

Increasing the CO₂ Reduction Activity of Cobalt Phthalocyanine by Modulating the σ -donor Strength of Axially Coordinating Ligands

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ABSTRACT

Axial coordination of a pyridyl moieties to CoPc (either exogenous or within poly-4-vinylpyridine polymer) dramatically increases the complex's activity for the CO₂ reduction reaction (CO₂RR). It has been hypothesized that axial coordination to the Co active site leads to an increase in the Co dz² orbital energy, which increases the complex's nucleophilicity and facilitates CO₂ coordination compared to the parent CoPc complex. The magnitude of the energy increase in the Co dz² orbital should depend on the σ -donor strength of the axial ligand—a stronger σ -donating ligand (L) will increase the overall CO₂RR activity of axially coordinated CoPc(L) and vice versa. To test this, we have studied a series of CoPc(L) complexes where the σ -donor strength of L is varied. We show that CoPc(L) reduces CO₂ with an increased activity as the σ -donor ability of L is increased. These observed electrochemical activity trends

are correlated with computationally-derived CO₂ binding energy and charge transfer terms as a function of σ -donor strength. The findings of this study supports our hypothesis that the increased CO₂RR activity observed upon axial coordination to CoPc is due to the increased energy of the dz^2 orbital, and highlight an important design consideration for macrocyclic MN₄-based electrocatalysts.

KEYWORDS

CO₂ reduction reaction, cobalt phthalocyanine, electrocatalysis, axial coordination, sigma donor strength, density functional theory, absolutely localized molecular orbitals, charge transfer analysis.

INTRODUCTION

The electrochemical conversion of CO₂ into energy-rich fuels and chemicals in the CO₂ reduction reaction (CO₂RR) has gained significant interest as a promising strategy for effectively storing intermittent renewable energy from sources such as solar and wind.¹⁻¹⁴ There is an increasing drive to discover new catalysts that reduce CO₂ into value-added products or chemical feedstocks, including CO,¹²⁻¹⁴ with high activity and selectivity at low overpotentials. Physically-motivated design principles help guide the rational design of new CO₂RR catalyst systems, and these design principles can be tested and verified by evaluation of how systematic modifications of catalyst model systems influence electrocatalytic activity and selectivity. Surface-immobilized cobalt phthalocyanine (CoPc) is a model system of particular interest to the electrocatalysis community; it was first reported as an active catalyst for the CO₂RR over 35 years ago,^{15,16} and since then its catalytic activity has been extensively explored under various aqueous conditions.¹⁷⁻⁴²

Our group's interest in CoPc lay in its utility as a model system for studying the effects of polymer encapsulation on electrocatalytic performance. In particular, we are interested in probing how changing

a catalyst's coordination environment due to polymer encapsulation influences its electrocatalytic activity and selectivity for multielectron small-molecule transformations such as the CO₂RR.⁴³ Physisorbed CoPc on carbon electrodes with no polymer binder shows only modest activity for the CO₂RR in aqueous phosphate and citrate solutions with appreciable H₂ generated from the competitive hydrogen evolution reaction (HER).^{15,17,18,21,22} Work by our group^{17,19} and others^{21,22} has shown that encapsulating CoPc within coordinating polymers like poly-4-vinylpyridine (P4VP) results in CoPc-P4VP composite materials that operate with increased activity and selectivity for the electrochemical reduction of CO₂ to CO compared to the parent physisorbed CoPc complexes. This increase in activity and selectivity has been attributed to three distinct effects of the polymer on the coordination environment of the Co active site in CoPc-P4VP: i) axial-coordination of pyridyl in P4VP to the Co center increases the catalyst's nucleophilicity for CO₂ binding, ii) H-bonding interactions that stabilize reactive CO₂ intermediates, and iii) control of H⁺ delivery through a multi-site proton relay mechanism involving the polymer's pyridyl residues (Figure 1).^{17,18,22}

Axial coordination plays an important role in modulating the activity^{17,18,23,29} and mechanism^{18,19} of the CO₂RR by CoPc systems. Previous *in situ* electrochemical X-ray Absorption Near Edge Spectroscopy (XANES) measurements of deposited films of CoPc(L3) (L3 = pyridine) and CoPc-P4VP on carbon electrodes show evidence of axial coordination to the pyridyl moieties resulting in a 5-coordinate Co center.¹⁹ Importantly, the steady-state coordination environment about the Co center determined from these XANES measurements does not change under applied potential and electrocatalytic operating conditions for the CO₂RR.¹⁹ The formation of CoPc(L3) complex leads to a change in the rate-determining step of the CO₂RR mechanism from an initial CO₂ coordination step for CoPc to a subsequent protonation event for CoPc(L3).¹⁸

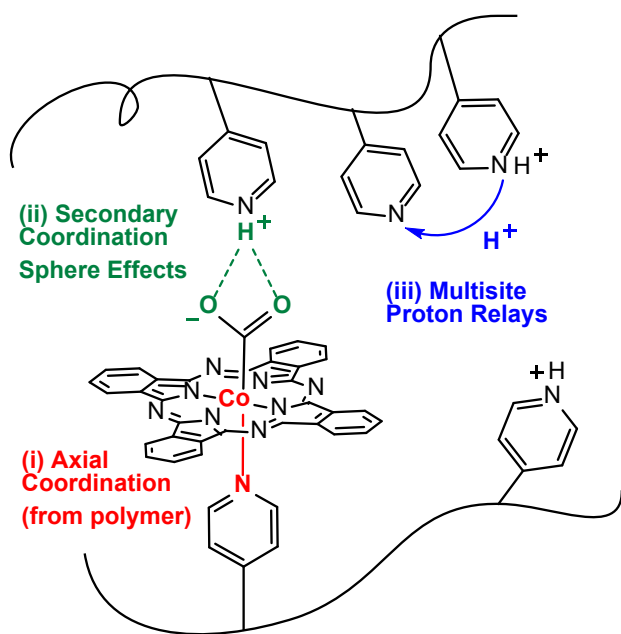


Figure 1. Illustration of CoPc-P4VP highlighting the i) axial-coordination to the pyridyl residues from the P4VP polymer in the primary coordination sphere, ii) the H-bonding stabilization of reduced CO₂ intermediates in the secondary coordination sphere, and iii) the controlled H⁺ delivery via a multisite proton relay in the outer coordination sphere. Adapted from Ref. 19 with permission from The Royal Society of Chemistry.

Based on the studies described above, we hypothesize that axial coordination of σ -donating ligands (L) to CoPc increases the nucleophilicity of the Co center in the CoPc(L) complexes by increasing the energy of the d_{z^2} orbitals. This, in turn, increases the ability of the Co center to coordinate and activate the Lewis acidic carbon of CO₂. Based on this hypothesis, we expect that increasing the σ -donating ability of the axial ligand on CoPc(L) should further increase the relative energy of the d_{z^2} orbital resulting in higher activity for the CO₂RR. In this report, we test this hypothesis by conducting a combined electrochemical and computational study on CoPc(L) complexes with various axial ligands (L). A list of the axial ligands used in this study are shown in Figure 2 in order of increasing σ -donor strength from

lowest (L1) to highest (L8) as indicated by the reported coupling constant between the donor group and the methylmercury (II) cation.⁴⁴⁻⁴⁸ We demonstrate that as the σ -donor strength of the axial ligand increases, there is a corresponding increase in the catalytic activity of the CoPc(L) for the CO₂RR. Using density functional theory analysis, we show that axial coordination does, in fact, increase the energy of the d_{z^2} orbital, and that the experimentally-observed increase in CO₂RR activity with increasing σ -donor strength of the axial ligand is correlated to an increased extent of charge transfer from the reduced CoPc(L) complex to CO₂ adduct. Overall, our work highlights the role of axial coordination in modulating activity for the CO₂RR and provides important insights into the design of highly active CoPc-based systems for selective CO₂ conversion.

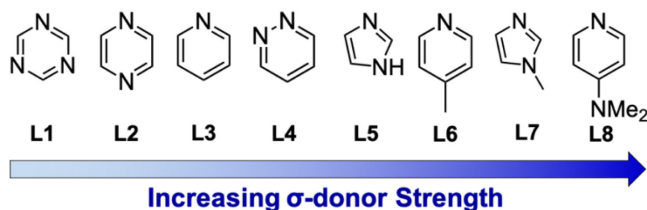


Figure 2. Axially-coordinating ligands (L) investigated in this study, listed in order of increasing σ -donor strength from lowest (L1) to highest (L8). L1 = 1,3,5-triazine, L2 = pyrazine, L3 = pyridine, L4 = pyridazine, L5 = imidazole, L6 = 4-methylpyridine, L7 = 1-methyl imidazole, and L8 = 4-dimethylaminopyridine.

EXPERIMENTAL

Materials and Supplies

All purchased chemicals were used as received unless otherwise specified. Cobalt phthalocyanine (CoPc, 97%), N,N-dimethylformamide (DMF, ACS grade), dimethyl sulfoxide (DMSO, ACS grade), ferrocene (Fc, 98%), 1,3,5-triazine (L1, 97%), pyrazine (L2, >99%), pyridine (L3, ACS grade, $\geq 99\%$),

pyridazine (L4, 98%), imidazole (L5, 99%), 4-methyl pyridine (L6, 99%), 1-methyl imidazole (L7, $\geq$ 99%), 4-dimethylaminopyridine (L8, $\geq$ 99%), sodium phosphate monobasic (BioXtra, $> 99.0\%$), ferrocenecarboxylic acid (97%), and sodium hydroxide (NaOH, BioXtra, $\geq 98\%$) were purchased from Sigma Aldrich. Nitric acid (HNO_3 , TraceMetal grade, 67-70%) was purchased from Fisher Scientific. Tetrabutylammonium hexafluorophosphate ($n\text{Bu}_4\text{NPF}_6$, $>98.0\%$) was purchased from TCI America and recrystallized from ethanol/ H_2O ($v/v = 8/1$) before use. Ethanol (EtOH , 200 proof, anhydrous) was purchased from Decon Laboratories, Inc., and used without further purification. Nafion-117 cation exchange membranes (Nafion) were purchased from the Fuel Cell Store. Nitrogen (N_2) was boil-off gas from a liquid nitrogen source. Compressed carbon dioxide gas (CO_2 , bone dry grade, 99.8%) and argon gas (Ar, pre-purified grade, 99.998%) were purchased from Cryogenic Gases. All water used in this study was ultrapure water ($18.2 \text{ M}\Omega\cdot\text{cm}$ resistivity) purified with a Thermo Scientific GenPure UV-TOC/UF xCAD-plus water purification system.

