

Catalysis by the Non-Heme Iron(II) Histone Demethylase PHF8 Involves Iron Center Rearrangement and Conformational Modulation of Substrate Orientation

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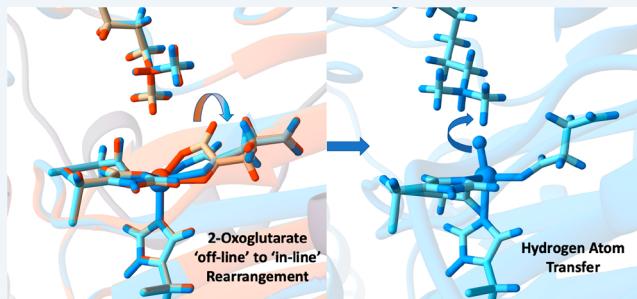
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Supporting Information

ABSTRACT: PHF8 (KDM7B) is a human non-heme 2-oxoglutarate (2OG) JmjC domain oxygenase that catalyzes the demethylation of the di/mono- N^e -methylated K9 residue of histone H3. Altered PHF8 activity is linked to genetic diseases and cancer; thus, it is an interesting target for epigenetic modulation. We describe the use of combined quantum mechanics/molecular mechanics (QM/MM) and molecular dynamics (MD) simulations to explore the mechanism of PHF8, including dioxygen activation, 2OG binding modes, and substrate demethylation steps. A PHF8 crystal structure manifests the 2OG C-1 carboxylate bound to iron in a nonproductive orientation, i.e., *trans* to His247. A ferryl–oxo intermediate formed by activating dioxygen bound to the vacant site in this complex would be nonproductive, i.e., “off-line” with respect to reaction with N^e -methylated K9. We show rearrangement of the “off-line” ferryl–oxo intermediate to a productive “in-line” geometry via a solvent exchange reaction (called “ferryl-flip”) is energetically unfavorable. The calculations imply that movement of the 2OG C-1 carboxylate prior to dioxygen binding at a five-coordination stage in catalysis proceeds with a low barrier, suggesting that two possible 2OG C-1 carboxylate geometries can coexist at room temperature. We explored alternative mechanisms for hydrogen atom transfer and show that second sphere interactions orient the N^e -methylated lysine in a conformation where hydrogen abstraction from a methyl C–H bond is energetically more favorable than hydrogen abstraction from the N–H bond of the protonated N^e -methyl group. Using multiple HAT reaction path calculations, we demonstrate the crucial role of conformational flexibility in effective hydrogen transfer. Subsequent hydroxylation occurs through a rebound mechanism, which is energetically preferred compared to desaturation, due to second sphere interactions. The overall mechanistic insights reveal the crucial role of iron-center rearrangement, second sphere interactions, and conformational flexibility in PHF8 catalysis and provide knowledge useful for the design of mechanism-based PHF8 inhibitors.



KEYWORDS: histone lysine demethylation, 2-oxoglutarate, non-heme iron enzyme, QM/MM, metal-center rearrangement, enzyme mechanism

1. INTRODUCTION

Genetic information in eukaryotic cells is stored in chromatin, which packs DNA and histone proteins in an organized and compact structure.¹ The basic unit of chromatin, the nucleosome, is usually built from a histone octamer containing two copies each of H2A, H2B, H3, and H4 histones, around which about two turns of DNA are wrapped to form an octamer.¹ Unstructured N-terminal histone tails extend from the nucleosome and undergo post-translational modifications (PTMs), including methylation, acetylation, and phosphorylation.^{2,3} PTMs of histones affect chromatin structure and function with consequences for compaction, transcription, DNA replication, and DNA damage repair.^{2,3} Histone N-

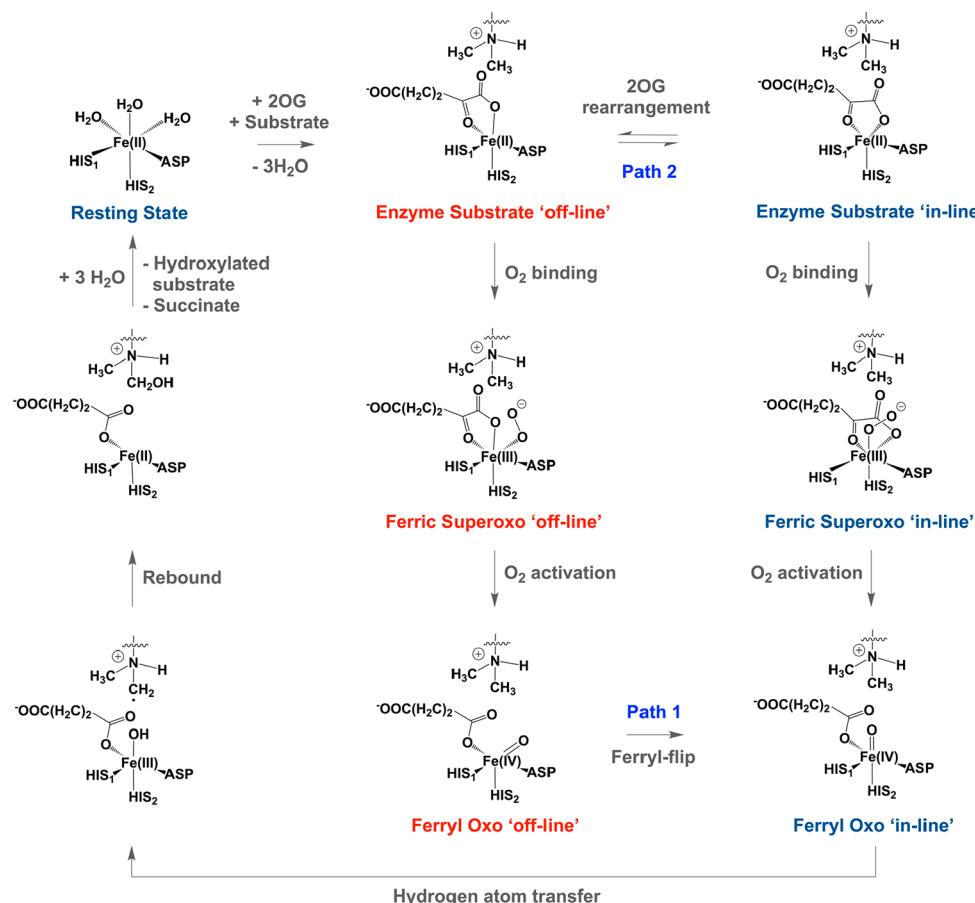
methylation occurs on arginine- and lysine-residues and affects a wide range of biological processes. Misregulation of N -methylation is linked with diseases including cancer.^{4–6} Histone lysines can be mono-, di- and trimethylated (me1, me2, and me3);⁷ changes in methylation status do not significantly change the charge on the N^e -nitrogen. Depending on its position, lysine N^e -methylation can be either transcriptionally repressing or activating, e.g., methylation of K9 or K27 of histone H3 can activate transcription, whereas methylation

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Scheme 1. Proposed Mechanism for the Catalytic Cycle of 2OG Oxygenases Involving either a Ferryl-Flip after the Formation of the “Off-Line” Ferryl–Oxo Intermediate (Path 1) or a Change in the 2OG C-1 Carboxylate Binding Mode from “Off-Line” to “In-Line” Prior to O_2 Binding (Path 2)



of K4 and K36 of the same histone protein can repress transcription.⁸ Histone methylation was once widely thought to be irreversible, but this view was reversed by the relatively recent identification of histone demethylases (KDMs).^{9,10}

There are two families of histone KDMs: (i) the flavin adenine dinucleotide (FAD)-dependent enzymes (KDM1s) and (ii) the Jumonji C (JmjC) domain-containing enzymes (KDM2-7 subfamilies in humans). The JmjC proteins are Fe (II) and 2-oxoglutarate- (2OG-) dependent dioxygenases that couple substrate oxidation to decarboxylation of 2OG to form succinate, CO₂, and a ferryl–oxo intermediate (Scheme 1). The latter performs hydrogen abstraction from the substrate, followed by hydroxylation and nonenzymatic formation of formaldehyde.^{11,12} The catalytic JmjC domain has a conserved distorted double-stranded β -helix (DSBH) fold, which has a funnel-like structure formed from conserved antiparallel β -sheets and additional elements, where one end of DSBH harbors the active site including the iron center.¹³ The 2OG cosubstrate chelates to Fe(II) via its C1 carboxylate and C2 keto groups.^{11,12} Three conserved protein residues (His₁-X-Asp/Glu-X-His₂ triad), present in most 2OG oxygenases, bind to one face of the octahedrally coordinated metal ion. Most characterized 2OG oxygenases employ a conserved mechanism involving sequential binding of 2OG, substrate, then O₂.^{11,12,14,15} 2OG oxygenases play vital roles in the industrial fermentation of vitally important drugs, such as the penicillins and cephalosporins, and drug precursors.^{8a} Their engineering

to directly ferment commercially useful products is of interest.^{8a} Hence, there is considerable interest in understanding their mechanisms from a biocatalytic perspective.

