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URVA and Local Mode Analysis of an Iridium Pincer Complex Efficiently Catalyzing the Hydrogenation of Carbon Dioxide

Marek Freindorf [†] and Elfi Kraka ^{*,†}

Chemistry Department, Southern Methodist University, 3215 Daniel Avenue, Dallas, TX 75275-0314, USA

* Correspondence: ekraka@smu.edu

† These authors contributed equally to this work.

Abstract: The catalytic effects of iridium pincer complexes for the hydrogenation of carbon dioxide were investigated with the Unified Reaction Valley Approach (URVA), exploring the reaction mechanism along the reaction path traced out by the reacting species on the potential energy surface. Further details were obtained with the Local Mode Analysis performed at all stationary points, complemented by the Natural Bond Orbital and Bader's Quantum Atoms in Molecules analyses. Each of the five reaction paths forming the catalytic cycle were calculated at the DFT level complemented with DLPNO-CCSD(T) single point calculations at the stationary points. For comparison, the non-catalytic reaction was also investigated. URVA curvature profiles identified all important chemical events taking place in the non-catalyzed reaction and in the five reactions forming the catalytic cycle, and their contribution to the activation energy was disclosed. The non-catalytic reaction has a large unfavorable activation energy of 76.3 kcal/mol, predominately caused by HH bond cleave in the H₂ reactant. As shown by our study, the main function of the iridium pincer catalyst is to split up the one-step non-catalytic reaction into an energy efficient multistep cycle, where HH bond cleavage is replaced by the cleavage of a weaker IrH bond with a small contribution to the activation energy. The dissociation of the final product from the catalyst requires the cleavage of an IrO bond, which is also weak, and contributes only to a minor extent to the activation energy. This, in summary, leads to the substantial lowering of the overall activation barrier by about 50 kcal/mol for the catalyzed reaction. We hope that this study inspires the community to add URVA to their repertoire for the investigation of catalysis reactions.

Keywords: unified reaction valley approach; local mode analysis; iridium-pincer complexes; catalysis; hydrogenation of carbon dioxide



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1. Introduction

One of the key causes of global warming is our excessive production of CO₂ [1–4]. Therefore, a lot of efforts have been put into the exploration of the use of this free carbon source for the commercial synthesis of chemicals and for fuel production [5–11], as well as the transformation of CO₂ into functional organic molecules [12–23]. Formic acid is one of the target molecules obtained via CO₂ hydrogenation, because of its importance in organic synthesis and its applications in hydrogen storage. Recent research has shown that formic acid can serve as a H₂ storage reservoir via its decomposition into CO₂ and H₂ and the reverse transformation [24–27]. A substantial number of investigations on CO₂ hydrogenation in homogenous catalysis have been reported over the last two decades [28–42].

The Nozaki group identified six-coordinate Ir(III)–PNP trihydride (PNP = 2,6-bis(diisopropylphosphinomethyl)pyridine) as a highly active catalyst with high turn over frequencies (TOF) and turnover number (TON), reflecting the stability of the active site for the CO₂ hydrogenation [41,43,44]. The original mechanism of the CO₂ hydrogenation to formic acid was reported to include three major steps, (i) CO₂ addition to the catalyst, (ii) formic acid dissociation, and (iii) H₂ addition in order to regenerate the catalyst [43]. In