Preparation of CoPc(L)-modified Edge Plane Graphite Electrodes

Deposition solutions for CoPc and CoPc(L) were prepared based on an adaptation of previously described procedure.¹⁸ 0.05 mM CoPc/DMF deposition solutions were prepared by dissolving 0.13 mg of CoPc in 45 mL DMF and sonicating for at least 30 min until the CoPc was fully dispersed. To prepare CoPc(L)/DMF deposition solutions, 2.25 mmol of the axial ligand L was added to 45 mL of the 0.05 mM CoPc/DMF deposition solution, and the resulting solution was sonicated for at least 30 min until the CoPc and L were fully dispersed. Note that for the liquid ligands CoPc(L3), CoPc(L4), CoPc(L6) and CoPc(L7) in DMF deposition solution, the CoPc and Ligands were dissolved with 10mL of DMF, it was vortexed for 1 min and sonicated for 10 min prior the addition of the remaining 35mL DMF to favor the complex formation. For the solid ligands CoPc(L1), CoPc(L2), CoPc(L5) and CoPc(L8) in DMF deposition

solutions, ~60 min of sonication time was used to ensure full dispersion of CoPc and L. The final concentrations in the CoPc(L)/DMF deposition solutions was 0.05 mM CoPc and 50 mM L. The large excess of the axial ligand (~1,000×) used in the deposition solutions was to ensure equilibrium favored the axially-coordinated species.⁴⁹

Prior to deposition, 5 mm diameter edge plane graphite (EPG) disk electrodes (3.81 mm EPG disk encapsulated in epoxy, 0.114 cm² effective surface area, Pine Research Instrumentation) were manually polished with 600 grit SiC grinding paper (Buehler CarbiMet) followed by sonication in ultrapure water for ~ 1 min. The CoPc- and CoPc(L)-modified CoPc EPG working electrodes were then prepared by dropcasting 5 µL of the CoPc/DMF deposition solution or the CoPc(L)/DMF deposition solution onto the EPG electrode. The disk electrodes were then placed in a drying oven at ~60-70 °C for ~15 min to allow the solvent to evaporate. The resulting loading of CoPc or CoPc(L) on the electrode surface is calculated as 2.19×10^{-9} mol cm⁻². Note that we have previously demonstrated that CoPc(L3) films prepared as described above from operate with identical activity and mechanism as films prepared from independently prepared and isolated 5-coordinate CoPc(L3) complexes.¹⁸ Moreover, we have previously shown that CoPc(L3) films prepared as described above have similar XANES spectra as independently synthesized and characterized CoPc(L3).¹⁹ This evidence suggests that the CoPc(L3) films prepared in this study, and by analogy the CoPc(L) films, are most likely comprised of 5-coordinate CoPc(L) complexes.

Electrochemical Measurements

Electrochemical measurements were conducted using a Bio-Logic SP200 potentiostat, and data were recorded using the Bio-Logic EC-Lab software package. Reference electrodes were commercial saturated calomel electrodes (SCE, CH Instruments) externally referenced to ferrocenecarboxylic acid in 0.2 M phosphate buffer at pH 7 (0.284 V vs. SCE),⁵⁰ and auxiliary electrodes were carbon rods (99.999 %, Strem

Chemicals Inc.). Unless otherwise noted, all electrochemical measurements were conducted at least three times with independently prepared electrodes, and reported values are the averages of these repetitions with standard deviations included as the reported errors. The electrolyte solution used for all electrochemical studies was a pH 5 phosphate solution prepared from a 0.1 M NaH_2PO_4 solution adjusted to pH 5 by the addition of aqueous 1 M NaOH. The solution pH was confirmed using a Fisher Scientific Accumet AB200 pH meter with an Accumet pH/ATC Epoxy Body Combination Electrode calibrated with a 3-point calibration curve at pH = 4.01, 7.00, and 10.01.

Rotating Disk Electrode Voltammetry and Chronoamperometry Experiments. For rotating disk electrode voltammetry (RDEV) experiments and rotating disk electrode chronoamperometric step (RDE-CA) measurements, the working electrodes were mounted into a Pine Research Instrumentation E6-series Change Disk rotating disk electrode (RDE) assembly attached to an MSR rotator. RDEV and RDE-CA measurements were conducted at room temperature in a custom two-compartment glass H-cell as previously described.¹⁸ The working and reference electrodes were submerged in ~30 mL solution in the first compartment, and the auxiliary electrode was submerged in ~15 mL solution in the second compartment. The two compartments were separated by a Nafion cation exchange membrane. Both compartments were sparged with CO_2 for ~30 min prior to each set of measurements, and the headspace was blanketed with CO_2 during the measurements. The CO_2 used was first saturated with electrolyte solution by bubbling through a gas washing bottle filled with the same electrolyte solution used in the cell to minimize electrolyte evaporation in the cell during the measurements. RDEV measurements were measured at 1600 rpm rotation rate and a scan rate of 1 mV s^{-1} . RDE-CA measurements were conducted at 1600 rpm with 2-min potential steps from -0.95 V to -1.35 V vs. SCE at 0.05 V increments. The 1600 rpm rotation rate was meant to ensure steady-state delivery of substrate to our surface to allow for accurate comparisons of catalytic rates. Note that 1600 rpm does not imply kinetically-limiting conditions—mass

transport to catalyst sites in non-uniform catalyst-polymer composite films is not governed by simple Koutecký-Levich kinetics.⁵¹⁻⁵⁴ The uncompensated resistance of the cell (R_u) was measured with a single-point high-frequency impedance measurement, and RDEV and RDE-CA measurements were corrected for IR-drop at 85% through positive feedback using the Bio-Logic EC-Lab software. In general, our electrochemical cell for RDEV and RDE-CA measurements had $R_u \approx 100 \Omega$ when filled with our pH 5 phosphate electrolyte solution.

Controlled Potential Electrolysis Experiments. Controlled potential electrolyses (CPE) were conducted at room temperature in a custom, gas-tight two compartment U-cell as previously described.¹⁸ The working electrode was held in a RDE internal hardware kit (Pine Research Instrumentation) and mounted into a custom-designed PEEK sleeve. The first compartment of the U-cell held the working electrode, reference electrode, and a stir bar in 25 mL of electrolyte, and the second compartment held the auxiliary electrode in ~15 mL electrolyte. The two compartments were separated by a Nafion cation exchange membrane. The total volume of the first chamber was measured after each experiment by measuring the volume of H₂O necessary to completely fill the chamber when the cell was fully assembled with the working and reference electrodes, and the headspace volume for the CPE measurements was calculated by subtracting the electrolyte volume of 25 mL from the total volume of the first chamber. Prior to each experiment, both chambers were sparged with CO₂ for ~30 min and then the main chamber was sealed under CO₂ atmosphere. The uncompensated resistance of the cell was measured with a single-point high-frequency impedance measurement. In general, our electrochemical cell for CPE had $R_u = \sim 200 \Omega$ in pH 5 phosphate solution. The CPE measurements were conducted with no iR compensation, and the reported electrolysis potentials are the actual applied potentials.

Product detection and quantification were conducted as previously described.¹⁸ CO and H₂ were detected on a Thermo Scientific Trace 1310 Gas Chromatography system with two analyzer channels for

the detection of H₂ and C1-C2 products. After each electrolysis, a 5 mL aliquot of the headspace in the first compartment of the cell was collected using a Pressure-Lok gastight syringe (10 mL, Valco VICI Precision Sampling, Inc.), and the aliquot was injected into a 3 mL sample loop on the gas chromatograph. Using a custom valve system, column configuration, and method provided by Thermo Scientific, gases were separated such that H₂ was detected on the first channel using an Ar carrier gas and thermal conductivity detector (TCD), and all other gases were detected on the second channel using a He carrier gas and a TCD. The GC system was calibrated using calibration gas mixtures (SCOTTY Specialty Gas) at H₂ = 0.02, 0.05, 0.5, and 1% v/v, and CO = 0.02, 0.05, 0.5, 1, and 7% v/v. Chromatographs were analyzed using the Chromeleon Console Workstation software. To determine the concentration of dissolved products in the electrolyte solution, 1 mL aliquots of the post-electrolysis solutions were analyzed using a Thermo Scientific UltiMate 3000 HPLC system equipped with a refractive index detector (RFD), a 5 cm Thermo Scientific™ HyperREZ™ XP Carbohydrate H⁺ LC guard column and a 30 cm Thermo Scientific™ HyperREZ™ XP Carbohydrate H⁺ LC analytical column in series using a 5 mM H₂SO₄ aqueous mobile phase at a constant temperature of 50 °C. The detection limit of the HPLC for formic acid was determined to be 0.1 mM. No formic acid was observed after the CO₂RR CPEs.