PHF8 (KDM7B) catalyzes the demethylation of the dimethylated or monomethylated lysine 9 residue in histone H3 (H3K9me2/me1), in a manner actively promoted by binding of its N-terminal plant homeodomain (PHD) to H3K4me3.^{16,17} The PHF8 gene is overexpressed in many cancers including prostate cancer,^{18,19} breast cancer,²⁰ laryngeal cancer,²¹ hypopharyngeal cancer,²¹ nonsmall cell lung cancer,²² and esophageal cancer.²³ Cancer cells in which PHF8 is depleted manifest reduced cell growth^{19,22,23} and migration^{18,23} and increased apoptosis.^{18,19,22} Substitution of phenylalanine 279 to serine in PHF8 (F279S) is linked with X-linked mental retardation.²⁴⁻²⁷ Thus, PHF8 is a current medicinal chemistry target. Crystallographic analysis of PHF8 in complex with a H3K4me3K9me2 fragment^{16,17} informs on interactions between the substrate and the JmjC and PHD domains. A PHD-JmjC linker is responsible for the productive orientation of the JmjC domain to bind H3K9me2/me1 productively, as shown by experimental and computational studies.^{17,28}

Similarly to other 2OG dependent hydroxylases,^{11,12} the proposed reaction mechanism for JmjC KDMs involves binding of O₂ to Fe(II)-2OG-substrate complex.^{11,12} The subsequently formed ferric–superoxide radical complex (Fe(III)-OO[·]) reacts with C2 of 2OG, leading to decarboxylation

and formation of succinate and the ferryl–oxo ($\text{Fe(IV)}=\text{O}$) species (Scheme 1). The reactive ferryl–oxo complex formed can then abstract a hydrogen atom from the lysine N^{e} -methyl group and form a ferric–hydroxyl species which performs rebound hydroxylation to form a hemiaminal,^{11,12} which, at least in studied cases, rapidly fragments to give formaldehyde and the demethylated lysine-residue.^{11,12,29}

Depending on the particular protein environment in 2OG oxygenases, the 2OG C1 carboxylate can adopt two different active site Fe-coordination modes.¹² In one mode, the 2OG C1 carboxylate ligates *trans* to His₁, leaving the position *trans* to His₂ for O₂ binding (water may need to be displaced). Binding of O₂ in this “in-line” mode would lead to the reactive ferryl–oxo intermediate being adjacent to the substrate C–H bond (Scheme 1).^{12,30} Evidence for this binding mode comes from crystallographic studies on clavaminic acid synthase (CAS),³⁰ taurine dioxygenase (TauD),³¹ factor inhibiting hypoxia (FIH),³² and other 2OG oxygenases. In the second, “off-line” mode, the 2OG C1 carboxylate binds *trans* to His₂, leaving the position *trans* to His₁ for O₂ binding, which is approximately perpendicular to the substrate, as observed in crystallographic studies on carbapenem synthase (CarC),³³ anthocyanidin synthase (ANS),³⁴ alkylated base repair protein (AlkB),³⁵ and some other 2OG oxygenases. Thus, based on the coordination position of the 2OG C1 carboxylate, two possible pathways consistent with crystallographic and kinetic studies have been proposed for the formation of the ferryl–oxo complex³⁰ (Scheme 1): (i) The 2OG C1 carboxylate binds *trans* to His₂ in “off-line” binding mode leading to O₂ binding *trans* to His₁. In order for the “off-line” mode to produce a catalytically productive ferryl–oxo intermediate, it has been proposed that in some cases the initially formed ferryl intermediate may “flip”, potentially via oxygen atom exchange with a solvent water molecule, to a position adjacent to the substrate (path 1, Scheme 1).³⁰ Experimental and computational studies for a ferryl-flip via oxygen atom exchange using a water molecule have been reported for natural as well as synthetic non-heme iron enzymes.^{36–40} A ferryl-flip mechanism involving hydration is also consistent with the less than stoichiometric incorporation of an oxygen atom from dioxygen into hydroxylated products in the case of some, but not all, 2OG dependent hydroxylases (by contrast there are consistently high levels of incorporation of single oxygen from dioxygen into succinate).⁴¹ (ii) Alternatively, in the presence of substrate when the water molecule from the Fe-center is displaced, for O₂ binding, the 2OG C1 carboxylate can rearrange to the position *trans* to His₁ and O₂ binds in the position *trans* to His₂. As the O₂ binding and subsequent ferryl–oxo intermediate forming are *trans* to His₂, i.e., “in-line” geometry, no ferryl-flip is required, and the mechanism is simpler.³⁰ Consistent with the second mechanism, both “in-line” and “off-line” binding modes have been observed for PHF8. A PHF8_{1–447}·Fe(II)·H3_{1–14}K4me3K9me2·NOG (NOG, *N*-oxalylglycine is a close 2OG analog) crystal structure manifests the NOG C1 carboxylate *trans* to His₂ (His247), i.e. in an “off-line” binding mode,¹⁷ while a PHF8_{86–447}·Fe·2OG crystal structure shows the 2OG C1 carboxylate *trans* to His₁ (His319), i.e., in an “in-line” binding mode.⁴² Also, a crystal structure of PHF8 with a Fe-chelating inhibitor daminozide shows an “in-line” binding mode.⁴³

Despite the biophysical insights on PHF8, there is a lack of knowledge on its demethylation mechanism, including with respect to the effects of conformational dynamics on catalysis

and the different possible binding modes of 2OG. Alternative mechanistic proposals for the hydrogen abstraction step (HAT), which is often rate-limiting in 2OG oxygenase catalysis, have been proposed based on model compounds;⁴⁴ however, these have not been tested within a protein environment. While the 2OG iron–O₂ derived intermediates (e.g., a superoxo complex) has not been trapped so far, studies with other iron–oxygen intermediates indicate they might have a sufficiently long lifetime to cause conformational changes in proteins. The enzyme ferric–superoxo intermediate has been characterized by crystallographic studies with both non-heme⁴⁵ and heme enzymes, suggesting it has a sufficiently long lifetime to enable conformational changes.⁴⁶ The mean lifetime of a ferryl–oxo intermediate has been measured to be 2.1 s in another 2OG dependent oxygenase, TauD,⁴⁷ hence, it is viable to propose it might cause potential conformational changes in the protein. To explore such conformational changes, we performed 1 μ s molecular dynamics (MD) simulations of PHF8 ferric–superoxo and ferryl–oxo intermediates in PHF8 followed by QM/MM reaction path calculations. The results provide insight into the effects of different 2OG binding modes and the role of the protein environment in promoting demethylation performed by PHF8.

2. METHODS

System Preparation. A crystal structure of the PHF8_{1–447} in complex with H3_{1–14} (Lys4-trimethylated (H3K4me3) and Lys9-dimethylated (H3K9me2)) and the 2OG analog *N*-oxalylglycine was used for modeling (PDB ID: 3KV4).¹⁷ Missing residues (1–2, 65–77, including the linker region between the PHD and the JmjC domains) were modeled using Modeler.⁴⁸ The crystal structure manifests Fe-center in “off-line” mode bidentate coordination with 2OG (Scheme 1), and monodentate coordination with two histidines, one aspartic acid, and one water molecule.¹⁷ To generate parameters for (a) the ferric–superoxo complex, the water molecule coordinated to Fe in the crystal structure (PDB ID: 3KV4) was replaced with O₂ using GaussView 5.0.⁴⁹ The 2OG analog *N*-oxalylglycine was modeled to 2OG by replacing its NH with a methylene group using GaussView 5.0.⁴⁹ Hydrogen atoms were added to 2OG using the leap module of AMBER 16.⁵⁰ The protonation states of ionizable side chains of the proteins were assessed using Amber routines.⁵¹ The Amber topology for 2OG, Lys4-trimethylated (H3K4me3), and Lys9-demethylated (H3K9me2) were developed using the GAFF tool in Antechamber.⁵² The Amber parameters for the Fe–superoxo complex active site containing iron (Fe(III) high spin (HS) $S = 2$, $M = 5$) and the coordinating ligands (2OG, bidentate; O₂, bonded in end-on fashion; H247, H319, and D249, all monodentate) were prepared using the Metal Centre Parameter Builder (MCPB) using MCPB.py v3.0.⁵³ Bond and angle force constants were derived using the Seminario method, and the point charge parameters for the electrostatic potentials were obtained by RESP charge fitting using the ChgModB method at the B3LYP/6-31G* level of theory. The description of the zinc ion and its coordinating ligands in the PHD domain were described using the Zinc Amber force field (ZAFF) method.⁵⁴ The remainder of the protein was modeled using the Amber FF14SB force field.⁵⁵