the following years, the mechanism was extended including, e.g., basic conditions leading to a more complex catalytic cycle and a larger number of intermediates confirmed by DFT calculations [44–46]. Because the insertion of CO₂ into the metal hydride is the crucial step of the catalysis [47,48], possible CO₂ insertion into a five-coordinate Iridium(III) dihydride complex as alternative was experimentally investigated [49], showing the formation of κ^2 -bound formate monohydride products by selective electro-catalytic reduction of CO₂. The mechanism of the CO₂ insertion into five-coordinate Iridium(III) dihydrides and four-coordinate Iridium(I) hydrides was also analyzed theoretically [50], exploring the different reactivity of these complexes with different iridium oxidation states. The catalytic reduction of CO₂ to methane with silanes using Brookhart's cationic Ir(III) pincer complex was investigated theoretically [51], focusing on the transfer of silane hydrogens to CO₂ and the formation of silylformate, bis(silyl)acetal, methoxysilane and methane. The importance of a hydrogen bond donor in the secondary coordination sphere of the Ir(III)-PNP catalyst was analyzed theoretically [52], finding that the isolation of the formate product from the reaction is straight forward and therefore makes this reaction favorable. CO₂ hydrogenation was also investigated using a series of modifications of the Iridium complexes bearing sophisticated N^{*}N-bidentate ligands [53] showing importance of hydroxy groups as proton-responsive substituents in CO₂ hydrogenation reactions. Other modifications of the iridium complexes included amide-based ligands, showing that the electron-donating effects of an anionic nitrogen atom and the presence of an OH group near the metal center improves the catalytic activity for CO₂ hydrogenation [54]. Although these studies provide valuable insights, a clear understanding of the mechanistic details of each reaction step of the CO₂ hydrogenation catalytic cycle and how the catalyst actually lowers the activation barrier has been missing so far.

Therefore, we investigated in this study all major chemical events taking place during the CO₂ hydrogenation to formic acid, for the original catalytic cycle proposed by Nozaki [43], using as computational tools the Unified Reaction Valley Approach (URVA) and the Local Modes Analysis (LMA) developed in our group, which are both based on vibrational spectroscopy. URVA requires the accurate and precise determination of the reaction path starting at the transition state (TS) and moving far down into both the entrance and exit channels, a time consuming enterprise. Therefore, we simplified the original Ir(III)-PNP complex in our study by replacing the isopropyl groups with hydrogen atoms. The catalytic cycle involves the five reaction steps **R1–R5** shown in Figure 1. For comparison we also included the non-catalyzed reaction **R0**. Reaction movies illustrating the movements of all atoms of the complexes along the reaction path together with the Cartesian coordinates of the stationary points for each of the five reaction paths **R1–R5** of the catalytic cycle and the non-catalyzed reaction **R0**, both investigated in the gas phase, are presented in the Supplementary Materials.

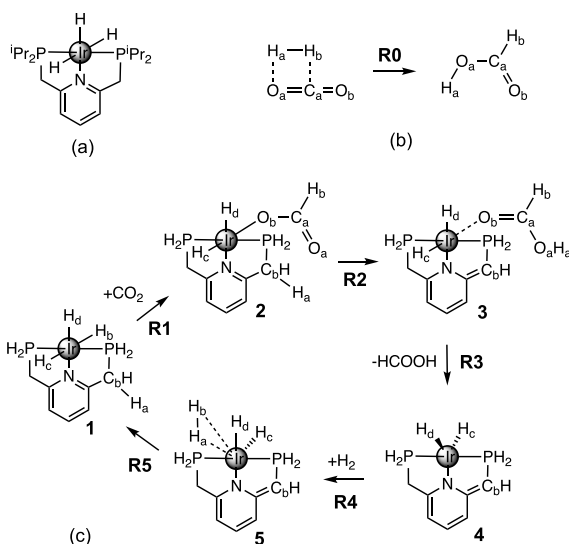


Figure 1. (a) Sketch of the Ir(III)-PNP catalyst; (b) non-catalyzed CO₂ hydrogenation to formic acid (R0); (c) Catalytic cycle of CO₂ hydrogenation reactions (R1–R5) with the model Ir(III)-PNP catalyst investigated in this work.

2. Materials and Methods

The theoretical background of URVA is provided in a comprehensive review article [55], therefore only some highlights are summarized. URVA explores the reaction complex (RC, i.e., the union of reacting molecules) along the path it traces out on the potential energy surface (PES) starting from the TS down into the entrance and exit channels, described by a large amplitude motion [56–59] as well as in the close vicinity of the reaction path, the so-called *reaction valley* being spanned by the vibrations perpendicular to the reaction path. [60,61]. The major focus of URVA is on the curvature of the reaction path. Since the reaction path is a curved line in N_{vib} dimensional space with $N_{vib} = 3N - L$ internal coordinates ($L = 6$ for a non-linear N -atomic RC and 5 for an N -atomic linear RC), its direction and curvature can be derived with the Frenet-Serret formalism [62]. The reaction path direction at a path point s is given by the unit vector $\eta(s)$:

$$\eta(s) = \frac{d\tilde{\mathbf{x}}(s)}{ds} = -\frac{\tilde{\mathbf{g}}(\tilde{\mathbf{x}}(s))}{c(s)} \quad (1)$$

where the derivative of the mass-weighted reaction coordinate $\tilde{\mathbf{x}}(s)$ with regard to s is the normalized mass-weighted gradient vector $\tilde{\mathbf{g}}(s) \equiv \tilde{\mathbf{g}}(\tilde{\mathbf{x}}(s)) = \mathbf{M}^{1/2}\mathbf{g}(s)$ and \mathbf{M} is a diagonal matrix of atomic masses. $c(s)$ is the normalization constant being equal to the length of the gradient vector $\|\tilde{\mathbf{g}}(s)\|$. The curvature vector $\kappa(s)$ is given by [63,64]

$$\kappa(s) = \frac{d^2\tilde{\mathbf{x}}(s)}{ds^2} = \frac{d\eta(s)}{ds} = \frac{-1}{\|\tilde{\mathbf{g}}(s)\|} \left(\tilde{\mathbf{f}}^x(s)\eta(s) - \left[\left(\eta(s) \right)^\dagger \tilde{\mathbf{f}}^x(s)\eta(s) \right] \eta(s) \right) \quad (2)$$

where $\tilde{\mathbf{f}}^x(s)$ is the mass-weighted Hessian matrix in Cartesian coordinates \mathbf{x} .

Any electronic structure change, e.g., bond breaking/forming processes, but also re-hybridization, charge polarization and transfer, etc. leads to distinct curvature peaks, which are directly reflected in the scalar reaction path curvature $\kappa(s) = \|\kappa(s)\|$ in contrast to regions of minimal electronic change, which are reflected by curvature minima as sketched in Figure 2 showing the scalar curvature as a function of the reaction path parameter s . We call the region from one curvature minimum to the next embedding a curvature peak a *reaction phase*. Each chemical reaction has a unique pattern of curvature maxima and minima and as such a different number of reaction phases leading to a specific reaction

profile, the so-called *fingerprint* of the reaction [55]. Deeper insights into the nature of a chemical event reflected by a curvature peak are obtained from the decomposition of the scalar reaction curvature into individual components [65] such as bond lengths, bond angles, and dihedral angles or other components such as puckering coordinates [66] denoted by the colored lines in Figure 2. Interesting to note is that even for complex catalytic reactions a curvature peak is most likely composed of a few components, which facilitates the analysis of the mechanistic details considerably. The sign of a component indicates if the parameter in question supports the chemical event (positive sign) or if it resists the chemical change (negative sign) [65]. For a detailed mathematical derivation, its connection to the work of Miller, Handy and Adams [67] and Kato and Morokuma [68] on the reaction path Hamiltonian, and recent advances of URVA, interested readers are referred to Refs [55,69]. Applications of URVA can be found, e.g., in Refs [69–84].

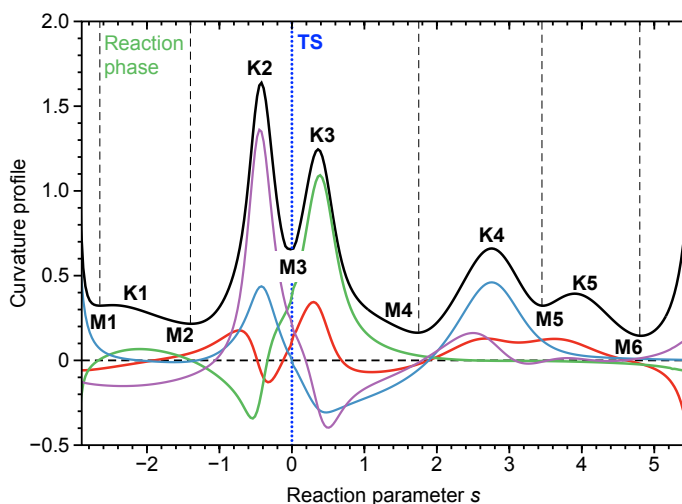


Figure 2. Schematic representation of the *reaction profile* for a model reaction defined as the scalar curvature $\kappa(s)$ (solid black line) given as a function of the reaction parameter s . Curvature minima M and curvature maxima K are shown. The location of the TS is denoted by a dotted blue line. Start and end of each reaction phase are denoted by dashed black lines. In addition the decomposition of the scalar curvature $\kappa(s)$ into four components is shown, blue, red, green and purple colored lines.

