience greater diffusion coefficients than pure lithium. Recently, Li–Mg alloys were used to maintain interfacial contact and increase Li utilization with no external applied pressure in SSBs.⁴⁹ Furthermore, Li–In alloys were shown to be chemically stable against Li_3PS_4 when the alloy remained in the two-phase region (In)–InLi.⁵⁰ As described previously, the mechanical properties of Li metal (e.g., creep behavior) may not be sufficient to avoid delamination and pore formation at high current densities.^{2,3} The addition of Li-alloy interfacial layers in contact with Li-metal anodes could allow for tuning of their mechanical and Li transport properties, which could aid in maintaining uniform interfacial contact at the SE interface. The use of high-capacity Li alloys as anodes instead of Li metal is another potential route that could lead to high-energy SSBs. Many Li alloys may be more stable in contact with SEs because of their higher potentials compared to the Li/Li⁺ redox couple.⁵⁰ However, further research needs to be carried out to understand the properties and structure of interphases between SEs and Li alloys, especially as delithiation kinetics and Li diffusivities at the interface will be different in alloys compared to pure Li metal.

Engineered interfacial layers between Li metal and SEs can aid in improving cell performance. Much of the work in this area has focused on reducing interfacial impedance and improving physical contact between Li and the SE. The two primary routes to achieve these goals involve either introducing a physical interlayer or controlling the surface chemistry of a specific material (electrode or electrolyte).^{21,51} Interfacial layers (interlayers), such as polymer coatings, can act as a barrier film to decrease chemical decomposition at solid/solid interfaces during electrochemical cycling and to promote uniform Li deposition/stripping.⁵² Several coating strategies are reported within the literature. A soft and flowable coating can prevent the formation of cracks or pinholes in the SEI layer.⁵³ A highly elastic coating can regulate interface uniformity by exerting high strength at high-strain areas.⁵⁴ Additionally, a polymer coating with chemistries that can mitigate the reactivity between the electrolyte and Li is beneficial.^{55–57} These coatings are formed either *ex situ* or *in situ*. Other strategies include adding a very soft polymer layer underneath the current collector and Li instead of between Li and the SE. The soft polymer layer can effectively diminish the plating stress and prevent stress-driven dendrite growth.⁵⁸

Besides interfacial layers, another strategy in which ceramic ion conductors are combined with polymer ion conductors to form hybrid membranes has been shown to enable more compliant and tunable interfaces.^{59,60} In the hybrid approach, the ceramic SE could provide a high mechanical modulus and high ionic conductivity, and the polymer electrolyte could provide ease of manufacture as well as improved adhesion and stability with the electrodes.⁶¹ However, this approach requires careful design and proper processing methods to be able to take advantage of the advantageous properties of both the polymer and the ceramic.^{61–64} In particular, the interface between the polymer and the ceramic electrolytes needs to be optimized.^{65,66}

Alternatively, it may be feasible to design the SE to decompose into a beneficial interphase without the need for an additional interlayer. Ideally, such engineered interphases

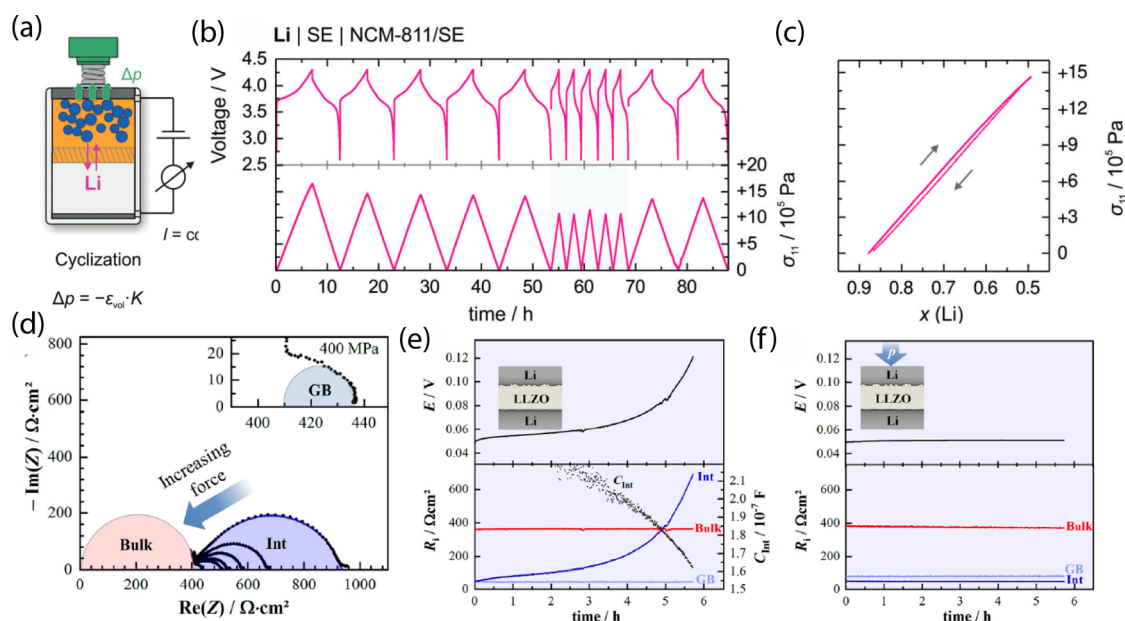


Figure 2. (a) Schematic of a measurement setup for pressure monitoring during solid-state battery operation.⁷³ (b and c) Volume expansion and contraction of the lithium metal anode during cycling results in large pressure changes and pressure oscillations.⁷³ Reproduced with permission from ref 73. Copyright 2018 The Royal Society of Chemistry. (d) Pressure-dependent impedance response of Li/LLZO/Li cells. At high enough preforming pressure, the interface resistance between Li and LLZO can be mitigated. However, while preformation is necessary, unless there is a constant applied pressure on the cell contact loss will occur during stripping experiments. Hence, high pressures are needed for preformation; however, some external pressure is still needed to avoid pore formation.⁷⁴ Panels d–f are reprinted from ref 74. Copyright 2019 American Chemical Society.

would prevent (electro)-chemical reaction between the SE and Li while still allowing for Li-ion transport. This requires materials that conduct ions, block electrons, and are (kinetically) stable in contact with Li. While a variety of polymers, ceramics, and even metals have been reported to improve the stability of reactive SEs,^{35,67,68} characterizing the atomic structure of the interphase that forms upon cycling is challenging, and further efforts are necessary to elucidate the interphase evolution in the presence of such protection layers.

Ultimately, strategies to mitigate deleterious interphase formation need to be extended to roll-to-roll (R2R) scales.^{69–71} Li metal forms a passivating surface film (even in dry environments), and this film can lead to a high interfacial resistance that can result in cell degradation. High interfacial resistance must be minimized during manufacturing, if large Li-metal foils are used in place of traditional anodes. Vapor deposition or melt processing techniques can be used to create a uniform layer of Li metal on top of the SE or interfacial layers.⁷² However, depositing Li metal in this manner can drive up manufacturing costs. These costs may be offset by the inclusion of a smart manufacturing environment, where inline quality control and environmental monitoring are used to reduce manufacturing defects and increase overall cell yield. In addition to manufacturing, the electrode architecture, cell design, and packaging must be re-engineered for future SSBs while considering the implications for both manufacturing and high-performance cells.

Pressure during Operation and Formation in a Solid-State Battery. SSBs rely on intimate contact between solid phases, i.e., the active material and the SE, as well as possible additives (Figure 2a). Charging and discharging leads to volume changes in the anode and cathode active materials. In the case of the anode, the volumetric expansion can be up to 300% for extreme cases such as pure silicon, and even graphite

undergoes volume expansion of approximately 10% upon full lithiation. The pressure oscillations due to volume change of a Li-metal anode in a thioSSB are shown in Figure 2b,c, in which solid–solid contacts will exhibit localized stresses. These local tensile stresses lead to cracking and bending of SSBs if no external pressure is used, ultimately resulting in contact loss and poor battery performance. Thiophosphate-based SEs are generally softer and more elastic ($K/G > 1.75$, where K is the bulk modulus and G is the Shear modulus in GPa) than other inorganic solid electrolytes (e.g., LLZO). The elasticity of the SE will be important for decreasing fracture events in all-inorganic cathodes⁷³ and will likely affect the critical state pressure.