Faradaic Efficiencies (FE) were calculated by dividing the moles of each product detected by the total moles of electrons based on the charge passed during the CPE according to Equation 1:

$$FE = \frac{\frac{V_{HS} \times C \times nF}{V_m}}{Q} \times 100\% \quad (1)$$

Here, V_{HS} is the volume of the headspace in the first chamber of the electrolysis cell, $V_m = 24.5 \text{ L mol}^{-1}$ is the molar volume of gas at 25 °C and 1.0 atm, C is the volume percent of the CO or H₂ product detected by GC, $n = 2$ is the number of electrons passed per reaction for the production of CO or H₂, $F = 96,485 \text{ C mol}^{-1}$ is Faraday's constant, and Q is the amount of charge passed during the CPE measurement.

Turnover Frequency for CO (TOF_{CO}) Calculations from RDE-CA Measurements. TOF_{CO} was calculated as previously described.¹⁸ Briefly, TOF_{CO} was determined by multiplying the measured current from the RDE-CAs by the Faradaic efficiency determined from CPE measurements at the same potential, and normalizing the resulting value for the coverage of catalyst as shown in Equation 2:

$$\text{TOF}_{\text{CO}} = \frac{|j| \times A \times \text{FE}_{\text{CO}}}{2F \times n_{\text{cat}}} \quad (2)$$

Here, $|j|$ is the absolute value of the current density for each RDE-CA step (A/cm²) at -1.25 V vs SCE, $A = 1.14$ cm² is the geometric surface area of the EPG electrode (0.114 cm²), FE_{CO} is the measured Faradaic efficiency of CO from CPE measurements at the same potential, $2F$ is two electrons multiplied by Faraday's constant, and $n_{\text{cat}} = 2.19 \times 10^{-9}$ mol cm⁻² is the loading of CoPc(L) on the surface. Note that there are inherent assumptions and limitations for any estimate of turnover frequency—in this case, the j and FE are taken from two different type of measurements at the same reduction potential (RDE-CA and CPE, respectively). However, we believe this measurement of TOF_{CO} provides a reasonable estimate of relative activity for adsorbed CoPc(L) species and related systems under hydrodynamic control. TOF_{CO} is the primary activity metric reported in the manuscript.

For completeness, we also calculated the turnover frequency for CO production based solely on CPE data (TOF_{CO, CPE}) as described in Equation 3.

$$\text{TOF}_{\text{CO, CPE}} = \frac{Q \times \text{FE}_{\text{CO}}}{2F \times t \times n_{\text{cat}}} \quad (3)$$

Here, Q is the charge passed during the CPE measurement (C) conducted at -1.25 V vs SCE, $A = 1.14$ cm² is the geometric surface area of the EPG electrode (0.114 cm²), FE_{CO} is the measured Faradaic efficiency of CO from CPE measurements at the same potential, $2F$ is two electrons multiplied by Faraday's constant, t is the time of the CPE measurement, and $n_{\text{cat}} = 2.50 \times 10^{-10}$ mol is the amount of CoPc(L) dropcast onto the surface. This definition of turnover frequency has the benefit of involving only one type of measurement (CPEs) in its calculation. However, a major limitation of the TOF_{CO, CPE}

definition of turnover frequency is that mass transport in the CPE is based on turbulent flow from solution stirring. This means that the mass transport is not well controlled, and so the local concentration of CO₂ near the working electrode surface will fluctuate and likely decreases over time. Thus, Q is not an accurate representation of the activity of the catalyst under hydrodynamic control, but rather is likely limited by inefficient CO₂ delivery over time, and so TOF_{CO, CPE} is likely a less accurate representation of the catalyst activity for CO production compared to TOF_{CO}.

Although the absolute numbers for TOF_{CO} and TOF_{CO, CPE} are slightly different due to the different ways they were determined, importantly the observed *trends* in activity with changing L using each metric are qualitatively similar as summarized in Tables S1 and S2. Note that TOF_{H₂} and TOF_{H₂, CPE} are calculated identically to TOF_{CO} and TOF_{CO, CPE}, respectively, but using FE_{H₂} instead of FE_{CO}.

Nonaqueous Cyclic Voltammetry Measurements of CoPc in DMSO solutions. CoPc/DMSO solutions were prepared by adding 0.0023 g of CoPc (4.0 μmol) to 10 mL of 0.1 M *n*Bu₄NPF₆ in DMSO in an aluminum foil-jacketed polypropylene conical centrifuge tube (Basix™) and sonicating for one hour. CoPc/L/DMSO solutions were prepared by adding 0.0023 g of CoPc (4.0 μmol) and 4 mmol of the axial ligand L to 10 mL of a 0.1 M *n*Bu₄NPF₆ in DMSO in an aluminum foil-jacketed polypropylene conical centrifuge tube (Basix™), vortexing for 1 minute at ~3000 rpm, and sonicating for one hour. After preparation, the CoPc/DMSO and CoPc/L/DMSO solutions were transferred to a 20 mL scintillation vial for the electrochemical measurements.

Cyclic voltammetry (CV) measurements of CoPc/DMSO and CoPc/L/DMSO electrolyte solutions were conducted in quiescent solutions using a 0.071 cm² glassy carbon disk working electrode (CH Instruments), a custom Ag/AgNO₃(1.0 mM in DMSO with 0.1 M *n*Bu₄NPF₆) nonaqueous reference electrode separated from the solution by a CoralPor glass frit (Bioanalytical Systems, Inc), and a Pt wire auxiliary electrode (99.99%, Surepure Chemetals, LLC). The electrochemical cell was a 20 mL

scintillation vial. The reference electrode was externally referenced to the ferrocenium/ferrocene redox couple ($\text{Fc}^{+/0}$) before each set of experiments.

Prior to each experiment, the electrolyte solution was sparged with Ar for at least 30 minutes, and the headspace was blanketed with Ar during each measurement. To avoid electrolyte evaporation, the Ar gas was saturated with DMSO before use by first bubbling it through a gas washing bottle filled with DMSO. The uncompensated solution resistance ($R_u \approx 350 \, \Omega$) was measured using a single-point impedance measurement at 100 kHz with a 20 mV amplitude about the open circuit potential before each set of measurements, and CVs were automatically corrected for iR drop at 85% through positive feedback using the Bio-Logic EC-Lab software.

UV-Vis Spectroscopy

0.01 mM CoPc/DMF and CoPc(L)/DMF solutions were analyzed using PerkinElmer Lambda 265 UV-Vis Spectrophotometer with fast mode and baseline correction. The 0.01 mM solutions were prepared by diluting the corresponding 0.05 mM CoPc/DMF and CoPc(L)/DMF deposition solutions.

Density Functional Theory Calculations

All quantum-chemical calculations utilized density functional theory as implemented in the Q-Chem 5.1 software package.⁵⁵ The B3LYP density functional with the 6-31G** basis function was employed for the gas-phase geometry optimization of the CoPc(L) systems. Single point calculations were employed to calculate the orbitals energies using B3LYP in combination with 6-311+G** basis on all atoms except def2-TZVP on the Co atom. To obtain the binding energy (BE) of CO_2 to the neutral and reduced CoPc(L) complexes structures were optimized with $\omega\text{B97X-D}$ and the 6-31G** basis, which accurately captures non-covalent interactions.⁵⁶ In the optimization of the reduced CoPc(L) complexes, the axial coordinating

ligands were fixed in the XY plane to prevent π -stacking interactions between the ligand and the phthalocyanine rings. The reported energies were obtained using ω B97X-D and the 6-31+G** basis to better describe negatively charged ions.³⁸ Solvation energies were calculated using SMD⁵⁷ to model the implicit water solvent effect.

The Absolutely Localized Molecular Orbital-Charge Transfer Analysis (ALMO-CTA)⁵⁸ was used to quantify electron donation within the CoPc unit and its ligands. The systems were divided into two fragments to study the forward charge transfer from: (1) ligands to the CoPc, (2) five-coordinate CoPc(L) complexes to the CO₂ adduct, and (3) five-coordinate complexes to the protonated CO₂ adduct. Reported forward CTA terms for neutral, reduced, CO₂ adsorbed CoPc(L) are obtained at the ω B97X-D/def2-SVP level of theory, while the protonated CO₂ adduct intermediate was obtained with B3LYP/def2-SVP. Additionally, the forward charge transfer from the reduced CoPc(L) species to the CO₂ adduct was calculated to obtain quantitative insight on the catalytically active species nucleophilic capacities.

RESULTS AND DISCUSSION

The core hypothesis we are testing in this study is that axial coordination of ligands L to CoPc makes the Co center more nucleophilic in the resulting CoPc(L) complexes, and that this increased nucleophilicity translates into an increase in activity for the CO₂RR compared to the four coordinate CoPc. A corollary to this hypothesis is that the nucleophilicity of the Co center in CoPc(L), and therefore the CO₂RR activity, should be correlated to the σ -donor strength of the axial ligand. To test this hypothesis, we have conducted a combined experimental and computational study that explores how the catalytic activity and electronic structure of CoPc(L) changes as we modulate the σ -donor strength of L.

Effect of σ -donor Strength of Axial Ligands on CO₂RR Activity by CoPc(L)

To experimentally quantify the relative σ -donor strength of the axial ligands in CoPc(L) catalysts, we measured how axial ligation influences the λ_{max} of the Q band of CoPc(L). For the four-coordinate CoPc complex, the measured Q band has $\lambda_{\text{max}} = 659$ nm (Figure S1). When adding an axial ligand L to form CoPc(L), there is a red shift in the λ_{max} of the Q band attributed to the increased electron density on the central metal atom,^{49,59-61} and the extent of the red shift should be directly correlated to the electron-donating ability of the ligand L (Figure S1).⁶¹ The λ_{max} of the red-shifted Q bands for CoPc(L) solutions with axial ligands L1-L8 are shown in Figure 3a. The shift in the Q-band is consistent with the postulated ordering of the σ -donor strength of the ligands L1-L8 shown in Figure 2.

