

Flux Crystal Growth, Structure, and Optical Properties of the New Germanium Oxysulfide $\text{La}_4(\text{GeS}_2\text{O}_2)_3$

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Keywords. flux crystal growth, oxysulfide, wide-gap insulator, germanium

Abstract

Single crystals of a new germanium oxysulfide $\text{La}_4(\text{GeS}_2\text{O}_2)_3$ were grown out of a $\text{BaCl}_2\text{-CaCl}_2$ eutectic flux and the crystal structure was determined via single-crystal X-ray diffraction. Ivory-white $\text{La}_4(\text{GeS}_2\text{O}_2)_3$ crystallizes in the centrosymmetric space group $R\bar{3}$ with lattice constants of $a = b = 16.8283(3) \text{ \AA}$, and $c = 8.4140(2) \text{ \AA}$. $\text{La}_4(\text{GeS}_2\text{O}_2)_3$ exhibits complex three-dimensional anion order in the framework composed of unusual GeS_2O_2 tetrahedra and three types of La-centered polyhedra. Moreover, the triangular arrangement of the GeS_2O_2 tetrahedra around a columnar structure of alternating face-sharing LaO_{12} and LaS_6O_6 units resembles that of GeO_4 tetrahedra in the apatite germanate $\text{La}_{9.33}(\text{GeO}_4)_6\text{O}_2$, which features face-sharing LaO_9 columns. The combination of UV-Vis absorption measurements and first-principle calculations revealed the existence of an indirect optical band gap ($E_g = 3.67 \text{ eV}$) between the valence band maximum composed of S $3p$ orbitals and the conduction band minimum composed of La $3d$, La $4f$, S $3p$, and Ge $4s$ orbitals.

Introduction

Mixed-anion compounds, in which two or more anionic species are incorporated into one crystal structure, have provided a playground for solid-state chemists to explore unique properties. Differences in the essential characteristics of anions, i.e., ionic radius, electronegativity, oxidation state, and polarizability, often lead to short or long-range ordering of the anions, resulting in the formation of new anion sublattices and building blocks.¹⁻³ Mixed-anion structures containing oxide ions and different period 3–5 anions, such as oxide-sulfide and oxide-arsenide combinations, exclusively exhibit long-range anion ordering, particularly anion segregation in two dimensional (2D) layers.⁴⁻⁷ These anion segregations can be explained by the hard-soft acid-base (HSAB) theory,⁸ in which highly electronegative/polarizable cations prefer to bond with highly electronegative/polarizable anions. However, exceptions to the HSAB theory exist when differences in electronegativity (or polarizability) are small or when the concentration of a given anion is high compared to that of the other.

Oxychalcogenide compounds have been extensively studied because of their interesting optical properties. The incorporation of chalcogenide ions (*Ch*: S, Se, Te) into oxide sublattices raises the valence band maximum (VBM) and narrows the optical band gap, allowing for the tuning of the optical band gaps.⁹⁻¹⁰ These band-gap engineered properties are useful for designing visible-light active photocatalysts, as observed in many layered oxysulfides^{9, 11-13}. The high polarizability of the chalcogenide ions also makes non-centrosymmetric oxychalcogenides promising non-linear optical materials exhibiting large second-harmonic generation (SHG) responses. Recently, our group reported a new polar zinc oxysulfide, SrZn₂S₂O, that crystallizes in the space group *Pmn*2₁ (no. 31) and that exhibited a large SHG response with phase matchability in contrast to non-phase-matching CaZnSO, space group *P*6₃*mc* (no. 186).¹⁴⁻¹⁵ Moreover, SrZn₂S₂O was found

to be capable of reducing and oxidizing water under ultraviolet light irradiation with greater photocatalytic activity and photo-corrosion resistance than found for the mono-anionic Zn compounds, ZnO and ZnS.¹⁶ The structure of SrZn₂S₂O contains close-packed corrugated double layers of ZnS₃O tetrahedra that are present in two distinct orientations, in which sulfide and oxide ions form a unique 2D arrangement in the same layer. CaZnSO possesses a similar ZnS₃O building unit, however, the connectivities of these units differ: each of these two anions is isolated into separate 2D layers. Although the origins of the phase matchability and high photocatalytic activity in SrZn₂S₂O is not yet fully understood, it appears that the new anion-order design and arrangement of the anionic building blocks plays a crucial role in generating the unusual optical properties.

The flux crystal growth method has been recognized as a useful technique for synthesizing new mixed-anion compounds, most of which do not readily occur in single-crystal form.¹⁷⁻²⁰ In particular, alkali or alkaline earth halides, which can be easily removed with water, are widely used as fluxes for the synthesis of a wide range of new oxychalcogenides, including non-linear optical materials^{14, 21}, low-dimensional magnets²²⁻²³, and superconductors²⁴. Following the discovery of SrZn₂S₂O, we applied this approach to develop new optical oxysulfide materials based on *d*¹⁰ metal-centered tetrahedra. Herein, we report the flux crystal growth of the new germanium-containing oxysulfide La₄(GeS₂O₂)₃ from a BaCl₂-CaCl₂ eutectic mixture. Its crystal structure, characterized by unusual GeS₂O₂ tetrahedra, electronic structure, and optical properties, are discussed.