Applying pressure to the SSB often occurs in two steps. First, high pressure (which may also be accompanied by an elevated temperature) can be applied during cell fabrication to initially contact all components. Subsequently, a lower external pressure is used during cycling to avoid contact loss.³ The initial pressure required for formation depends on the mechanical properties of the SE and active materials, as well as the interfacial wetting properties with the Li anode. In contrast, the external pressure during cycling only counteracts the detrimental volume changes. Panels d–f of Figure 2 show examples of these different scenarios. When using an Li/LLZO/Li cell, the interface resistance between Li and LLZO decreases and ultimately becomes negligible at 400 MPa of an external preforming pressure.⁷⁴ However, when using the Li/LLZO/Li cell during stripping experiments, a reduced pressure of 35 MPa is sufficient to retain good contact and prevent void formation and morphological instabilities during stripping.⁷⁴ For polymer electrolytes, a decrease in interfacial impedance has also been observed with increasing stack pressure, until a “critical stack pressure” is reached.⁷² Overall, the needed pressure conditions during forming, the maintenance of stack

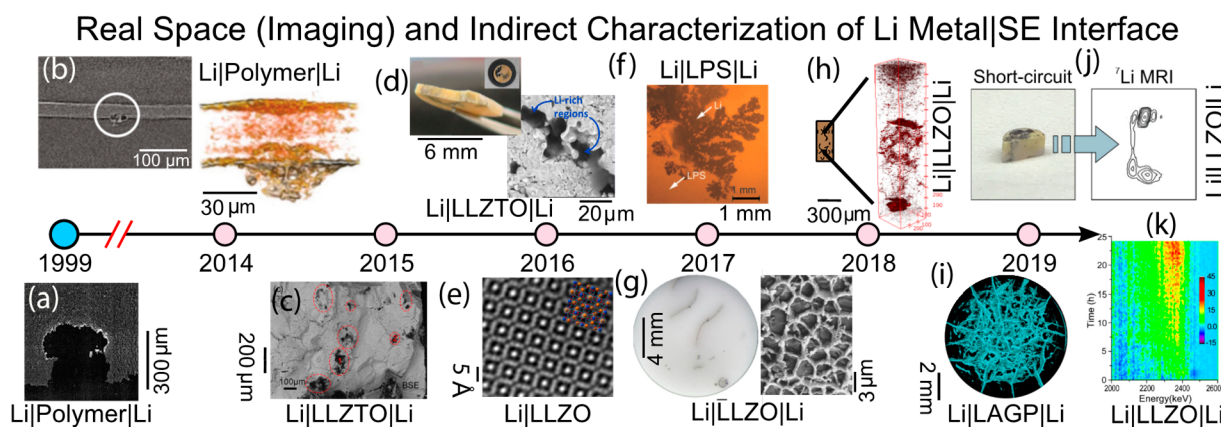


Figure 3. Timeline of imaging and characterization of Li-metal/SE interfaces. Optical techniques (a) (Copyright 1999, Elsevier) and X-ray tomography (b) reveal morphology of Li filaments at a polymer SE interfaces. (Reprinted by permission from ref 19. Copyright 2013 Springer Nature.) *Ex situ* scanning electron microscopy revealed subsurface deposition in inorganic electrolytes (c and d). (Reproduced with permission from ref 75. Copyright 2015 Elsevier.) Transmission electron microscopy experiments reveal structural transformation at LLZO interfaces in contact with Li metal. (Reproduced from ref 76. Copyright 2016 American Chemical Society.) Optical techniques demonstrate dendritic growth of β - Li_3PS_4 polycrystals (Reproduced with permission from ref 26. Copyright 2017 Wiley) (f). Lithium metal extruding out of intergranular regions (g) in LLZO (Copyright 2015, Elsevier) and microstructural representation of Li metal forming in pores from X-ray tomography experiments (h). (Reproduced from ref 31. Copyright 2018 American Chemical Society.) Large mechanical fracture of LAGP solid electrolyte (i) and evidence of lithium penetration across a LLZO electrolyte imaged with NMR (j). (Reproduced from ref 77. Copyright 2019 American Chemical Society.) Neutron depth profiling enabling tracking of lithium at subsurface regions in a SE (k). (Reproduced from ref 78. Copyright 2017 American Chemical Society.) Figure reproduced from refs 18, 19, 26, 31, 47, 75, and 77–80.

pressure under different operation conditions, and the implications of these needed pressures on manufacturing concerns are still open questions.

Diagnostics and Characterization Tools for Li/SE Interfaces. Diagnostics tools and characterization techniques that can probe subsurface phenomena are critical to improve our understanding of the mechanistic origins of Li filament growth and degradation pathways at solid/solid interfaces (Figure 3). There are different opportunities for different characterization techniques with respect to (1) technique spatial and temporal resolution, (2) sample preparation and experiment characteristic length scale, and (3) working environment. There are also several mechanisms and processes that occur at different length scales within a battery, and the resolution of the technique (both temporal and spatial) provides boundaries for experiments. Finally, the experimental working environment can introduce transient and nonequilibrium conditions, especially if the technique is destructive.

Over the last several years, there has been considerable interest in characterizing physical and chemical transformations at solid/solid interfaces that contribute to interfacial evolution and metal filament formation over a wide range of time and length scales.^{26,75,79} While the theoretical basis for the formation of passivating interphases is clear from the previous discussion, the actual structure, chemistry, and transport properties of interphases that form on the variety of SSE materials are largely unknown.⁸¹ Recently, SE chemical decomposition pathways when in contact with Li was discerned using *in situ* X-ray photoelectron spectroscopy (XPS) experiments^{82–84} and *ex situ* scanning electron and optical microscopy has revealed filament formation within the bulk SE (Figure 3a–d,f,g). *In situ* transmission microscopy (TEM) (Figure 3e) has also revealed important information about interfacial chemical and structural changes.^{52,76,81} In cubic LLZO, for instance, a few-nanometer region near the interface converts to the tetragonal phase because of Li insertion.⁷⁶ The NASICON-type LAGP material was found to

react and amorphize in contact with Li,⁸¹ with none of the thermodynamically predicted compounds observed via diffraction. Solid-state nuclear magnetic resonance (NMR) spectroscopy has been employed to gain an understanding of the composition⁸⁵ of amorphous decomposition products present at electrode/electrolyte interfaces and how these species impact Li-ion transport in SSBs.⁸⁶ Expanding the use of other *in situ/operando* techniques well-suited to characterizing amorphous phases will enable the investigation of structural changes in a wider variety of SE materials. Because of the highly heterogeneous and dynamic nature of interphase formation it is desirable to couple theory and computation with characterization to aid in experimental interpretation and decrease experimental uncertainty. This is key to understanding interphase dynamics and will also inform the engineering of artificial interphases with beneficial properties.

Beyond interphase characterization, there is a considerable need to understand mesoscale material transformations (delamination, filament growth, etc.). NMR techniques are powerful tools for understanding Li microstructural growth and Li-ion dynamics within SSBs.⁷⁵ ^7Li NMR is able to distinguish between ^7Li resonances from the bulk Li electrode and Li microstructures (filaments and/or dead Li).⁸⁷ Furthermore, ^7Li NMR combined with magnetic resonance imaging (MRI) can enable 3D spatial tracking of Li filament growth using ^7Li chemical shift imaging (CSI). Marbella et al. used *ex situ* ^7Li CSI to correlate Li microstructural growth with instabilities in galvanostatic cycling⁷⁷ (Figure 3j). These experiments showed heterogeneity at both interfaces, suggesting the formation of local “hot spots” during Li stripping and deposition. While the imaging dimension in these experiments is limited to approximately 300 μm , NMR can provide additional chemical resolution as well as enable *in situ/operando* experimentation. Recently, *in situ* 3D ^7Li MRI experiments enabled tracking of local Li concentration gradients within the bulk of $\text{Li}_{10}\text{GeP}_2\text{S}_{12}$ (LGPS).⁸⁸ These experiments directly corroborated increases in interfacial

resistances with Li concentration depletion at the anode interface. NMR will continue to play a crucial role in SSB characterization because of the unique ability of NMR to identify disordered structural features⁸⁹ that are often present at interfaces, as well as to monitor Li-ion dynamics.⁹⁰

X-ray computed tomography (XCT), like 3D ⁷Li MRI, is also a 3D imaging technique but can achieve higher spatial resolutions of <1 μ m (micro-XCT) and <50 nm (nano-XCT). Synchrotron XCT can be leveraged to carry out non-destructive, *in situ/operando* measurements because of fast acquisition times (Figure 3i,j).^{15,17,19,31} The key drawback of this technique is the lack of chemical specificity toward Li. Thus, only indirect observation of filament growth can be achieved. Neutron depth profiling (NDP) (Figure 3k) is another tool to investigate Li-metal deposits near the anode electrolyte interface. NDP is a near-surface analysis technique that employs cold neutrons to excite Li to emit α particles and protons which are subsequently detected.³ NDP is a nondestructive tool that can offer insight into Li concentrations near the interfacial region with micrometer-level resolution. However, NDP is currently limited in its ability to track deposition within the bulk and in its ability to track local structural characteristics.