The same procedure was used to generate parameters for (b) the ferryl–oxo complex with H3_{1–14}K4me3K9me2 histone substrate and succinate. Spectroscopic studies of another 2OG oxygenase (SyrB2) implies the ferryl–oxo intermediate is a

five-coordinate complex;⁵⁶ thus, the ferryl–oxo complex of PHF8 (Fe(IV)=O high spin $S = 2$, $M = 5$) was modeled to be five-coordinate with monodentate succinate, two histidines (H247 and H319), and an aspartate (D249). Previous molecular dynamics simulations run using these parameter generation procedures have successfully reproduced the geometry of non-heme Fe complexes.^{28,57}

MD Simulations. The Leap module in AMBER 16 was used to add counterions (Na^+) for the neutralization of the protein systems. The system was immersed into a rectangular box with TIP3Pwater⁵⁸ molecules present up to a minimum of 10 Å from the farthest protein boundary. Periodic boundary conditions were employed in all simulations. Long-range electrostatic interactions were calculated using the particle mesh Ewald (PME) method⁵⁹ with a direct space and vdw cutoff of 10 Å.

With a restraint of 100 kcal mol⁻¹ Å² on solute molecules, the systems were subjected to energy minimization first using the steepest descent (5000 steps) followed by the conjugate gradient (5000 steps). This was followed by the full minimization of the entire system with both steepest descent (5000 steps) and a conjugate gradient (5000 steps) to relax the system. The systems were then subjected to controlled heating from 0 to 300 K at constant volume using Langevin thermostat⁶⁰ with a collision frequency of 1 ps⁻¹ using a canonical ensemble (NVT) MD simulation for 250 ps. The solute molecules were restrained using a harmonic potential of 50 kcal mol⁻¹ Å² during the heating process. The SHAKE algorithm⁶¹ was used to constrain bonds involving hydrogen. For ferryl–oxo, an initial MD was performed with harmonic restraint between substrate H3K9me2 and the ferryl–oxo for 50 ns. All systems were then equilibrated at 300 K in an NPT ensemble without restraints for 50 ns; the pressure was maintained at 1 bar using the Berendsen barostat.⁶² A production MD run from equilibrated structures with explicit solvent for 1 μ s each with a time step of 2 fs was performed in NPT ensemble with a target pressure set at 1 bar and constant pressure coupling of 2 ps. All productive MD simulations were performed with the GPU version of the AMBER 16 package.⁵⁰ Hydrogen bond analysis was done using CPPTRAJ⁶³ with a cutoff of 3 Å between donor to the acceptor and 135° angle between donor-hydrogen-acceptor unless otherwise mentioned. The binding energy values were calculated using MMGBSA method⁶⁴ by taking 10000 frames in account from the respective 1 μ s MD simulation. Dynamic cross correlation analysis (DCCA) and principal component analysis (PCA) was performed on C- α atoms of PHF8 using trajectory from 500 to 1000 ns with Bio3D.⁶⁵

QM/MM Calculations. Equilibrated structures with productive distance (e.g., distal oxygen and C2 atom of 2OG for O_2 activation) and an angle (e.g., $\angle \text{Fe}-\text{O}-\text{H}$ for hydrogen atom transfer) were taken from MD simulations and used for the QM/MM calculations. Excess water molecules were truncated using CPPTRAJ,⁶³ such that protein has a water solvation layer of a maximum 12 Å surrounding it. The resulting total size of the systems was ~30000 atoms. QM/MM calculations were performed using ChemShell software,⁶⁶ with a combination of Turbomole⁶⁷ for the QM part and DL_POLY⁶⁸ for the MM part. The non-heme Fe center, the first coordinating sphere residues, and substrate-bound in the active site were included in the QM region (Figure 1). An unrestricted B3LYP functional was used to represent the QM region as it has been used for several non-heme⁶⁹ and 2OG-

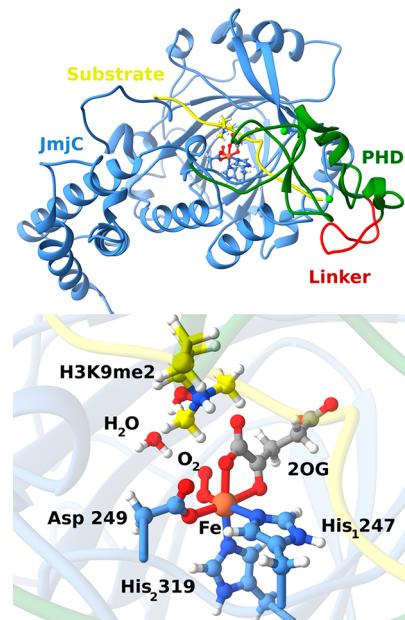


Figure 1. PHF8 ferric–superoxo intermediate with the H3_{1–14}K4me3K9me2 substrate bound at the active site. The QM region used for PHF8 QM/MM calculations is shown in ball and stick format. Note that the 2OG C1 carboxylate is in an “off-line” position in this view (Scheme 1).

enzymes with better accuracy in comparison to other functionals for calculation for hydrogen atom transfer (HAT) barrier.^{39,70–73} The protein region within 8 Å from the QM region was defined as the MM region, and the rest of the system was fixed. The Amber FF14SB force field was used for the MM region. QM/MM boundaries were treated with hydrogen link atom with a charge shift model.⁷⁴ The polarizing effect of the protein environment (MM region) was accounted on the QM region using an electrostatic embedding scheme.⁷⁵ For geometry optimization and frequency calculations, the def2-SVP basis set [QM(B1)/MM] was used.⁷⁶

Optimized reactant complexes (RC) were used to search for transition states (TS) along with the reaction coordinate by performing a relaxed potential energy scan (adiabatic mapping)⁷⁷ with a step size of 0.1 Å. The highest energy geometry in this potential energy surface (PES) was optimized using the P-RFO optimizer in HDLC code without any constraints.⁷⁸ Frequency calculations of all optimized geometries were carried out to confirm the minima and transition states. To improve calculated energies, single-point energy calculations on the optimized geometry were done using the large all-electron def2-TZVP basis set [QM(B2)/MM].⁷⁶ The zero-point energy (ZPE) calculations were performed for all geometries, and the energies are reported as QM(B2+ZPE)/MM. The D3 Grimme’s dispersion correction⁷⁹ is applied in single-point calculations, and energies are reported as B2+D3+ZPE in the Supporting Information. Experimental studies have shown the quintet spin state ($S = 2$) as a ground spin state for studied non-heme enzymes.^{80–82} Several computational studies on O_2 activation and HAT pathway in non-heme 2OG dependent enzymes have shown the reaction proceeds through a HS quintet spin state.^{39,69–72,83,84} Therefore, we performed reaction mechanism calculation on the quintet ($S = 2$, $M = 5$) spin state of PHF8 ferric–superoxo and ferryl–oxo complex. The geometries, spin densities, charges,