Experimental

Synthesis. Single crystals of $\text{La}_4(\text{GeS}_2\text{O}_2)_3$ were grown out of a eutectic $\text{BaCl}_2/\text{CaCl}_2$ flux. First, 0.5 mmol of La_2S_3 (Alfa Aesar 99%), 1.0 mmol of GeO_2 (Rare Metallic 99.99%), 3.13 mmol of BaCl_2 (Rare Metallic 99.9%), and 3.13 mmol of CaCl_2 (Rare Metallic 99.99%) were loaded into a magnesia crucible in an Ar-filled glovebox. The crucible was sealed inside a silica tube that was evacuated to 10^{-4} Pa. The tube containing the starting materials was heated in a muffle furnace to 800 °C at 150 °C/h, held for 24 h, cooled to 500 °C at 5 °C/h, and then cooled to room temperature by turning the heater off. The product was washed in sonicated water and extracted from the flux. Ivory-white block $\text{La}_4(\text{GeS}_2\text{O}_2)_3$ crystals, along with colorless transparent platelet LaOCl crystals, were collected via vacuum filtration. The typical dimensions of the $\text{La}_4(\text{GeS}_2\text{O}_2)_3$ single crystals were $0.3 \times 0.3 \times 0.3 \text{ mm}^3$ (Figure 1). Elemental analysis by energy-dispersive X-ray spectroscopy (EDS) indicated a La:Ge:S atomic ratio of approximately 4.0:2.9:5.4, which is consistent with the chemical composition of the germanium oxysulfide phase. The crystal structure of $\text{La}_4(\text{GeS}_2\text{O}_2)_3$ was determined via single-crystal X-ray diffraction.

Single crystal structure determination. X-ray intensity data from a yellowish block crystal were collected at 301(2) K using a Bruker D8 QUEST diffractometer equipped with a PHOTON-100 CMOS area detector and an Incoatec microfocus (Mo $K\alpha$ radiation, $\lambda = 0.71073 \text{ \AA}$). The data collection covered 100% of the reciprocal space to $2\theta_{\text{max}} = 72.6^\circ$, with an average reflection redundancy of 14.7 and $R_{\text{int}} = 0.057$ after absorption correction. The raw area-detector data frames were reduced and corrected for absorption effects using the SAINT+ and SADABS programs.^{25 26} Final unit cell parameters were determined by least-squares refinement of 9908 reflections taken from the data set. An initial structural model was obtained using SHELXT.²⁷

Subsequent difference Fourier calculations and full-matrix least-squares refinement against F^2 were performed with SHELXL-2018 using the ShelXle interface.²⁸

Powder X-ray diffraction. Single crystals of $\text{La}_4(\text{GeS}_2\text{O}_2)_3$ were crushed with an agate mortar and pestle to obtain fine powders for collecting synchrotron X-ray powder diffraction (SXRD) patterns and UV-Vis diffuse reflectance spectra. SXRD measurements were performed at room temperature using a one-dimensional detector installed on the NIMS BL15XU beamline at SPring-8 in Japan. The synchrotron radiation X-rays were monochromatized to a wavelength of 0.65298 Å. The powdered sample of $\text{La}_4(\text{GeS}_2\text{O}_2)_3$ was loaded into a glass capillary tube (internal diameter: 0.2 mm). The diffraction data were collected in 0.003° increments over the 2–60° range and analyzed by Rietveld refinement using the RIETAN-FP program.²⁹

UV-Vis diffuse reflectance spectra and SEM-EDS. Diffuse reflectance spectroscopy measurements were performed at room temperature using a Shimadzu UV-2600 spectrophotometer equipped with an ISR-2600Plus integration sphere. The diffuse reflectance data were internally converted to absorbance by the instrument using the Kubelka-Munk function. Single crystals of $\text{La}_4(\text{GeS}_2\text{O}_2)_3$ were analyzed by energy dispersive X-ray spectroscopy (EDS) equipped with a scanning electron microscopy (TESCAN Vega-e SBU). All crystals were mounted on carbon tape, and the analysis was carried using an accelerating voltage of 20 kV.

Density functional theory calculations.

To be described by Kuwabara san.

First-principles total energy calculations of $\text{La}_4(\text{GeS}_2\text{O}_2)_3$ were performed using the projector augmented wave method³⁰⁻³¹ as implemented in the Vienna Ab-initio Simulation Package (VASP).³²⁻³⁴ In the present study, the cut-off energy for the plane wave basis was 550 eV. The

exchange-correlation electron interaction potentials were considered via PBEsol-type generalized gradient approximation (GGA).³⁵

Results and discussion

The single-crystal structure analysis revealed that $\text{La}_4(\text{GeS}_2\text{O}_2)_3$ crystallizes in the space group $R\text{-}3$ (no. 148) with lattice constants of $a = b = 16.8283(3)$ Å and $c = 8.4140(2)$ Å. No significant deviation from full occupancy was detected for any atom. Thus, their occupancies were fixed at unity at the final stage of the structure refinements. The details of the structural refinement are listed in Table 1. The crystallographic parameters, including the atomic coordinates and atomic displacement parameters, are summarized in Table 2; anisotropic displacement parameters are listed in Table 3.

Figure 2 shows the SXRD patterns collected from a powder sample obtained by finely grinding hand-picked single crystals of $\text{La}_4(\text{GeS}_2\text{O}_2)_3$. The structural model determined by the single-crystal structure analysis was used for the Rietveld refinement. The calculated pattern well reproduced the observed patterns as the refinement readily converged with good reliability indexes, $R_{\text{wp}} = 4.90\%$, $R_{\text{B}} = 6.88\%$, and $R_{\text{F}} = 3.78\%$. Table S1 summarizes the final refined crystallographic data.