Most of the characterization techniques used to investigate chemical, physical/mechanical, and microstructural transformations at Li/SE interfaces have focused on a singular aspect of the system. Experimental results have conclusively

Experimental results have conclusively shown a strong codependence between material transformations and electrochemical performance. Thus, it is vital to pursue multimodal characterization in order to achieve complementary datasets on coupled phenomena. Coupling techniques like atomic force microscopy with spectroscopy can provide simultaneous information regarding mechanical, structural, and chemical properties.

shown a strong codependence between material transformations and electrochemical performance. Thus, it is vital to pursue multimodal characterization in order to achieve complementary datasets on coupled phenomena. Coupling techniques like atomic force microscopy with spectroscopy can provide simultaneous information regarding mechanical, structural, and chemical properties. Synchrotron techniques also offer high versatility in terms of combining techniques to probe multiple phenomena at varying length scales; for example, transmission X-ray microscopy coupled with X-ray absorption near edge structure (XANES), or microtomography coupled with X-ray diffraction/absorption spectroscopy. *Operando* optical video microscopy can be synchronized with voltage signatures during cycling to provide insights into electrochemical signatures of Li filament propagation.⁷² In the future, *operando* and *in situ* measurements coupled with theory and computation will continue to offer new insights into the dynamic, coupled phenomena occurring in these systems.

Modeling. Advanced modeling of SSBs is critical for the interpretation of experiments, probing transport mechanisms, and describing the phenomenological origin of interfacial transformations. Mechanical stresses at the Li/SE interface are intrinsic to SSBs, which point toward adopting rigid barriers as a possible strategy for dendrite mitigation. The complexities of ion transport in the electrolyte, reaction kinetics, and mechanics due to stress generation at the Li/SE interface need to be understood to correlate with changes in SSB performance (Figure 4a).

Transport behavior in all-inorganic and all-polymer electrolytes has been explored at the continuum scale, yet formulations for stress-induced modifications to ion transport have yet to be identified. The effects of stress on transport in the SE are likely important, which may be analogous to stress effects on Li diffusion during intercalation events.⁹¹ Accurate experimental determination of ionic conductivity as a function of external pressure as well as transport properties are necessary for building effective models for all inorganic or polymer SSBs. Furthermore, *operando* and *in situ* experimentation which combines real space imaging techniques with electrochemistry and transport measurements will aid in the development of accurate mesoscale models (Figure 4b–d). While transport mechanisms at the continuum level are well understood for homogeneous SEs, less is known about transport in hybrid composite SEs. The ion transport behavior in inorganic/polymer composites is complex and provides a rich design space to explore and optimize.^{63,64,92} Developing a molecular-scale understanding of ion transport at the interface of polymer–ceramic composites could help to develop high-performing composite electrolytes that can manifest the benefits of inorganic and polymeric materials.⁹² Recently, effective mean field theory (EMFT) approaches combined with experimentally obtained ionic conductivity measurements were used to map ion transport pathways in hybrid SEs.⁶³ While EMFT provides a pathway for discerning between anionic and cationic transport pathways at complex inorganic/organic interfaces, it does not include solvation/desolvation mechanisms and polymer physics. Advanced modeling and experiments are necessary to deconvolve transport mechanisms between polymer and ceramic ion conductors.

Monroe and Newman proposed the first model of stress-kinetics coupling by correlating the exchange current density from Butler–Volmer kinetics at the Li–polymer interface with the electrochemical potential change through interfacial hydrostatic/deviatoric stresses and molar volume mismatch.^{6,7} Furthermore, linear elastic perturbation analysis on a static metal–polymer system incorporating the stress-kinetic formulations was used to determine a threshold shear modulus of SEs for stable deposition, precluding transport effects. Barai and Srinivasan extended this model to incorporate transport, plastic Li deformation, and external pressure effects.^{8,93} Model reformulation for inorganic SEs to delineate stability regimes, as well as accounting for the polycrystalline microstructure of the SE, has been explored as well.^{11,94–96} First-principles studies have also proposed a mechanism for Li-metal penetration through polycrystalline solid electrolytes based on grain boundary softening.⁹⁷ However, studies have yet to ascertain a singular root cause of the CCD observed across all classes of inorganic electrolytes, whether single-crystal or polycrystalline. Recently, Griffith's theory based on linear elastic fracture mechanics has been proposed to understand the mechanism of lithium metal penetration through inorganic

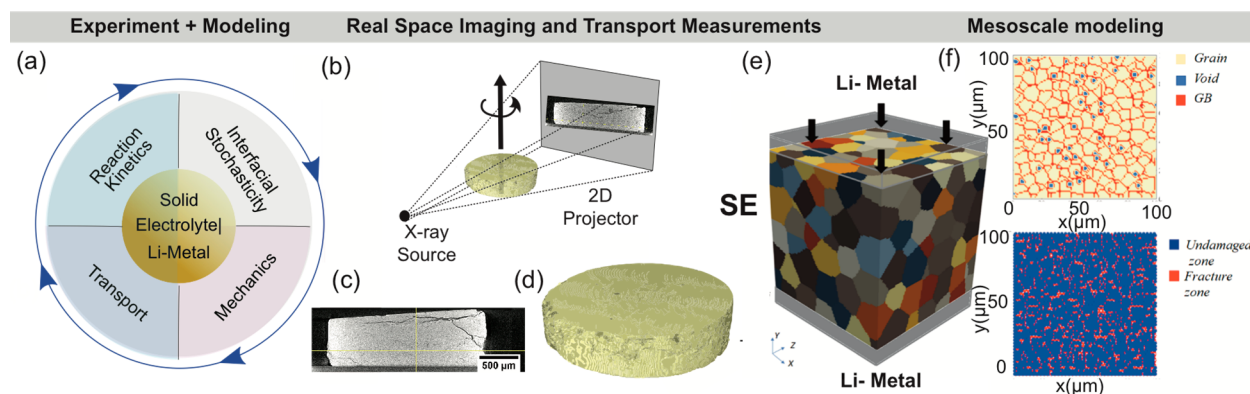


Figure 4. Advanced mesoscale models that combine experimentally obtained properties (transport, kinetics, interfacial properties, and mechanics) are necessary for discerning fundamental transformations in all-solid-state batteries (a). *Operando* experiments that combine real space imaging techniques (b–d) can provide a basis for modeling domains (Reproduced from ref 98, copyright 2019 American Chemical Society). Nonequilibrium meso-structural transformation is important for building effective models. Schematic of polycrystalline SE sandwiched between Li-metal electrodes in symmetric cell configuration with (e) corresponding microstructure phase map consisting of grains, voids, and grain boundaries (GB) and (f) damage map exhibiting fracture concentration along the grain boundaries with electrodeposition under the constrained configuration.

solid electrolytes, correlating the lithium plating overpotential to the surface defect size and fracture toughness of the electrolyte.²⁶ This theory gives contrasting results to the Monroe–Newman model; high Young’s modulus solid electrolyte metal is associated with lower CCDs through higher magnitudes stress-based crack opening. It is the authors’ opinion that a reconciliation of the two theories can be achieved through incorporation of pertinent interfacial physics to the models, including chemical/mechanical wettability and contact/fracture mechanics resulting from inherent interfacial surface roughness.

It is imperative to note that there is an ongoing debate in the scientific community with regards to the molar volume inside the inorganic solid electrolyte. Zero molar volume has been reported for LLZO in the literature based on negligible volumetric change with Li⁺ incorporation, alongside a wide range of values for different SEs.^{94,95} Ahmad et al. has delineated the stability regime of inorganic solid electrolytes as a function of solid electrolyte to Li-metal molar volume and shear modulus ratio and has shown that inorganic solid electrolytes with low molar volume and high shear moduli ratios (such as LPS or LLZO) are prone to current instabilities.¹¹ A bridge needs to be built between the molecular scale definitions of molar volume with those used in continuum models; in particular, there is a lack of consensus as to whether to correlate molar volume to the cationic/anionic radii or to consider it as an inherent part of the crystal structure.