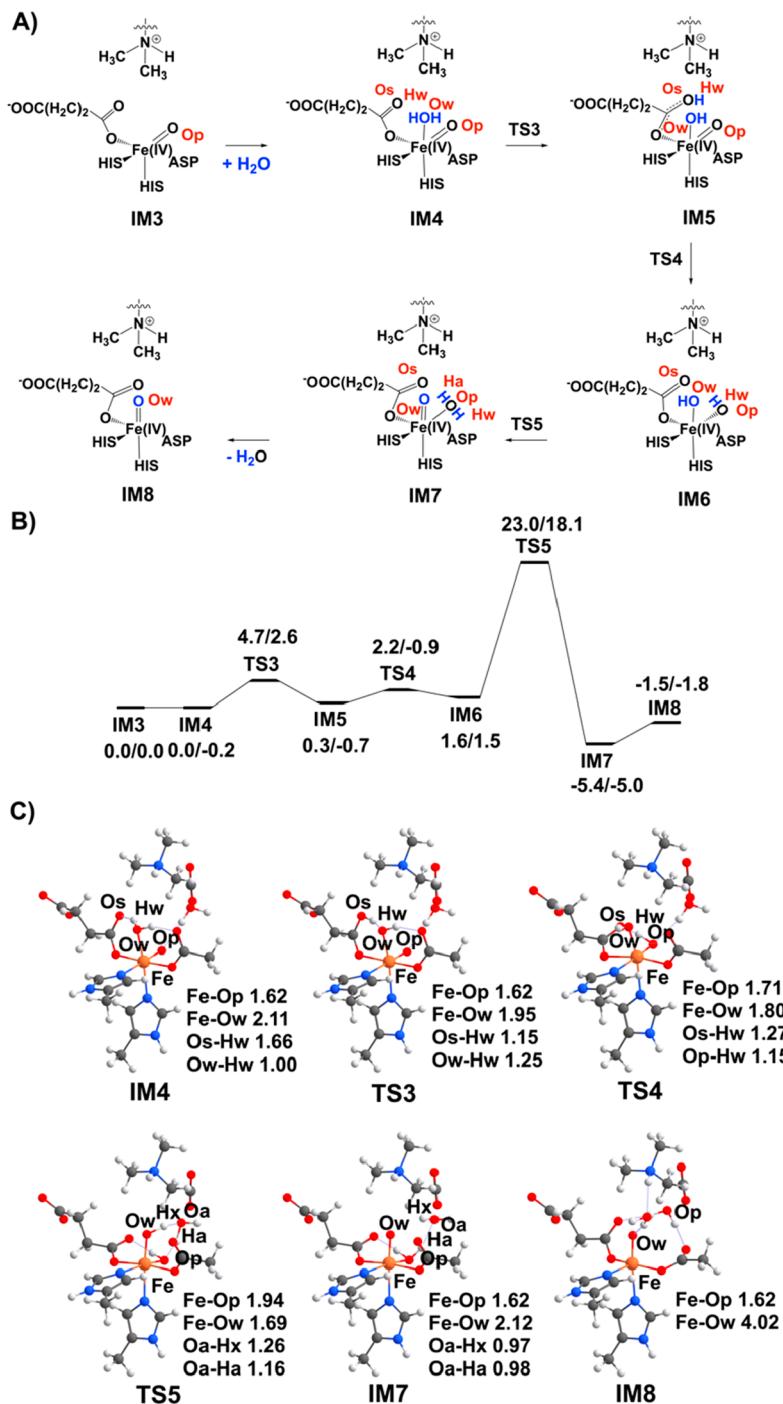


Figure 2. Ferryl-flip mechanism via oxygen atom exchange with a water molecule in PHF8. (A) Proposed mechanistic pathway of a ferryl-flip in an "off-line" PHF8-ferryl-oxo complex. (B) Potential energy surface for the ferryl-flip in the "off-line" PHF8-ferryl-oxo complex. Energies are in kcal/mol as calculated at QM(B2)/MM and QM(B2+ZPE)/MM level of theory. (C) QM active site geometries of transition states and selected intermediates in the ferryl-flip mechanism of the "off-line" PHF8-ferryl-oxo complex. Bond lengths are in angstroms.

and dispersion corrected energies of all intermediates and transition states are presented in Supporting Information.

3. RESULTS AND DISCUSSION

3.1. Conformational Flexibility of the Fe(III)-Superoxide Complex. Enzymes are flexible structures that typically exhibit a wide variety of internal motions ranging from femtoseconds to hours,^{85,86} some of which might be relevant to catalysis.⁸⁷ Structural studies imply the ferric-superoxide

intermediates in non-heme⁴⁵ and heme⁴⁶ oxygenase catalysis may have sufficiently long lifetimes to enable conformational changes. To investigate whether conformational changes are associated with O_2 binding in PHF8 and to obtain equilibrated structures for QM/MM calculations, we performed 1 μs MD simulations of the PHF8-Fe(III)-OO⁻-2OG(off-line).H3₁₋₁₄ substrate (subsequently referred to as the PHF8-ferric-superoxide complex). The MD simulations show a well equilibrated PHF8-ferric-superoxide complex with an average

root mean squared deviation (RMSD) of 1.97 Å for PHF8 C- α atoms with respect to the minimized crystal structure (Figure S1).¹⁷ Hydrogen bonds with protein residues and the solvent stabilize 2OG binding, the conformation of the substrate, and the orientation of the Fe-coordinated ligands (Figure S4). The non-Fe coordinating 2OG C5 carboxylate is stabilized by stable hydrogen bonding with Y257 (present in 80% of MD snapshots) and K264 (78%). T253 is positioned to make a hydrogen bond (82%) with the backbone oxygen of the iron-coordinating aspartic acid (D249). The Fe-coordinating histidines (H247 and H319) form hydrogen bonds with each other (72%) and with D245 (76%) and W317 (15%), respectively.

Enzymes can exhibit complex correlated motions; dynamic cross correlation analysis (DCCA) can be applied to explore correlated motions in protein dynamics, while principal component analysis (PCA) can be used to analyze motions in flexible regions of proteins.^{88,89} The PHF8–ferric–superoxide complex shows little correlation between motions of the linker and PHD domain with other parts of protein unlike in the PHF8-2OG·H3_{1–14} substrate (enzyme–substrate) complex,²⁸ suggesting the significant impact of the dioxygen binding on the network of correlated motions (Figure S5). In comparison with PHF8 enzyme–substrate complex dynamics (Scheme 1),²⁸ correlated motions are increased in the α 9 regions of the JmjC domain in the PHF8–ferric–superoxide complex. The increased motions in the α 9 regions of PHF8 are not manifest in the PHF8 enzyme–substrate complex dynamics²⁸ and hence appear related to O₂ binding. Principal component analysis (PCA) reveals that the α 9 regions move toward the active site, apparently preparing the complex for catalysis (Figure S6).

3.2. O₂ Activation Mechanism from “Off-Line” and “In-Line” 2OG Binding Modes. The reaction mechanism for the oxygen activation mechanism in PHF8·Fe(III)-OO[–]·2OG·H3_{1–14} substrate starting from both “off-line” RC1, or “in-line” RC1’ binding mode of 2OG was studied using QM/MM calculations. The intermediates and transition states obtained for both the “in-line” and “off-line” 2OG binding mode in PHF8 are similar to those identified for other 2OG dependent enzymes (Figures S7–S13).^{39,69–72,83,90–92} The mechanism proceeds via attack of the distal oxygen (Od) onto C2 of 2OG, leading to a loss of CO₂ to form a decarboxylated succinyl-peroxide intermediate IM1/IM1’, which has been recently characterized crystallographically for a 2OG dependent enzyme VioC.⁹⁰ The energy barrier for decarboxylation, which is the rate-determining step in the O₂ activation process, is 12.7 and 10.2 kcal/mol at the QM(B2+ZPE)/MM level of theory for “off-line” and “in-line” modes of 2OG, and slightly favors “in-line” binding of 2OG for O₂ activation (Figure S7–S8). These energy barriers compare reasonably well to experimental kinetic data on TauD, a similar 2OG dependent enzyme (15.2 kcal/mol, calculated from rate constant value of 42 s^{–1})⁹³ and computational studies on other 2OG oxygenases.^{70–72} The succinyl-peroxide IM1/IM1’ then undergoes homolytic cleavage of the O–O bond to generate a ferryl–oxo intermediate IM3/IM3’. The ferryl–oxo formed from “off-line” binding mode of 2OG IM3 has the oxo group *trans* to His247, i.e., away from the substrate (Figure S7) while the ferryl–oxo formed from “in-line” binding mode of 2OG IM3’ has the oxo group *trans* to His319, i.e., close to substrate (Figure S13). The geometries and spin densities of all intermediates and transition state in the O₂

activation mechanism are shown in Supporting Information (Table S1 and S7) and are consistent with computational studies on similar 2OG enzymes.^{39,69–72,83,90}