Similar as URVA, LMA is based on vibrational spectroscopy and it refines the use of normal vibrational force constants and frequencies derived from normal vibrational modes to characterize chemical bonds and/or weak chemical interactions. Normal vibrational modes of a polyatomic molecule are generally delocalized, as already stated by Wilson in 1941 via his proof that the associated normal mode coordinates are a linear combination of internal coordinates \mathbf{q} or Cartesian coordinates \mathbf{x} [85]. Therefore, normal mode stretching frequencies and associated stretching force constants are of limited use as individual bond strength descriptors. Konkoli and Cremer solved this problem via the transformation of normal vibrational modes into their local mode counterparts [86–90] (for details the reader may refer to two comprehensive review articles [91,92]). A local vibrational mode \mathbf{a}_n is defined as

$$\mathbf{a}_n = \frac{\mathbf{K}^{-1} \mathbf{d}_n^+}{\mathbf{d}_n \mathbf{K}^{-1} \mathbf{d}_n^+} \quad (3)$$

implying the important result that all what is needed for the local mode analysis are the diagonal normal mode force constant matrix \mathbf{K} in normal mode coordinates and the normal mode vectors \mathbf{d}_n in internal coordinates, which can be obtained from a standard vibrational frequency calculation via the Wilson GF formalism [93–95], a routine part of most modern quantum chemistry packages.

Once the local mode vector \mathbf{a}_n is known, one can define molecular properties corresponding to this local motion, such as local mode force constant, local mass and local frequency [86–90].

The corresponding local force constant k_n^a of local mode n (superscript a denotes an adiabatically relaxed, i.e., local mode) [96] can be expressed as

$$k_n = \mathbf{a}_n^\dagger \mathbf{K} \mathbf{a}_n = \frac{1}{\mathbf{d}_n \mathbf{K}^{-1} \mathbf{d}_n^\dagger}. \quad (4)$$

As has been shown by Zou et al., local mode stretching force constants reflect the intrinsic strength of a chemical bond and/or weak chemical interaction [97].

The associated local mass is defined as

$$m^a = \frac{1}{\mathbf{B}_n \mathbf{M}^{-1} \mathbf{B}_n^\dagger} = \frac{1}{G_{nn}} \quad (5)$$

where G_{nn} is a diagonal element of Wilson's inverse kinetic energy matrix \mathbf{G} matrix and the Wilson \mathbf{B} matrix provides the important relationship between internal and Cartesian coordinates via the first derivatives of the internal coordinates q_n ($n = 1, 2, 3 \dots N_{vib}$) with regard to the Cartesian coordinates x_i ($i = 1, 2, 3 \dots 3N$) [93],

$$\mathbf{B}_n = \frac{\delta q_n(\mathbf{x})}{\delta x_i}. \quad (6)$$

From the local mode force constant k^a and local mode mass m^a the local mode frequency ω_n can be calculated

$$(\omega^a)^2 = 1/(4\pi^2 c^2) \frac{k_n^a}{m^a} = 1/(4\pi^2 c^2) G_{nn} \mathbf{a}_n^\dagger \mathbf{K} \mathbf{a}_n \quad (7)$$

with c being the speed of light. In addition, the local mode infrared intensity has been defined which can be related to bond dipole moments [98].

LMA has been successfully applied to characterize covalent bonds [83,97,99–108] and weak chemical interactions such as hydrogen bonding [109–118], halogen bonding [119–125], pnictogen bonding [125–128], chalcogen bonding [107,125,129], tetrel bonding [130], metal bonding [115,121,131–136] as well as protein–ligand interactions [79,118,137,138].

It is convenient to transform local mode stretching force constants k^a into relative bond strength order (BSO) via a power relationship $BSO = A * (k^a)^B$ being based on the generalized Badger rule [83,139]. The dimensionless parameters A and B are obtained from two reference molecules with known BSO and k^a values for the bond in question and the requirement that for a zero k^a the BSO is zero too. We used different reference molecules for different bond types investigated in our study, for the Ir–ligand bonds we used Mayer's bond orders [140–142]. Details of the reference bonds and reference molecules are summarized in Table 1. The resulting A and B parameters for each bond are documented in the corresponding plots of the results and discussion part.