Detailed insights into the performance of SSBs through modeling requires a dynamic model that is capable of temporal tracking of deposition/stripping at the interface which includes the interactions emanating from elastic–plastic stresses, ion transport, and electrochemistry. Furthermore, incorporation of chemical (surface tension) and mechanical wettability (external pressure), imperfect contact/peeling, and microstructural heterogeneities (e.g., voids, grains, grain boundaries for a polycrystalline electrolyte) into the modeling paradigm will help elucidate the scenarios that limit performance (Figure 4e,f). The presence of an interphase/interlayer between the metal and SE will add to the model complexity, requiring stresses and transport to be solved in an additional layer, while

accounting for characteristics of the layer (brittle/ductile deformation, Maxwell/Ohm transport law, Butler–Volmer/Tafel kinetics, etc.). First-principles studies catalyzed by machine learning approaches will also be useful in rapid estimation of the transport, mechanical, and thermodynamic properties of the SE materials.⁹⁹ First-principles atomistic calculations utilizing density functional theory catalyzed by machine learning approaches will also be useful in rapid estimation of the transport (ionic conductivity/diffusivity), mechanical (stiffness tensor), kinetic (exchange current density), and thermodynamic (molar volume) properties of the SE materials.⁹⁹ In the literature, density functional theory has proven useful in estimating the ionic conductivity and Young’s modulus of both amorphous and polycrystalline solid electrolyte materials.^{100,101} Our conjecture and proposition is that dynamic mesoscale interactions,^{102–105} including the interplay between reaction kinetics, interfacial stochasticity, transport, and mechanics, are essential to provide fundamental mechanistic insights into Li-metal-based SSBs (Figure 4e).

Outlook and Future Directions. To date, the majority of work on Li filament formation in SSBs has focused on characterizing Li plating/stripping in Li/Li symmetrical cells. Interfacial phenomena that are key to filament nucleation and growth may differ substantially in full cells, emphasizing the need to expand these characterization techniques to more realistic systems. Furthermore, the formation of heterogeneous interphases can lead to stress nonuniformities that impact the local accumulation of stripping and/or deposition products. While there are many potential pathways toward mitigating lithium filaments (e.g., Li alloys, pressure, and SE strength), maintaining uniform contact between the electrode and SE is paramount. Characterization and control over the atomic structure of the interphase will be important for addressing chemo-mechanical challenges at solid/solid interfaces. Advanced techniques that combine real and reciprocal space experimental techniques with modeling may provide a means for probing the variety of length scales that exist in a solid-state battery at realistic temporal and spatial length scales. There is also an emerging need for standardization of experimental methods across the community. Communication of protocols and operating conditions (Li-metal thickness,

temperature, pressures, etc.) will be paramount for interpretation of results within this rapidly growing community, with the ultimate goal of achieving lithium–solid electrolyte interfaces for solid-state battery operation.

AUTHOR INFORMATION

Corresponding Author

Kelsey B. Hatzell — Department of Mechanical Engineering, Interdisciplinary Department of Material Science, and Department of Chemical and Biomolecular Engineering, Vanderbilt University, Nashville, Tennessee 37202, United States; orcid.org/0000-0002-5222-7288; Email: Kelsey.b.hatzell@vanderbilt.edu

Authors

Xi Chelsea Chen — Chemical Sciences Division, Oak Ridge National Laboratory, Oak Ridge, Tennessee 37830, United States; orcid.org/0000-0003-1188-7658

Corie L. Cobb — Department of Mechanical Engineering, University of Washington, Seattle, Washington 98195, United States

Neil P. Dasgupta — Department of Mechanical Engineering, University of Michigan, Ann Arbor, Michigan 48103, United States; orcid.org/0000-0002-5180-4063

Marm B. Dixit — Department of Mechanical Engineering and Interdisciplinary Department of Material Science, Vanderbilt University, Nashville, Tennessee 37202, United States

Lauren E. Marbella — Department of Chemical Engineering, Columbia University, New York, New York 10027, United States; orcid.org/0000-0003-1639-3913

Matthew T. McDowell — G. W. Woodruff School of Mechanical Engineering, Atlanta, Georgia 30332, United States; School of Materials Science and Engineering, Atlanta, Georgia 30332, United States; orcid.org/0000-0001-5552-3456

Partha P. Mukherjee — School of Mechanical Engineering, Purdue University, West Lafayette, Indiana 47907, United States

Ankit Verma — School of Mechanical Engineering, Purdue University, West Lafayette, Indiana 47907, United States

Venkatasubramanian Viswanathan — Department of Mechanical Engineering, Carnegie Mellon University, Pittsburgh, Pennsylvania 15213, United States; orcid.org/0000-0003-1060-5495

Andrew S. Westover — Chemical Sciences Division, Oak Ridge National Laboratory, Oak Ridge, Tennessee 37830, United States

Wolfgang G. Zeier — Institute of Physical Chemistry and Center for Materials Research (LaMa), Justus-Liebig-University Giessen, D-35392 Giessen, Germany; orcid.org/0000-0001-7749-5089

Complete contact information is available at:

<https://pubs.acs.org/10.1021/acseenergylett.9b02668>

Notes

The authors declare no competing financial interest.

Biographies

Kelsey B. Hatzell is an assistant professor at Vanderbilt University with appointments in Mechanical and Chemical and Biomolecular departments. Hatzell has received numerous awards, including the NSF CAREER award, ECS Toyota Young Investigator Award, and MRS Nelson “Buck” Robinson Award. Her research interests focus on understanding transport in solid-state batteries.

X. Chelsea Chen is a staff scientist in the Energy Storage and Membrane Materials Group at Oak Ridge National Lab. Her fields of expertise are polymer physics, ion transport in polymers, and structural and morphological characterizations. Her current research focuses on developing polymer–ceramic composite electrolytes and understanding polymer/ceramic interfaces.

Corie L. Cobb is the Washington Research Foundation Innovation Professor in Clean Energy and Associate Professor of Mechanical Engineering at the University of Washington. She is a recent recipient of a 2019 DARPA Young Faculty Award and a 2019 3M Non-Tenured Faculty Award. Her current research focuses on computational design and manufacturing solutions for 3D and solid-state Lithium-ion batteries.

Neil Dasgupta is an Assistant Professor in the Department of Mechanical Engineering at the University of Michigan. He has received numerous awards, including the NSF CAREER award, DARPA Young Faculty Award, and AFOSR Young Investigator Award (YIP). His research focuses on the intersection of materials chemistry, electrochemical energy storage, and nanomanufacturing.

Marm B. Dixit is a doctoral candidate of Mechanical Engineering at Vanderbilt University. He is working on multimodal solid-state batteries using synchrotron techniques. He has published nearly 20 journal articles on electrochemical systems.

Lauren E. Marbella is an Assistant Professor of Chemical Engineering at Columbia University. Her group uses NMR and MRI to understand interfacial phenomena and structure–performance relationships in systems for electrochemical energy storage and conversion.

Matthew T. McDowell is an assistant professor at Georgia Tech with appointments in the G. W. Woodruff School of Mechanical Engineering and the School of Materials Science and Engineering. McDowell has over 70 publications and has received numerous awards, including the Presidential Early Career Award for Scientists and Engineers (PECASE), the Sloan Fellowship, the NSF CAREER Award, and the NASA Early Career Faculty Award.

Partha P. Mukherjee is currently an Associate Professor of Mechanical Engineering at Purdue University. Prior to that, he held positions at Texas A&M University and Oak Ridge and Los Alamos National Laboratories. His research is focused on mesoscale physics and stochastics of transport, chemistry, and microstructure interactions in energy storage and conversion.

Ankit Verma is a doctoral candidate of Mechanical Engineering at Purdue University. He is working on multimodal degradation mechanisms and interactions in lithium intercalation and metal batteries. He has authored/coauthored more than 10 journal papers on batteries and energy storage.

