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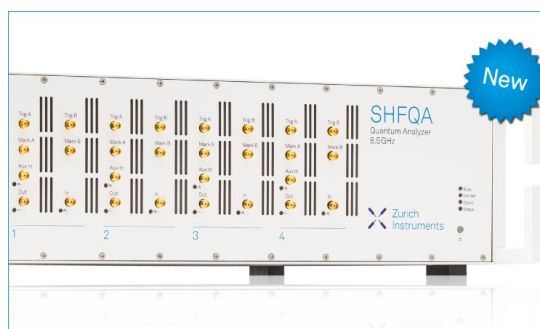
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ABSTRACT

Many chemical reactions in Nature, the laboratory, and chemical industry occur in solvent mixtures that bring together species of dissimilar solubilities. Solvent mixtures are visually homogeneous, but are not randomly mixed at the molecular scale. In the all-important binary water-hydrotrope mixtures, small-angle neutron and dynamic light scattering experiments reveal the existence of short-lived (<50 ps), short-ranged (~1 nm) concentration fluctuations. The presence of hydrophobic solutes stabilizes and extends such fluctuations into persistent, mesoscopic (10–100 nm) inhomogeneities. While the existence of inhomogeneities is well established, their impacts on reactivity are not fully understood. Here, we search for chemical signatures of inhomogeneities on the surfaces of W:X mixtures (W = water; X = acetonitrile, tetrahydrofuran, or 1,4-dioxane) by studying the reactions of Criegee intermediates (CIs) generated *in situ* from O₃(g) addition to a hydrophobic olefin (OL) solute. Once formed, CIs isomerize to functionalized carboxylic acids (FC) or add water to produce α -hydroxy-hydroperoxides (HH), as detected by surface-specific, online pneumatic ionization mass spectrometry. Since only the formation of HH requires the presence of water, the dependence of the R = HH/FC ratio on water molar fraction x_w expresses the accessibility of water to CIs on the surfaces of mixtures. The finding that R increases quasi-exponentially with x_w in all solvent mixtures is consistent with CIs being preferentially produced (from their OL hydrophobic precursor) in X-rich, long-lived OL:X_mW_n interfacial clusters, rather than randomly dispersed on W:X surfaces. R vs x_w dependences therefore reflect the average (m, n) composition of OL:X_mW_n interfacial clusters, as weighted by cluster reorganization dynamics. Water in large, rigid clusters could be less accessible to CIs than in smaller but more flexible clusters of lower water content. Since mesoscale inhomogeneities are intrinsic to most solvent mixtures, these phenomena should be quite general.

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I. INTRODUCTION

Both in the laboratory and in large-scale operations, most chemical reactions are performed in solvent mixtures tailored to bring together reactants of dissimilar solubilities.¹ Important natural processes also take place in aqueous organic media of complex compositions.^{2–4} Solvents not only allow reactant molecules to get in close contact but also influence reactions in various ways and at different scales.⁵ At one time, solvent effects in chemical reactions were rationalized in terms of transition state theory and continuous models based on macroscopic descriptors, such as solvent viscosity and dielectric constants.¹ This perspective, however, has had

mixed success over the years. A deeper understanding of the solvent effects emerged from the results of small and wide-angle X-ray⁶ and small angle neutron scattering experiments, which revealed concentration fluctuations in the bulk and the surface of macroscopically homogeneous, visually clear, “internally mixed” binary solvent mixtures.^{7–10} These concentration (or structural) fluctuations are brief (tens of picoseconds) inhomogeneities in the ~1 nm scale that happen to peak in the concentration ranges that give rise to thermodynamic and transport properties anomalies in binary mixtures.^{11–17}

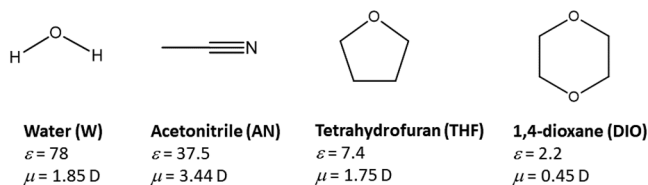
The addition of solutes to binary mixtures generates the larger (in the 10–100 nm mesoscale), more persistent (from

milliseconds to years) non-covalent aggregates detected by dynamic light scattering experiments in many systems.¹⁸⁻²² These aggregates also appear in the mass spectra of liquid jets produced by adiabatic expansion of solutions of hydrophobic species in aqueous organic mixtures.²³⁻²⁶ The universality of this phenomenon opens up the possibility that molecular environments could be engineered in solvent mixtures for steering the course of chemical reactions in unanticipated ways.^{12,27-29} It also leads to the realization that spontaneous compartmentalization at the mesoscale in complex liquid mixtures, by potentially supporting disjoint but coupled chemical processes in apparently homogeneous systems, is an alternative mechanism to the reaction-diffusion Turing model of the onset of complexity and pattern formation in Nature.³⁰⁻³²

