COMMUNICATION

Hydrogenation and electrocatalytic reduction of carbon dioxide to formate with a single Co catalyst

Received 00th January 20xx,
Accepted 00th January 20xx

DOI: 10.1039/x0xx00000x

Fang Wang, Austin Cannon, Moumita Bhattacharya, Robert Baumgarten, Ryan T. Vanderlinden, and Caroline T. Saouma*

A cobalt(I) complex is shown to be capable of both electrocatalytic reduction and hydrogenation of CO_2 to formate. Several proposed intermediates are characterized and thus form the basis for a proposed mechanism that allows for the dual reactivity: reduction of CO_2 via H_2 addition, and H^+/e^- equivalents. The work makes use of a novel tris(phosphino) ligand. When a pendent amine is atttached to the ligand, no change in catalytic reactivity is observed.

Reduction of CO_2 to fuels and/or fuel precursors is integral to minimize global warming and advance future energy schemes.[1] One approach is to use H_2 to hydrogenate CO_2 to formic acid (FA) or MeOH, though challenges include transportation of the gas, the necessity of high pressures of H_2 and/or elevated temperatures required for many catalysts, and improving catalyst performance. Lifecycle analysis for CO_2 hydrogenation to FA (using a homogeneous catalyst) suggests that this approach can decrease the net greenhouse gas emissions when compared to FA production from CO.[2] Despite showing an improvement, this analysis also indicates that H_2 production accounts for a significant amount of the emissions.[3]

An alternative method is the solar-derived electrochemical reduction of CO_2.[4] While the $2e^-/2H^+$ reduction of CO_2 to CO is well-established, reduction to FA has proven more challenging at homogenous systems due to competing H_2 production.[5] This latter reaction can be thought of as an electrochemical hydrogenation. It necessitates a proton source capable of generating a metal hydride, and that the subsequent insertion of CO_2 be favoured over loss of H_2; both reactions have similar thermodynamic driving forces.[4] Berben's group showed that selectivity for FA over H_2 can achieved by exclusion of a pendant proton shuttle, which alters the kinetics of proton transfer to the active-site.[6] Recently, the groups of Kubiak[7] and Yang[5] have shown how H_2 production can be circumvented on thermodynamic arguments if the product is formate and not FA.

Given the widespread utility of hydrogenations, advancement of electrochemical alternatives may have significant impact. Waymouth showed that a Ru transfer hydrogenation catalyst can serve as an electrocatalyst for the oxidation of alcohols to ketones;[8] in this system a cationic solvent species is proposed as an intermediate. With regards to CO_2 conversion to formate, Meyer and Brookhart reported that the $2e^-/1H^+$ reduction of a PCP-ligated $IrH(MeCN)_2^{+}$[9] gives a species capable of inserting CO_2, with subsequent formate release. This system also necessitates a labile solvent molecule to avoid an 18-electron species that cannot be reduced. The limited literature examples of electrocatlytic hydrogenations re-enforce the need for better understanding how the two mechanistic pathways intersect, as well as establishing catalyst design criteria that allows for the desired reactivity.

Herein we describe a new family of Co complexes that perform both hydrogenation and electrocatalytic reduction of CO_2 to formate with excellent product selectivity. To our knowledge, this is the first system capable of this dual reactivity: reaction of CO_2 with H_2 to give FA, and reaction of CO_2 with H^+ and e^- equivalents to selectively give FA (over competing H_2 production). Mechanistic studies indicate how the mechanisms are related. The ligand features a pendent amine that does not impact either reaction type.

To explore the dual reactivity of electrocatalytic reduction[10, 11] and hydrogenation[12-15] of CO_2, phosphine-ligated Co complexes were prepared. Electrocatalytic generation of Co-H is known to occur for proton reduction catalysts;[16] one example[17] employs a tris(phosphino) ligand that has also been shown to catalytically hydrogenate CO_2 to MeOH using Co.[18] To explore the role that pendent proton-relays may have on both catalytic pathways, a tris(phosphino) scaffold was developed that features a single pendent amine and is flexible in *mer*/*fac* coordination to the metal.