As shown in Figure 3, $\text{La}_4(\text{GeS}_2\text{O}_2)_3$ crystallizes in a unique three-dimensional crystal structure consisting of La- and Ge-centered polyhedra. Three crystallographically non-equivalent La1, La2, and La3 atoms occupy $18f$, $3b$, and $3a$ Wyckoff positions, respectively; in contrast, there is only one single germanium site occupying the $18f$ position. Two anions, namely O^{2-} and S^{2-} , are fully ordered in the structure. Figure 4 shows the local coordination environments around the La

and Ge atoms. Of these three lanthanum sites, La1 and La2 are surrounded by two different anions. La1 is coordinated by three oxygen atoms at an average distance of $d_{\text{La-O1, La-O2}} \sim 2.45 \text{ \AA}$ and six sulfur atoms at an average distance of $d_{\text{La-S1}} \sim 2.99 \text{ \AA}$ and $d_{\text{La-S2}} \sim 3.28 \text{ \AA}$, forming a distorted gyroelongated square pyramid. Although these bond distances vary within 7 to 12% of the reference values derived from Shannon's ionic radii³⁶, the bond-valence-sum (BVS) calculation³⁷ for La1 yielded 2.91, which is consistent with the nominal oxidation state of the lanthanum ions. La2 is icosahedrally coordinated by six oxygen atoms at $d_{\text{La-O1}} = 2.6123(19) \text{ \AA}$ and six sulfur atoms at $d_{\text{La-S2}} = 3.4262(6) \text{ \AA}$. The La–O1 bond distance is 5% shorter than that of La–S2, but the La–S2 bond distance is 7% longer than the sum of the ionic radii of La^{3+} and $\text{O}^{2-}/\text{S}^{2-}$. In contrast to these two lanthanum atoms, La3 is 12-fold coordinated with only oxygen atoms, forming a relatively symmetric icosahedron ($d_{\text{La-O1}} = 2.7159(19) \text{ \AA}$, $d_{\text{La-O2}} = 2.7524(16) \text{ \AA}$). The bond distances between lanthanum and oxygen atoms agree well, as they are within 1% of the sum of their ionic radii. The BVS values calculated for La2 and La3 were 2.54 and 2.63, respectively, much lower than the expected values. Typically, a BVS value lower than the ideal valence is often interpreted as indicating the existence of underbonding (or tensile stress) between the metal center and the surrounding ligands. However, this understanding is not applicable to the bond distances determined by XRD analysis. However, the reason for the low BVS values for La2 and La3 remains unclear. The germanium atom was tetrahedrally coordinated by two oxygen atoms at $d_{\text{Ge-O1}} = 1.7721(15) \text{ \AA}$ and $d_{\text{Ge-O2}} = 1.7908(18) \text{ \AA}$ and two sulfur atoms at $d_{\text{Ge-S1}} = 2.1600(10) \text{ \AA}$ and $d_{\text{Ge-S2}} = 2.1707(7) \text{ \AA}$. These bond distances are consistent with those calculated from the sum of Ge^{4+} and La^{3+} ionic radii with O^{2-} and S^{2-} ionic radii, and the BVS calculation for Ge gave a reasonable value of 4.15. The S–Ge–S bond angle

of $124.56(3)^\circ$ is much larger than the O–Ge–O bond angle of $99.89(8)^\circ$. This can be rationalized by considering the electrostatic repulsion between the sulfide ions with the larger ionic radii.

Germanium ions coordinated tetrahedrally by two sulfide and two oxide ions are rare compared to GeS_3O tetrahedra, which is often stabilized in oxysulfides having the melilite structures[ref]. To the best of our knowledge, only two other germanium-based compounds with a similar coordination geometry around the Ge center have been synthesized:

$\text{Na}_2\text{GeS}_2(\text{OH})_2 \cdot 5\text{H}_2\text{O}$ ³⁸ and AGeS_2O ($A = \text{Sr}, \text{Ba}$)²¹. The Ge–S and Ge–O distances and the S–Ge–S and O–Ge–O bond angles of these compounds are similar to those in $\text{La}_4(\text{GeS}_2\text{O}_2)_3$ (Table S2), but the GeS_2O_2 tetrahedron in $\text{Na}_2\text{GeS}_2(\text{OH})_2 \cdot 5\text{H}_2\text{O}$ is slightly elongated with smaller S–Ge–S and O–Ge–O bond angles.

Figure 5 shows the linkage between the La- and Ge-centered polyhedra. $\text{La}_2\text{S}_6\text{O}_6$ and La_3O_{12} polyhedra share a common O_3 face to form one-dimensional (1D) chains along the c direction. Each La-based 1D chain is surrounded by $\text{La}_1\text{S}_6\text{O}_3$ and GeS_2O_2 polyhedra in a complex manner. The $\text{La}_2\text{S}_6\text{O}_6$ icosahedron are connected to six $\text{La}_1\text{S}_6\text{O}_3$ polyhedra and six GeS_2O_2 tetrahedra by edge-sharing through pairs of oxygen and sulfur atoms. These La_1 -based polyhedra are separated by Ge and La_2 atoms. In contrast, the La_3O_{12} icosahedron shares O_3 faces with six $\text{La}_1\text{S}_6\text{O}_3$ polyhedra and O_2 edges with six GeS_2O_2 tetrahedra, both of which have two different independent z coordinates. Moreover, these six $\text{La}_1\text{S}_6\text{O}_3$ polyhedra are linked with each other by sharing edges formed by a pair of oxygen and sulfur atoms. All GeS_2O_2 tetrahedra share edges with $\text{La}_1\text{S}_6\text{O}_3$ and $\text{La}_3\text{O}_{12}/\text{La}_2\text{S}_6\text{O}_6$ polyhedra.

It is interesting to compare the present oxysulfide with the germanium oxyapatite $\text{La}_{9.33}(\text{GeO}_4)_6\text{O}_2$ that forms in the hexagonal space group $P6_3/m$.³⁹⁻⁴⁰ Despite the long-range anion-ordered structure and the completely different chemical composition of $\text{La}_4(\text{GeS}_2\text{O}_2)_3$, the

oxysulfide phase of this compound exhibits Ge- and La-centered polyhedra arrangements similar to those found in the oxyapatite structure. As shown in **Figure 7**, the La_2O_9 polyhedra in the structure of $\text{La}_{9.33}(\text{GeO}_4)_6\text{O}_2$ form an infinite 1D chain along the c direction by sharing their common O_3 faces as in $\text{La}_4(\text{GeS}_2\text{O}_2)_3$. Moreover, these 1D chains are surrounded by six GeO_4 tetrahedra and six La1O_7 polyhedra with two different z coordinates, which are linked with La1O_7 through shared oxygen edges and vertexes, respectively. In both $\text{La}_4(\text{GeS}_2\text{O}_2)_3$ and $\text{La}_{9.33}(\text{GeO}_4)_6\text{O}_2$, the La_2/La_3 atoms are in the geometric center of the regular triangles formed by the Ge or La_1 atoms. However, the Ge- or La_1 -based triangles form a regular hexagon in the oxysulfide, as seen in the top-down view of these crystal structures shown in **Figure 6**; this is in contrast with the distorted hexagon in the oxyapatite structure.