Water is a major component of natural systems; however, it is rarely present as a neat water, but as a solvent or distributed in diverse materials. Arguably, aqueous interfaces are where most interesting things happen: solutes of dissimilar characteristics meet at aqueous interfaces with hydrophobic media, and chemical reactions proceed much faster and sometimes along different pathways from bulk media.³³⁻³⁷ Most biochemical reactions *in vivo* take place within self-organized enzymatic cavities and related micellar arrays, where water-insoluble species react with aqueous solutes at the intervening interfaces.

Aqueous organic interfaces are widespread in the environment,^{38,39} most conspicuously as the organic films that cover the oceans, and on atmospheric aqueous organic aerosols.^{40,41} These systems, which generally consist of water plus amphiphilic and hydrophobic organic species, can be internally mixed, phase-separated into aqueous cores covered by organic shells, or phase-separated into segregated aqueous and organic domains, depending on the inorganic solutes, temperature, relative humidity, and the abundance and functionality of organic species.^{2,42-50} It is apparent that the rates and products of the chemical processes associated with the formation, growth and aging of organic aerosols, and the exchange of gases between the atmosphere and the oceans may critically depend on whether organics are present as solutes in internally mixed or in phase-separated mixtures. In this context, phase separation refers to the meso/macroscopic scale. The length scale at which chemical reactions occur, however, is much smaller. The bulk and the surfaces of segregated solvent mixtures, and even those of “internally mixed” mixtures can be inhomogeneous at the molecular level. Self-clustering of hydrophobic molecules into discrete islands or domains on the surface of aqueous solutions,⁵¹⁻⁵³ in conjunction with vertical concentration and solvation gradients through air-liquid interfacial layers, could significantly affect chemistry in such media.⁵⁴⁻⁵⁹

In the past, liquid interfaces remained invisible to conventional photon or particle scattering techniques. The advent of molecular dynamics studies, surface sensitive techniques, such as small-angle X-ray and neutron scattering, neutron and X-ray reflectometry, second harmonic generation



SCHEME 1.

spectroscopy, and photoelectron microscopy has produced a wealth of structural information.^{52,60,61} A valid question is whether the features revealed by these studies on the surfaces of binary solvent mixtures could assist us in understanding and predicting the products and rates of reactions, which necessarily involve additional components. One caveat is that liquid structures are essentially dynamic, and therefore, the length and time scales of the entities detected by physical techniques may not be directly relevant to chemical reactions. The concentration fluctuations detected by the scattering techniques in binary mixtures are not the persistent mesoscale aggregates generated upon the addition of reactant solutes. These are important issues in several fields, which we address herein in experiments where the surfaces of liquid mixtures are probed for molecular aggregates by reacting molecules.⁶²

The reacting molecules in our experiments are carbonyl oxide Criegee intermediates (CIs) on the gas-liquid interfaces of W:X mixtures (W = water; X = acetonitrile, AN; tetrahydrofuran, THF; and 1,4-dioxane, DIO) (Scheme 1). CIs are generated *in situ* by reacting $\text{O}_3(\text{g})$ with β -caryophyllene (a C_{15} -diolefin, denoted as β -C or OL from now on) on the surface of its solutions in W:X mixtures. These solvents are fully miscible, but are known to display bulk microheterogeneity.²⁸ A previous sum-frequency spectroscopy study showed that the surface of W:AN mixtures undergoes a sharp transition into a structurally uniform interfacial layer below water molar fraction $x_w \approx 0.93$.^{51,53,63} Related molecular dynamics calculations refined such scenario and revealed that the outermost interfacial layers are actually microheterogeneous, being populated by disjointed W:AN clusters in proportions that depend on x_w in a non-trivial way.^{64,65} In our experiments, the products of CI reactions are monitored in real time via surface-specific online pneumatic ionization mass spectrometry (o-PI-MS) (see below). More specifically, we track the relative yields of the products of competing reactions of the CIs as functions of x_w . We report that relative yields are strong exponential functions of x_w in all mixtures and analyze the implications of these results.

