

ARTICLE

Synthesis and Reactivity of a tris(carbene) Zinc Chloride Complex

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The $^{\text{Ph}}\text{B}(\text{tBulm})_3^{1-}$ ligand has gained increased attention since it was first reported in 2006 due to its ability to stabilize highly reactive first row transition metal complexes. In this work, we explore the coordination chemistry of this ligand with redox-inert zinc to understand how a zinc metal center behaves in such a strong coordinating environment. The $^{\text{Ph}}\text{B}(\text{tBulm})_3\text{ZnCl}$ (**1**) complex can be formed via deprotonation of $[\text{PhB}(\text{tBulm})_3][\text{OTf}]_2$ followed by the addition of ZnCl_2 . Salt metathesis reaction with nucleophilic $n\text{-BuLi}$ yields the highly carbon-rich zinc coordination complex $^{\text{Ph}}\text{B}(\text{tBulm})_3\text{ZnBu}$ (**2**) with three carbene atom donors and one carbanion donor. In contrast, reaction of complex **1** with a less nucleophilic polysulfide reagent, $[\text{K.18-C-6}]_2[\text{S}_4]$, leads to the formation of a tetrahedral zinc tetrasulfido complex via protonation of one carbene donor to form $^{\text{Ph}}\text{B}(\text{tBulm})_2(\text{tBulmH})\text{Zn}(\kappa^2\text{-S}_4)$ (**3**).

Introduction

Studies of zinc coordination chemistry generally focus on the implications of its role as a redox-inert, abundant, and bioavailable metal. Zinc is an attractive metal for biological systems because of its predictable coordination chemistry, which favors a four-coordinate tetrahedral geometry. Furthermore, the ability to ligate both larger polarizable as well as smaller non-polarizable atoms such as S, and O, respectively, allows Nature to utilize zinc in a wide array of functions. These range from structural motifs in zinc finger proteins¹ to Lewis acidic sites in the enzyme carbonic anhydrase that activate water molecules for CO_2 hydration.² These characteristics have allowed synthetic chemists to investigate the coordination chemistry of zinc and discover a wide array of reactivity, including small molecule activation,³ secondary coordination sphere effects,⁴⁻⁶ catalysis,^{7, 8} and metal-ligand cooperative reactivity.⁹⁻¹¹

One area of zinc coordination chemistry that has gained increased attention involves its reactivity with reactive sulfur species (RSS) due to the common coordination of thiolate ligands to zinc in biology. Tsui and coworkers have utilized the redox-inert properties of zinc to investigate the redox chemistry of ligand-based thiolates with RSS. Reaction of a *bis*(carboxamide)pyridine zinc complex featuring two flanking thiolate chelates with elemental sulfur demonstrated sulfane sulfur insertion into the Zn-S bond yielding an unusual tetrasulfido zinc complex, and which was shown to be reversible in the presence of 3 equiv of a PPh_3 reductant.¹² The coordination chemistry of anions such as thionitrite (SNO^-) and

perthionitrite (SSNO^-), which are subject to rapid reactivity in the presence of redox-active reagents, have also taken advantage of the redox-inert properties of zinc to form the first structurally characterized transition metal complexes of SNO^- and SSNO^- .¹³