Figure 8 shows the UV-Vis absorption spectrum of a powdered $\text{La}_4(\text{GeS}_2\text{O}_2)_3$ sample. The absorption data were obtained by converting the sample's reflectance spectrum using the Kubelka-Munk transformation.⁴¹ The absorption curve presents two steep increases near 2.5 eV and 3.5 eV, and the corresponding absorption edge energies, which were estimated by extrapolating the linear portions of the absorption curve to the x axis, were $E_{g1} = 2.23$ eV and $E_{g2} = 3.67$ eV, respectively. Given that the sample is ivory-white, the absorption at E_{B1} is probably due to defects from anion vacancies, although the XRD analysis could detect neither anion nor cation vacancies. The larger absorption edge at E_{G2} should correspond to the intrinsic optical band gap, which will be discussed below.

First-principles calculations were performed to gain a better understanding of the relationship between the electronic structure and optical properties of $\text{La}_4(\text{GeS}_2\text{O}_2)_3$. **Figure 9** shows the total and partial density of states (DOS) and the band dispersions of the oxysulfide compound. The total DOS reveals an energy gap with $E_g = 4.04$ eV, which is close to the E_{g2} value but

significantly larger than the E_{g1} value. This result strongly suggests that the absorption edge observed at E_{g2} is intrinsic to the electronic structure of the oxysulfide phase. The valance bands located between -6 and 0 eV are mainly derived from the O $2p$ and S $3p$ orbitals, which are weakly hybridized with the La $3p$, $3d$, and $4f$ orbitals. The S $3p$ bands dominate the upper energy region in the valence band maximum (VBM) compared with the O $2p$ bands in the lower energy region, as is often observed in oxyfluoride compounds.^{9-10,42} The conduction band minimum (CBM) is mainly composed of La $3d$ and $4f$, Ge $4s$, and S $3p$ orbitals. The band dispersion shown in **Figure 10** indicates that the VBM is located at point Γ and the CBM between points T and P. The optical absorption above E_{g2} in the UV-Vis spectrum results from the transition from the VBM to the CVM. The Ge $4s$ and $4p$ bands are located in the narrow energy regions between -8 and -7 eV and -6 and -4 eV, respectively, and are highly hybridized with O $2p$ and S $3p$ orbitals. The localized nature of these Ge bands is in contrast to the widely dispersive La bands located in the valence bands, resulting from the isolated linkages between GeS_2O_2 tetrahedra.

Conclusion

A new germanium oxysulfide $\text{La}_4(\text{GeS}_2\text{O}_2)_3$ with an indirect optical band gap of 3.67 eV was synthesized from a BaCl_2 - CaCl_2 eutectic molten salt. Single-crystal structure analysis demonstrated a unique structure comprised of unusual GeS_2O_2 tetrahedra linked with La-centered polyhedra through corner or edge sharing. The linkage between the Ge- and La-centered polyhedra is similar to that observed in the germanium oxyapatite, $\text{La}_{9.33}(\text{GeO}_4)_6\text{O}_2$, although the oxysulfide phase does not possess the so-called “apatite channel” in which oxide ions at $(0\ 0\ z)$ are mobile at high temperatures. Recently, several new classes of sulfide materials have shown promise for practical use as solid-state electrolytes for Li-ion batteries. In the structures of such

materials, highly polarizable sulfur ions play an important role in promoting high Li-ion conductivity. In $\text{La}_4(\text{GeS}_2\text{O}_2)_3$, each LaS_6O_3 polyhedron is connected to sulfide ions by edge-sharing, resulting in the formation of the sulfide-rich regions outside the alternating chains of face-sharing LaO_{12} and LaS_6O_6 units. Future studies should seek to determine if this three-dimensional anion order in $\text{La}_4(\text{GeS}_2\text{O}_2)_3$ creates Li-ion pathways via the substitution of Li for La atoms, as found in the $\text{La}_{1-x}\text{Li}_y\text{TiO}_3$ perovskite⁴³.

ASSOCIATED CONTENT

Supporting Information.

The following files are available free of charge.

Rietveld refinement details, Selected bond distances and bond angles (PDF)

Crystallographic information (CIF)

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Notes

The authors declare no competing financial interest.

ACKNOWLEDGMENT

We gratefully acknowledge a research grant from UofSC (NSF grant DMR-1806279). YT and AK acknowledge Grant-in-Aid for Scientific Research from Japan Society of for the Promotion

Science (JSPS) (no. 16H06438, 16H06439, 16H06440, 17H05493, 19H04711). KY acknowledges a research grant for Innovative Science and Technology Initiative for Security, from ATLA, Japan.

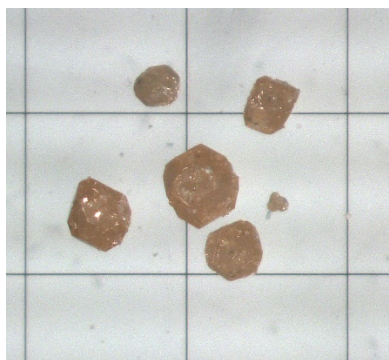


Figure 1. Photograph of single crystals of $\text{La}_4(\text{GeS}_2\text{O}_2)_3$ on a 1 mm grid-cell plate.