II. EXPERIMENTAL SECTION

In our experiments, (β -C + NaCl) solutions in W:X mixtures are injected (at $100 \mu\text{l min}^{-1}$) as liquid microjets ($v_j = 21 \text{ cm/s}$, see Fig. S1) into the spraying chamber of the mass spectrometer (ES-MS, Agilent 6130 Quadrupole LC/MS Electrospray System at NIES, Japan) through an electrically grounded stainless-steel syringe needle ($100 \mu\text{m}$ bore). The

spraying chamber is flushed with $N_2(g)$ and maintained at 1 atm, at 298 K. The liquid microjets are fragmented (within $\tau_R < 10 \mu s$)^{66,67} into charged microdroplets^{68,69} by a high velocity (160 m/s) nebulizer gas [$N_2(g)$] flowing through a sheath coaxial with the syringe needle.⁷⁰ In these events, the kinetic energy of the nebulizing gas provides energy for creating the surfaces of the microdroplets and for charging the microdroplets by separating the anions and cations already present in the microjets. Nebulization is the single event that generates the net charges that ultimately appear as mass spectral signals. That the droplets of mists produced by nebulizing liquids are charged was known for more than a century.^{69,71} This charging process does not require,^{72,73} although may be enhanced by, the presence of an electric field. In the [supplementary material](#) (see Validation Experiments), we show that in our setup the creation of net charge is mainly due to pneumatic nebulization rather than to the applied electric field needed to deflect charged microdroplets towards the mass analyzer (see Fig. S1). For this reason, we label our technique online pneumatic ionization mass spectrometry o-PI-MS, rather than electrospray ionization mass spectrometry. In this regard, it is important to realize that “electrospray ionization mass spectrometers” are not all alike. The main differences are whether the needle injector is grounded and whether the microjets are directly projected towards mass analyzer or not.⁷⁴ Our approach to studying interfacial processes via o-PI-MS is not a new idea. In their seminal 1984 paper, Yamashita and Fenn had already suggested that electrospray ionization mass spectrometry could be used to investigate liquid interfaces.⁷⁵

CIs are produced upon exposing to $O_3(g)$ on the surfaces of liquid microjets containing β -C. We chose β -C as the CIs precursor for its exceptional reactivity towards $O_3(g)$, which leads to significant conversions during the short $\tau_R < 10 \mu s$ lifetimes of the intact microjets.^{54,76–78} Reaction products are detected without any manipulation by o-PI-MS in less than a 1 ms. NaCl is added to bind the neutral α -hydroxy-hydroperoxide (HH) products into detectable Cl^- -adducts, which are unambiguously identified by the characteristic $M/M+2 = 3/1$ ratio imparted by natural abundance of $^{35}Cl/^{37}Cl$. We verified that Cl^- is inert towards $O_3(g)$ under present conditions.⁵⁹ Additional experiments (see Validation Experiments in the [supplementary material](#)) show that the detected ions originate from species produced in the outermost interfacial layers of intact liquid microjets, rather than on the microdroplets produced by nebulization, or in gas-phase reactions between $O_3(g)$ and evaporated β -C(g). See the [supplementary material](#) and Figs. S2–S4 for additional validation experiments. Further experimental details may be found in previous publications.^{37,59,70,77,79,80}

Ozone was generated by flowing ultrapure $O_2(g)$ ($>99.999\%$) through a silent discharge ozonizer (KSQ-050, Kotohira, Japan) and quantified via online ultraviolet-visible (UV-Vis) absorption spectrophotometry (Agilent 8453) before entering the spraying chamber. Reported $[O_3(g)]$ values correspond to concentrations on the surface of the liquid microjets, which we estimate by multiplying the concentrations