Table 1. Bond length R , local mode force constant k^a , bond strength order BSO, and energy density H_ρ evaluated at the bond critical point $\rho(r)$ of the reference bonds and reference molecules used in our study. B3LYP/6-31G(d,p)/SDD(Ir) level of theory.

Bond	$R(\text{\AA})$	k^a (mDyn/ \AA)	BSO ¹	H_ρ (Hr/Bohr ³)	Molecule
CO	1.418	4.905	1.000	−0.3719	CH ₃ OH
	1.207	13.607	2.000	−0.6883	CH ₂ O
CH, HH, OH	1.150	1.203	0.500	−0.1951	F ₂ H ^{−1}
	0.925	9.420	1.000	−0.7275	HF
IrO	2.034	2.829	1.030	−0.0300	Ir(CO) ₅ OH
	1.822	5.340	1.469	−0.1035	Ir(CO) ₄ O
IrH	2.095	0.370	0.197	−0.0110	Ir(CO) ₅ H
	1.857	0.764	0.313	−0.0219	Ir(CO) ₅ H ₂
CO ²	1.169	16.404	2.271	−0.7595	CO ₂
HH ³	0.743	5.921	0.855	−0.3360	H ₂

¹ The BSO values of bonds involving Ir atoms in the metal complexes are based on Mayer's bond orders [140–142].

² Based on the CH₃OH and CH₂O references. ³ Based on the on the HF and F₂H[−] references.

Bader's quantum theory of atoms-in-molecules (QTAIM) characterizes of chemical bonds based on the topological features of the total electron density $\rho(r)$ [143–146]. According to QTAIM a chemical bond between two atoms A and B is characterized by a bond path and a bond critical point ρ . The covalent character of the bond AB can be determined via the Cremer–Kraka criterion of covalent bonding [147–149]. According to this criterion, a chemical bond has covalent character when the energy density H_ρ , i.e., the energy density $H(r)$ taken at the at the bond critical point ρ is negative, while a positive value of H_ρ indicates an electrostatic character. $H(r) = G(r) + V(r)$; the positive kinetic energy density $G(r)$ describes electron density depletion and the negative potential energy density $V(r)$ describes electron density accumulation. Therefore, a chemical bond has covalent character when the potential energy density dominates.

We used in this study as a reaction path the intrinsic coordinate (IRC) path of Fukui [150,151] combined with the improved reaction path following procedure of Hratchian and Kraka, enabling the tracking of a chemical reaction far out into the entrance and exit channel [151]. We employed an IRC step size of 0.03 amu^{1/2}Bohr and an ultrafine grid for numerical DFT integrations. IRC calculations and the investigation of all stationary points were performed with the B3LYP density functional [152–154] utilizing Pople's 6-31G(d,p) basis set [155–157] and the Stuttgart–Dresden (SDD) effective core potential for iridium [158]. The energetics of the investigated chemical reactions was additionally recalculated via single energy point calculations with the coupled cluster DLPNO-CCSD(T) method [159] utilizing the def2-TZVP basis set [160] and the def2-ECP pseudo potential for iridium [158]. The DLPNO-CCSD(T) calculations were based on the DFT geometries and applying thermochemical corrections from the DFT frequency calculations. DFT calculations were performed with Gaussian program package [161], and DLPNO-CCSD(T) calculations with ORCA [162]. URVA was carried out using the program pURVA [163], and LMA analysis was performed using the LModeA program [164]. The Natural Bond Orbital (NBO) analysis [165–168] was applied for calculating NBO atomic charges along the reaction path. The QTAIM analysis was performed with the AIMALL program package [169].