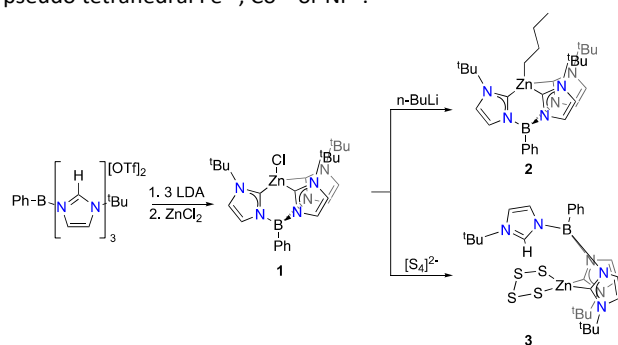
In this work, we investigate the coordination chemistry of zinc using a highly donating *tris*(carbene) ligand (phenyltris(3-tert-butylimidazol-2-ylidene)borato, $[\text{PhB}(\text{tBulm})_3]^{1-}$). The $[\text{PhB}(\text{tBulm})_3]^{1-}$ ligand platform has been employed by others to synthesize and stabilize highly reactive molecules including some with unusual geometries. For example, Smith and coworkers synthesized the first Fe^{V} nitride and noted that the highly σ -donating carbene ligands interacts with iron d_{xz} , d_{yz} and d_{z^2} orbitals, which are the same orbitals involved in multiple bonding of the nitride ligand.¹⁴ Using an adamantyl rather than tert-butyl version of this ligand, Anderson and coworkers synthesized a rare Co-oxo species¹⁵ and demonstrated the biomimetic C-H activation properties of Co-oxo complexes using an adamantyl derivative.¹⁶ Nickel complexes featuring the $[\text{PhB}(\text{tBulm})_3]^{1-}$ ligand were also used to stabilize a bridging peroxo complex that was synthesized via O_2 activation, which further demonstrates the impacts of the strongly donating ligand platform towards small molecule reactivity.¹⁷ In our own work, we found that the $^{\text{Ph}}\text{B}(\text{tBulm})_3\text{NiCl}$ can react with RSS to form square planar *bis*(hydrosulfide) or disulfide complexes with $[\text{NBu}_4][\text{SH}]$ or $[\text{K.18-C-6}]_2[\text{S}_4]$ reagents, respectively.¹⁸

Here we incorporate the $[\text{PhB}(\text{tBulm})_3]^{1-}$ ligand into Zn coordination complex to investigate the reactivity of this electron-rich zinc complex in ligand substitution reactions. Carbene based complexes of zinc have gained popularity in recent years, however the incorporation of multiple carbene based ligands is more uncommon and to the best of our knowledge zinc complexes with three carbene based ligands have not been reported. We demonstrate that treatment of ZnCl_2 with $[\text{PhB}(\text{tBulm})_3]^{1-}$ forms a $^{\text{Ph}}\text{B}(\text{tBulm})_3\text{ZnCl}$ complex, and that salt metatheses reactions result in $[\text{PhB}(\text{tBulm})_3]^{1-}$ ligand protonation, which is consistent with work with other metals.

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Results and Discussion

Complex **1** was synthesized by deprotonation of $[\text{PhB}(\text{tBulm})_3][\text{OTf}]_2$ with 3 equiv. of lithium diisopropyl amide (LDA) generated *in situ* in THF, to generate the $\text{PhB}(\text{tBulm})_3\text{Li}$ intermediate, followed by salt metathesis with ZnCl_2 to provide $\text{PhB}(\text{tBulm})_3\text{ZnCl}$ (**1**) as a white powder in 37% yield. The formation of **1** was confirmed from analysis of the ^1H NMR spectrum of the product, which showed a single peak at 1.61 ppm corresponding to the ^tBu peaks of the $\text{PhB}(\text{tBulm})_3$ ligand indicating a single C_3 symmetric product. Furthermore, the absence of the H-Im peak of the free ligand as well as the shifted resonances compared to the $\text{PhB}(\text{tBulm})_3\text{Li}$ intermediate provided strong evidence for the formation of **1**. In addition, the ^{13}C carbene resonance shifts significantly from 135.3 ppm in $\text{PhB}(\text{tBulm})_3\text{Li}$ ¹⁹ to 186.70 ppm in **1**. The structure of **1** was confirmed by single crystal X-ray diffraction studies using colorless single crystals grown from a concentrated THF solution layered with hexanes and placed in the $-25\text{ }^\circ\text{C}$ freezer overnight. The molecular geometry about Zn in complex **1** is best described as pseudo-tetrahedral with three carbon donor atoms from the *tris*(carbene) borate ligand and one chloride ligand. Zn- C_{Im} bond lengths in **1** vary between 2.045(2) and 2.086(2) Å, which is in the standard range of other Zn-carbene species.²⁰ When compared to other metal centers bearing the same $[\text{PhB}(\text{tBulm})_3]^{1-}$ and chloride ligands, the Zn- C_{Im} distances are longer than expected despite Zn having the smallest atomic radius. The average M- C_{Im} bond lengths where for M = Fe, Co, Ni and Zn are 2.093(3) Å,²¹ 2.035(2) Å,¹⁹ 2.003(2) Å,²² and 2.069(2) Å, respectively. We initially expected that that the Zn- C_{Im} distances would decrease with respect to atomic radii, however the observed elongation may be due to the decreased covalent nature of the Zn- C_{Im} bond when compared to more covalency of M- C_{Im} bonds due to the partially filled d-orbitals of pseudo tetrahedral Fe^{2+} , Co^{2+} or Ni^{2+} .