determined by absorption spectrophotometry by a $\sim 1/12$ factor due to dilution by the nebulizing gas flow. Conditions used these experiments were as follows: drying nitrogen gas flow rate: 12 l min^{-1} ; drying nitrogen gas temperature: 340°C ; capillary inlet voltage: $+3.5 \text{ kV}$ relative to the grounded injector; and fragmentor voltage value: 60 V . Fresh solutions were prepared daily using ultrapure water (resistivity $\geq 18.2 \text{ M}\Omega \text{ cm}$ at 298 K) from a Millipore Milli-Q water purification system. The following chemicals were used as received: β -caryophyllene ($\geq 98.5\%$, Sigma-Aldrich), acetonitrile ($\geq 99.8\%$, Wako), tetrahydrofuran ($\geq 99.8\%$, stabilizer free, Wako), 1,4-dioxane ($\geq 99.8\%$, stabilizer free, Wako), NaCl ($\geq 99.999\%$, Sigma-Aldrich), L-ascorbic acid ($\geq 99\%$, Sigma-Aldrich), and octanoic acid ($\geq 97\%$, Wako).

III. RESULTS AND DISCUSSION

Figure 1 shows negative ion mass spectra of the products generated on the surface of (1 mM β -C + 0.2 mM NaCl) in $x_w = 0.03, 0.42$, and 0.74 W:AN liquid microjets exposed to $O_3(g)$ for $\tau_R < 10 \mu s$. The intense peaks at m/z 305 (307) correspond to^{35,37} Cl^- -adducts of α -hydroxy-hydroperoxides (HH).^{59,81–83} HH are formed in the reaction of H_2O (MW = 18) with the Criegee intermediates (CIs) produced (via a primary ozonide POZ) from the addition of O_3 (MW = 48) to the endo C=C bond of β -C (MW = 204).^{54,84} Neutral HH subsequently bind Cl^- upon desolvation.^{85,86} Combined, these processes lead to 305 (307) = $204 + 48 + 18 + 35$ (37) (Scheme 2).^{59,82,83} The attachment of hydrated $^{35,37}Cl^-$ to HH into stable, detectable gas-phase HH- Cl^- adducts takes place upon desolvation in the heated capillary section (not shown in Fig. S1) of the mass analyzer. The m/z 251 peak corresponds to anions of functionalized-carboxylates (FC) from the isomerization of CIs to carboxylic acids, followed by deprotonation: $251 = 204 + 48 - 1$ (Scheme 2).^{54,87} We also detected the signals of secondary species resulting from the reactions of primary products with O_3 . For example, m/z 267 signals correspond to species resulting from O-addition to the primary FC: $267 = 251 + 16$.⁵⁴ These m/z 267 signals appeared in all solvent mixtures (see Figs. S5 and S6).

The m/z 321 (323) signals correspond to the chloride-adducts of a secondary neutral product of stoichiometry, formally amounting to the addition of 4 O-atoms plus a H_2O molecule: 321 (323) = $204 + 64 + 18 + 35$ (37).^{59,81–83} However, water does not participate in their formation because the intensities of 321 (323) signals decrease at larger x_w values, and their positions did not shift in experiments carried out in $H_2^{18}O$ (Figs. S7–S10).⁵⁹

Figures S7–S9 are the plots of product yields: $Y(P_i) = I_{P_i}/\Sigma I_{P_j}$ (I_{P_i} is the intensity of the mass signal of product P_i) as functions of x_w [at constant $O_3(g)$ exposures] in W:AN, W:THF, and W:DIO mixtures. Figure S10 shows I_{P_i} vs. x_w plots at various $O_3(g)$ exposures in W:AN mixtures. The main result is that only the intensities of the HH 305 (307) signals increased, while all others decreased with x_w . Together, these observations imply that HH is produced from CIs + H_2O