3.3. The Ferryl-Flip Mechanism via Solvent Oxygen Atom Exchange. The “off-line” ferryl–oxo intermediate (IM3) generated after O₂ activation has to be oriented in the direction of the substrate for productive catalysis, i.e., adopt an “in-line” geometry. One possible mechanism for such reorientation is via oxygen atom exchange with a water molecule (Figure 2A).^{12,30} This process could begin by binding of a water molecule to the open coordination site of the Fe-center at the IM3 intermediate stage (Figure 2B). The so formed intermediate (IM4) is a six-coordinate “off-line” ferryl–oxo with a water molecule facing in the direction of the substrate (Figure 2C). The bond length between Fe and the oxygen of water (Ow) is 2.11 Å; the Fe–Op bond distance is the same as in IM3 (i.e., 1.62 Å) and agrees well with another computational study of oxygen atom exchange in a 2OG oxygenase (AsqJ).³⁹ The proton (Hw) of the Fe coordinated water is transferred to the succinate oxygen (Os label on Figure 2A/2C) to form an intermediate (IM5) with a low barrier of 2.6 kcal/mol at QM(B2+ZPE)/MM level of theory via a transition state (TS3). IM5 then rearranges to give the ferryl-dihydroxyl species IM6 with a very low energy barrier (of 1.9 at B2 and –0.2 kcal/mol at QM(B2+ZPE)/MM). In IM6, the Fe–Op and Fe–Ow bond distances are 1.73 and 1.79 Å, respectively, which compares well with a computational study on AsqJ.³⁹ To complete the ferryl-flip, the second proton of the ligating water molecule (Hx) is transferred to Op via a proton-relay transfer with the help of a nearby water molecule. The second proton transfer proceeds with a transition state (TS5), which is 16.6 kcal/mol higher in energy than the previous intermediate at the QM(B2+ZPE)/MM level of theory. The intermediate (IM7) formed is the flipped six-coordinated “in-line” complex. This intermediate (IM7) can then release a water molecule and give a final flipped “in-line” 5-coordinate ferryl–oxo complex (IM8). Although the flipped 5-coordinate complex (IM8) is slightly higher in energy (by 3.2 kcal/mol) than the 6-coordinate “in-line” ferryl–oxo, modeling studies show that ferryl–oxo species can be either five- or six-coordinate;^{94–96} spectroscopic studies with another 2OG oxygenase (SyrB2) imply the ferryl–oxo is five-coordinate.⁵⁶ The total barrier for the ferryl-flip via oxygen atom exchange is calculated to be 18.1 kcal/mol at the QM(B2+ZPE)/MM level of theory, which is higher than a previously reported barrier of 10.9 kcal/mol for the 2OG oxygenase AsqJ.³⁹ The higher reaction barrier for the ferryl-flip in the case of PHF8 might be due to its more compact active site,^{17,44} and the steric hindrance between the water molecule needed for proton-relay in TS5 and the methyl groups of the substrate, which is not the case in AsqJ. Since the total barrier for the ferryl-flip is relatively high in PHF8, it may not be an energetically favored pathway for formation of the “in-line” ferryl–oxo intermediate, but it cannot be completely excluded as the barrier is still less than the experimental barrier of 21.3 kcal/mol (calculated from a rate constant of 5.4 h^{–1}) for the hydrogen abstraction step of PHF8, which is believed to be the rate-determining step in the entire reaction mechanism of PHF8.¹⁷

3.4. 2OG Rearrangement before O₂ Binding. To explore the change in the binding mode of 2OG from “off-line” to “in-line” in PHF8, we calculated a potential energy surface by changing the dihedral angle for N^e (nitrogen of

His_1)—Fe—O2—O1. We first optimized the five-coordinate PHF8 enzyme–substrate complex at the QM/MM level. The five-coordinate PHF8 enzyme–substrate complex (**ES1**) has iron in the Fe(II) high spin $S = 2$, $M = 5$ state, with the 2OG bound in an “off-line” bidentate fashion, and two histidines (H247 and H319) and an aspartate (D249) bound in monodentate fashion along with the H3_{1–14}K4me3K9me2 substrate. The bond lengths of iron and the two coordinating oxygens of 2OG (Fe—O1 and Fe—O2) are increased from 1.97 and 2.31 Å in **RC1** to 2.03 and 2.48 Å, respectively, in **ES1** (Figure 3). The dihedral angle for N^e (nitrogen of His₁)—Fe—

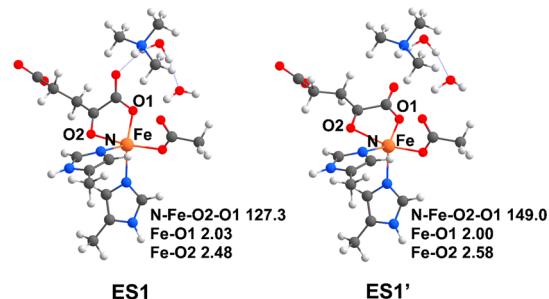


Figure 3. QM region geometry of QM/MM optimized **ES1** and **ES1'** for calculations involving a change in the binding mode of 2OG from “off-line” in **ES1** to “in-line” in **ES1'**. Angles and bond lengths are in degrees and angstroms, respectively.

O2—O1 also changes from 100° in **RC1** to 127° in **ES1**. The potential energy surface was calculated starting from **ES1**, using a reaction coordinate of the dihedral angle N^e (nitrogen of His₁)—Fe—O2—O1 to give the “in-line” five-coordinate enzyme–substrate complex **ES1'** (Figure 3). The dihedral rotation which leads to the “in-line” five-coordinate PHF8 enzyme–substrate complex (**ES1'**) proceeds with a low barrier of 1.48 kcal/mol (at the QM(B2+ZPE)/MM level of theory) principally because there are no protein residues interacting with the C1 carboxylate of 2OG in **ES1**. The transition state (**ES2**) has a dihedral angle of 144° for N^e (nitrogen of His₁)—Fe—O2—O1, and bond lengths between iron and the two coordinating oxygens of 2OG (Fe—O1 and Fe—O2) of 2.02 and 2.52 Å, respectively. The “in-line” five-coordinate PHF8 enzyme–substrate complex (**ES1'**) has a relative energy of −0.98 kcal/mol with respect to **ES1** at the QM(B2+ZPE)/MM level of theory. The dihedral angle for N(nitrogen of His₁)—Fe—O2—O1 in **ES1'** increases to 149°, with bond lengths between iron and the two coordinating oxygens of 2OG (Fe—O1 and Fe—O2) of 2.0 and 2.6 Å, respectively.

Thus, the change from “off-line” to “in-line” coordination mode of 2OG in PHF8 proceeds with a relatively low barrier, suggesting both possible 2OG C1 carboxylate geometries might coexist at room temperature. Consistent with the result of a change in the binding mode of 2OG in PHF8, the CAS.Fe(II).2OG.substrate.NO crystal structure (NO is used as an analog for O₂) shows a changed binding mode in the 2OG C1 carboxylate geometry upon NO binding in comparison to the CAS.Fe(II).2OG.substrate crystal structure.³⁰ Crystal structures of KDM2A-Ni-H3K36me1–3·2OG/NOG (Ni and NOG are used as analogs for Fe and 2OG, respectively) show evidence for both “in-line” and “off-line” binding modes of the 2OG/NOG C1 carboxylate, but the KDM2A.Fe.H3K36me1–succinate.NO structure (note the 2OG in the crystal structure was reported to be oxidized to succinate even though the

H3K36 substrate remained methylated) was refined to show only the “in-line” binding mode; i.e., NO was bound *trans* to His₂.⁹⁷

3.5. Conformational Dynamics and Substrate Stabilization in PHF8 Fe(IV)=O Intermediate. The ferryl–oxo species is a key intermediate in 2OG oxygenase catalysis that performs hydrogen abstraction from the substrate. This process might be sensitive to the dynamic orientation between the substrate and the iron, which is likely controlled by conformational changes.^{98–100} The mean lifetime of a ferryl–oxo intermediate has been measured to be 2.1 s for the 2OG dependent TauD,⁴⁷ and hence appears stable enough to enable potential conformational changes. To explore this possibility, we performed 1 μ s MD simulations of the PHF8-ferryl–oxo-inline complex (**IM8**) with succinate and the H3_{1–14}K4me3K9me2histone substrate. The MD simulations indicate a well equilibrated PHF8-ferryl–oxo-inline complex with an average root mean squared deviation of 2.03 Å for C- α atoms of the protein from the QM/MM optimized starting structure **IM8** (Figure S15). The protein stabilizes the H3K9me2 substrate and the iron center via multiple hydrogen bonds (Figure S18). The noncoordinating (C4) carboxyl group of succinate is stabilized by hydrogen bonding interactions with Y257 (83%), N189 (61%), K264 (60%), and T244 (41%), and solvent-mediated hydrogen bonds with the iron-ligated D249 (50%) and T253 (52%). The H3K9me2 is supported by hydrogen bonding with F250 (72%), R460 (70%), and a steric interaction with I191. T253 (75%) and G252 (12%) make hydrogen bonds to support the backbone oxygen of the Fe-coordinating D249, while the noncoordinating oxygen of the D249 carboxyl group makes a strong interaction with N333. The coordinating histidines (H247, H319) form a hydrogen bond with each other (89%) and with residues D245 (70%) and W317 (28%), respectively.