3. Results and Discussion

3.1. Energetics

Table 2 presents the activation and reaction energies E^a , E_R and activation and reaction enthalpies H^a , H_R of reactions **R0–R5** calculated at the B3LYP/6-31G(d,p)/SDD(Ir) and DLPNO-CCSD(T)/def2-TZVP/def2-ECP(Ir) levels of theory. The energy and enthalpy values in Table 2 are taken for each reaction relative to the energy of first point on the IRC (i.e., first point in the entrance channel corresponding to a van der Waals complex). The following discussion is based on the DLPNO-CCSD(T) results. According to Table 2,

the activation energy E^a of the non-catalyzed reaction with a value of 76.3 kcal/mol is considerably larger than that for all reactions of the catalytic cycle, ranging between 1.7 kcal/mol for reaction **R4**, and 23.9 kcal/mol for reaction **R2**. Figure 3 displays the overall energetics of the catalytic cycle being composed of reactions **R1**–**R5**. All energies are relative to the energy of the van der Waals complex formed between the catalyst and CO_2 **1** in the entrance channel. All energies in Figure 3 are based on the DLPNO-CCSD(T) calculations performed at the end points of the DFT IRC calculations and for simplicity they have been calculated by adding the activation and reaction energy of the next reaction step of the cycle, to the product energy of the previous reaction step.

Table 2. The activation and reaction energy E^a , E_R and the activation and reaction enthalpy H^a , H_R of the reactions **R0**–**R5** investigated in this study ¹. The thermochemistry corrections of the DLPNO-CCSD(T) method is based on the DFT optimal geometry and frequency calculations. The B3LYP/6-31G(d,p)/SDD(Ir) and DLPNO-CCSD(T)/def2-TZVP/def2-ECP(Ir) levels of theory.

Reaction	DFT				DLPNO-CCSD(T)			
	E^a	E_R	H^a	H_R	E^a	E_R	H^a	H_R
R0	72.2	3.6	71.8	9.2	76.3	5.1	75.9	10.7
R1	14.9	−12.2	16.2	−9.3	17.6	−9.3	18.9	−6.3
R2	23.9	22.3	22.8	21.7	23.9	22.9	22.7	22.3
R3	12.3	8.7	11.3	8.0	12.2	10.5	11.3	9.7
R4	1.2	−10.3	0.6	−8.5	1.7	−13.6	1.2	−11.7
R5	20.9	−15.3	18.6	−13.8	23.5	−11.9	21.3	−10.4

¹ Energy and enthalpy values (kcal/mol) from the end points of IRC and relative to the reactants of the corresponding reaction.

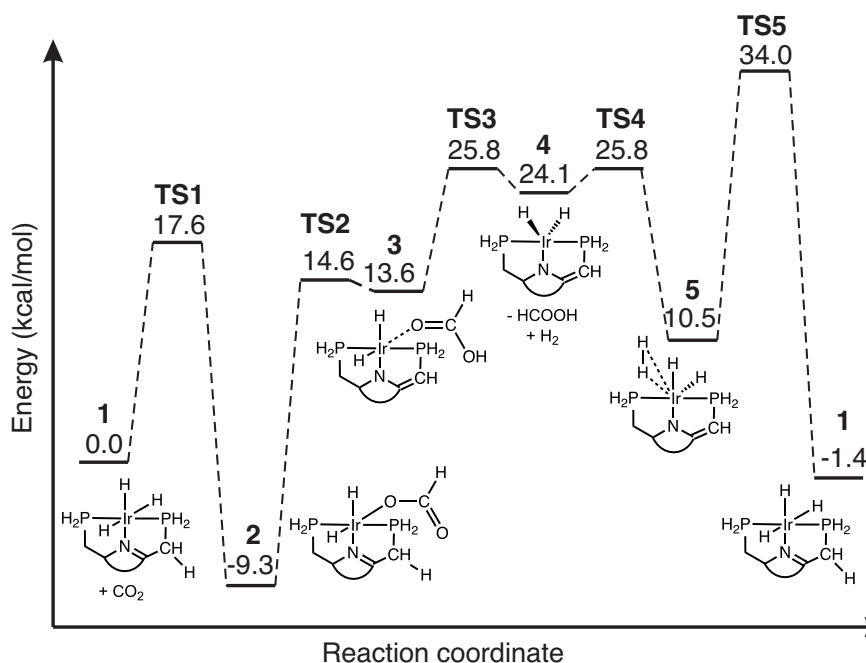


Figure 3. Energy diagram for the catalyzed reactions **R1**–**R5**, relative to the initial reactant van der Waals complex of **1**. For simplicity, the energies have been calculated by adding the activation and reaction energy of the next reaction step of the cycle, to the product energy of the previous reaction step. DLPNO-CCSD(T)/def2-TZVP/def2-ECP(Ir) level of theory performed at the end points of the DFT IRC calculations with B3LYP/6-31G(d,p)/SDD(Ir).