Venkat Viswanathan is an Associate Professor of Mechanical Engineering at Carnegie Mellon University. He is a recipient of numerous awards including the Alfred P. Sloan Fellowship, Office of Naval Research Young Investigator Award, and National Science Foundation CAREER award. He leads an interdisciplinary group working on advanced batteries for electric vehicles, trucks, and plane

Andrew Westover is a Staff Scientist in the Chemical Sciences Division at Oak Ridge National Laboratory. His research focuses on the development of solid-state batteries with an emphasis on Li-metal anodes, interfaces, and solid-state electrolytes.

Wolfgang G. Zeier is a junior group leader at the Justus-Liebig University Giessen, within the framework of an Emmy-Noether research group. His research interests encompass the fundamental

structure–property relationships in solids, with a focus on thermo-electric and ion-conducting materials, as well as solid–solid interfacial chemistry in all-solid-state batteries.

■ ACKNOWLEDGMENTS

K.B.H. and M.B.D. acknowledge support from National Science Foundation grant No. 1847029. P.P.M. and A.V. acknowledge support from Grant No. 1805656. K.B.H., N.P.D., L.E.M., M.T.M., P.P.M., and V.V. acknowledge support from the Research Corporation for Scientific Advancement (Scialog program in Energy Storage). X.C.C. acknowledges support from U.S. Department of Energy (DOE), Office of Energy Efficiency and Renewable Energy for the Vehicle Technologies Office's Advanced Battery Materials Research Program as well as Office of Sciences, Basic Energy Sciences, Materials Science and Engineering Division. V.V. acknowledges support from the Advanced Research Projects Agency Energy (ARPA-E) under Grant DE-AR0000774. A.S.W. acknowledges support from the ARPA-E IONICS program Award No. DE-AR0000775. N.P.D. acknowledges support from an ECS Toyota Young Investigator Fellowship.

■ REFERENCES

- (1) Wood, K. N.; Noked, M.; Dasgupta, N. P. Lithium metal anodes: toward an improved understanding of coupled morphological, electrochemical, and mechanical behavior. *ACS Energy Letters* **2017**, *2* (3), 664–672.
- (2) Kasemchainan, J.; Zekoll, S.; Spencer Jolly, D.; Ning, Z.; Hartley, G.; Marrow, T.; Bruce, P. Critical stripping current leads to dendrite formation on plating in lithium anode solid electrolyte cells. *Nat. Mater.* **2019**, *18*, 1105–1111.
- (3) Wang, M. J.; Choudhury, R.; Sakamoto, J. Characterizing the Li–Solid–Electrolyte Interface Dynamics as a Function of Stack Pressure and Current Density. *Joule* **2019**, *3*, 2165–2178.
- (4) Albertus, P.; Babinec, S.; Litzelman, S.; Newman, A. Status and challenges in enabling the lithium metal electrode for high-energy and low-cost rechargeable batteries. *Nature Energy* **2018**, *3* (1), 16.
- (5) Betz, J.; Bieker, G.; Meister, P.; Placke, T.; Winter, M.; Schmuck, R. Theoretical versus Practical Energy: A Plea for More Transparency in the Energy Calculation of Different Rechargeable Battery Systems. *Adv. Energy Mater.* **2019**, *9* (6), 1803170.
- (6) Monroe, C.; Newman, J. The effect of interfacial deformation on electrodeposition kinetics. *J. Electrochem. Soc.* **2004**, *151* (6), A880–A886.
- (7) Monroe, C.; Newman, J. The impact of elastic deformation on deposition kinetics at lithium/polymer interfaces. *J. Electrochem. Soc.* **2005**, *152* (2), A396–A404.
- (8) Barai, P.; Higa, K.; Srinivasan, V. Lithium dendrite growth mechanisms in polymer electrolytes and prevention strategies. *Phys. Chem. Chem. Phys.* **2017**, *19* (31), 20493–20505.
- (9) Onsager, L. Theories and problems of liquid diffusion. *Ann. N. Y. Acad. Sci.* **1945**, *46* (5), 241–265.
- (10) Bai, P.; Li, J.; Brushett, F. R.; Bazant, M. Z. Transition of lithium growth mechanisms in liquid electrolytes. *Energy Environ. Sci.* **2016**, *9* (10), 3221–3229.
- (11) Ahmad, Z.; Viswanathan, V. Stability of electrodeposition at solid–solid interfaces and implications for metal anodes. *Phys. Rev. Lett.* **2017**, *119* (5), 056003.
- (12) Li, S.; Mohamed, A. I.; Pande, V.; Wang, H.; Cuthbert, J.; Pan, X.; He, H.; Wang, Z.; Viswanathan, V.; Whitacre, J. F.; et al. Single-ion homopolymer electrolytes with high transference number prepared by click chemistry and photoinduced metal-free atom-transfer radical polymerization. *ACS Energy Letters* **2018**, *3* (1), 20–27.
- (13) Zhang, H.; Li, C.; Piszcz, M.; Coya, E.; Rojo, T.; Rodriguez-Martinez, L. M.; Armand, M.; Zhou, Z. Single lithium-ion conducting solid polymer electrolytes: advances and perspectives. *Chem. Soc. Rev.* **2017**, *46* (3), 797–815.
- (14) Stone, G.; Mullin, S.; Teran, A.; Hallinan, D.; Minor, A.; Hexemer, A.; Balsara, N. Resolution of the modulus versus adhesion dilemma in solid polymer electrolytes for rechargeable lithium metal batteries. *J. Electrochem. Soc.* **2012**, *159* (3), A222–A227.
- (15) Harry, K. J.; Higa, K.; Srinivasan, V.; Balsara, N. P. Influence of electrolyte modulus on the local current density at a dendrite tip on a lithium metal electrode. *J. Electrochem. Soc.* **2016**, *163* (10), A2216–A2224.
- (16) Khurana, R.; Schaefer, J. L.; Archer, L. A.; Coates, G. W. Suppression of lithium dendrite growth using cross-linked polyethylene/poly (ethylene oxide) electrolytes: a new approach for practical lithium-metal polymer batteries. *J. Am. Chem. Soc.* **2014**, *136* (20), 7395–7402.
- (17) Maslyn, J. A.; Loo, W. S.; McEntush, K. D.; Oh, H. J.; Harry, K. J.; Parkinson, D. Y.; Balsara, N. P. Growth of Lithium Dendrites and Globules through a Solid Block Copolymer Electrolyte as a Function of Current Density. *J. Phys. Chem. C* **2018**, *122* (47), 26797–26804.
- (18) Brissot, C.; Rosso, M.; Chazalviel, J.-N.; Lascaud, S. Dendritic growth mechanisms in lithium/polymer cells. *J. Power Sources* **1999**, *81*, 925–929.
- (19) Harry, K. J.; Hallinan, D. T.; Parkinson, D. Y.; MacDowell, A. A.; Balsara, N. P. Detection of subsurface structures underneath dendrites formed on cycled lithium metal electrodes. *Nat. Mater.* **2014**, *13* (1), 69.
- (20) Schausser, N. S.; Harry, K. J.; Parkinson, D. Y.; Watanabe, H.; Balsara, N. P. Lithium dendrite growth in glassy and rubbery nanostructured block copolymer electrolytes. *J. Electrochem. Soc.* **2015**, *162* (3), A398–A405.
- (21) Sharafi, A.; Kazyak, E.; Davis, A. L.; Yu, S.; Thompson, T.; Siegel, D. J.; Dasgupta, N. P.; Sakamoto, J. Surface chemistry mechanism of ultra-low interfacial resistance in the solid-state electrolyte Li₇La₃Zr₂O₁₂. *Chem. Mater.* **2017**, *29* (18), 7961–7968.
- (22) Ansell, R. The chemical and electrochemical stability of beta-alumina. *J. Mater. Sci.* **1986**, *21*, 365–379.
- (23) Sharafi, A.; Meyer, H. M.; Nanda, J.; Wolfenstine, J.; Sakamoto, J. Characterizing the Li–Li₇La₃Zr₂O₁₂ interface stability and kinetics as a function of temperature and current density. *J. Power Sources* **2016**, *302*, 135–139.
- (24) De Jonghe, L. C.; Feldman, L.; Beuchele, A. Slow degradation and electron conduction in sodium/beta-aluminas. *J. Mater. Sci.* **1981**, *16* (3), 780–786.
- (25) De Jonghe, L. C.; Feldman, L.; Beuchele, A. Failure modes of Na-beta alumina. *Solid State Ionics* **1981**, *5*, 267–269.
- (26) Porz, L.; Swamy, T.; Sheldon, B. W.; Rettenwander, D.; Frömling, T.; Thaman, H. L.; Berendts, S.; Uecker, R.; Carter, W. C.; Chiang, Y. M. Mechanism of lithium metal penetration through inorganic solid electrolytes. *Adv. Energy Mater.* **2017**, *7* (20), 1701003.
- (27) Swamy, T.; Park, R.; Sheldon, B. W.; Rettenwander, D.; Porz, L.; Berendts, S.; Uecker, R.; Carter, W. C.; Chiang, Y.-M. Lithium metal penetration induced by electrodeposition through solid electrolytes: example in single-crystal Li₆La₃ZrTaO₁₂ garnet. *J. Electrochem. Soc.* **2018**, *165* (16), A3648–A3655.
- (28) Westover, A. S.; Dudney, N. J.; Sacci, R. L.; Kalnaus, S. Deposition and Confinement of Li Metal along an Artificial Lipon–Lipon Interface. *ACS Energy Letters* **2019**, *4* (3), 651–655.
- (29) Han, F.; Westover, A. S.; Yue, J.; Fan, X.; Wang, F.; Chi, M.; Leonard, D. N.; Dudney, N. J.; Wang, H.; Wang, C. High electronic conductivity as the origin of lithium dendrite formation within solid electrolytes. *Nature Energy* **2019**, *4* (3), 187.
- (30) Krauskopf, T.; Dippel, R.; Hartmann, H.; Peppeler, K.; Mogwitz, B.; Richter, F. H.; Zeier, W. G.; Janek, J. Lithium–Metal Growth Kinetics on LLZO Garnet-Type Solid Electrolytes. *Joule* **2019**, *3* (8), 2030–2049.
- (31) Shen, F.; Dixit, M. B.; Xiao, X.; Hatzell, K. B. Effect of pore connectivity on Li dendrite propagation within LLZO electrolytes observed with synchrotron X-ray tomography. *ACS Energy Letters* **2018**, *3* (4), 1056–1061.