Additionally, we expect the redox potential of the putative Co^{2+/+} couple ($E_{\text{Co}^{2+/+}}^0$) should shift more negative as a function of the increasing σ -donor ability of the axial ligand L. Ideally, we would use cyclic voltammetry (CV) to measure $E_{\text{Co}^{2+/+}}^0$ for physisorbed CoPc and CoPc(L) on edge plane graphite electrode surfaces exposed to aqueous electrolytes analogous to those used in the catalytic CO₂ reduction studies described below. Unfortunately, CoPc tends to aggregate when physisorbed to carbon surfaces,^{18,25,29,30,36,62} and this aggregation results in poorly defined redox features that complicate accurate assessment and comparison of redox potentials. Therefore, we instead conducted CVs of nonaqueous solutions containing 0.4 mM CoPc and 400 mM of L in DMSO with 0.1 M *n*BuPF₆ supporting electrolyte. Here, the ratio of CoPc to L is analogous to that in the deposition solution, although the overall concentrations have been increased to enable the electrochemical measurements.

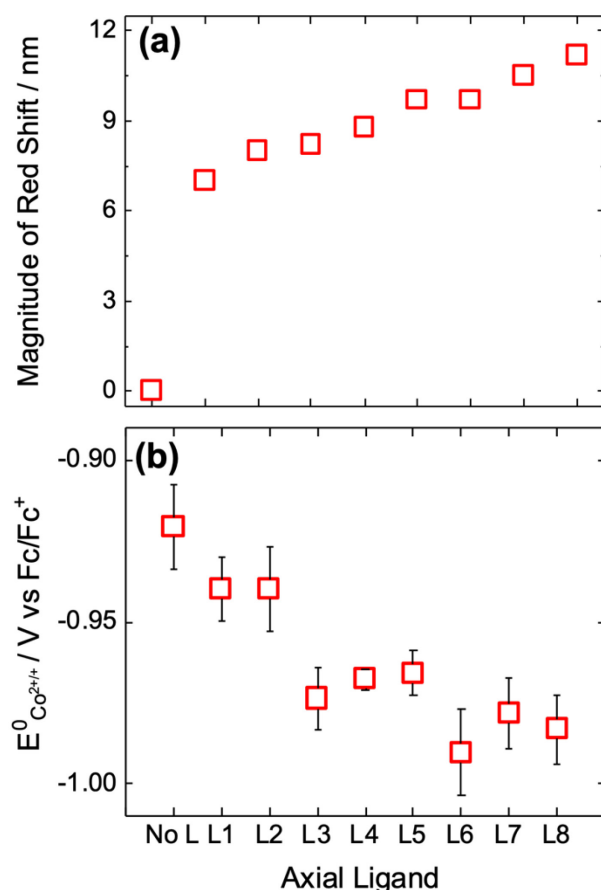


Figure 3. (a) Magnitude of the red shift of the Q-band for diluted deposition solutions containing 0.01 mM CoPc with 10 mM axially coordinating ligands L1-L8 compared to the $\lambda_{max} = 659$ nm of the CoPc without any axially-coordinating ligand (No L). The individual Q-band spectra are shown in Figure S1. (b) The redox potential of the putative $Co^{2+/+}$ couple for CoPc ($E_{Co^{2+/+}}^0$) for DMSO electrolyte solutions containing 0.4 mM CoPc with 400 mM axially coordinating ligands L1-L8 compared to that of CoPc without any axially-coordinating ligand (no L). The reported potentials are averages from at least three experiments in independently prepared sample solutions, and the error bars represent standard deviations. Representative CVs of each CoPc/L/DMSO solution are shown in Figure S2.

Representative voltammograms of the CoPc/L/DMSO solutions are shown in Figure S2. The measured $E_{\text{Co}^{2+}/+}^0$ for CoPc in the DMSO electrolyte is ~ -0.92 V vs $\text{Fc}^{+/0}$, similar to the potential measured for this redox couple in previous reports.^{18,19,63,64} Upon the addition of L to the DMSO electrolyte solution, there is a shift in the $E_{\text{Co}^{2+}/+}^0$ to more negative potentials. Moreover, as the σ -donor ability of L increases in the CoPc/L/DMSO solutions, there is a general negative shift in the redox potential of the putative $\text{Co}^{2+}/+$ couple (Figure 3b): from $E_{\text{Co}^{2+}/+}^0 \approx -0.93$ V vs $\text{Fc}^{+/0}$ for solutions of CoPc with L1 to $E_{\text{Co}^{2+}/+}^0 \approx -0.98$ V vs $\text{Fc}^{+/0}$ for solutions of CoPc with L8. These electrochemical results are qualitatively consistent with our spectroscopic observations. Increasing σ -donor ability of the axially-coordinated ligand L increases the electron density on Co as evidenced by the red shifts in λ_{max} of the Q-band shown in Figure 3a, and this increased electron density on Co results in more negative $\text{Co}^{2+}/+$ redox potentials as shown in Figure 3b. Overall, the spectroscopic and electrochemical studies above are consistent with our hypothesis that axial coordination of L increases the nucleophilicity of the Co center in CoPc(L), and the extent of this increase in nucleophilicity is correlated to the σ -donor ability of L.

To investigate the influence of σ -donor strength of the axial ligand on CO_2RR activity, we evaluated the CO_2RR activity physisorbed CoPc(L) catalysts on edge-plane graphite (EPG) surfaces using rotating disk electrode voltammetry (RDEV) at 1600 rpm and 1 mV s^{-1} scan rate. This scan rate is slow enough to ensure steady-state behavior at the electrode surface, and the rotation rate is sufficiently fast to aid in product removal and limit bubble formation from evolved CO and H_2 at the electrode surface. The resulting average RDEVs of CoPc, CoPc(L1) and CoPc(L8) for the CO_2RR are shown in Figure 4, and the average RDEVs of all complexes are shown in Figure S2. To confirm that the RDEV measurements are a true measure of steady state behavior, a series of rotating disk electrode chronoamperometric (RDE-CA) steps were measured for each catalyst (Figure S3). In these experiments, the potential was held

constant at given applied potentials for 2 min at 1600 rpm rotation rate, and the current was allowed to decay to a steady state value. Average steady-state currents for the CO₂RR by each CoPc(L) complex investigated as determined from RDE-CA measurements show good agreement when overlaid with the RDEV measurements (Figure 4), suggesting that the RDEV measurements are valid approximations of steady-state behavior. Note that the lack of change in the onset and half wave potential with the increasing σ -donor ability of L likely arises from the complicated dependence of these parameters on not only the catalyst redox potential, but also on the rate of the various chemical steps such as CO₂ binding and subsequent protonation in the multistep, multielectron mechanism.⁶⁵

For each CoPc(L) complex investigated, there is a plateau in the catalytic RDEV at ~ -1.25 V vs. SCE, and we attribute this plateau current to the maximum CO₂RR activity under the experimental conditions. Therefore, the activity and selectivity at -1.25 V vs. SCE was used as the primary metric to compare the CO₂RR performance for the CoPc(L) catalysts. The average current densities measured from the catalytic RDE-CA measurements are summarized in Table S1. A plot of the catalytic current density at -1.25 V for each CoPc(L) catalyst as a function of the magnitude of the red shift of the complexes' Q bands, a proxy for the σ -donor strength of the axial ligands, is shown in Figure 5a. As the σ -donor strength increases and the magnitude of the red shift increases, there is a corresponding increase in the catalytic current density. This activity trend is consistent with our hypothesis that increasing the σ -donor strength of the axial ligand in CoPc(L) correspondingly increases the nucleophilicity of the Co center, facilitating CO₂ activation and reduction.

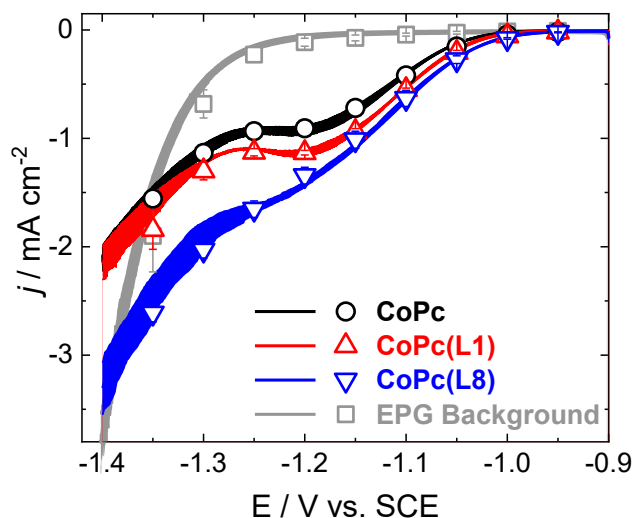


Figure 4. Rotating disk electrode voltammograms (RDEVs) of the CO₂RR by CoPc, CoPc(L1), and CoPc(L8) catalysts at 1 mV s⁻¹ scan rate and 1600 rpm in CO₂-saturated pH 5 phosphate solutions. The solid lines represent the averages RDEVs from at least three measurements on independently prepared samples, and the shaded region represents the standard deviation for these measurements. The average results of 2 min rotating disk electrode chronoamperometric (RDE-CA) steps from at least three measurements on independently prepared samples are shown by open symbols, and the error bars represent standard deviations for these measurements. The close overlay of the RDEV and RDE-CA data suggests that the RDEV measurements are a good approximation of steady-state conditions. The RDEV and RDE-CA data for a bare EPG electrode under identical experimental conditions is included for comparison.