After 300 ns, the H3K9me2 substrate equilibrates to a conformation wherein one of the substrate methyl groups is closest to the ferryl–oxo unit, followed by another methyl group and hydrogen of NH (Figure 4). These observations suggest that the first hydrogen abstraction happens from the methyl group of the substrate. Similarly to observations with

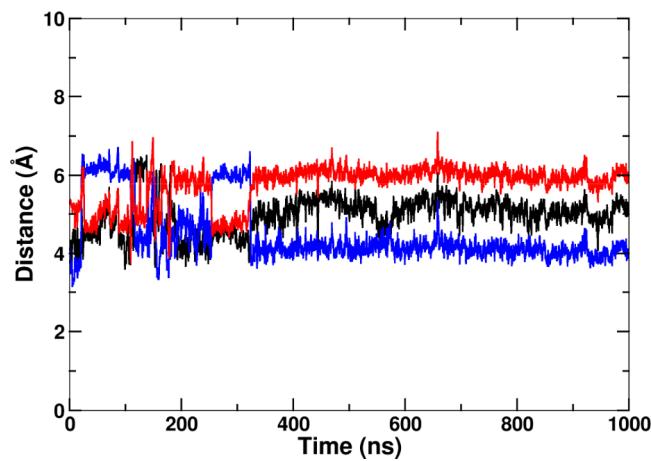


Figure 4. Distance between the ferryl–oxo oxygen atom and carbon atom of the first methyl group (black), carbon atom of the second methyl group (blue), and hydrogen of the amino group (red) of H3K9me2 substrate in the dynamics of PHF8-ferryl–oxo-inline complex.

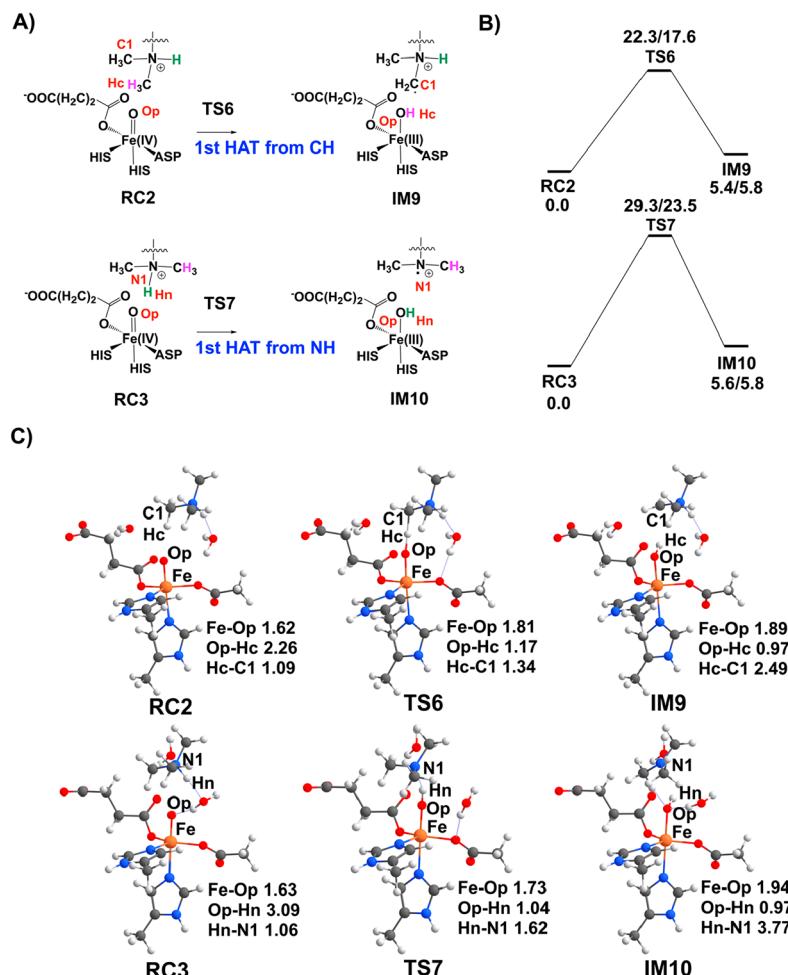


Figure 5. (A) Possible first hydrogen abstraction pathways from CH (RC3-TS6-IM9) and NH (RC4-TS7-IM10). (B) Potential energy surface for hydrogen abstraction step for hydrogen abstraction from carbon (CH) and for hydrogen abstraction from nitrogen (NH). Energies are reported in kcal/mol as calculated at the QM(B2)/MM and QM(B2+ZPE)/MM level of theory. (C) QM active site geometries of transition states and selected intermediates in the hydrogen atom transfer mechanism of PHF8 from QM/MM calculations. Bond lengths are in angstroms.

the PHF8–ferric–superoxide MD simulations, the PHD domain and the linker show reduced correlated motions, which could provide a stable enzyme conformation for catalysis (Figure S19). The motions in $\alpha 9$ – $\alpha 12$ are increased in comparison with the PHF8–ferric–superoxide MD and show motion in the direction of the active site, thus further compressing the overall PHF8 conformation (Figure S20). Such conformational changes will improve binding between PHF8 and the substrate, as seen in the calculated binding energies of the PHF8–Fe(IV)–oxo·succinate·H3_{1–14}K4me3K9me2 complex (−139.0 kcal/mol) and PHF8–Fe(III)–superoxide·2OG·H3_{1–14}K4me3K9me2 complex (−80.6 kcal/mol).

3.6. Hydrogen Atom Abstraction (HAT) Takes Place from C–H Rather Than N–H. It is generally accepted that in non-heme enzyme catalyzed demethylation, the oxygen atom of the Fe(IV)=O species abstracts a hydrogen atom from the methyl group of the substrate to form a Fe(III)–OH group and a substrate-based radical.^{11,12} For PHF8, since the H3K9 substrate is dimethylated, the Fe(IV)=O could also (potentially) attack a hydrogen atom on the N atom of H3K9me2 substrate. It has further been proposed that hydrogen atom abstraction can occur either from the methyl group or from the NH group, depending on the orientation of H3K9me2 with respect to the Fe–O bond (Figure 5A).⁴⁴

When the hydrogen atom of the NH group of the substrate points away from the iron-center, it is denoted as an “H-out” conformation (shown as reactant RC2 in Figure 5C), and if the hydrogen atom of the NH group points toward the iron-center, it has been denoted as an “H-in” conformation (shown as reactant RC3 in Figure 5C). Previous QM calculations on non-heme models of methylated lysine have indicated that for the dimethylated substrate, hydrogen abstraction from a methyl group and NH group are both feasible.⁴⁴ The MD simulations of the ferryl–oxo complex in PHF8 show the NH group in the “H-out” conformation, which makes NH abstraction less energetically favorable than C–H abstraction (Figure 4). To verify further the MD prediction of the ferryl–oxo PHF8 structure, we carried out a QM/MM potential energy surface for both C–H and N–H abstraction steps.

Two snapshots from PHF8–ferryl–oxo MD were chosen for QM/MM analysis, one snapshot with the “H-out” conformation which is apparently favorable for C–H abstraction (RC2), and another snapshot with “H-in” conformation, which is apparently favorable for N–H abstraction (RC3) (Figure 5C). The starting MD snapshots of reactants, RC2 and RC3, have the ferryl–oxo at distances of 2.3 and 2.6 Å from the hydrogen atom of carbon and nitrogen (O–HC/O–HN), respectively; the Fe–O–HC/Fe–O–HN angles are 163° and 139°,

Table 1. Table Showing the $\angle\text{Fe}-\text{O}-\text{H}$ Angles for QM/MM Optimized Multiple Snapshots of the First C–H Abstraction Step, with Spin Densities on the Methyl Carbon Atom (σ_{CH}) of the Substrate and Energy Barriers for the Transfer Listed

snapshot number	energy barrier [QM(B2+ZPE)/MM]	$\angle\text{Fe}-\text{O}-\text{H}$		spin density on C-atom in TS
		QM/MM optimized RC	QM/MM optimized TS	
1	24.57	163	164	-0.50
2	21.43	154	142	-0.47
3	17.65	167	150	-0.49
4	21.48	144	152	-0.51
5	17.64	127	164	-0.53

respectively. QM/MM optimization of both snapshots changes the O–HC/O–NH distances to 2.2/3.1 Å and Fe–O–HC/Fe–O–HN angles to 167°/149°, respectively, for **RC2/RC3**. The potential energy surfaces for hydrogen atom transfer from **RC2** and **RC3** are shown in Figure 5B. The activation barriers of the two HAT reactions with **RC2** and **RC3** are 17.6 and 23.5 kcal/mol, respectively, at the QM(B2+ZPE)/MM level of theory, suggesting that HAT involving C–H abstraction is the more energetically preferable reaction. The D3 dispersion corrected energies also reproduce the same trend and are shown in Supporting Information with spin densities (Table S9). The QM/MM calculations support the proposal of unfavorable N–H abstraction on the basis of MD studies. Thus, the substrate conformation and the protein environment around the active site of PHF8 support H-abstraction from the methyl group of the substrate to form a ferric–hydroxyl species and a substrate-based radical.