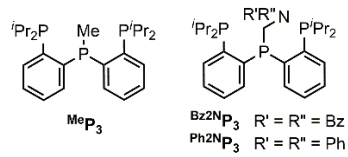


Chart 1. Ligands and abbreviations used in this study.

Department of Chemistry, University of Utah, 315 S. 1400 E., Salt Lake City, UT, 84112, USA. Email: caroline.saouma@utah.edu

Electronic Supplementary Information (ESI) available. See DOI: 10.1039/x0xx00000x. Crystal structures have been deposited to the CCDC, with deposition numbers 2010199-2010206, 2010332.

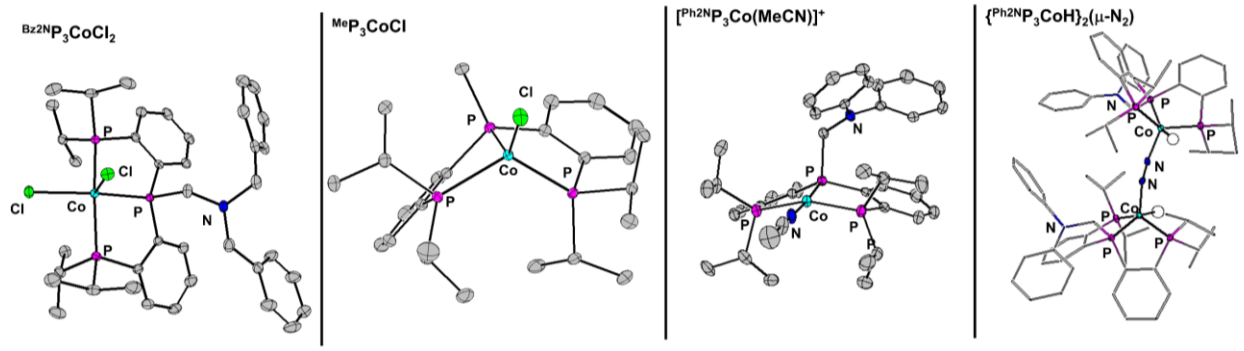
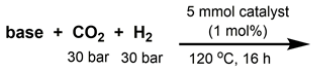


Figure 1. 50% Thermal ellipsoid plot of several complexes. All hydrogen atoms not located in the difference map are removed for clarity. Only the cation of **[(^{Ph2N}P$_3$Co(MeCN)][BArF]** is shown.

Tris(phosphino) ligands with a central phosphine that can be functionalized were prepared (Chart 1). The ligands that feature a pendent amine are readily prepared by addition of a suitable amine and paraformaldehyde to the precursor secondary phosphine (see ESI).[19] For this study, two tertiary amines (^{Bz2N}**P$_3$**, ^{Ph2N}**P$_3$**) were chosen for the pendent amine. A ligand with no amine, ^{Me}**P$_3$** was also synthesized.[20]

Metalation of the ligands is achieved by stirring equimolar ligand with $CoCl_2$ or $Co(PPh_3)_3Cl$ to give (**^{R}P$_3$)CoCl$_2$** or (**^{R}P$_3$)CoCl** respectively. Solid-state structures of (**^{Bz2N}P$_3$)CoCl$_2$**, (**^{Ph2N}P$_3$)CoCl$_2$**, (**^{Bz2N}P$_3$)CoCl**, (**^{Ph2N}P$_3$)CoCl** and (**^{Me}P$_3$)CoCl** were obtained, and exemplary structures shown in Figure 1. All of the Co(II) species feature two inner-sphere chloride ions, and have distorted square pyramidal geometry (τ ~ 0.15).[21] The Co(I) species are 4-coordinate and are best described as distorted tetrahedral (τ ~ 0.75).[22] In no instances does the amine nitrogen coordinate the metal centre.