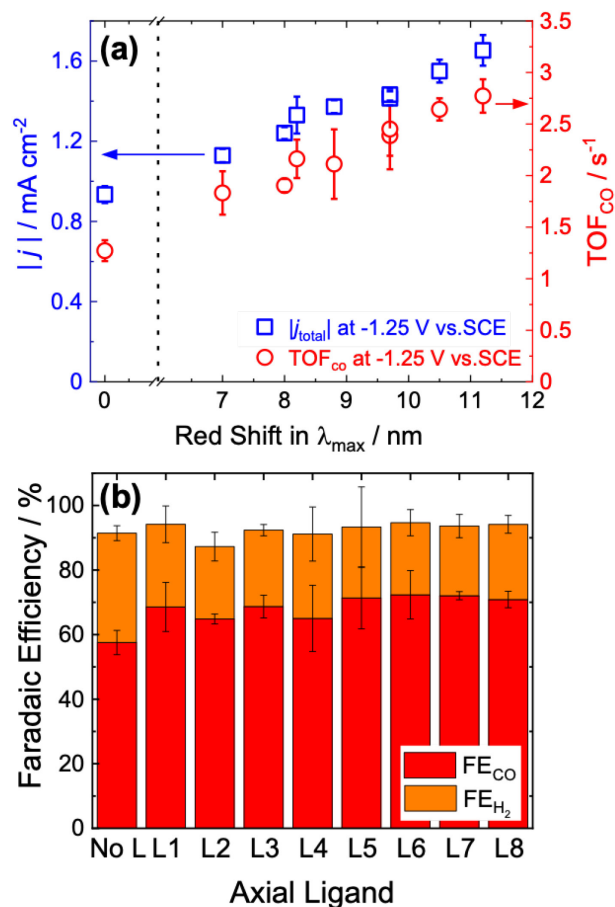


Figure 5. (a) A plot of the absolute value of the average current density from RDE-CA measurements (left y-axis) and turnover frequency (TOF_{CO}) measured at -1.25 V vs SCE (right y-axis) as a function of the red shift in λ_{max} for CoPc(L) . The magnitude of the red shift is related to the σ -donor ability of the axial ligands, L. CoPc is reported here at red shift in $\lambda = 0$ nm, and all CoPc(L) complexes with axial ligands are red shifted from CoPc as reported in Figure 3. Note the vertical dotted line highlights a break in the x-axis at $\lambda_{\text{max}} = 1$ nm, and the scale continues after the break at $\lambda_{\text{max}} = 6$ nm. (b) A plot of the average Faradaic efficiencies for CO and H₂ production from CPE measurements at -1.25 V vs SCE. Note that the label “No L” on the x-axis refers to CoPc on EPG with no added axial ligand. For both (a) and (b), the data shown for $|j|$, FE_{CO} , and FE_{H_2} are averages from at least three experiments on independently

prepared samples, and the error bars represent standard deviations. For the TOF_{CO} values, the error bars represent standard errors propagated from the standard deviations $|j|$ and FE_{CO} measurements.

Controlled potential electrolysis (CPE) were performed at -1.25 V vs. SCE to assess Faradaic efficiencies for CO_2 reduction to CO (FE_{CO}) by CoPc and CoPc(L) modified electrodes. Electrolyses were conducted in a sealed electrochemical cell in CO_2 atmosphere for 2 h, and CO and H_2 evolved were quantified from headspace analysis using gas chromatography. Representative examples of the CPE current traces are shown in Figures S4-S12 for each catalyst, and the resulting Faradaic efficiencies for H_2 (FE_{H_2}) and CO (FE_{CO}) generation are summarized in Figure 5b and Table S2. In general, the only products produced were CO and H_2 . All CoPc(L) catalysts reduce CO_2 to CO with $\text{FE}_{\text{CO}} \approx 70\%$, and CoPc with no axial ligand produces CO with $\text{FE}_{\text{CO}} \approx 60\%$. Note that control CPE experiments of the EPG background under identical conditions showed only H_2 as the sole product, confirming that the adsorbed CoPc(L) complexes are the active species for the CO_2RR . The increased FE for CoPc(L) compared to CoPc is consistent with our previous findings that axial coordination of ligands such as pyridine to the CoPc complex facilitate CO_2 coordination by shifting the rate determining step from an initial CO_2 binding step to a subsequent protonation event.¹⁸ The fact that all the CoPc(L) systems show the same FE for CO production suggests that this change in the rate-determining step upon axial coordination occurs for all CoPc(L) systems investigated independent of σ -donor ability of the axial ligand. A plot of the electrocatalytic turnover frequencies for CO production (TOF_{CO}) as function of the magnitude of the red shift of the complexes' Q bands shows increasing activity with increasing σ -donor strength of the axial ligand, qualitatively similar to the trend observed from the RDE-CA measurements (Figure 5a).

Note that the rising region in the RDEV and RDE-CA measurements in Figure 4 after the plateau at potentials < -1.25 V is attributed to the onset of competitive HER from the EPG electrode, which occurs at ~ -1.3 V vs SCE. CPE experiments conducted at -1.35 V for CoPc, CoPc(L3), and CoPc(L8) show a large increase in H_2 production with $FE_{H_2} \geq 75\%$, consistent with the postulated increase in the HER at these more negative potentials (Table S3).

The Effect of σ -donor Strength of Axial Ligands on Electronic Structure of CoPc(L)

To better understand the influence of the σ -donor strength of axially-coordinating ligands on the electronic structure of CoPc(L), density functional theory (DFT) was used to examine molecular orbitals for the various CoPc(L) complexes. The orbitals with primarily dz^2 character (dz^2 -MOs) and the highest occupied molecular orbitals (HOMOs) for CoPc and CoPc(L) are shown in Figure 6 for CoPc, CoPc(L1), and CoPc(L8), and for all complexes in Figure S13. For all CoPc and CoPc(L) complexes, the HOMOs are primarily composed of π -orbitals of the Pc rings. The dz^2 orbital of CoPc is relatively low energy (i.e. the HOMO-5 orbital), but upon axial ligation it increases in energy by ~ 2.1 eV and becomes the HOMO-1 for CoPc(L). The relative increase of the dz^2 orbital energy increases with increasing σ -donor strength of L as shown in Figure 7a and summarized in Table S4.

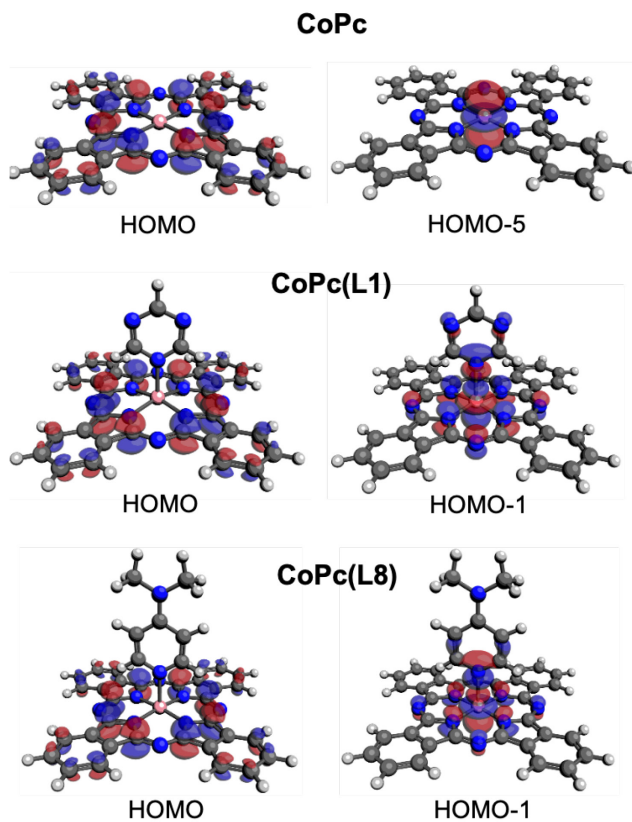


Figure 6. Graphical representations of the HOMO (left) and dz^2 orbitals (right) for CoPc, CoPc(L1), and CoPc(L8). Calculated charge distributions (red and blue shaded orbitals) are based on DFT calculations (pink ball = Co, blue ball = N, grey ball = C, white ball = H). Isovalue is 0.10 a.u.

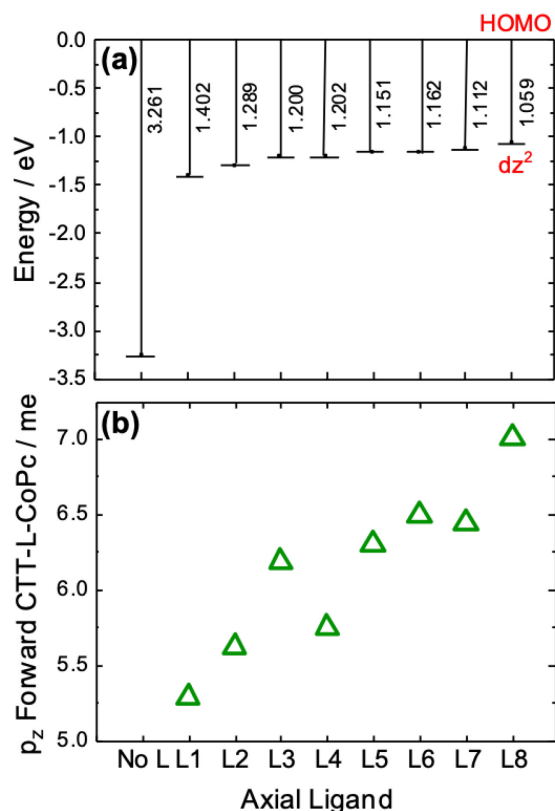


Figure 7. (a) Calculated energy differences between the HOMO and dz^2 -MO for CoPc(L) with various axial ligands L. (b) The forward charge transfer term from the ligand L nitrogen p_z to the CoPc complex (CTT-L-CoPc) from ALMO-CTA. CTT-L-CoPc are in millielectrons (me).