Scheme 1. Synthesis of complexes **1**, **2**, and **3**.

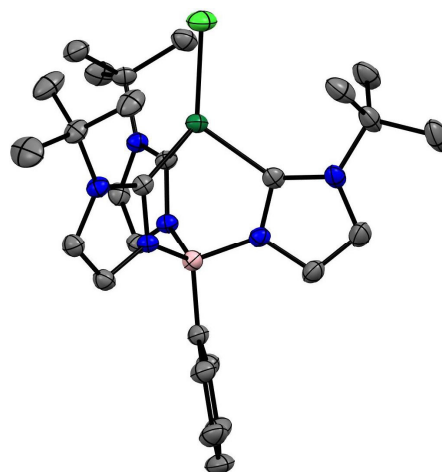


Figure 1. Solid-state structure of **1**. Ellipsoids are shown at 50% probability levels. Pink, gray, blue, neon green, and forest green ellipsoids represent B, C, N, Cl, and Zn atoms, respectively. Hydrogen atoms are omitted for clarity.

Addition of $n\text{-BuLi}$ to **1** resulted in formation of Cl exchange to form alkyl analog $\text{PhB}(\text{tBulm})_3\text{ZnBu}$ (**2**). Treatment of a solution of **1** in THF with 2.0 equiv. of a 1.6 M $n\text{-BuLi}$ solution in hexanes afforded **2** as a white powder in 70% yield. We chose to use 2.0 equiv. of $n\text{-BuLi}$ because we noticed that when only 1 equiv. of $n\text{-BuLi}$ was added that the reaction proceeded very slowly to completion. The ^1H NMR spectrum of **2** showed four new resonances at 2.31, 2.02, 1.43 and 0.93 ppm corresponding to the butyl ligand in addition to minor shifts of the $[\text{PhB}(\text{tBulm})_3]^{1-}$ ligand. The structure of **2** was confirmed by single crystal X-ray diffraction studies using colorless single crystals grown from a concentrated Et_2O solution that was placed in the $-25\text{ }^\circ\text{C}$ freezer overnight. The molecular geometry about Zn in complex **2** is best described as pseudo-tetrahedral with three carbon donor atoms from the *tris*(carbene) borate ligand and one alkyl butyl ligand. Most notable changes in the metal ligand bond distances include elongation of the average Zn- C_{Im} distances from 2.069(2) Å in **1** to 2.138(2) Å in **2**. As expected, the strong σ -donating alkyl ligand has a Zn-C bond length of 2.050(3) Å, significantly shorter compared to the Zn-Cl bond length of 2.2492(5) Å observed in **1**. This change agrees with the elongation of Zn- C_{Im} bond distances observed in **2** due to the greater electron density residing on the Zn center. Complex **2** is comprised of a Zn center with four Zn-C bonds, which to the best of our knowledge is the first Zn complex with three carbene and one alkyl-based ligands. We exclusively handled **2** under an inert atmosphere and did not observe any significant stability issues despite the electron rich nature of this complex.

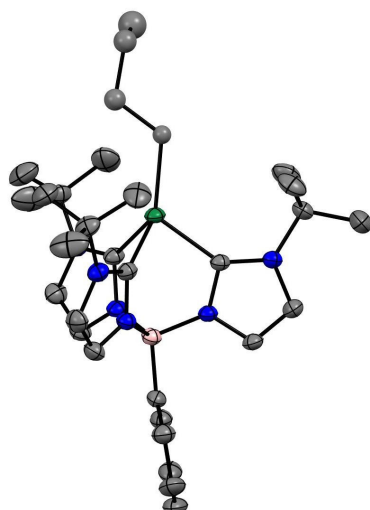


Figure 2. Solid-state structure of **2**. Ellipsoids are shown at 50% probability levels. Pink, gray, blue, and forest green ellipsoids represent B, C, N, and Zn atoms, respectively. Hydrogen atoms are omitted for clarity. The terminal butyl group in **2** is disordered over three positions, and the C atoms in this group were refined with isotropic thermal parameters.