Table 1.



Entry	Catalyst	Base	Equiv. of Base	Formate TON[c]
1[a]	**(^{Ph2N}P$_3$)CoCl**	K_3PO_4	100	12
2[a]	**(^{Ph2N}P$_3$)CoCl**	KO^{t}Bu	100	65 (±6)
3[a]	**(^{Ph2N}P$_3$)CoCl**	DBU	100	51
4[b]	**[(^{Ph2N}P$_3$)Co(MeCN)][BArF]**	KO^{t}Bu	100	111 (±4)[d]
5[b]	**(^{Bz2N}P$_3$)CoCl**	KO^{t}Bu	100	124 (±9)[d]
6[b]	**(^{Bz2N}P$_3$)CoCl**	KO^{t}Bu	200	37 (±8)
7[b]	**(^{Bz2N}P$_3$)CoCl**	KO^{t}Bu	2000	242
8[b]	**(^{Me}P$_3$)CoCl**	KO^{t}Bu	100	122 (±6)[d]
9[b]	**(^{Bz2N}P$_3$)CoCl**	--	--	9
10[b,e]	--	KO^{t}Bu	100	20

[a]Reactions run in 3 mL THF, at 150 °C for 20 h. [b]Reactions run in 10 mL THF with the conditions provided in the scheme. [c]Average of two runs with standard deviation in parenthesis. If no standard deviation, single run. [e]Run for 20 h.

Hydrogenation of CO_2 under basic conditions was then explored with the Co(I) complexes (Table 1). No MeOH was observed by GC analysis, and the only product detected was formate.

Entries 1-3 of Table 1 indicate that the base strength impacts catalysis; increasing the base strength from K_3PO_4 to KO^{t}Bu gives higher turnover number (TON), suggesting that a deprotonation event may limit the catalysis. Two of the Co(I)Cl species gave ~ quantitative TON with respect to base (entries 5, 8) when 100 eq KO^{t}BU is employed. Increasing the amount of base diminishes catalysis (entries 5-7); a color change is noted when large amounts of base are added to the catalyst solution, suggestive of catalyst degradation. Recycling studies indicate that a viable catalyst is present at the end of catalysis, though the paramagnetic nature of the complexes makes it difficult to ascertain the identity (see ESI).

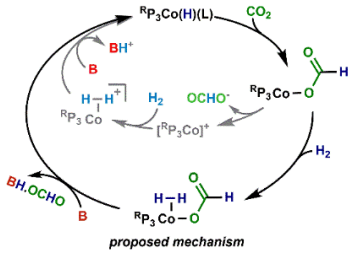
To determine if the chloride is pertinent to catalysis, **^{Ph2N}P$_3$CoCl** was treated with NaBArF to give cationic **[(^{Ph2N}P$_3$Co(MeCN)][BArF]** (BArF = tetrakis[3,5-bis(trifluoromethyl)phenyl]borate), the structure of which is shown in Figure 1. Using the cation as a catalyst for the hydrogenation of CO_2 improves the TON compared to that of the chloride (entries 2 and 4) and indicates that ~quantitative conversion is possible using complexes of all ligands examined.

A mechanism that includes CO_2 insertion into a Co-H (to give Co-OCHO) seems plausible and hence this reactivity was explored. Addition of 2 eq of $NaBHEt_3$ to a stirring THF solution of **^{Ph2N}P$_3$CoCl** at -70 °C results in formation of a new species. The ^{31}P NMR spectrum shows two singlets at 101.6 and 98.9 ppm, suggesting that a single diamagnetic Co(I) species has formed. The corresponding ^{1}H NMR spectrum shows a doublet of triplet at -11.35 ppm, consistent with a Co-H and IR analysis shows a stretch at 2082 cm^{-1}. Vapor diffusion of benzene into heptane gave crystals suitable for diffraction, and the solid-state structure indicates the formation of a dimeric species, **{^{Ph2N}P$_3$CoH}$_2$(μ-N$_2$)** (Figure 1). Each Co is 5-coordinate with the hydride in the plane of the three phosphines and the N_2 coordinating in the apical position.