- (32) Mo, F.; Ruan, J.; Sun, S.; Lian, Z.; Yang, S.; Yue, X.; Song, Y.; Zhou, Y. N.; Fang, F.; Sun, G.; et al. Inside or Outside: Origin of Lithium Dendrite Formation of All Solid-State Electrolytes. *Adv. Energy Mater.* **2019**, *9*, 1902123.
- (33) Tian, H.-K.; Xu, B.; Qi, Y. Computational study of lithium nucleation tendency in Li₇La₃Zr₂O₁₂ (LLZO) and rational design of interlayer materials to prevent lithium dendrites. *J. Power Sources* **2018**, *392*, 79–86.
- (34) Tian, H.-K.; Liu, Z.; Ji, Y.; Chen, L.-Q.; Qi, Y. Interfacial Electronic Properties Dictate Li Dendrite Growth in Solid Electrolytes. *Chem. Mater.* **2019**, *31*, 7351.
- (35) LePage, W. S.; Chen, Y.; Kazyak, E.; Chen, K.-H.; Sanchez, A. J.; Poli, A.; Arruda, E. M.; Thouless, M.; Dasgupta, N. P. Lithium Mechanics: Roles of Strain Rate and Temperature and Implications for Lithium Metal Batteries. *J. Electrochem. Soc.* **2019**, *166* (2), A89–A97.
- (36) Masias, A.; Felten, N.; Garcia-Mendez, R.; Wolfenstine, J.; Sakamoto, J. Elastic, plastic, and creep mechanical properties of lithium metal. *J. Mater. Sci.* **2019**, *54* (3), 2585–2600.
- (37) Herbert, E. G.; Hackney, S. A.; Thole, V.; Dudney, N. J.; Phani, P. S. Nanoindentation of high-purity vapor deposited lithium films: A mechanistic rationalization of diffusion-mediated flow. *J. Mater. Res.* **2018**, *33* (10), 1347–1360.
- (38) Herbert, E. G.; Hackney, S. A.; Thole, V.; Dudney, N. J.; Phani, P. S. Nanoindentation of high-purity vapor deposited lithium films: A mechanistic rationalization of the transition from diffusion to dislocation-mediated flow. *J. Mater. Res.* **2018**, *33* (10), 1361–1368.
- (39) Xu, C.; Ahmad, Z.; Aryanfar, A.; Viswanathan, V.; Greer, J. R. Enhanced strength and temperature dependence of mechanical properties of Li at small scales and its implications for Li metal anodes. *Proc. Natl. Acad. Sci. U. S. A.* **2017**, *114* (1), 57–61.
- (40) Luntz, A. C.; Voss, J.; Reuter, K. *Interfacial challenges in solid-state Li ion batteries*; ACS Publications: 2015.
- (41) Lotsch, B. V.; Maier, J. Relevance of solid electrolytes for lithium-based batteries: A realistic view. *J. Electroceram.* **2017**, *38* (2–4), 128–141.
- (42) Zhu, Y.; He, X.; Mo, Y. First principles study on electrochemical and chemical stability of solid electrolyte–electrode interfaces in all-solid-state Li-ion batteries. *J. Mater. Chem. A* **2016**, *4* (9), 3253–3266.
- (43) Nolan, A. M.; Zhu, Y.; He, X.; Bai, Q.; Mo, Y. Computation-accelerated design of materials and interfaces for all-solid-state lithium-ion batteries. *Joule* **2018**, *2* (10), 2016–2046.
- (44) Wenzel, S.; Leichtweiss, T.; Weber, D. A.; Sann, J.; Zeier, W. G.; Janek, J. r. Interfacial reactivity benchmarking of the sodium ion conductors Na₃PS₄ and sodium β -alumina for protected sodium metal anodes and sodium all-solid-state batteries. *ACS Appl. Mater. Interfaces* **2016**, *8* (41), 28216–28224.
- (45) Wu, F.; Fitzhugh, W.; Ye, L.; Ning, J.; Li, X. Advanced sulfide solid electrolyte by core-shell structural design. *Nat. Commun.* **2018**, *9* (1), 4037.
- (46) Lewis, J. A.; Tippens, J.; Cortes, F. J. Q.; McDowell, M. T. Chemo-Mechanical Challenges in Solid-State Batteries. *Trends in Chemistry* **2019**, *1* (9), 845–857.
- (47) Tippens, J.; Miers, J.; Afshar, A.; Lewis, J.; Cortes, F. J. Q.; Qiao, H.; Marchese, T. S.; Di Leo, C. V.; Saldana, C.; McDowell, M. T. Visualizing Chemo-Mechanical Degradation of a Solid-State Battery Electrolyte. *ACS Energy Letters* **2019**, *4* (6), 1475–1483.
- (48) Chae, S.; Choi, S.-H.; Kim, N.; Sung, J.; Cho, J. Integration of Graphite and Silicon Anodes for the Commercialization of High-Energy Lithium-Ion Batteries. *Angew. Chem., Int. Ed.* **2020**, *59* (1), 110–135.
- (49) Krauskopf, T.; Mogwitz, B.; Rosenbach, C.; Zeier, W. G.; Janek, J. Diffusion Limitation of Lithium Metal and Li–Mg Alloy Anodes on LLZO Type Solid Electrolytes as a Function of Temperature and Pressure. *Adv. Energy Mater.* **2019**, *9*, 1902568.
- (50) Santhosha, A.; Medenbach, L.; Buchheim, J. R.; Adelhelm, P. The Indium–Lithium Electrode in Solid-State Lithium-Ion Batteries: Phase Formation, Redox Potentials, and Interface Stability. *Batteries & Supercaps* **2019**, *2* (6), 524–529.
- (51) Han, X.; Gong, Y.; Fu, K. K.; He, X.; Hitz, G. T.; Dai, J.; Pearce, A.; Liu, B.; Wang, H.; Rubloff, G.; et al. Negating interfacial impedance in garnet-based solid-state Li metal batteries. *Nat. Mater.* **2017**, *16* (5), 572.
- (52) Cheng, Q.; Li, A.; Li, N.; Li, S.; Zangiabadi, A.; Huang, W.; Li, A. C.; Jin, T.; Song, Q.; Xu, W.; et al. Stabilizing Solid Electrolyte–Anode Interface in Li-Metal Batteries by Boron Nitride-Based Nanocomposite Coating. *Joule* **2019**, *3* (6), 1510–1522.
- (53) Zheng, G.; Wang, C.; Pei, A.; Lopez, J.; Shi, F.; Chen, Z.; Sendek, A. D.; Lee, H.-W.; Lu, Z.; Schneider, H.; Safont-Sempere, M. M.; Chu, S.; Bao, Z.; Cui, Y. High-Performance Lithium Metal Negative Electrode with a Soft and Flowable Polymer Coating. *ACS Energy Letters* **2016**, *1* (6), 1247–1255.
- (54) Li, N.-W.; Shi, Y.; Yin, Y.-X.; Zeng, X.-X.; Li, J.-Y.; Li, C.-J.; Wan, L.-J.; Wen, R.; Guo, Y.-G. A Flexible Solid Electrolyte Interphase Layer for Long-Life Lithium Metal Anodes. *Angew. Chem., Int. Ed.* **2018**, *57* (6), 1505–1509.
- (55) Gao, Y.; Zhao, Y.; Li, Y. C.; Huang, Q.; Mallouk, T. E.; Wang, D. Interfacial Chemistry Regulation via a Skin-Grafting Strategy Enables High-Performance Lithium-Metal Batteries. *J. Am. Chem. Soc.* **2017**, *139* (43), 15288–15291.
- (56) Zhou, W.; Wang, S.; Li, Y.; Xin, S.; Manthiram, A.; Goodenough, J. B. Plating a Dendrite-Free Lithium Anode with a Polymer/Ceramic/Polymer Sandwich Electrolyte. *J. Am. Chem. Soc.* **2016**, *138* (30), 9385–9388.
- (57) Ates, T.; Keller, M.; Kulisch, J.; Adermann, T.; Passerini, S. Development of an all-solid-state lithium battery by slurry-coating procedures using a sulfidic electrolyte. *Energy Storage Materials* **2019**, *17*, 204–210.
- (58) Wang, X.; Zeng, W.; Hong, L.; Xu, W.; Yang, H.; Wang, F.; Duan, H.; Tang, M.; Jiang, H. Stress-driven lithium dendrite growth mechanism and dendrite mitigation by electroplating on soft substrates. *Nature Energy* **2018**, *3* (3), 227–235.
- (59) Chen, L.; Li, Y.; Li, S.-P.; Fan, L.-Z.; Nan, C.-W.; Goodenough, J. B. PEO/garnet composite electrolytes for solid-state lithium batteries: From “ceramic-in-polymer” to “polymer-in-ceramic”. *Nano Energy* **2018**, *46*, 176–184.
- (60) Liu, W.; Lee, S. W.; Lin, D.; Shi, F.; Wang, S.; Sendek, A. D.; Cui, Y. Enhancing ionic conductivity in composite polymer electrolytes with well-aligned ceramic nanowires. *Nature Energy* **2017**, *2*, 17035.
- (61) Dixit, M. B.; Zaman, W.; Hortance, N.; Vujic, S.; Harkey, B.; Shen, F.; Tsai, W.-Y.; De Andrade, V.; Chen, X. C.; Balke, N. Nanoscale Mapping of Extrinsic Interfaces in Hybrid Solid Electrolytes. *Joule* **2020**, *4*, 207.
- (62) Pandian, A. S.; Chen, X. C.; Chen, J.; Lokitz, B. S.; Ruther, R. E.; Yang, G.; Lou, K.; Nanda, J.; Delnick, F. M.; Dudney, N. J. Facile and scalable fabrication of polymer-ceramic composite electrolyte with high ceramic loadings. *J. Power Sources* **2018**, *390*, 153–164.
- (63) Dixit, M. B.; Zaman, W.; Bootwala, Y.; Zheng, Y.; Hatzell, M. C.; Hatzell, K. B. Scalable manufacturing of hybrid solid electrolytes with interface control. *ACS Appl. Mater. Interfaces* **2019**, *11* (48), 45087–45097.
- (64) Zaman, W.; Hortance, N. M.; Dixit, M. B.; De Andrade, V.; Hatzell, K. B. Visualizing percolation and ion transport in hybrid solid electrolytes for Li-metal batteries. *J. Mater. Chem. A* **2019**, *7*, 23914–23921.
- (65) Chen, X. C.; Sacci, R. L.; Osti, N. C.; Tyagi, M.; Wang, Y.; Palmer, M. J.; Dudney, N. J. Study of segmental dynamics and ion transport in polymer–ceramic composite electrolytes by quasi-elastic neutron scattering. *Molecular Systems Design & Engineering* **2019**, *4* (2), 379–385.
- (66) Chen, X. C.; Liu, X.; Samuthira Pandian, A.; Lou, K.; Delnick, F. M.; Dudney, N. J. Determining and Minimizing Resistance for Ion Transport at the Polymer/Ceramic Electrolyte Interface. *ACS Energy Letters* **2019**, *4*, 1080–1085.