Charge transfer analysis (see computational details) determined the primary interaction between the CoPc complex and axial ligand L in CoPc(L) is σ -donation from the p_z orbital of the coordinating N atom to the dz^2 orbital of the Co atom. This interaction between N p_z and Co dz^2 accounts for $> 80\%$ of the forward charge transfer between the ligands L and the CoPc. The second significant interaction is a π -interaction between the p_x orbitals of the conjugated C atoms in L and the p orbitals of the outer C-N-C atoms in the Pc, accounting for $\sim 10\%$ of the forward charge transfer. The additional electron density on the Co center compared to the parent complex CoPc suggests that the Co will have increased

nucleophilicity. Moreover, the magnitude of the forward charge transfer term from the ligand L to the CoPc complex (CTT-L-CoPc) increases with increasing σ -donor strength of L as shown in Figure 7b. The DFT findings are therefore consistent with the UV-Vis measurements showing a red shift in the Q-band as a function of σ -donor strength of L (Figure 3 and Figure S14). These results support the hypothesis that axial coordination increases the energy of the Co dz^2 orbital, and thus the nucleophilicity of the Co center, in proportion to the σ -donor strength of the axial ligand.

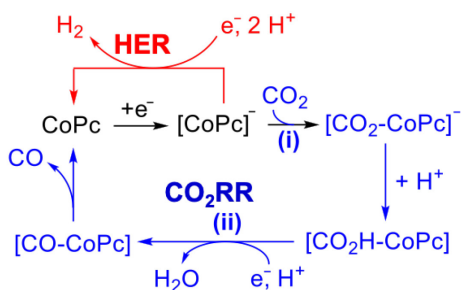
Mechanistic Implications of Axial Coordination from First Principles Calculations

The above DFT analysis focused on the electronic structure of the uncharged CoPc(L) species serves as a baseline for understanding the influence of axial coordination on the electronic structure. To best inform the electrocatalytic mechanism of the CoPc(L) system, it is important to understand the influence of axial coordination on the catalytically-active reduced CoPc and CoPc(L) species. Here, we introduce probable mechanisms for CO₂RR by CoPc in order to identify relevant catalytically active species for further computational study.

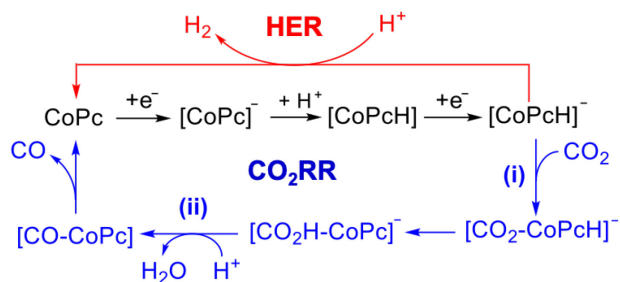
Although the mechanism for the CO₂RR by CoPc is likely condition-dependent and remains an active point of discussion,^{21,22,25,39} two mechanisms have emerged that are largely consistent with experimental measurements in aqueous solution and are depicted in Scheme 1. In Mechanism 1 (Scheme 1a), the active species for CO₂ activation is the singly-reduced [CoPc]⁻ intermediate, which can either coordinate and reduce CO₂ to generate CO in the CO₂RR pathway or undergo reductive protonation to generate H₂ along the competitive HER pathway. In Mechanism 2 (Scheme 1b), CoPc is reduced by 2e⁻ and protonated once (presumably on the Pc ring) to form [CoPcH]⁻. This species is the selectivity-determining intermediate that can either react with another H⁺ to form H₂ in an HER pathway or react with CO₂ to form the CO₂ adduct [CoPcH-CO₂]⁻. The latter intermediate undergoes a putative intramolecular H⁺ transfer from the

Pc ring to the CO₂ adduct followed by a subsequent solution protonation step to eventually generate CO. Both mechanisms are consistent with KIE studies that show axial coordination changes the rate determining step from the initial CO₂-coordination event (step i) for CoPc to a subsequent protonation event (step ii) for CoPc(L).¹⁸

(a) Mechanism 1



(b) Mechanism 2



Scheme 1. Two proposed mechanisms for the CO₂RR by CoPc. (a) In Mechanism 1, CO₂ coordination occurs at the singly-reduced catalyst.^{25,39} (b) In Mechanism 2, CO₂ coordination occurs at the doubly-reduced and singly-protonated catalyst.^{15,21,22} Both mechanisms are consistent with KIE studies that show a change in the rate-determining step upon axial coordination from CO₂ coordination (step i) to a subsequent protonation event (step ii).¹⁸ Postulated pathways for competitive HER are shown in red for each mechanism.

To help distinguish between the two proposed mechanisms for CO₂RR by CoPc(L) catalysts, we compared the experimental activity measurements to computed frontier orbitals, CO₂ binding energies (BE), and charge transfer terms (CTT) for the CoPc(L) CO₂ adduct. We propose that a close correlation of increasing BE and CTT for the pertinent mechanistic intermediates in Mechanisms 1 or 2 with increasing σ -donor strength of the axial ligand suggests that the mechanism is consistent with experimental results (Figure 8). Such an agreement would then correlate BE and CTT of CoPc(L) to the CO₂ adduct with the measured catalytic activity, which also increases with σ -donor strength of the axial ligand. These first-principles studies will relate the electronic properties of CoPc(L) to the measured electrocatalytic activity for the CO₂RR.

The orbital ordering for the intermediates preceding CO₂ coordination was examined for each mechanism, [CoPc(L)][−] for Mechanism 1 and [CoPcH(L)][−] for Mechanism 2 (Figure 9 a,b). In the case of the singly-reduced [CoPc][−] active species from Mechanism 1, the dz²-MO is HOMO-3 (Figure 9a). Upon axial coordination to reach [CoPc(L)][−], the dz²-MO increases in energy to become the HOMO for all L1-L8 (Figure 9a, S16). This shows that the dz²-MO of [CoPc(L)][−] is better poised to activate and coordinate the electrophilic C of CO₂ compared to that of [CoPc][−], assuming Mechanism 1 is the operative mechanism for CO₂RR. We note that the calculated orbital ordering for [CoPc][−] and [CoPc(L)][−] show a buried d_{xy} orbital inconsistent with a simple crystal-field theory description of a square planar complex. In CoPc, the dxy orbital has a low level of covalency with the aza-nitrogens in the Pc ring, decreasing the relative energy of the dxy orbital. The relative position of the computed dxy orbital compared to the other d-orbitals is supported by various spectroscopic studies.⁶⁶⁻⁶⁸

For the active species from Mechanism 2, [CoPcH][−], the dz²-MO is also a buried HOMO-3 similar to the intermediate in Mechanism 1 (Figure 9b). For the axially-coordinated [CoPcH(L)][−] active species, the dz²-MO becomes HOMO-2 and the HOMO is still composed primarily of π -orbitals of the Pc rings

(Figure 9b, S18). These calculations suggest that Co center of $[\text{CoPcH(L)}]^-$ is more nucleophilic than that of $[\text{CoPcH}]^-$, but that the buried dz^2 -MO of $[\text{CoPcH(L)}]^-$ is not as well positioned energetically to interact with CO_2 compared to the singly-reduced $[\text{CoPc(L)}]^-$ in Mechanism 1. This may suggest that Mechanism 2 is a less likely mechanism for the CO_2RR compared to Mechanism 1.

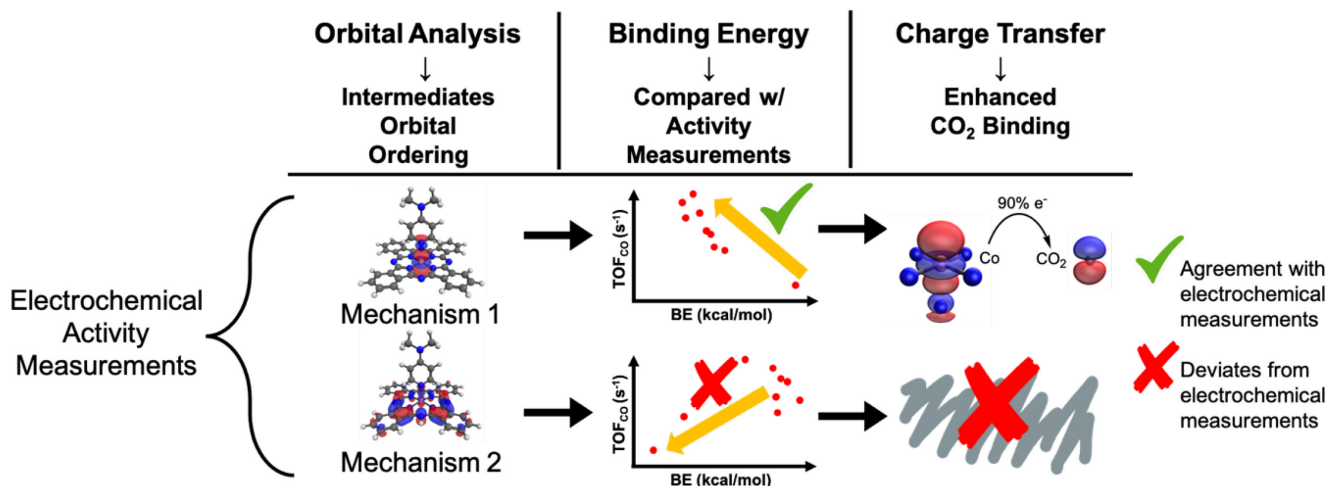


Figure 8. Overview of DFT analysis for the proposed CO_2RR mechanisms for CoPc(L) catalysts. These calculations were used to look for trends that can explain the comparative measured electrocatalytic activities for the CoPc(L) systems. First, we determined the HOMO orbital character for the postulated catalytically-active intermediates. Next, we compared trends in the CO_2 binding energy for the intermediates in Mechanism 1 and 2 to trends in the measured catalytic activities (TOF_{CO}). Finally, we measured the influence of the σ -donor strength of the axial ligands L on the extent of charge transfer between CoPc(L) and the CO_2 adducts in the relevant intermediates for both mechanisms. If we assume that CO_2 -binding prior to the rate determining step thermodynamically-controls CO_2RR activity, and the energy of the mechanistically-pertinent $\text{CoPc(L)}\text{-CO}_2$ intermediates influence catalytic turnover, then these analyses are more consistent with Mechanism 1 being the operative mechanism for CO_2RR under our experimental conditions.