To determine if the hydride is sufficiently hydridic to insert CO_2, 0.85 atm of CO_2 was added to a solution of **{^{Ph2N}P$_3$CoH}$_2$(μ-N$_2$)**. NMR analysis shows complete conversion of the diamagnetic hydride to a new paramagnetic species. Now, the IR spectrum shows disappearance of the hydride resonance and a new peak at 1628 cm^{-1}, consistent with formation of a species such as **^{Ph2N}P$_3$Co-OCHO**. The related species, $(PPh_3)_3Co(H)(N_2)$, inserts CO_2 to give $(PPh_3)_3Co(OCHO)$,[23] the corresponding formate stretch at 1620 cm^{-1}.

A proposed mechanism is shown in Scheme 1. Entry into the catalytic cycle occurs from the reaction of **^{R}P$_3$CoCl** with H_2 and base, which would give a 5-coordinate species such as

$^{R}P_3Co(H)(L)$ (L = N_2 or solvent). Subsequent insertion of CO_2 gives **$^{R}P_3Co$-OCHO**. The resulting 16-electron species **$^{R}P_3Co$-OCHO** may then coordinate H_2 to give the proposed 18-electron **$^{R}P_3Co(H_2)$-OCHO**. Base-mediated deprotonation of the bound H_2 coupled to formate loss regenerates **$^{R}P_3Co(H)(L)$**. Another mechanism would be deprotonation of the bound H_2 by the inner-sphere formate to generate **$^{R}P_3Co(H)(L)$** and formic acid (B = formate in Scheme 1). Indeed, this may explain the > 100 TON (entries 4, 5, 8) when only 100 equivalents of base is added, as well as the low TON obtained in the absence of base (entry 9). This is proposed to be a minor pathway that is viable in the absence of base, that proceeds with slower kinetics.



Scheme 1. Proposed mechanism for hydrogenation of CO_2 to formate, with alternate mechanism shown in grey.

An alternative mechanism that has been proposed for related complexes is that the formate dissociates from the Co to give **$[^{Ph2N}P_3Co(MeCN)]^+$** and free formate (Scheme 1, inner pathway). The cation then coordinates H_2 and base-mediated deprotonation of the bound H_2[12] or oxidative addition product ensues.[13, 14] Given that both **$[^{Ph2N}P_3Co(MeCN)]^+$** and **$^{R}P_3Co$-OCHO** are stable, this seems unlikely. Moreover, in related work on Ru we have shown that binding of formate to a cationic Ru centre is favorable.[24]

Though it was envisioned that the pendent amine may facilitate deprotonation of H_2 via hydrogen-bonding,[25] the similar catalytic performances amongst all the catalysts indicates that such an effect, if present, is irrelevant in the catalytic hydrogenation.

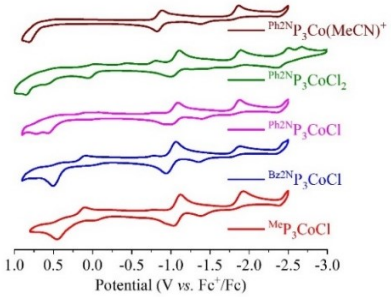


Figure 2. Cyclic voltammograms of the various Co complexes. Conditions: GC working electrode, 0.1 M $TBAPF_6$ electrolyte in MeCN. Initial scan is in the positive direction.