- (67) Liu, Y.; Li, C.; Li, B.; Song, H.; Cheng, Z.; Chen, M.; He, P.; Zhou, H. Germanium Thin Film Protected Lithium Aluminum Germanium Phosphate for Solid-State Li Batteries. *Adv. Energy Mater.* **2018**, *8* (16), 1702374.
- (68) Hao, X.; Zhao, Q.; Su, S.; Zhang, S.; Ma, J.; Shen, L.; Yu, Q.; Zhao, L.; Liu, Y.; Kang, F.; et al. Constructing Multifunctional Interphase between Li₁.4Al₀.4Ti₁.6(PO₄)₃ and Li Metal by Magnetron Sputtering for Highly Stable Solid-State Lithium Metal Batteries. *Adv. Energy Mater.* **2019**, *9* (34), 1901604.
- (69) Ding, X.; Liu, J.; Harris, T. A. A review of the operating limits in slot die coating processes. *AIChE J.* **2016**, *62* (7), 2508–2524.
- (70) Parsekian, A.; Harris, T. A. Scalable, Alternating Narrow Stripes of PVA Support and Unmodified PEDOT: PSS with Maintained Conductivity using a Single-Step Slot Die Coating Approach. *ACS Appl. Mater. Interfaces* **2020**, *12*, 3736.
- (71) Parsekian, A. W.; Jeong, T.-J.; Harris, T. A. A process model for slot coating of narrow stripes. *Journal of Coatings Technology and Research* **2019**, *16*, 1653.
- (72) Gupta, A.; Kazyak, E.; Craig, N.; Christensen, J.; Dasgupta, N.; Sakamoto, J. Evaluating the Effects of Temperature and Pressure on Li/PEO-LiTFSI Interfacial Stability and Kinetics. *J. Electrochem. Soc.* **2018**, *165* (11), A2801–A2806.
- (73) Koerver, R.; Zhang, W.; de Biasi, L.; Schweidler, S.; Kondrakov, A. O.; Kolling, S.; Brezesinski, T.; Hartmann, P.; Zeier, W. G.; Janek, J. Chemo-mechanical expansion of lithium electrode materials—on the route to mechanically optimized all-solid-state batteries. *Energy Environ. Sci.* **2018**, *11* (8), 2142–2158.
- (74) Krauskopf, T.; Hartmann, H.; Zeier, W. G.; Janek, J. r. Toward a Fundamental Understanding of the Lithium Metal Anode in Solid-State Batteries—An Electrochemo-Mechanical Study on the Garnet-Type Solid Electrolyte Li₆.25Al₀.25La₃Zr₂O₁₂. *ACS Appl. Mater. Interfaces* **2019**, *11* (15), 14463–14477.
- (75) Aguesse, F.; Manalastas, W.; Buannic, L.; Lopez del Amo, J. M.; Singh, G.; Llordés, A.; Kilner, J. Investigating the dendritic growth during full cell cycling of garnet electrolyte in direct contact with Li metal. *ACS Appl. Mater. Interfaces* **2017**, *9* (4), 3808–3816.
- (76) Ma, C.; Cheng, Y.; Yin, K.; Luo, J.; Sharafi, A.; Sakamoto, J.; Li, J.; More, K. L.; Dudney, N. J.; Chi, M. Interfacial stability of Li metal–solid electrolyte elucidated via in situ electron microscopy. *Nano Lett.* **2016**, *16* (11), 7030–7036.
- (77) Marbella, L. E.; Zekoll, S.; Kasemchainan, J.; Emge, S. P.; Bruce, P. G.; Grey, C. P. ⁷Li NMR Chemical Shift Imaging To Detect Microstructural Growth of Lithium in All-Solid-State Batteries. *Chem. Mater.* **2019**, *31* (8), 2762–2769.
- (78) Wang, C.; Gong, Y.; Dai, J.; Zhang, L.; Xie, H.; Pastel, G.; Liu, B.; Wachsman, E.; Wang, H.; Hu, L. In situ neutron depth profiling of lithium metal–garnet interfaces for solid state batteries. *J. Am. Chem. Soc.* **2017**, *139* (40), 14257–14264.
- (79) Ren, Y.; Shen, Y.; Lin, Y.; Nan, C.-W. Direct observation of lithium dendrites inside garnet-type lithium-ion solid electrolyte. *Electrochem. Commun.* **2015**, *57*, 27–30.
- (80) Cheng, E. J.; Sharafi, A.; Sakamoto, J. Intergranular Li metal propagation through polycrystalline Li₆.25Al₀.25La₃Zr₂O₁₂ ceramic electrolyte. *Electrochim. Acta* **2017**, *223*, 85–91.
- (81) Lewis, J. A.; Cortes, F. J. Q.; Boebinger, M. G.; Tippens, J.; Marchese, T. S.; Kondekar, N.; Liu, X.; Chi, M.; McDowell, M. T. Interphase morphology between a solid-state electrolyte and lithium controls cell failure. *ACS Energy Letters* **2019**, *4* (2), 591–599.
- (82) Wenzel, S.; Leichtweiss, T.; Krüger, D.; Sann, J.; Janek, J. Interphase formation on lithium solid electrolytes—An in situ approach to study interfacial reactions by photoelectron spectroscopy. *Solid State Ionics* **2015**, *278*, 98–105.
- (83) Hartmann, P.; Leichtweiss, T.; Busche, M. R.; Schneider, M.; Reich, M.; Sann, J.; Adelhelm, P.; Janek, J. r. Degradation of NASICON-type materials in contact with lithium metal: formation of mixed conducting interphases (MCI) on solid electrolytes. *J. Phys. Chem. C* **2013**, *117* (41), 21064–21074.
- (84) Wood, K. N.; Steirer, K. X.; Hafner, S. E.; Ban, C.; Santhanagopalan, S.; Lee, S.-H.; Teeter, G. Operando X-ray photoelectron spectroscopy of solid electrolyte interphase formation and evolution in Li₂SP₂S₅ solid-state electrolytes. *Nat. Commun.* **2018**, *9* (1), 2490.