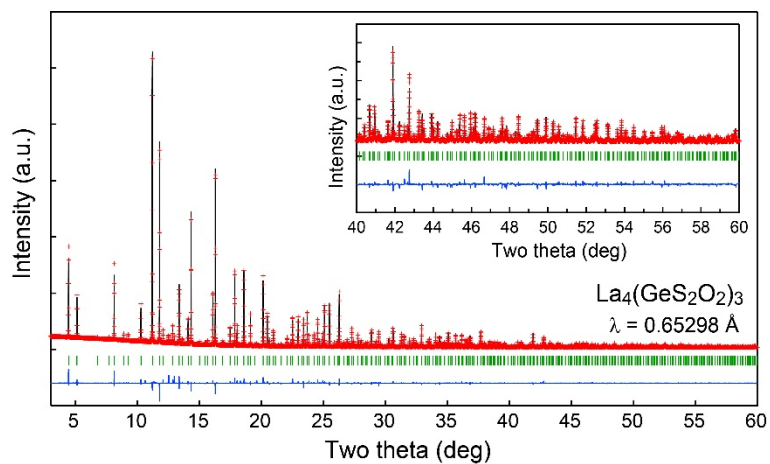


Figure 2. SXR D patterns of $\text{La}_4(\text{GeS}_2\text{O}_2)_3$ measured at room temperature. Obtained, calculated, and difference data are presented by cross marks (red), upper solid lines (black), bottom solid lines (blue), respectively. The vertical lines (green) represented the expected Bragg peak positions of $\text{La}_4(\text{GeS}_2\text{O}_2)_3$.

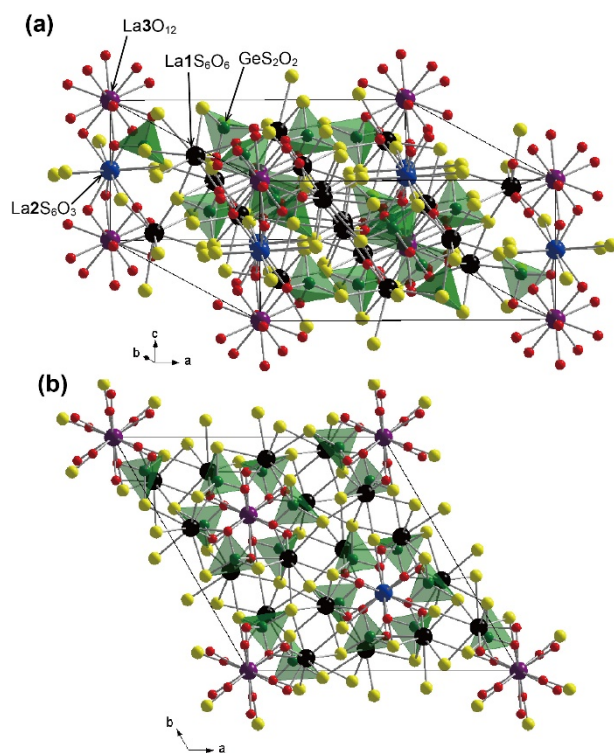


Figure 3. Crystal structure of $\text{La}_4(\text{GeS}_2\text{O}_2)_3$. Black, blue, purple, green, red, and yellow spheres stand for La1, La2, La3, Ge, O, and S atoms, respectively.

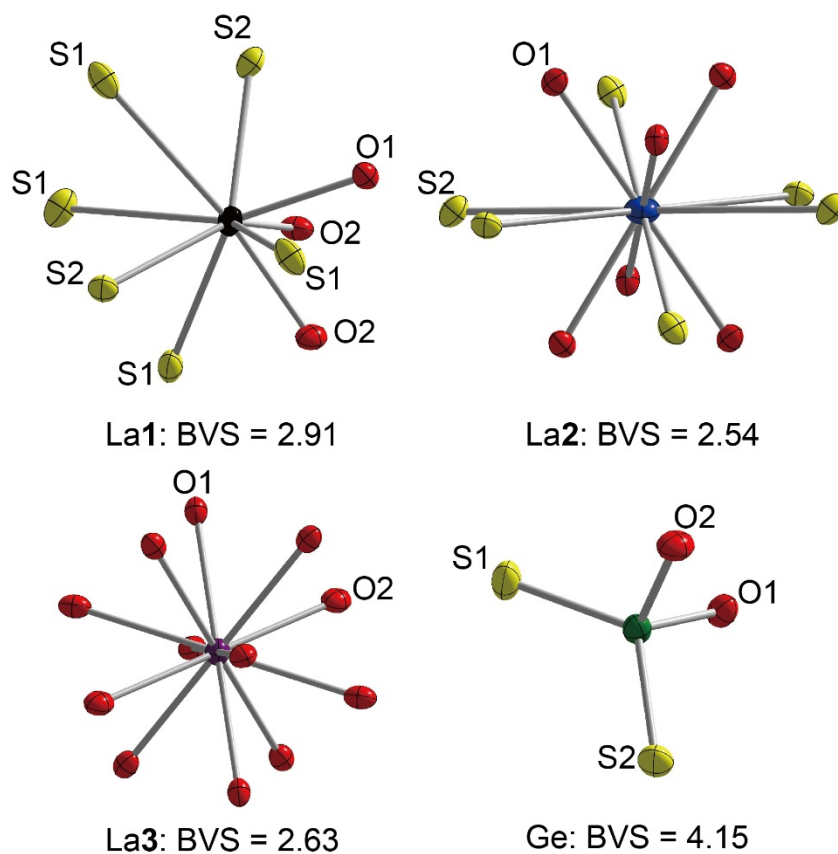


Figure 4. Local coordination environment around La1, La2, La3 and Ge atoms. Displacement ellipsoids are shown at the 90% level.