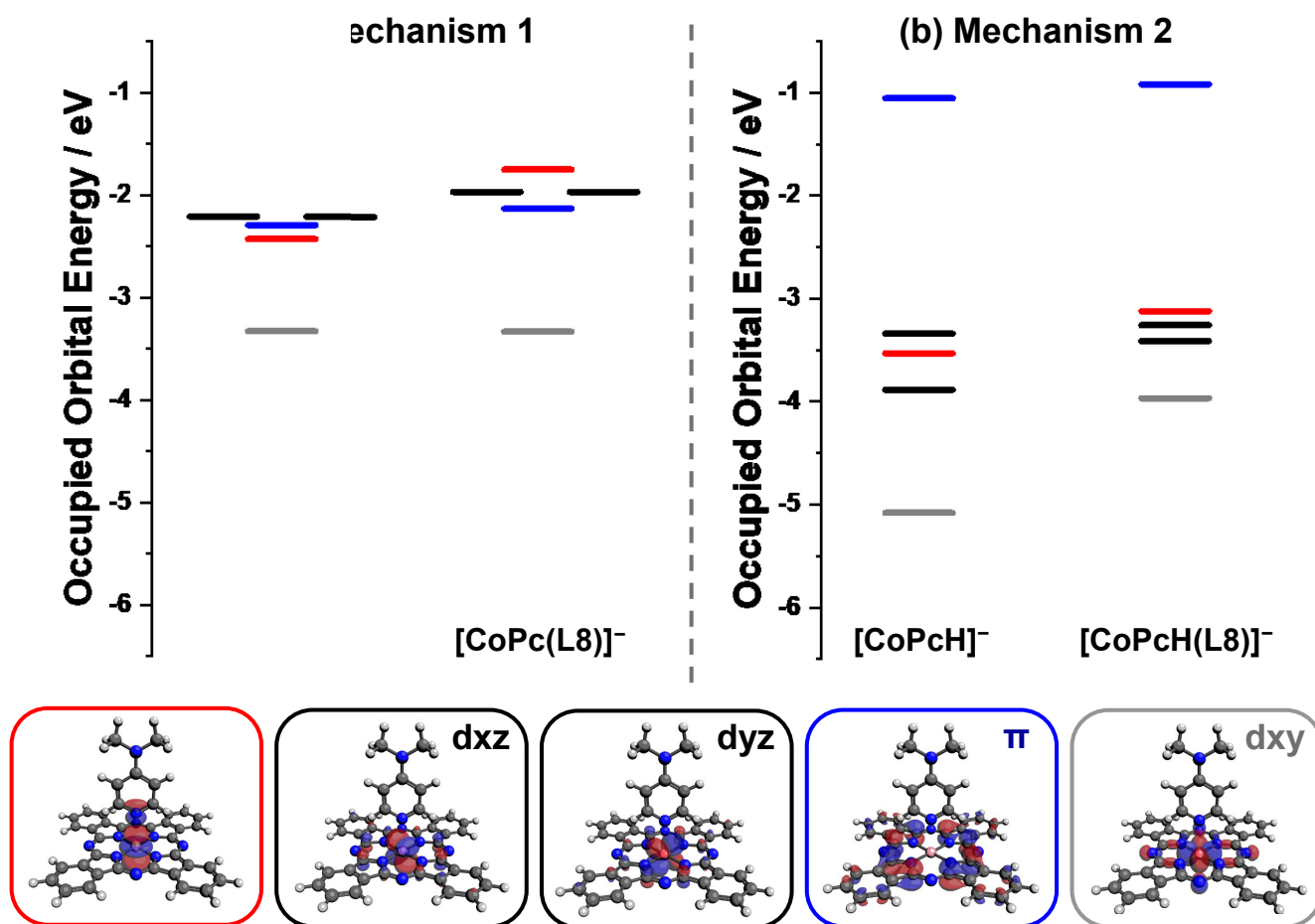


Figure 9. Energies of the occupied frontier molecular orbitals obtained for the singlet state singly reduced $[\text{CoPc}(\text{L})]^-$ and the doublet state of the doubly reduced, singly protonated $[\text{CoPcH}(\text{L})]^-$ with and without axially coordinating L ligands. Note that the reported occupied molecular orbital energies for each $\text{CoPc}(\text{L})$ complex are calculated in the gas-phase, and therefore each orbital energy is relative to vacuum. The colors of the energy levels correspond to the specific molecular orbitals: red is d_{z^2} , black is d_{xz} and d_{yz} , grey is d_{xy} , and blue is the π orbital of the Pc ring. Graphical representations of the molecular orbitals for $[\text{CoPc}(\text{L8})]^-$ with color-coordinated labels are shown for reference, and the graphical representations for the other orbitals are shown in Figures S15 and S17. Summaries of the orbital energies and assignments are listed in Table S5 and Table S6.

The CO₂ coordination via the [CoPc(L)-CO₂]⁻ and [CoPc(L)-CO₂H]⁻ intermediates were then evaluated based on the assumption that these intermediates are the thermodynamically controlling species prior to the rate-limiting CO₂ conversion steps. According to Mechanism 1, CO₂ coordination should occur at the singly-reduced [CoPc(L)]⁻. The computed BE of CO₂ in [CoPc(L)-CO₂]⁻ is much more negative (larger magnitude) than that of the comparable CO₂-adduct of the four-coordinate CoPc species in [CoPc-CO₂]⁻. This suggests that axial coordination facilitates the binding of CO₂ at the singly-reduced state in Mechanism 1. This is consistent with our experimental observations that show axial coordination changes the rate-determining step from CO₂ coordination to a subsequent mechanistic step.¹⁸ Moreover, the BE of the CO₂ adduct increases with σ-donor ability of L as measured by the red-shift in the Q-band for the CoPc (L) complex (Figure 10a). In addition, there is a direct correlation between BE and TOF_{CO} for the CoPc(L) complexes (Figure 10c). The DFT and experimental observations suggest that axial coordination by strong σ-donor ligands facilitates CO₂ coordination, increasing the rate of CO₂ reduction to CO by the CoPc(L) complexes.

According to Mechanism 2, CO₂ coordination should occur at the doubly-reduced singly-protonated [CoPcH(L)]⁻, producing [CoPc(L)-CO₂H]⁻. However, unlike with Mechanism 1, there is no clear correlation between the BE of the CO₂ adduct in the intermediate [CoPc(L)-CO₂H]⁻ with either the σ-donor ability of L (Figure 10b) or TOF_{CO} (Figure 10d). Overall, these results suggest that Mechanism 2 is less likely than Mechanism 1 to be operative under our experimental conditions.

To provide further understanding of the observed trends in BE, the extent of CTTs from the reduced CoPc(L) species to the CO₂ adducts were computed. In the case of Mechanism 1, the primary interaction contributing to the charge transfer between the complex and CO₂ is the σ-donation from the Co dz² to the p_z orbitals in the CO₂ adduct (Table S7). For the intermediate [CoPc(L)-CO₂]⁻, there is a general increase in CTT between [CoPc(L)]⁻ and CO₂ with increasing σ-donor strength of the ligand L (Figure 11a), and

this increased CTT is correlated to the increased activity for the CO₂RR (Figure 11b). The extent of CTT for the [CoPc(L)]⁻ species to the protonated CO₂ adducts for [CoPc(L)-CO₂H]⁻ was calculated for Mechanism 2 (Table S8), but there was a lack of correlation between the CTT and the experimentally-measured activity (Figure S21). The trends in CTT to CO₂ with L were less predictive of electrocatalytic activity for the key intermediate of Mechanism 2 compared to Mechanism 1.