With the feasibility to hydrogenate CO_2 to formate established, we sought to determine if electrocatalytic hydrogenation of CO_2 is also accessible. The cyclic voltammograms (CVs) of several complexes are shown in Figure 2; the initial scan direction was positive. Cationic **$[^{Ph2N}P_3Co(MeCN)]^+$** shows a reversible reduction at -0.863 V (vs $Fc^{+/0}$) that corresponds to the Co(II/I) couple. A second reduction event occurs at ~ -1.9 V, with a return oxidation at ~ -1.3 V. This tentatively is assigned to the Co(I/0) couple. Consistent with this, the reaction of **$^{Ph2N}P_3CoCl$** with Na/Hg gives **$^{Ph2N}P_3Co(N_2)$** (see ESI); the irreversible nature of the couple is attributed to N_2 coordination upon reduction. The irreversibility may also be attributed to different numbers and types of L-type ligands upon reduction (L = MeCN or N_2). The CVs of **$^{Ph2N}P_3CoCl_2$** and **$^{Ph2N}P_3CoCl$** are similar, and show a quasi-reversible reduction at -1.050 V. Both CVs show peaks that correspond to **$[^{Ph2N}P_3Co(MeCN)]^+$**, consistent with chloride loss upon reduction to Co(0). Notably, the peaks that correspond to the second reduction are super-imposable with those in the CV of **$[(^{Ph2N}P_3Co(MeCN)]^+$**. Modest changes in the reduction potentials is anticipated as the R group on the central phosphine is varied.[10] Indeed, the Co(II/I) potential of **$^{Bz2N}P_3CoCl$** occurring at -1.013 V, and that of **$^{Me}P_3CoCl$** at -1.089 V.

In the presence of 50 eq water, no noticeable changes to the CVs are observed for all **$^{R}P_3CoCl$** (Figure 3 and ESI). However, upon addition of CO_2, a catalytic current is observed, suggesting formation of CO or formate. The current increases further with 1617 eq H_2O (3% by volume, see ESI) at potentials close to the $Co^{I/0}$ couple; other Co electrocatalysts reduce CO_2 at potentials well shifted from the redox couples of the catalyst.[10] Indeed, rapid current enhancement at ~ -2.5 V suggests that there may be two pathways for reduction.

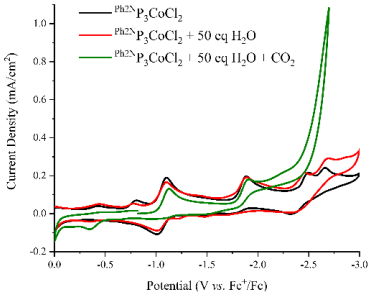


Figure 3. Cyclic voltammograms of 1 mM **$^{Ph2N}P_3CoCl_2$** under various conditions. **(black):** under N_2; **(red):** in the presence of 50 eq H_2O; **(green):** in the presence of 50 eq H_2O and CO_2. Conditions: GC working electrode, 0.1 M $TBAPF_6$ electrolyte in MeCN solvent, scan rate of 0.06 V/s, initial scan is positive

Table 2. Product distribution of the controlled potential electrolysis (CPE) experiments.

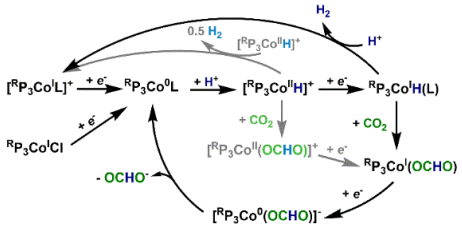
Complex	Time (s)	Charge Passed (C)	FE[b] H_2	FE[b] CO	FE[b] $HCOO^-$
$^{Me}P_3CoCl$	1062 (±127)	1.6 (±0.5)	2.5 (±2.8)	nd[c]	58 (±15)
$^{Bz2N}P_3CoCl$	1005 (±89)	1.9 (±0.3)	3.2 (±0.4)	nd[c]	36 (±18)

[a]Conditions: CPE carried out at -2.1 V *vs.* Fc^+/Fc with 1 mM catalysts using graphite rod working electrode in 0.1 M $TBAPF_6$/MeCN solvent with 380 rpm stirring; [b]Faradaic efficiency; average of four runs, standard deviation given in parenthases. [c]None detected.