- (85) Groh, M. F.; Sullivan, M. J.; Gaultois, M. W.; Pecher, O.; Griffith, K. J.; Grey, C. P. Interface Instability in LiFePO₄–Li₃₊ x P_{1–x} Si_x O₄ All-Solid-State Batteries. *Chem. Mater.* **2018**, *30* (17), 5886–5895.
- (86) Yu, C.; Ganapathy, S.; de Klerk, N. J.; Roslon, I.; van Eck, E. R.; Kentgens, A. P.; Wagemaker, M. Unravelling li-ion transport from picoseconds to seconds: bulk versus interfaces in an argyrodite li₆ps₅cl–li₂s all-solid-state li-ion battery. *J. Am. Chem. Soc.* **2016**, *138* (35), 11192–11201.
- (87) Trease, N. M.; Zhou, L.; Chang, H. J.; Zhu, B. Y.; Grey, C. P. In situ NMR of lithium ion batteries: Bulk susceptibility effects and practical considerations. *Solid State Nucl. Magn. Reson.* **2012**, *42*, 62–70.
- (88) Chien, P.-H.; Feng, X.; Tang, M.; Rosenberg, J. T.; O'Neill, S.; Zheng, J.; Grant, S. C.; Hu, Y.-Y. Li Distribution heterogeneity in solid electrolyte Li₁₀GeP₂S₁₂ upon electrochemical cycling probed by ⁷Li MRI. *J. Phys. Chem. Lett.* **2018**, *9* (8), 1990–1998.
- (89) Pecher, O.; Carretero-González, J.; Griffith, K. J.; Grey, C. P. Materials' methods: NMR in battery research. *Chem. Mater.* **2017**, *29* (1), 213–242.
- (90) Bottke, P.; Rettenwander, D.; Schmidt, W.; Amthauer, G.; Wilkening, M. Ion dynamics in solid electrolytes: NMR reveals the elementary steps of Li⁺ hopping in the garnet Li₆.5La₃Zr₁.75Mo₀.25O₁₂. *Chem. Mater.* **2015**, *27* (19), 6571–6582.
- (91) Zhang, X.; Shyy, W.; Marie Sastry, A. Numerical simulation of intercalation-induced stress in Li-ion battery electrode particles. *J. Electrochem. Soc.* **2007**, *154* (10), A910–A916.
- (92) Fu, C.; Venturi, V.; Ahmad, Z.; Ells, A. W.; Viswanathan, V.; Helms, B. A. Universal Chemomechanical Design Rules for Solid-Ion Conductors to Prevent Dendrite Formation in Lithium Metal Batteries. *arXiv* **2019**, 1901.04910.
- (93) Barai, P.; Higa, K.; Srinivasan, V. Impact of external pressure and electrolyte transport properties on lithium dendrite growth. *J. Electrochem. Soc.* **2018**, *165* (11), A2654–A2666.
- (94) Ahmad, Z.; Viswanathan, V. Role of anisotropy in determining stability of electrodeposition at solid-solid interfaces. *Physical Review Materials* **2017**, *1* (5), 055403.
- (95) Barai, P.; Higa, K.; Ngo, A. T.; Curtiss, L. A.; Srinivasan, V. Mechanical Stress Induced Current Focusing and Fracture in Grain Boundaries. *J. Electrochem. Soc.* **2019**, *166* (10), A1752–A1762.
- (96) Raj, R.; Wolfenstine, J. Current limit diagrams for dendrite formation in solid-state electrolytes for Li-ion batteries. *J. Power Sources* **2017**, *343*, 119–126.
- (97) Yu, S.; Siegel, D. J. Grain boundary softening: A potential mechanism for lithium metal penetration through stiff solid electrolytes. *ACS Appl. Mater. Interfaces* **2018**, *10* (44), 38151–38158.
- (98) Dixit, M. B.; Regala, M.; Shen, F.; Xiao, X.; Hatzell, K. B. Tortuosity Effects in Garnet-Type Li₇La₃Zr₂O₁₂ Solid Electrolytes. *ACS Appl. Mater. Interfaces* **2019**, *11* (2), 2022–2030.
- (99) Sendek, A. D.; Yang, Q.; Cubuk, E. D.; Duerloo, K.-A. N.; Cui, Y.; Reed, E. J. Holistic computational structure screening of more than 12000 candidates for solid lithium-ion conductor materials. *Energy Environ. Sci.* **2017**, *10* (1), 306–320.
- (100) Dawson, J. A.; Canepa, P.; Famprikis, T.; Masquelier, C.; Islam, M. S. Atomic-scale influence of grain boundaries on Li-ion conduction in solid electrolytes for all-solid-state batteries. *J. Am. Chem. Soc.* **2018**, *140* (1), 362–368.
- (101) Mathiesen, N. R.; Yang, S.; García-Lastra, J. M.; Vegge, T.; Siegel, D. J. Charge Transport in Alkali-Metal Superoxides: A Systematic First-Principles Study. *Chem. Mater.* **2019**, *31* (21), 9156–9167.
- (102) Hao, F.; Mukherjee, P. P. Mesoscale Analysis of the Electrolyte-Electrode Interface in All-Solid-State Li-Ion Batteries. *J. Electrochem. Soc.* **2018**, *165* (9), A1857–A1864.

- (103) Hao, F.; Wang, W.; Mukherjee, P. P. Electrochemical-Reaction-Driven Interfacial Stress in a Solid-Solid Layered Architecture. *Phys. Rev. Appl.* **2019**, *11* (3), 034038.
- (104) Mistry, A.; Fear, C.; Carter, R.; Love, C. T.; Mukherjee, P. P. Electrolyte confinement alters lithium electrodeposition. *ACS Energy Letters* **2019**, *4* (1), 156–162.
- (105) Mistry, A.; Mukherjee, P. P. Molar Volume Mismatch: A Malefactor for Irregular Metallic Electrodeposition in Solid Electrolytes. *ChemRxiv* **2019**, DOI: [10.26434/chemrxiv.10266635.v1](https://doi.org/10.26434/chemrxiv.10266635.v1).