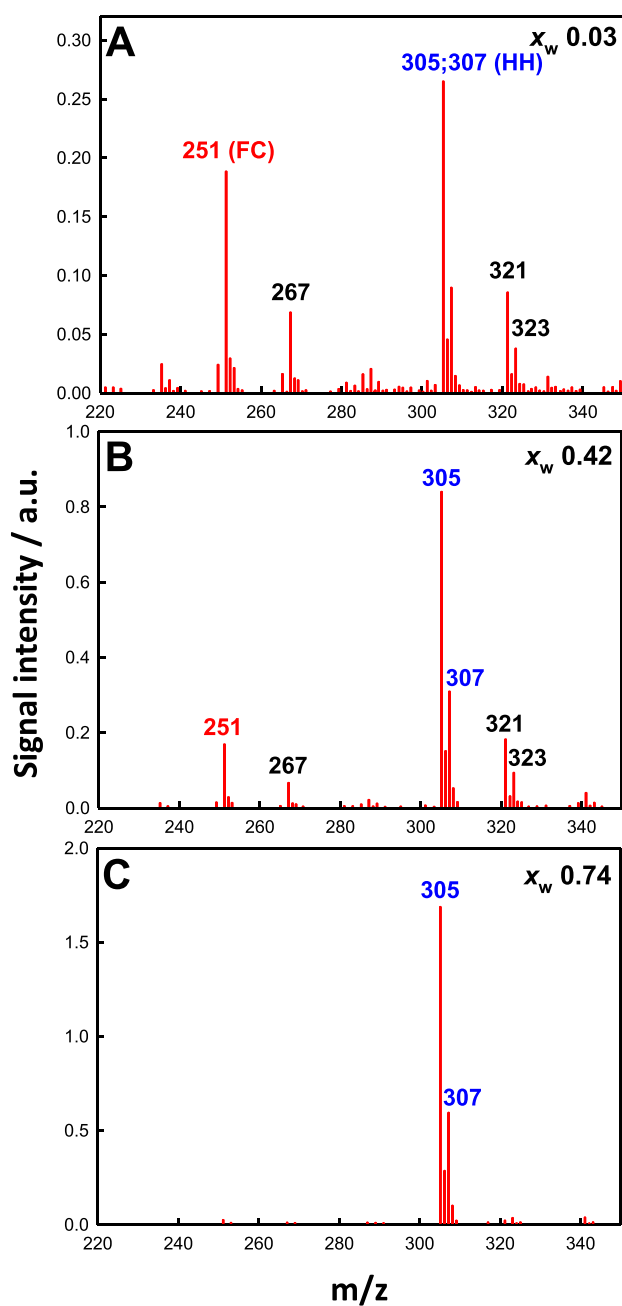


FIG. 1. Negative ion mass spectra of the products generated in the interfacial layers of (1 mM β -C + 0.2 mM NaCl) in W:AN liquid microjets exposed to $E = 5.3 \times 10^{11}$ $\text{O}_3(\text{g})$ molecules cm^{-3} s. Water molar fractions in bulk W:AN mixtures: (a) $x_w = 0.03$, (b) $x_w = 0.42$, and (c) $x_w = 0.74$.

in competition with other CIs reaction channels, as shown in Scheme 2.

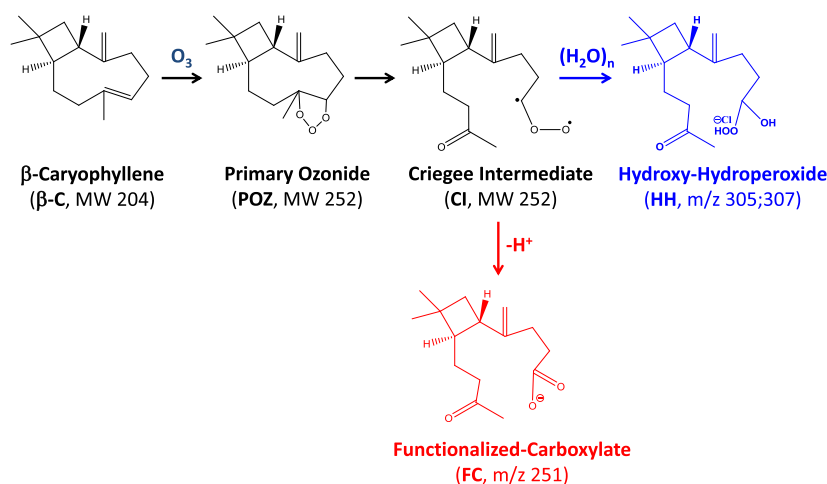
Figures 2(a) and 2(b) show linear and semi-logarithmic plots of $R = \text{HH}/\text{FC} = (305 + 307)/251$ ratios of mass

spectral signal intensities as functions of x_w in W:AN, W:THF, and W:DIO mixtures at constant $\text{O}_3(\text{g})$ exposure. The accessible x_w ranges were determined either by β -C solubility in each W:X mixture or by low signal-to-noise ratios. The data points in Figs. 2(a) and 2(b) are the averages of three separate experiments. The error bars do not exceed the size of symbols in most cases.