Controlled potential electrolysis with 3% water and 0.85 atm of CO_2 at -2.1 V vs. Fc/Fc^+ was conducted using **$^{Me}P_3CoCl$** and **$^{Bz2N}P_3CoCl$** as the catalyst (Table 2). This potential is chosen to see if catalysis does occur near the reduction potential of the complexes. In both instances, no CO was detected in the

headspace, and H_2 is only produced in small quantities with both catalysts. No other gaseous products are produced, and the solution phase shows the presence of formate and MeOH. The catalysts are stable, as ascertained by CVs after electrolysis and the steady current (see ESI).

A proposed mechanism is shown in Scheme 2. Upon reduction to $^{R}P_3Co^0L$ (L = N_2 or solvent), protonation ensues to give $[^{R}P_3Co^{II}H]^+$. This is reduced at the electrode to give $^{R}P_3Co^{I}H$, which then inserts CO_2 to give $^{R}P_3Co^{I}(OCHO)$; this sequence being identical to that in the hydrogenation reaction. Reduction to $^{R}P_3Co^0(OCHO)$ and subsequent formate release then regenerates $^{R}P_3Co^0L$. The lack of lability of the formate in $^{R}P_3Co^{I}(OCHO)$ and the cathodic potential of catalysis is consistent with this EC mechanism. While we cannot rule out initial CO_2 insertion to $[^{R}P_3Co^{II}H]^+$ followed be reduction, this reactivity is not known for this system. Finally, the pH of the solution increases during the course of catalysis, indicating that formate and not formic acid is lost.



Scheme 2. Proposed (black) & alternative (gray) mechanisms for CO_2 and H^+ reduction.

Regarding proton reduction, the following can be gleaned. As $^{R}P_3Co^{I}H$ is stable, we rule out a bimetallic mechanism that would generate $^{R}P_3Co^0L$. A bimetallic mechanism from $[^{R}P_3Co^{II}H]^+$ also seems unlikely, as it would not explain why H_2 is only produced in the presence of pendent amines; these species are also more sterically encumbering and hence should minimize this pathway on steric grounds. Protonation of $^{R}P_3Co^{I}H$ seems most plausible, and literature precedence is consistent with proton relays enhancing H_2 production over formate.[6]

The work presented here provides the first example of a well-defined catalyst that can hydrogenate CO_2 to formate and electrocatalytically reduce CO_2 to formate. Notably, the latter reaction occurs with good selectivity for formate, and H_2 is only produced in the presence of a pendent amine. The stability of several intermediates, including a cationic solvent species that seems essential for the dual reactivity, allows for further mechanistic insight. Optimization studies and detailed mechanistic work is ongoing.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

The authors gratefully acknowledge the NSF CAREER (1945646) and start-up funding. A.C. received funding from NSF REU (1659579) and NIH (1S10OD25241-01) for NMR. We acknowledge Dr. Stacey J. Smith and BYU for the use of their single crystal X-ray instrumentation