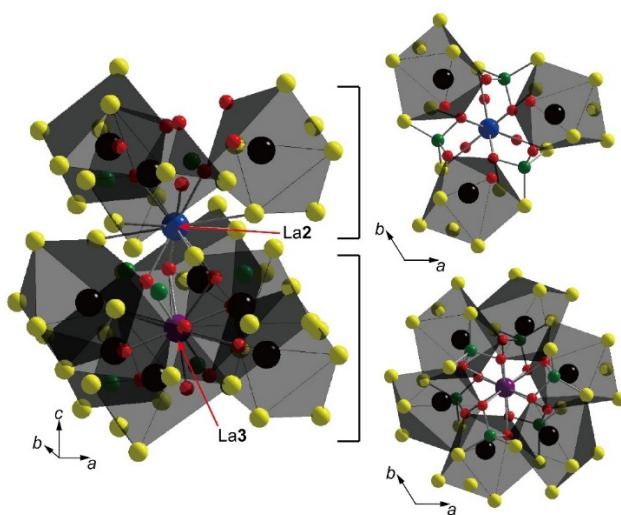


Figure 5. Linkage between La- and Ge-centered polyhedra in $\text{La}_4(\text{GeS}_2\text{O}_2)_3$ viewed along two different directions. La2 and La3-centered polyhedra are surrounded by six La1-centered polyhedra and six Ge-centered tetrahedra in a different arrangement.

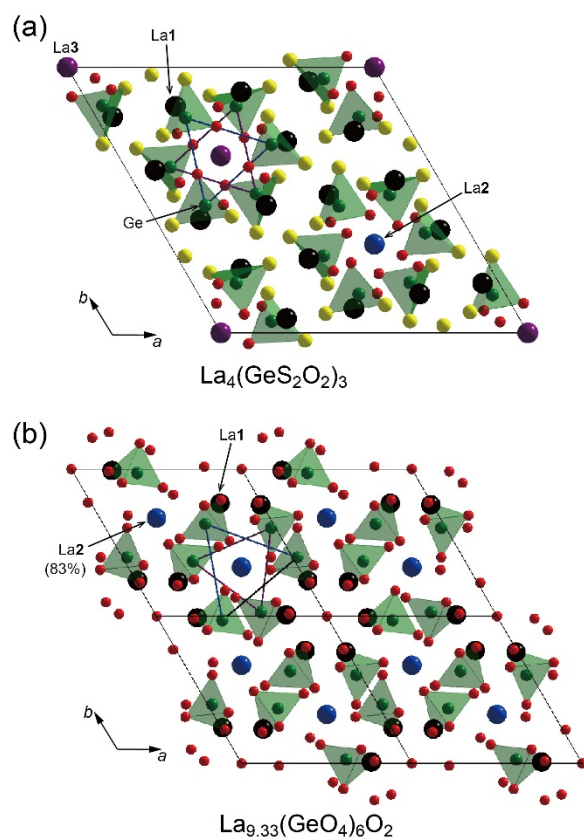


Figure 6. The crystal structure of $\text{La}_4(\text{GeS}_2\text{O}_2)_3$ (a) and $\text{La}_{9.33}(\text{GeO}_4)_6\text{O}_2$ (b) viewed along the c axis.

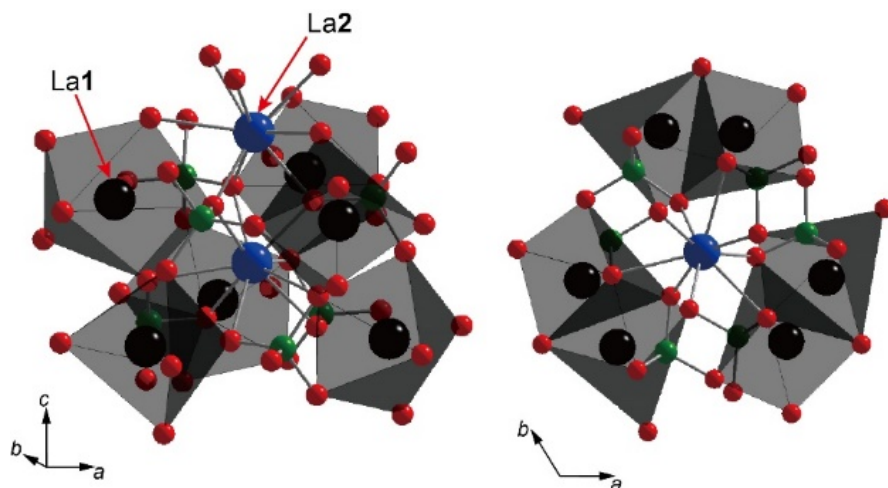


Figure 7. Linkage between La- and Ge-centered polyhedra in $\text{La}_{9.33}(\text{GeO}_4)_6\text{O}_2$ viewed along two different directions. La2-centered polyhedra are surrounded by six La1-centered polyhedra and six Ge-centered tetrahedra in a different arrangement.

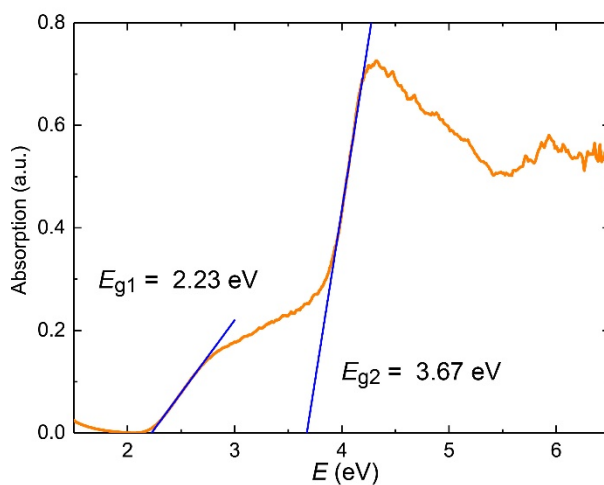


Figure 8. UV-vis optical absorption spectrum of $\text{La}_4(\text{GeS}_2\text{O}_2)_3$, which was converted from the diffuse reflectance data using the Kubelka–Munk function.