Below, we analyze the physicochemical underpinnings of the steep, positive R vs x_w dependences shown in Figs. 2(a) and 2(b). It should be borne in mind that our experiments probe the composition of the outermost interfacial layers rather than the bulk liquids. Therefore, R vs x_w trends could depend on the relative affinities of the neutral HH and functionalized carboxylic acid products for the liquid surfaces. In this regard, note that the 252 undissociated carboxylic acids and the 270 α -hydroxy-hydroperoxides are both neutral, of similar molecular sizes and amphiphilic, because both possess polar O-containing functional groups on a hydrophobic C_{15} hydrocarbon framework. Therefore, their propensities for air-liquid interfaces are expected to be similar and weakly dependent on x_w in aqueous organic mixtures.^{88,89} The degree of ionization of the 252 carboxylic acids into the detected $m/z = 251$ FC anions, however, should increase with x_w , by analogy with the x_w -dependences of the octanoate m/z 143 mass signal intensities from octanoic acid solutions both in W:AN and W:THF mixtures (Figs. S11 and S12). Ascorbate m/z 175 mass signal intensities in ascorbic acid dissolved in W:AN mixtures follow a similar trend (Fig. S13).

Therefore, if $R = \text{HH}/\text{FC}$ ratios were determined by relative surface affinities and acid dissociation, they should have decreased with x_w in all cases, in striking contrast with the results shown in Fig. 2. By assuming that the polarities of the CIs and the transition states for their isomerization into functionalized acids are similar, i.e., that CIs isomerization rates are weakly dependent on x_w , we infer that the strong positive R vs x_w dependences arise from the acceleration of $\text{CIs} + \text{H}_2\text{O} = \text{HH}$ reactions in water-rich mixtures. The implication is that the availability of water to CIs in these solvent mixtures is a strong non-linear function of their water content. This finding challenges expectations because if β -C and the cosolvents were randomly mixed at the molecular scale, R vs x_w plots should have been nearly linear functions of x_w .

Noticeably, the R vs x_w plots of Fig. 2 are not smooth, i.e., show discernible structure, and depend on X in W:X mixtures. Regarding the source of these differences, note that AN, THF, and DIO can accept H-bonds from water (AN can also donate a H-bond) but their molecular dipole moments are quite different. AN is highly polar: $\mu = 3.44$ D, THF and W are less polar and have similar dipole moments: $\mu = 1.75$ D and 1.85 D, respectively, whereas symmetrical DIO has a small dipole moment: $\mu = 0.45$ D, due to non-planarity (Scheme 1). In Fig. 2(b), it is possible to distinguish that R in W:AN mixtures sharply increases 2.7 times in the narrow $0.13 \leq x_w \leq 0.19$ range, followed by a rather flat section (only a $\sim 30\%$ increase)



SCHEME 2. Formation of α -hydroxy-hydroperoxides (HH, detected as chloride adducts at m/z 305; 307) and functionalized-carboxylates (FC, m/z 251) from the ozonolysis of β -C at air–liquid interfaces. Structures shown are the representatives of possible isomers.

in the $0.19 \leq x_w \leq 0.49$ interval, before increasing again ~ 10 times at $x_w = 0.74$. These transitions bring to mind a recent photoelectron ejection spectroscopic study, which reported that the AN population of the outermost interfacial layers of W:AN mixtures decreased in three stages as a function of x_w .⁶⁵ The same study also reported the orientations of AN molecules at the interface, but gave no indication about the composition or the dimensions of the interfacial W:AN domains in each region. The molecular dynamics calculations later suggested that AN on the surface of W:AN mixtures is present as discrete clusters rather than as continuous films.⁶⁴ Significantly, Fig. 2(b) shows that among these cosolvents, the R vs x_w dependence in aqueous mixtures of the least polar DIO is the steepest and displays transitions as well. These findings are apparently in line with X-ray diffraction, mass spectrometry, and NMR relaxation studies of bulk W:DIO mixtures that reported three structural regions in the $0.1 \leq x_w \leq 0.7$, $0.8 \leq x_w \leq 0.9$, and $0.93 \leq x_w \leq 1.0$ ranges.⁹⁰ The transitions observed in Fig. 2 are physically significant, but do not detract from the fact that the large R variations as a function of x_w are well represented ($\rho^2 = 0.98$) by exponentials: $R = a + b \exp(c x_w)$ or by strong power functions: $R = a + b x_w^n$ ($n \geq 7$) in all W:X mixtures.