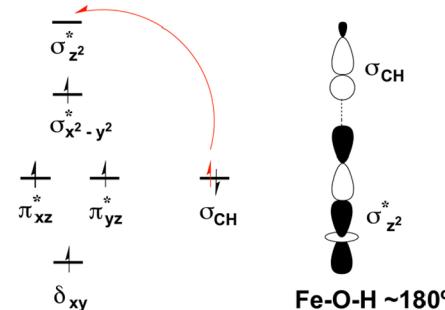
3.7. Conformational Flexibility, Molecular Orbitals, and Six-Coordinated Fe(IV)=O in the Hydrogen Atom Transfer. We performed multiple reaction path calculations starting from different structures of the ferryl–oxo complex generated by MD to explore the effects of conformational flexibility on the HAT. In all the HAT transition states, the iron-center is stabilized by a chain of hydrogen bonding water molecules. The orientation of the H3K9me2 substrate toward the ferryl–oxo is maintained by interactions from I191 and H3R8. The nonbonded oxygen of Fe-coordinated Asp249 makes hydrogen with N333. The Boltzmann-averaged reaction barrier for the rate-determining first hydrogen abstraction step using five snapshots was found to be 18.19 kcal/mol at QM(B2+ZPE)/MM level of theory (Table 1) and compares reasonably well with the experimental barrier of 21.3 kcal/mol (calculated from a rate constant of 5.4 h⁻¹)¹⁷ and other computational studies on similar 2OG oxygenases.^{95,96,101,102}

The ferryl–oxo complex **RC2** is considered “in-line”, as its Fe=O bond is oriented along the direction of the substrate (taken as the *z*-axis).^{98,99} In the quintet spin state of **RC2**, the Fe 3d orbitals are split into two π^* orbitals (π_{xz}^* and π_{yz}^*), two σ^* ($\sigma_{x^2-y^2}^*$ and σ_z^*) orbitals, and a nonbonding (δ_{xy}) orbital. The antibonding interactions between the 3d orbitals of Fe and the 2p orbitals of the oxo ligand result in the π_{xz}^* and π_{yz}^* orbitals, instead of π_{xy}^* and π_{yz}^* , as in **IM3** (Figure S14). The σ_z^* in **RC2** is the result of an antibonding interaction between the $3d_z^2$ orbital of the Fe and the $2p_z$ orbital of the oxo group; the $\sigma_{x^2-y^2}^*$ orbital arises from an antibonding interaction between $3d_{x^2-y^2}$ of Fe and ligands in the *x*–*y* plane (H247, D249, and succinate). Therefore, the orbital occupancy of quintet **RC2** is $\delta_{xy}^1 \pi_{xz}^1 \pi_{yz}^1 \sigma_{x^2-y^2}^1 \sigma_z^0$. The orbital arrangement in **RC2** leads to a stabilization of α -spin orbitals via an exchange enhanced reactivity, which lowers the energy of the α – σ_z^* orbital.^{98,99,103,104} Thus, there are two types of frontier molecular orbitals (FMOs), the σ_z^* and the

degenerate $\pi_{xz/yz}^*$ available for hydrogen atom abstraction.^{98,99} In the **RC2**, the σ_z^* orbital lies along the direction of the ferryl–oxo bond, while the π^* orbitals lie perpendicular to the direction of the Fe–O vector.

For the transition states of hydrogen transfer, the orbital overlap is optimized by a Fe–O–H angle close to 180° for σ^* electron transfer and a Fe–O–H angle close to 120° for π^* electron transfer (Figure 6).^{98,99} Several calculated transition states of HAT reactions by Fe(IV)=O species have shown to follow these orbital-selection rules.^{98–100}

A) σ^* transfer



B) π^* transfer

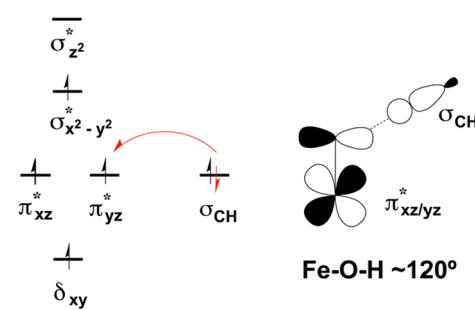


Figure 6. Orbital selection rules for the hydrogen abstraction process in high-spin ($S = 2$) non-heme iron–oxo intermediates.

To investigate how electron transfer in PHF8 occurs, we analyzed transition state geometries from the five QM/MM snapshots for the HAT from the C–H bond. The $\angle\text{Fe}-\text{O}-\text{H}$ in the transition state of all the five reaction paths remains in the range of 142°–164°, implying electron transfer via the σ^* pathway. In PHF8, a transition state $\angle\text{Fe}-\text{O}-\text{H}$ close to 180° is favored due to an interaction between the I191 side chain and the methyl groups of the substrate. The spin natural orbital (SNO) of the HAT transition state (**TS6**) shows electron transfer to the ferryl–oxo σ_z^* orbital (Figure 7). σ^* – α -electron transfer is supported by the residual negative spin density on the substrate carbon (Table 1). Thus, in PHF8

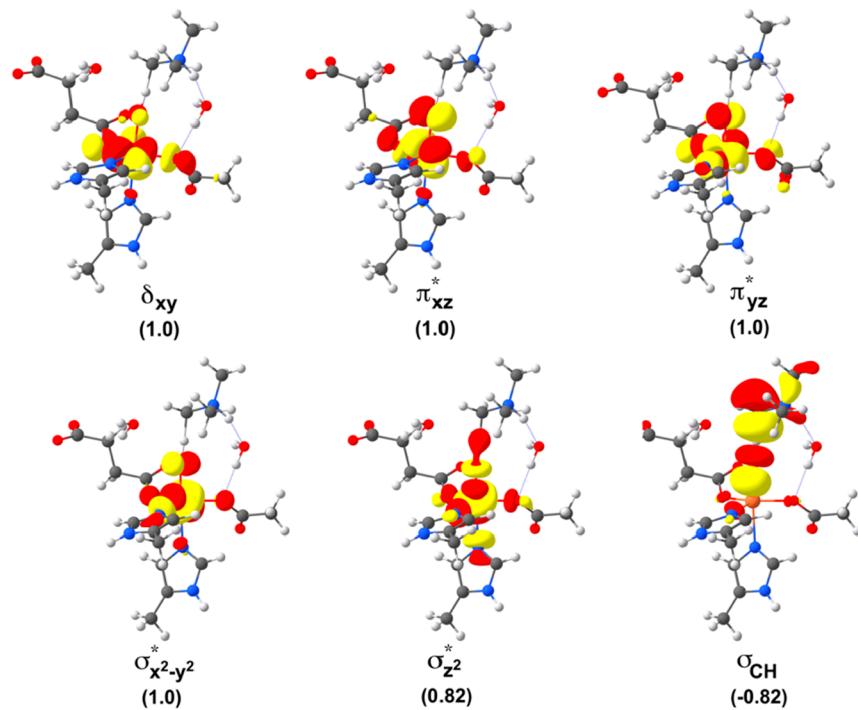


Figure 7. Transition state spin natural orbitals (SNO) of the HAT (TS6) from C–H with orbital occupancies in the bracket.