Table 1. Selected bond distances and angles for **1**, **2**, and **3**. For **1**, X = Cl; for **2**, X = C, and for **3**, X = S.

	1	2	3
$d(\text{Zn}-\text{C}_{\text{im}1})$ (Å)	2.045(2)	2.160(2)	2.064(10)
$d(\text{Zn}-\text{C}_{\text{im}2})$ (Å)	2.075(2)	2.160(2)	2.083(9)
$d(\text{Zn}-\text{C}_{\text{im}3})$ (Å)	2.086(2)	2.093(2)	N/A
$d(\text{Zn}-\text{X})$ (Å)	2.2492(5)	2.050(3)	2.366(3), 2.360(3)

^aIn complex **3** there are two independent molecules in the asymmetric unit, and the values shown are the average of these two values.

To expand the coordination chemistry of **1**, we also investigated the reaction with electron rich RSS. Stirring **1** with 1 equiv. of $[\text{K.18-C-6}]_2[\text{S}_4]$ in THF led to a color change from a colorless to light-yellow over the course of 16 h and formation of the tetrasulfido complex $^{\text{Ph}}\text{B}(\text{tBulm})_2(\text{tBulmH})\text{Zn}(\kappa^2\text{-S}_4)$ (**3**). Compound **3** was isolated in 54% yield, which was not improved by varying the $[\text{K.18-C-6}]_2[\text{S}_4]$ stoichiometry. Complex **3** is best described as tetrahedral and features two C donor atoms from the $^{\text{Ph}}\text{B}(\text{tBulm})_2(\text{tBulmH})$ ligand and two S donor atoms from the tetrasulfide ligand. The most notable difference in **3** in comparison to the molecular structures of **1** and **2** is that it contains one protonated imidazolyliene that is no longer coordinated to the metal center. The observation of the protonated imidazolyliene is supported by changes in the ^1H NMR spectrum. We observe a reduction in symmetry that is consistent with formation of a two-fold symmetric species, with the tBu resonances at 1.72 and 1.62 ppm integrating with a 2:1 stoichiometry, respectively. Furthermore, the Im-H proton is observed at 7.55 ppm. We have previously observed similar protonation reactivity of one carbene of the $^{\text{Ph}}\text{B}(\text{tBulm})_3^{1-}$ ligand in prior work when $^{\text{Ph}}\text{B}(\text{tBulm})_3\text{NiCl}$ was treated with $[\text{K.18-C-6}]_2[\text{S}_4]$ or $[\text{NBu}_4][\text{SH}]$ to form $^{\text{Ph}}\text{B}(\text{tBulm})_2(\text{tBulmH})\text{NiS}_2$ and $^{\text{Ph}}\text{B}(\text{tBulm})_2(\text{tBulmH})\text{Ni}(\text{SH})_2$, respectively.¹⁸ Mechanistic studies revealed that the most likely source of this advantageous

proton was from unidentified species reacting with either the $[\text{NBu}_4]^+$ or $[\text{K.18-C-6}]^+$.

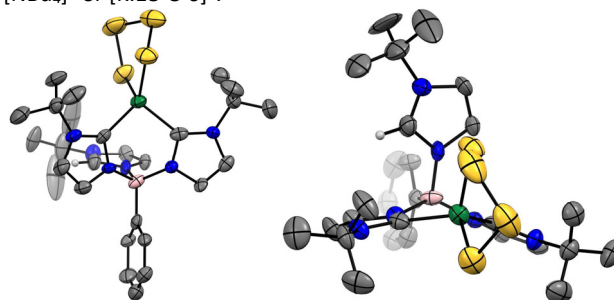


Figure 3. Side (left) and top (right) views of solid-state structure of **3**. Ellipsoids are shown at 50% probability levels. White, pink, gray, blue, yellow, and forest green ellipsoids represent H, B, C, N, S and Zn atoms, respectively. Hydrogen atoms, except Im-H, are omitted for clarity.