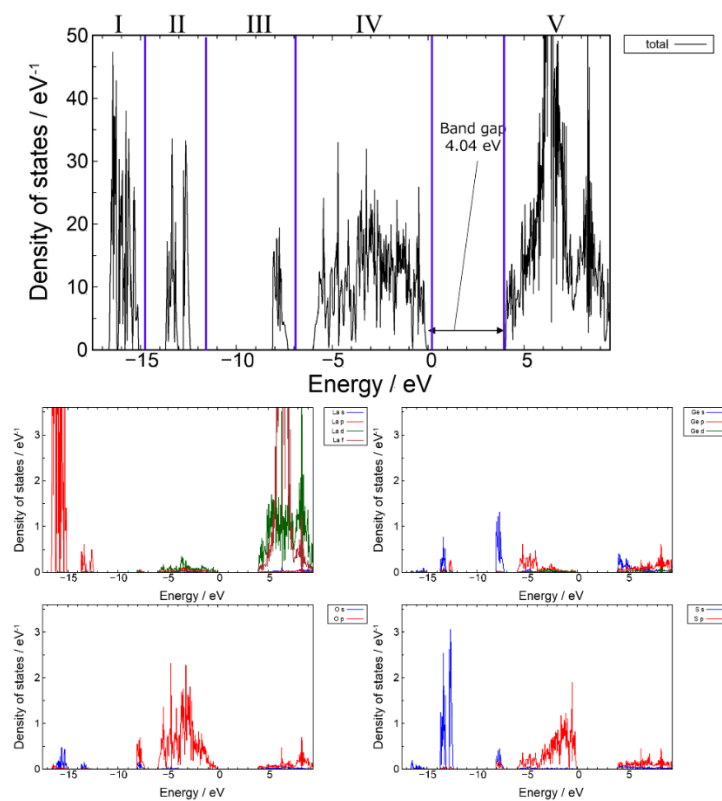


Figure 9. Total and partial density of states of $\text{La}_4(\text{GeS}_2\text{O}_2)_3$ obtained by first-principles calculations. (to be revised by Kuwabara san)

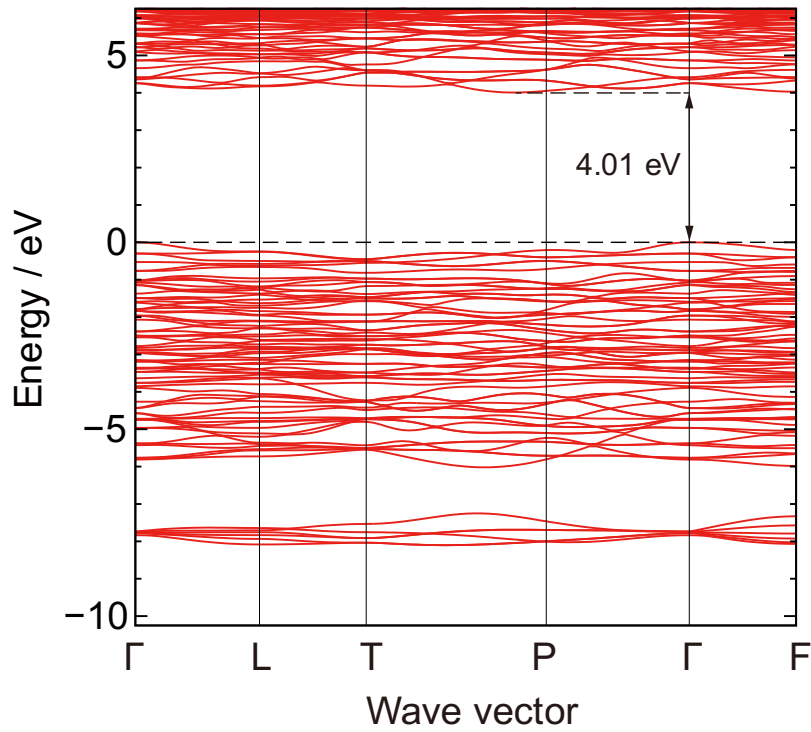


Figure 10. Band dispersion of $\text{La}_4(\text{GeS}_2\text{O}_2)_3$ obtained by first-principles calculations. (to be revised by Kuwabara san)

Table 1. Results of the Structure Refinement of $\text{La}_4(\text{GeS}_2\text{O}_2)_3$ Using Single-Crystal X-ray Diffraction

Formula	$\text{La}_4(\text{GeS}_2\text{O}_2)_3$
Formula weight	1061.77
Radiation	Mo $\text{K}\alpha$ ($\lambda=0.71073\text{\AA}$)
T (K)	301(2)
Crystal system	trigonal
Space group	$R\bar{3}$
a ($\text{\AA}$)	16.8283(3)
b ($\text{\AA}$)	16.8283(3)
c ($\text{\AA}$)	8.4140(2)

$V(\text{\AA}^3)$	2063.54(9)
Z	6
$D_{\text{calc}}(\text{g/cm}^3)$	5.126
F_{000}	2808
no. of measured reflns	32798
no. of unique reflns	2225
no. of observed reflns	1868
$(F^2 > 2\sigma(F^2))$	
$R_{\text{int}}(\%)$	5.70
$R[F^2 > 2\sigma(F^2)]/wR(F^2)$ (%)	2.22/3.45
GoF	1.062

Table 2. Crystallographic and Refinement Data Obtained from Single-Crystal Structure Analysis of $\text{La}_4(\text{GeS}_2\text{O}_2)_3$.