Based on our chemical experiments, we can only speculate about the origin of the significant differences between R vs x_w dependences in W:AN, W:THF, and W:DIO mixtures. The overarching message is that microheterogeneity, both in the bulk and at interfaces, is an intrinsic feature of water:hydrotrope mixtures.¹⁶ The existence of sharp transitions between different sub-phases in our ternary OL:W:X mixtures indicates that the competition between the stability of the H-bonding network of water, and the strength of W_iX_j vs X_k associations is a discontinuous function of cluster size and composition. The solvent effects on CIs reactions in our ternary mixtures are closely related to the x_w -dependences of the intensity, peak emission wavelength, and decay lifetime of 7-cyanoindole fluorescence in W:AN, W:THF, and W:DIO mixtures (see below).²⁸ The correlations between the degree of

dissociation of 2-naphthol and the rates of *tert*-butyl chloride hydrolysis with the composition of W_iAN_j clusters as functions of x_w in W:AN mixtures (Fig. S14)^{23,25,26,91,92} support our interpretation that the R vs x_w trends of Fig. 2 reflect the composition of β -C: X_mW_n clusters. Note that not only AN, THF, and DIO but also dimethylsulfoxide, methanol, ethanol, isopropanol, and *tert*-butanol all show evidence of microheterogeneity.²⁸ The implication is that microheterogeneity (both at shorter and longer time scales) may be a general phenomenon of water-hydrotrope mixtures and, therefore, that “molecularly homogeneous” aqueous mixtures may not exist. There is also thermodynamic evidence on microheterogeneity in aqueous mixtures.^{62,93}

Our findings clearly show that the hydrophobic β -C solute is not randomly dispersed at the air–liquid interfaces of mixtures of water with amphiphilic organic solvents.¹² The strongly non-linear R vs x_w dependences in the three W:X solvent mixtures further reveal that the hydration of CIs takes place on intact liquid microjets of well-defined x_w values (see Validation Experiments in the supplementary material). Since the hydrophobic C_{15} -diolefin should be preferentially associated with the X hydrotropes rather than with water, CIs are expected to be generated in X-rich β -C: X_mW_n mixed clusters.⁹⁴ The observed non-linear R vs x_w dependences therefore reflect the accessibility of water to CIs embedded in X_mW_n clusters. X_mW_n clusters are entities that persist longer than milliseconds, as revealed by the time dependence of the second relaxation mode of the autocorrelation function of scattered light.^{18–20,95} The lifetimes of X_mW_n clusters are therefore much longer than the time window of our experiments: $\tau_R < 10 \mu s$. The characteristic times of X and W exchange between X_mW_n clusters with the rest of the solution, however, are much shorter, as revealed by the fact that 7-cyanoindole fluorescence in W:AN, W:THF, and W:DIO mixtures decays as single exponentials in the 2–15 ns range.²⁸ Single exponential decays suggest that 7-cyanoindole fluoresces while embedded in X_mW_n clusters of average composition, rather than from an ensemble of X_mW_n clusters. Therefore,