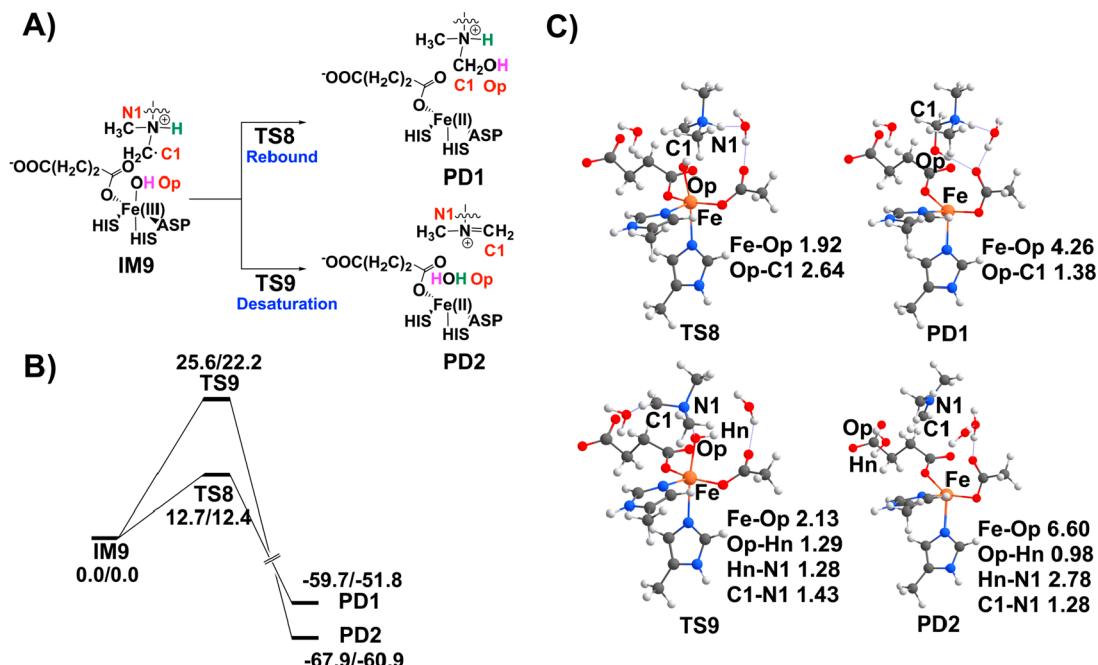


Figure 8. (A) Possible reaction pathways after HAT for the rebound (IM9-TS8-PD1) and desaturation (IM9-TS9-PD2) mechanisms. (B) Potential energy surface for the rebound and desaturation mechanisms. Energies are reported in kcal/mol at the QM(B2)/MM/QM(B2+ZPE)/MM level of theory. (C) QM active site geometries of selected transition states and products in the hydrogen atom transfer step of PHF8 catalysis from QM/MM calculations. Bond lengths are in angstroms.

(RC2), electron transfer occurs via a σ channel to the ferryl–oxo $\sigma^*_{z^2}$ orbital.

To explore the effect of additional water coordination at the ferryl–oxo stage, including on the hydrogen atom transfer reaction barrier, we performed a QM/MM potential energy scan starting from the six-coordinated ferryl–oxo IM7. The reaction barrier for hydrogen abstraction from methyl C–H of the substrate was found to be 21.7 kcal/mol at the

QM(B2+ZPE)/MM level of theory for the ferryl–oxo with additional water. The activation barrier for hydrogen atom transfer from the six-coordinated ferryl–oxo species is thus within the range of activation energy obtained from the five-coordinated ferryl–oxo intermediate.

3.8. Following HAT, the Reaction Proceeds through Rebound Rather than Desaturation. While the rate-limiting step for demethylation by PHF8 (at least once the

substrate is bound at the active site) is hydrogen abstraction, the following step raises another interesting mechanistic question. The product formed after HAT, i.e., the ferric-hydroxyl species (**IM9**), can either undergo a rebound reaction to form the hydroxylated product or produce a desaturated product by abstracting the second hydrogen from NH to give an iminium ion which could then undergo hydrolysis (Figure 8A). This mechanistic dilemma manifests itself in other 2OG demethylases and some other related enzymes (in some cases, both hydroxylation and desaturation are observed;¹² a synthetic model of the non-heme oxygenases has shown desaturation to be a feasible mechanistic option).¹⁰⁵ The factors governing the Fe(III)–OH species to act as hydroxylase or desaturase have been shown to depend on bond strength, coordination-sphere, and ligand sphere flexibility, and steric effects.^{105,106}

In the absence of experimental data on enzyme-bound hydroxylated intermediates in KDMs catalysis, we explored the effects of the protein environment on the selection of the reaction path after the HAT; we investigated the reactivity of the ferric-hydroxyl product formed after HAT (**IM9**) to undergo rebound hydroxylation and desaturation (Figure 8A). The potential energy profiles for rebound (**IM9-TS8-PD1**) and desaturation (**IM9-TS9-PD2**) reactions are shown in Figure 8B. The calculated activation energies are 12.4 and 22.2 kcal/mol for **TS8** and **TS9** at the QM(B2+ZPE)/MM level of theory for the rebound and the desaturation reaction, respectively. The D3 dispersion corrected energies also support the rebound over the desaturation mechanism (see Table S13 and Table S15 for spin densities). The calculations show a clear preference for a rebound process in the presence of the PHF8 protein environment. Steric effects are a major factor helping favor a rebound over a desaturation pathway, including those involving I191. As seen earlier for calculations on the hydrogen atom abstraction step (Figure 4), the substrate conformation does not make N–H easily accessible to the **IM9** intermediate because of steric constraints; in **IM9** it is stabilized via hydrogen bonding (with iron-bound aspartate D249), thus rendering desaturation unfavorable (Figure 8C). The hydroxyl group in **IM9** might have reduced electrophilicity compared to a ferryl–oxo group (the spin density on the oxygen atom of the hydroxyl group reduces to 0.22 after HAT from a spin density of 0.55), making it a bad electrophile for a second hydrogen abstraction step/desaturation from the NH bond. Radical delocalization is also an important factor in determining whether rebound or desaturation occurs; if the radical is relatively delocalized, desaturation is more preferred; if the radical is localized, then rebound is preferred.¹⁰⁵ The electron density of (−1.04) on the carbon atom in **IM9** shows the radical is localized on the carbon-atom, making rebound more likely. Thus, steric effects by the protein environment, hydrogen-bond stabilization of the NH species with iron-bound aspartate D249, and radical localization contribute to the rebound reaction being energetically preferred over desaturation in PHF8 catalysis.

4. CONCLUSIONS

We used multiple microsecond MDs and QM/MM simulations to explore the catalytic strategy of the non-heme Fe(II) and 2OG-dependent histone lysine demethylase PHF8, which is involved in key epigenetic processes. Similarly to other 2OG-dependent non-heme enzymes, PHF8 follows an O₂ activation pathway wherein a ferric–superoxo intermediate

reacts with the Fe-complexed C2 carbonyl group of 2OG leading to decarboxylation and formation of a ferryl–oxo complex and succinate. If generated from a starting complex in which the 2OG C1 carboxylate is complexed to the Fe *trans* to His247, the so generated ferryl–oxo complex has an “off-line” geometry, i.e., is unproductive for catalysis. This unproductive species has been proposed to undergo a ferryl-flip via oxygen atom exchange with a water molecule to form an “in-line” productive ferryl–oxo species (*trans* to His319); however, our calculations indicate that this process is energetically unfavorable. Instead, we propose that at a five-coordinate 2OG-bound state prior to dioxygen binding, 2OG flips from an “off-line” to an “in-line” coordination geometry, in a process with a minimal energy barrier, thus enabling the subsequent, direct production of “in-line” ferric-superoxo, succinyl–peroxy, and ferryl–oxo intermediates. The “in-line” ferryl–oxo intermediate is thus in a productive orientation with respect to the substrate for the subsequent reaction steps.

QM/MM calculations imply that hydrogen abstraction occurs from a C–H bond of substrate N^ε-methyl group; by contrast, hydrogen abstraction from N^ε–H is a highly energy-inefficient process, in part due to the substrate conformation relative to the ferryl–oxo species and steric repulsion from the second sphere residue I191. By analyzing multiple reaction paths for the HAT step, we show that hydrogen abstraction proceeds via a σ channel. The QM/MM calculations in the presence of the protein environment of PHF8 show that the rebound pathway is more energetically favored than desaturation due to the substrate conformation, second sphere interactions, and reduced reactivity of Fe–OH species generated after HAT.

Overall our studies provide key insights into the atomistic effects of the protein environment on the reaction path of PHF8 and distinguish between alternative mechanistic proposals for hydroxylation/demethylation of the substrate N^ε-methyl group. Presently, with rare exceptions (e.g., meldonium¹⁰⁷), most 2OG oxygenase inhibitors, including one recently approved for clinical use,^{108,109} are active site Fe-chelators that compete with 2OG. Such compounds, however, block rather than modulate 2OG oxygenase activity with consequent potential “on target” side effects in the presence of excess inhibitor. We hope that the detailed mechanistic insights presented here for PHF8, as well as ongoing work on other 2OG oxygenases, will help enable the design of mechanism-based inhibitors of 2OG oxygenases that modulate kinetics, e.g., by slowing the movement of the 2OG C1 carboxylate from an “off-line” to an “in-line” geometry.

■ ASSOCIATED CONTENT

■ Supporting Information

The Supporting Information is available free of charge at <https://pubs.acs.org/doi/10.1021/acscatal.9b04907>.

Analysis of molecular dynamics and QM/MM optimized QM geometries with spin densities and partial charges are included (PDF)

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Notes

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