atom	site	x	y	z	Occp.	$U_{\text{eq}}(\text{\AA}^2 \times 10^2)$
La1	18 <i>f</i>	0.18672(2)	0.24104(2)	0.15890(2)	1	0.00902(3)
La2	3 <i>b</i>	0	0	1/2	1	0.01243(7)
La3	3 <i>a</i>	0	0	0	1	0.00792(6)
Ge	18 <i>f</i>	0.18713(2)	0.04347(2)	0.24397(3)	1	0.00720(5)
S1	18 <i>f</i>	0.26044(4)	-0.02833(4)	0.19177(7)	1	0.01228(11)
S2	18 <i>f</i>	0.22741(4)	0.14889(4)	0.42442(7)	1	0.01035(11)
O1	18 <i>f</i>	0.06959(11)	-0.04136(11)	0.25780(19)	1	0.0082(3)
O2	18 <i>f</i>	0.18462(11)	0.10106(11)	0.06669(18)	1	0.0099(3)

Table 3. Anisotropic Displacement Parameters U_{ij} ($\text{\AA}^2 \times 10^2$) for $\text{La}_4(\text{GeS}_2\text{O}_2)_3$ at 301 K.

U_{11}	U_{22}	U_{33}	U_{23}	U_{13}	U_{12}
0.00858(6)	0.00651(6)	0.01144(6)	-0.00042(4)	-0.00159(5)	0.00338(5)
0.01475(10)	0.01475(10))	0.00781(14)	0	0	0.00737(5)
0.00733(9)	0.00733(9)	0.00910(13)	0	0	0.00366(4)
0.00696(11)	0.00689(11))	0.00789(10)	-0.00046(8)	-0.00042(8)	0.00356(9)
0.0096(3)	0.0139(3)	0.0158(3)	-0.0047(2)	-0.0019(2)	0.0077(2)
0.0099(3)	0.0099(3)	0.0106(2)	-0.00343(19)	-0.00225(19)	0.0046(2)
0.0068(7)	0.0071(7)	0.0096(7)	-0.0008(6)	0.0010(6)	0.0025(6)
0.0136(8)	0.0109(8)	0.0070(7)	0.0010(6)	0.0002(6)	0.0074(7)

Table S1. Atomic Coordinates and Isotropic Displacement Parameters U_{iso} Refined from SXRD Data Collected from $\text{La}_4(\text{GeS}_2\text{O}_2)_3$ at Room Temperature.

Atom	Site	x	y	z	Occp. ^a	U_{iso} ($\text{\AA}^2 \times 10^2$)
La1	18 <i>f</i>	0.18640(7)	0.24115(7)	0.15842(12)	1	0.780(15)
La2	3 <i>b</i>	0	0	1/2	1	1.11(6)
La3	3 <i>a</i>	0	0	0	1	0.43(4)
Ge	18 <i>f</i>	0.18670(9)	0.04372(13)	0.24143(17)	1	0.45(3)
S1	18 <i>f</i>	0.2602(3)	-0.0272(2)	0.1898(4)	1	1.26(8)
S2	18 <i>f</i>	0.2273(2)	0.1531(2)	0.4181(4)	1	1.54(9)
O1	18 <i>f</i>	0.0679(5)	-0.0441(6)	0.2503(9)	1	0.18(14) <i>b</i>

O2 18^f 0.1753(5) 0.0996(7) 0.0684(9) 1 0.18 ^b

The space group is *R*-3 (No. 148) with $a = b = 16.82677(1)$ Å, $c = 8.41591(1)$ Å. *R* indices are $R_{\text{wp}} = 4.90\%$, $R_{\text{B}} = 6.88\%$, $R_{\text{F}} = 3.78\%$. ^a Site occupancy for each atom was fixed to unity. ^b The U_{iso} values for O1 and O2 atoms were constrained to the same value.

Table S2. Selected Interatomic Distances (Å) and Bond Angles (°) of La₄(GeS₂O₂)₃

	bond distance (Å)		bond angle (°)
La1-O1	2.473(2)	O2-La1-O2	73.36(6)
La1-O1	2.4228(19)	O2-La1-O1	73.67(5)
	2.464(3)	O2-La1-O1	74.01(5)
La1-S1	3.1571(7)	O1-La2-O1×6	65.62(6)
	3.2076(7)	O1-La2-O1×6	114.38(6)
	3.3468(7)	O1-La2-O1×2	180.000
	3.4280(8)	O1-La2-S2×4	61.99(4)
La1-S2	2.9877(8)	O1-La2-S2×6	118.01(4)
	2.9905(7)	O1-La2-S2×6	115.82(4)
La2-O1×6	2.6123(19)	O1-La2-S2×6	64.18(4)
La2-S2×6	3.4262(6)	O1-La2-S2×6	62.06(4)
La3-O1×6	2.7159(16)	O1-La2-S2×6	117.94(4)
La3-O2×6	2.7524(16)	S2-La2-S2×5	63.363(6)
Ge1-O1	1.7720(16)	S2-La2-S2×5	116.637(6)
Ge1-O2	1.7908(16)	S2-La2-S2×3	180.000
Ge1-S1	2.1600(6)	O1-La3-O1×6	117.18(5)
Ge1-S2	2.1701(6)	O1-La3-O1×6	62.82(5)
		O1-La3-O1×3	180.00(8)
		O1-La3-O2×6	120.17(5)

O1-La3-O2×6	65.81(5)
O1-La3-O2×6	114.19(5)
O1-La3-O2×6	59.83(5)
O2-La3-O2×6	115.95(2)
O2-La3-O2×6	64.05(2)
O1-Ge1-O2	99.89(7)
O1-Ge1-S1	106.40(5)
O2-Ge1-S1	107.94(5)
O1-Ge1-S2	112.29(5)
O2-Ge1-S2	102.86(5)
S1-Ge1-S2	124.56(2)

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