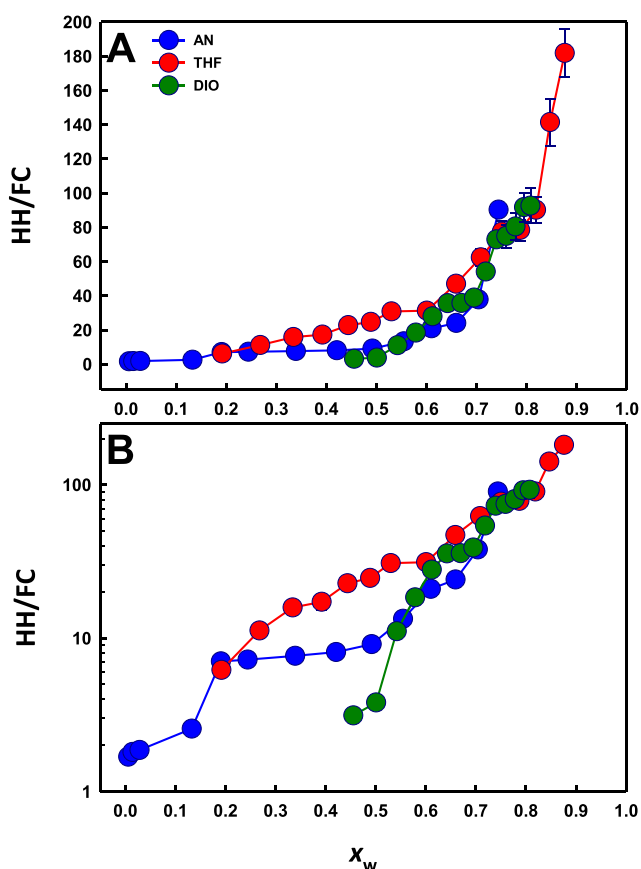


FIG. 2. Ratios $R = \text{HH/FC} = (305 + 307)/251$ of mass spectral signal intensities as linear (a) and semi-logarithmic (b) functions of the molar fraction of bulk water, x_w , in experiments involving exposing microjets of (1 mM β -caryophyllene + 0.2 mM NaCl) solutions in W:AN, W:THF or W:DIO to $E = 5.3 \times 10^{11}$ $\text{O}_3(\text{g})$ molecules $\text{cm}^{-3} \text{s}^{-1}$. The error bars are within the symbols size in most cases. The lines are visual guides. Note that the HH/FC vs x_w functions are not smooth and depend sensitively on the identity of the X cosolvent. However, the overall trends could be equally well-fitted ($\rho^2 = 0.98$) by exponentials: $R = \text{HH/FC} = a + b \exp(c x_w)$ or power functions: $R = \text{HH/FC} = a + b x_w^n$ ($n \geq 7$) in all W:X mixtures. The regression lines are not shown.

averaging takes place via X and W exchange in the sub-ns range scale. Since our thermalized C_{15} CIs should isomerize unimolecularly to FC in a longer time scale, we conclude that CIs reactions take place in persistent X_mW_n clusters of average composition. Therefore, the R vs x_w plots of Fig. 2 should be related to the average composition of persistent $\beta\text{-C}:\text{X}_m\text{W}_n$ clusters as a function of x_w . This relationship, however, may not be direct. Since hydrophobic $\beta\text{-C}$ is expected to be preferentially bound to X molecules, water accessibility could be limited by the reorganization dynamics of $\beta\text{-C}:\text{X}_m\text{W}_n$ clusters. Water in large, rigid clusters could take longer time to come in close contact with CIs than in smaller but more flexible clusters of lower water content. Thus, in principle, water accessibility to CIs should be determined by the convolution of average cluster composition and cluster reorganization rates.

In summary, our experiments reveal that chemical reactions on the surface of “internally mixed” aqueous organic solutions take place in clusters of compositions that are very different from the compositions of the bulk liquids.

This conclusion is particularly relevant to chemistry on the surface of aqueous organic aerosols. Atmospheric air-liquid interfaces function as substrates for the adsorption of organic gases and vapors, and as reaction media for their subsequent chemical transformations.^{41,56,96–102} Our findings show that surface microheterogeneity plays an unanticipated role in atmospherically relevant reactions on the surface of “internally mixed” aqueous/organic aerosol particles. Chemical transformations of amphiphilic, hydrotropic, and hydrophobic species on the surface of aqueous organic aerosols should be expected as a rule to take place in microheterogeneous rather than homogeneous media. The implication is that reaction rates and products could be relatively independent of aerosol water content and/or relative humidity over wide ranges.^{101,103–108}

IV. CONCLUSIONS

The products of fast, competing reactions on the surface of water:hydrotrope mixtures were monitored as functions of mixture composition via online surface-sensitive mass spectrometry. It was found that the relative yields of the functionalized carboxylates (FC) and α -hydroxy-hydroperoxides (HH) products of competing isomerization vs hydration reactions of Criegee intermediates depend quasi-exponentially rather than linearly on water molar fraction. Quasi-exponential dependences, however, are not smooth but display significant structure, which is consistent with the existence of distinct interfacial sub-phases suggested by other techniques. Our study demonstrates that the surfaces of “internally mixed” aqueous/organic solvent mixtures are not “well mixed,” and are indirectly related to the composition of the bulk solvents. The notion that microheterogeneity is an intrinsic feature of interfaces separating aqueous organic mixtures from hydrophobic phases and could dramatically influence interfacial reactions should become a key concept in various fields.

SUPPLEMENTARY MATERIAL

See [supplementary material](#) for validation experiments and additional experimental data. This material is available free of charge via the Internet.

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