Complex **3** is just one of three other zinc tetrasulfido complexes that has been structurally characterized, and emphasizes the utility of the $[\text{K.18-C-6}]_2[\text{S}_4]$ reagent as an effective method for the delivery of tetrasulfides. Other methods for zinc tetrasulfido synthesis rely on the unpredictable decomposition of zinc polysulfides^{23, 24} or by the reduction of larger polysulfides with PPh_3 .²⁵ The S-S bond distances in **3** range between 2.012(4) Å and 2.100(5) Å and are consistent with other ZnS_4 motifs.²⁵ We have recently commented on the stabilizing factors of C-H $\cdots$ S interactions and upon further inspection of the structure of complex **3** we find multiple short C-H $\cdots$ S contacts that meet the definition of a hydrogen bond.²⁶ Multiple contacts between the C-H groups of the flanking tBu of the $^{\text{Ph}}\text{B}(\text{tBulm})_2(\text{tBulmH})$ ligand point directly at the sulfur atoms of the S_4^{2-} chain with some contacts as short as 2.730 Å with a C-H $\cdots$ S bond angle near 160°, which is significantly shorter than the sum of the van der Waals radii of the H and S atoms (3.0 Å). In addition to crystallographic evidence for C-H $\cdots$ S interactions in complex **3**, we also performed variable temperature ^1H NMR experiments (233 – 323 K) and observed a downfield shift of the flanking tBu and imidazolium hydrogens at lower temperatures, and an upfield shift of these resonances at elevated temperatures (see SI Figures S7 and S8). Both observations further support H-bonding to the nearby S atoms.²⁷

Conclusions

We have demonstrated the reactivity and coordination chemistry of a zinc complex with a *tris*(carbene) ligand scaffold. The chloride analog, complex **1**, performs a salt metathesis reaction with nucleophilic *n*-BuLi yielding a highly carbon rich zinc coordination complex **2** with three carbene atom donors and one carbanion donor. In contrast, reaction of complex **1** with a less nucleophilic polysulfide reagent, $[\text{K.18-C-6}]_2[\text{S}_4]$, leads to the formation of a tetrahedral zinc tetrasulfido complex via protonation of one carbene donor to form complex **3**. Taken together this work demonstrates the reactivity of the donating $^{\text{Ph}}\text{B}(\text{tBulm})_3^{1-}$ ligand when coordinated to a redox-inert metal such as zinc.

Experimental Section

Materials and Methods All manipulations were performed under an inert atmosphere using an Innovative Atmospheres N₂-filled glove box unless otherwise noted. [PhB(^tBulm)₃][OTf]₂,¹⁹ and [K18-C-6]₂[S₄]²⁷ were synthesized from established procedures. C₆D₆ was degassed with N₂ and stored in an inert atmosphere glove box over 4 Å molecular sieves. d₈-THF was degassed with N₂, distilled from Na/benzophenone, and stored in an inert atmosphere glove box over 4 Å molecular sieves. All commercially available chemicals were used as received and purchased from Strem or Sigma Aldrich. C₆D₆ and d₈-THF were purchased from Cambridge Isotope Laboratories. UV/Vis spectra were acquired on an Agilent Cary 60 UV/Vis spectrophotometer equipped with a Quantum Northwest TC-1 temperature controller set at 25.0 ± 0.05 °C. NMR spectra were acquired on a Varian 500 MHz or a Bruker 600 MHz spectrometer. Chemical shifts are reported in parts per million (δ) and are referenced to an internal standard. We did not observe the ¹³C resonance corresponding to the carbon atom bound to boron in compounds **1-3**. We attribute this absence to the quadrupolar nuclear spin of boron and expected significant signal broadened of the carbon atom. Mass spectrometric measurements were performed by the University of Illinois, Urbana Champaign MS facility. IR spectra were acquired on a Nicolet 6700 IR Spectrometer as ATR or KBr pellet samples.