Notes and references

1 1. W.-H. Wang, Y. Himeda, J. T. Muckerman, G. F. Manbeck and E. Fujita, *Chem. Rev.*, 2015, **115**, 12936-12973.
2 2. A. Sternberg, C. M. Jens and A. Bardow, *Green Chemistry*, 2017, **19**, 2244-2259.
3 3. J. Artz, T. E. Müller, K. Thenert, J. Kleinekorte, R. Meys, A. Sternberg, A. Bardow and W. Leitner, *Chem. Rev.*, 2018, **118**, 434-504.
4 4. R. Francke, B. Schille and M. Roemelt, *Chem. Rev.*, 2018, **118**, 4631-4701.
5 5. B. M. Ceballos and J. Y. Yang, *Proceedings of the National Academy of Sciences*, 2018, **115**, 12686-12691.
6 6. N. D. Loewen, E. J. Thompson, M. Kagan, C. L. Banales, T. W. Myers, J. C. Fettinger and L. A. Berben, *Chemical Science*, 2016, **7**, 2728-2735.
7 7. K. M. Waldie, A. L. Ostericher, M. H. Reineke, A. F. Sasayama and C. P. Kubiak, *ACS Catalysis*, 2018, **8**, 1313-1324.
8 8. K. M. Waldie, K. R. Flajslik, E. McLoughlin, C. E. D. Chidsey and R. M. Waymouth, *J. Am. Chem. Soc.*, 2017, **139**, 738-748.
9 9. P. Kang, C. Cheng, Z. Chen, C. K. Schauer, T. J. Meyer and M. Brookhart, *J. Am. Chem. Soc.*, 2012, **134**, 5500-5503.
10 10. S. Roy, B. Sharma, J. Pécaut, P. Simon, M. Fontecave, P. D. Tran, E. Derat and V. Artero, *J. Am. Chem. Soc.*, 2017, **139**, 3685-3696.
11 11. F.-W. Liu, J. Bi, Y. Sun, S. Luo and P. Kang, *ChemSusChem*, 2018, **11**, 1656-1663.
12 12. A. Z. Spentzos, C. L. Barnes and W. H. Bernskoetter, *Inorg. Chem.*, 2016, **55**, 8225-8233.
13 13. M. S. Jeletic, M. T. Mock, A. M. Appel and J. C. Linehan, *J. Am. Chem. Soc.*, 2013, **135**, 11533-11536.
14 14. M. V. Vollmer, J. Ye, J. C. Linehan, B. J. Graziano, A. Preston, E. S. Wiedner and C. C. Lu, *ACS Catalysis*, 2020, **10**, 2459-2470.
15 15. Y. M. Badiei, W.-H. Wang, J. F. Hull, D. J. Szalda, J. T. Muckerman, Y. Himeda and E. Fujita, *Inorg. Chem.*, 2013, **52**, 12576-12586.
16 16. E. S. Rountree, D. J. Martin, B. D. McCarthy and J. L. Dempsey, *ACS Catalysis*, 2016, **6**, 3326-3335.
17 17. S. C. Marinescu, J. R. Winkler and H. B. Gray, *Proceedings of the National Academy of Sciences*, 2012, **109**, 15127-15131.
18 18. J. Schneidewind, R. Adam, W. Baumann, R. Jackstell and M. Beller, *Angew. Chem. Int. Ed.*, 2017, **56**, 1890-1893.
19 19. E. Payet, A. Auffrant, X. F. Le Goff and P. L. Floch, *J. Organomet. Chem.*, 2010, **695**, 1499-1506.
20 20. Y.-E. Kim, J. Kim and Y. Lee, *Chem. Commun.*, 2014, **50**, 11458-11461.
21 21. A. W. Addison, T. N. Rao, J. Reedijk, J. van Rijn and G. C. Verschoor, *J. Chem. Soc., Dalton Trans.*, 1984, DOI: 10.1039/DT9840001349, 1349-1356.
22 22. L. Yang, D. R. Powell and R. P. Houser, *Dalton Transactions*, 2007, DOI: 10.1039/B617136B, 955-964.
23 23. L. S. Pu, A. Yamamoto and S. Ikeda, *J. Am. Chem. Soc.*, 1968, **90**, 3896-3896.
24 24. C. L. Mathis, J. Geary, Y. Ardon, M. S. Reese, M. A. Philliber, R. T. VanderLinden and C. T. Saouma, *J. Am. Chem. Soc.*, 2019, **141**, 14317-14328.
25 25. M. Rakowski Dubois and D. L. Dubois, *Acc. Chem. Res.*, 2009, **42**, 1974-1982.