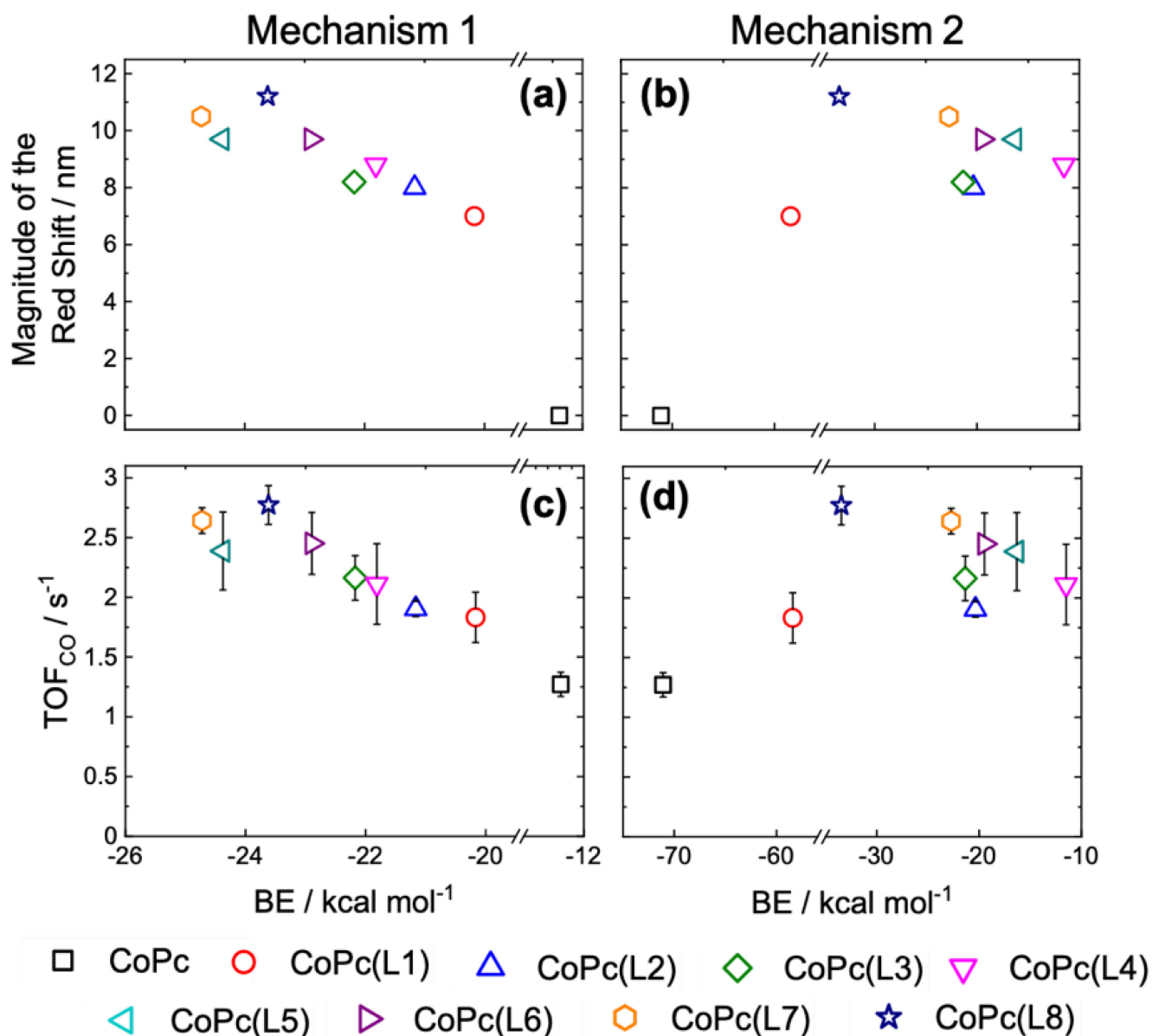


Figure 10. (a) The magnitude of the red shift in $\lambda_{\max}$ of the Q-band for the CoPc(L) complex (related to the σ -donor ability of L) is plotted versus the binding energy (BE) of CO₂ adduct in the [CoPc(L)-CO₂]⁻ intermediate. b) The magnitude of the red shift in $\lambda_{\max}$ versus BE of the CO₂H adduct in [CoPc(L)-CO₂H]⁻. (c-d) Same as (a) and (b), respectively, but with y-axis as Turnover Frequency for CO (TOF_{CO}). Graphical representations for the HOMOs and orbitals showing the interaction between the complex and CO₂ for [CoPc(L)-CO₂]⁻ and [CoPc(L)-CO₂H]⁻ are shown in Figures S19 and S20, respectively.

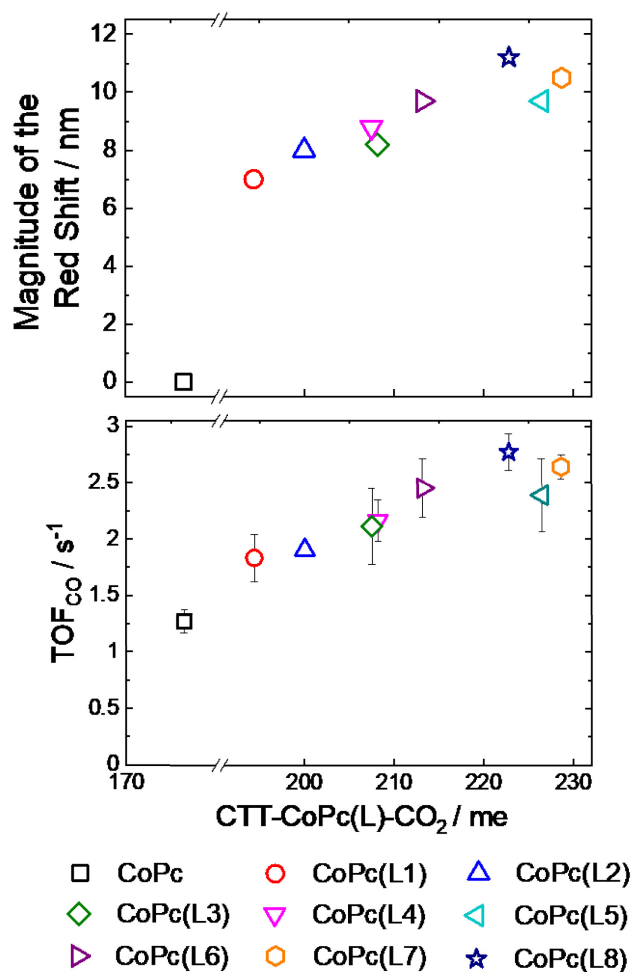


Figure 11. (a) The magnitude of the red shift in $\lambda_{\max}$ of the Q-band plotted versus the forward charge transfer term from the singly-reduced CoPc(L) to the coordinated CO₂ adduct in the [CoPc(L)-CO₂]⁻ intermediates (CTT-CoPc(L)-CO₂) (b) The activity as measured by the Turnover Frequency (TOF_{CO}) for the CoPc(L) complexes is plotted versus the same CTT. CTT were from ALMO-CTA analysis, and the units are millielectrons (me). TOF_{CO} here were determined by multiplying the RDE-CA measurements at -1.25 V vs SCE by the FE_{CO} from the CPE measurements at -1.25 V.

CONCLUSIONS

It has been previously shown that axial ligation of the CoPc complex by either a coordination ligand such as pyridine (L3) or a coordinating polymer such as P4VP results in increased activity and selectivity for the CO₂RR to CO in aqueous phosphate solutions.^{17,18,21,22} We hypothesized that axial coordination increases the energy of the dz^2 orbital, the orbital most likely involved in CO₂ coordination, and that this higher energy dz^2 orbital facilitates CO₂ reduction. The combined electrochemical and computational study reported here provides strong support for this hypothesis. We investigated the activity for the CO₂RR by a series of CoPc(L) complexes, where L increases in σ -donor strength from the relatively weak σ -donor triazine (L1) to the comparatively strong sigma donor 4-dimethylaminopyridine (L8). Our studies showed a correlation between the increasing σ -donor strength of L and the increased catalytic activity of the resulting CoPc(L) complexes for the CO₂RR. Using DFT analysis, we also showed that the relative energy of the dz^2 orbital increases with increasing σ -donor ability of the axial ligand, and that Co dz^2 orbital is the primary orbital that interacts with CO₂ in the reduced CoPc(L)-CO₂ intermediates. These findings suggest that the observed increases in CO₂RR activity for CoPc(L) with increasing σ -donor ability of L are correlated with the increased relative energy of the Co dz^2 orbitals in the CoPc(L) complexes.

We also compared first-principles calculations of CO₂ Binding Energy (BE) and Charge Transfer terms (CTT) between the CoPc(L) and CO₂ for active intermediates in two proposed mechanisms for the CO₂RR by CoPc. We showed that, in the case of Mechanism 1, where CO₂ binding occurs after the first reduction of the CoPc(L) system, the CTT and BE terms for the singly-reduced [CoPc(L)-CO₂]⁻ intermediate increased with increasing σ -donor ability of the ligand. The result is a direct correlation of the calculated CTT and BE terms with the experimentally measured CO₂RR activity, suggesting that Mechanism 1 is a possible operative mechanism for this reaction. However, for Mechanism 2, where CO₂ binding occurs after the second reduction and first protonation of the CoPc(L) system, there was a less

clear correlation between the calculated CTT and BE terms for the doubly-reduced, singly protonated $[\text{CoPc}(\text{L})\text{-CO}_2\text{H}]^-$ intermediate and the σ -donor ability of the ligand. Consequently, there was no clear correlation between the BE and CTT terms for Mechanism 2 and the measured catalytic activity, suggesting that Mechanism 2 is less likely to be an operative mechanism. From these studies, we conclude that Mechanism 1 is more likely than Mechanism 2 to explain the CO_2RR by the CoPc and CoPc(L) systems under our experimental conditions.

Overall, our studies provide strong support that axial ligation increases the energy of the dz^2 orbital in the CoPc(L) complexes, the extent of this increase in the dz^2 orbital energy is dependent on the σ -donor strength of the axial ligand, and that this increased dz^2 orbital energy for the CoPc(L) complexes is likely correlated to the observed increase in activity for the CO_2RR . We believe that these studies illustrating the influence of axial coordination on electronic structure and electrocatalytic activity provide important design considerations for future macrocyclic MN_4 -based electrocatalytic systems.

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Notes:

The authors declare no competing financial interest.

ASSOCIATED CONTENT

Supporting Information

Supporting Information is available free of charge on the ACS Publications website at DOI: 10.1021/XXX.

- UV-Vis spectra; representative nonaqueous cyclic voltammograms, rotating disk voltammograms and rotating disk chronamperometric steps;
- summary of results from controlled-potential electrolysis experiments; calculated highest occupied molecular orbitals, d_{z^2} orbitals, and other relevant molecular orbitals for CoPc and CoPc(L) species.
- The XYZ coordinates of the structures used in the DFT calculations for all species investigated: CoPc, CoPc(L), [CoPc]⁻, [CoPc(L)]⁻, [CoPc-CO₂]⁻, [CoPc(L)-CO₂]⁻, [CoPcH]⁻, [CoPcH(L)]⁻, [CoPc-CO₂H]⁻ and [CoPc(L)-CO₂H]⁻; L = L1 to L8.

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