X-Ray Crystallography Diffraction intensities for **1**, **2** and **3** were collected at 173 K (**1** and **2**) and 213 K (**3**) on a Bruker Apex2 CCD diffractometer using a *Incoatec* Cu *μ*S source, CuK_α radiation, 1.54178 Å. Space groups were determined based on systematic absences. Absorption corrections were applied by SADABS.²⁸ Structures were solved by direct methods and Fourier techniques and refined on *F*² using full matrix least-squares procedures. H atoms were refined in calculated positions in a rigid group model. The terminal –CH₂CH₂CH₂CH₃ group in **2** is disordered over three positions. Refinement of occupation factors shown that occupation factors for these positions are close to 0.45; 0.35 and 0.20. The final refinement was completed for a model where this terminal group is disordered over three positions and occupation factors for these positions have been fixed at the mentioned values. The C atoms in this group were refined with isotropic thermal parameters. This group was also refined with restrictions on its geometry; the standard C-C distances were used in the refinement as the targets for the corresponding C-C bonds. In the crystal structure of **3** there are three tetrahydrofuran solvent molecules filling empty spaces in the packing. Two THF molecules were found from the diffraction data and refined although there are significant elongations of thermal parameters for atoms in these solvent molecules. The third molecule is highly disordered and was treated by SQUEEZE.²⁹ Correction of the X-ray diffraction was 148 electron/cell; the required value is 160 electron/cell for four THF molecules in the full unit cell. X-ray diffraction from crystals of **3** at high angles was very weak due to the small sizes of crystals and the high disorder inside the

structure. Even by using a strong *Incoatec* Cu *μ*S source it was only possible to collect diffraction up to 2θ_{max} = 99.62°. Some atoms in the structure of **3** were refined with isotropic thermal parameters. The data collected provide the appropriate number of measured reflections per refined parameters, 7645 per 740. The structure of **3** has not been precisely determined but provides clear evidence for the chemical results. All calculations were performed by the Bruker SHELXL-2014/7 package.³⁰

Synthesis of PhB(^tBulm)₃ZnCl (1**)** A 20 mL scintillation vial containing solution of diisopropyl amine (0.251 g, 2.40 mmol) in 5 mL THF was cooled in a cold well at -78 °C in the glovebox. To this solution was added 1.55 mL (2.48 mmol) of a 1.6 M *n*-BuLi solution in hexanes to generate lithium diisopropylamide (LDA). The cooled LDA solution was added to a separate, cooled, 20 mL scintillation vial containing a slurry of [PhB(^tBulm)₃][OTf]₂ (0.567 g, 0.74 mmol) in 5 mL THF and let stir for one hour and allowed to warm to room temperature. Solid ZnCl₂ (0.10 g, 0.74 mmol) was added to the triply-deprotonated ligand solution. No immediate color change was observed, and the solution was left stirring for 1 hour. After 1 hour the THF solvent was removed under reduced pressure. The slightly yellow residue was extracted in benzene (3 mL), filtered over Celite, and the benzene was removed under reduced pressure. The residue was then heated under vacuum for 1 hour at 60 °C to remove THF coordinated to the LiOTf byproduct. The residue was then redissolved in benzene, and the purification process was repeated to remove the LiOTf(THF)_x byproduct until the ¹H NMR indicated that no more THF (in the form of LiOTf(THF)_x) was present. The fine white powder product was isolated in 37% yield (155 mg). Single crystals suitable for X-ray diffraction were grown from a concentrated THF solution layered with hexanes and placed in the -25 °C freezer overnight to yield colorless crystals of PhB(^tBulm)₃ZnCl. ¹H NMR (500 MHz, C₆D₆) δ: 8.10 (d, *J* = 6.9 Hz, 2H), 7.42 (m, 2H), 7.36 (t, *J* = 7.2 Hz, 1H), 6.97 (d, *J* = 1.7 Hz, 3H), 6.39 (d, *J* = 1.7 Hz, 3H), 1.61 (s, 27H). ¹³C NMR (126 MHz, C₆D₆) δ: 186.70, 135.91, 135.85, 122.47, 114.68, 114.54, 57.03, 31.27.

Synthesis of PhB(^tBulm)₃ZnBu (2**)** In a 20 mL scintillation vial, PhB(^tBulm)₃ZnCl (0.0114 g, 0.0203 mmol) was dissolved in THF and placed in a cold well at -78 °C in the glovebox. To this solution was added 0.025 mL (0.0400 mmol) of a 1.6 M *n*-BuLi solution in hexanes, and the solution was allowed to warm to room temperature while stirring for one hour. The solvent was removed under reduced pressure to afford a white residue that was extracted with benzene to remove salts and the filtered over a pad of Celite. The benzene was then removed under reduced pressure leaving the product in 70% yield (8.0 mg). Single crystals suitable for X-ray diffraction were grown from a concentrated Et₂O solution and placed in the -25 °C freezer overnight to yield colorless crystals of PhB(^tBulm)₃ZnBu. ¹H NMR (600 MHz, C₆D₆) δ: 8.19 (d, *J* = 7.3 Hz, 2H), 7.43 (t, *J* = 7.4 Hz, 2H), 7.36 (t, *J* = 7.2 Hz, 1H), 7.04 (s, 3H), 6.44 (s, 3H), 2.31 (dt, *J* = 16.5, 7.4 Hz, 2H), 2.02 (h, *J* = 7.2 Hz, 2H) 1.52 (s, 27H), 1.43 (t, *J* = 7.3 Hz, 3H) 0.93 (m, 2H). ¹³C NMR (151 MHz, C₆D₆) δ: 191.85,

136.17, 128.59, 127.43, 122.27, 113.99, 56.07, 34.25, 32.32, 31.34, 14.98, 14.73.

Synthesis of $\text{PhB}(\text{tBulm})_2(\text{tBulmH})\text{Zn}(\kappa^2\text{-S}_4)$ (3) In a 20 mL scintillation vial, $\text{PhB}(\text{tBulm})_3\text{ZnCl}$ (0.019 g, 0.035 mmol) was dissolved in THF. Solid $[\text{K}1.8\text{-C-6}]_2[\text{S}_4]$ (0.028 g, 0.038 mmol) was added to this solution, and it was allowed to stir at room temperature for 16 h. As the sparingly soluble orange $[\text{K}1.8\text{-C-6}]_2[\text{S}_4]$ was consumed, the reaction mixture changed from colorless to a light-yellow solution. The solvent was removed under reduced pressure and the solution was triturated with hexanes (20 mL) to remove the crown ether byproduct. The residue was then extracted in benzene and filtered over a pad of Celite to remove any salts. The benzene was removed under reduced pressure to provide a light-yellow residue that was recrystallized from a concentrated THF solution layered with hexanes and placed in the -25°C freezer overnight to yield yellow crystals of $\text{PhB}(\text{tBulm})_2(\text{tBulmH})\text{Zn}(\kappa^2\text{-S}_4)$ (15 mg, 54% yield). ^1H NMR (600 MHz, THF- d_8) δ : 7.55 (t, $J = 1.7$ Hz, 1H), 7.34 (m, 3H), 7.24 (d, $J = 1.8$ Hz, H), 7.11 (m, 2H), 6.79 (t, $J = 1.5$ Hz, 1H), 6.61 (d, $J = 1.8$ Hz, 1H), 6.50 (t, $J = 1.7$ Hz, 1H), 1.72 (s, 18H), 1.62 (s, 9H). ^{13}C NMR (151 MHz, THF) δ : 185.84, 134.35, 134.28, 129.20, 128.77, 125.07, 124.80, 121.35, 117.76, 59.62, 59.43, 32.69, 29.73. HRMS m/z $[\text{M} + \text{H}]^+$ calcd. For $[\text{C}_{27}\text{H}_{40}\text{BN}_6\text{S}_4\text{Zn}]^+$ 651.1582; found 651.1587.

Conflicts of interest

There are no conflicts to declare.

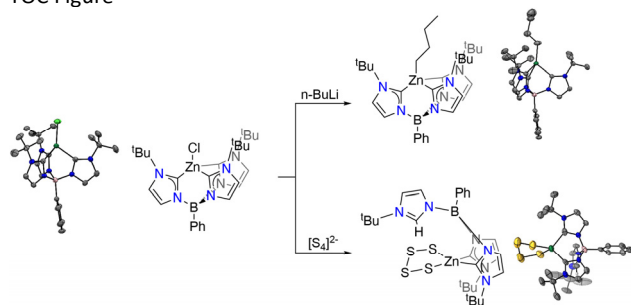
Acknowledgements

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TOC Figure



**Synthesis and Reactivity of a tris(carbene)
Zinc Coordination Complex**