According to Table 2 and Figure 3, reactions **R1**— CO_2 addition ($E^a = 17.6$ kcal/mol), **R2**—hydrogen transfer ($E^a = 23.9$ kcal/mol), and **R5**—catalyst regeneration ($E^a = 23.5$ kcal/mol),

have the largest activation energies, i.e., they are the decisive steps of the overall catalytic cycle. The values of the activation enthalpies H^a for reactions **R1** (18.9 kcal/mol), **R2** (22.7 kcal/mol), and **R5** (21.3 kcal/mol) show a similar trend as seen from the activation energies E^a of these reactions, which indicates small thermal energy contributions to the overall reaction energies. Although information about the energetics is useful, it does not disclose the actual mechanism taking place in each reaction step forming the catalytic cycle. This is explored in the next section.

3.2. Reaction Mechanism

Reaction R0 Figure 4a presents the energy profile of the non-catalyzed reaction **R0** along the reaction path. The decomposition of the reaction curvature into selected components is presented in Figure 4b, selected geometrical parameters in Figure 4c and atomic NBO charges along the reaction path are presented in Figure 4d. According to Figure 4b, the reaction starts with initial formation of the O_aH_a bond in reaction phase 2 (green line), with a resisting contribution of the $O_aC_aO_b$ bond angle (purple line). Bending of CO_2 molecule is also reflected in Figure 4c. The $O_aC_aO_b$ bond angle starts to decrease in phase 2. As shown in Figure 4a, both events the initial O_aH_a bond formation and the $O_aC_aO_b$ bending, contribute 6.8 kcal/mol to the activation energy in this reaction phase. As revealed in Figure 4b, breaking of the H_aH_b bond takes place in phase 4, which is confirmed by a strong supporting contribution of the H_aH_b component (blue line). HH bond cleavage is also reflected by the change of H_aH_b bond length (see Figure 4c), which starts to elongate in phase 4, making H_a positively charged (see Figure 4d). H_aH_b bond breaking is supported by the start of C_aH_b bond formation (red line in Figure 4b), which is also confirmed by changes of C_a and H_b atomic charges (see Figure 4d). As shown in Figure 4a, the chemical events happening in phase 4, in particular the energy demanding breakage of the hydrogen bond, contributing 40.8 kcal/mol to the activation energy is the major cause of the high barrier. Finalization of O_aH_a and C_aH_b bonds takes place in phases 5 and 6 after TS (see Figure 4b), which is also reflected by O_aH_a and C_aH_b bond length changes (Figure 4c) and the NBO charge changes (Figure 4d). Obviously, the catalyzed reactions circumvents direct HH bond breakage, which is explored next.

Reaction R1 Properties of reaction **R1** of the catalytic cycle along the reaction path are presented in Figure 5a–d. Reaction **R1** involves addition of CO_2 to the coordination sphere of iridium and hydrogen atom transfer from Ir to one of the CO_2 carbon atoms. According to Figure 5b the reaction starts in phase 3 with the initial stage of the IrO_b bond formation (blue line) first resisting, followed by the next stage in phase 4 where it becomes supporting, whereas the final step of IrO_b bond formation occurs after TS in phase 9. The different stages of IrO bond formation are in line with changes in IrO_b bond lengths shown in Figure 5c. In phase 3 the IrO_b bond distance starts to decrease, reaching its final value in phase 9. Figure 5d presents the corresponding changes of Ir and O_b NBO charges taking place between reaction phases 4 and 9. The accompanying transfer of the H atom from Ir to the C_a carbon of CO_2 implies IrH_b bond cleavage and C_aH_b bond formation. According to Figure 5b, the transfer process (red line) starts at the end of phase 4, where the formation of the new C_aH_b bond (green line) is still resisting, changing into a large supporting contribution in phase 5.

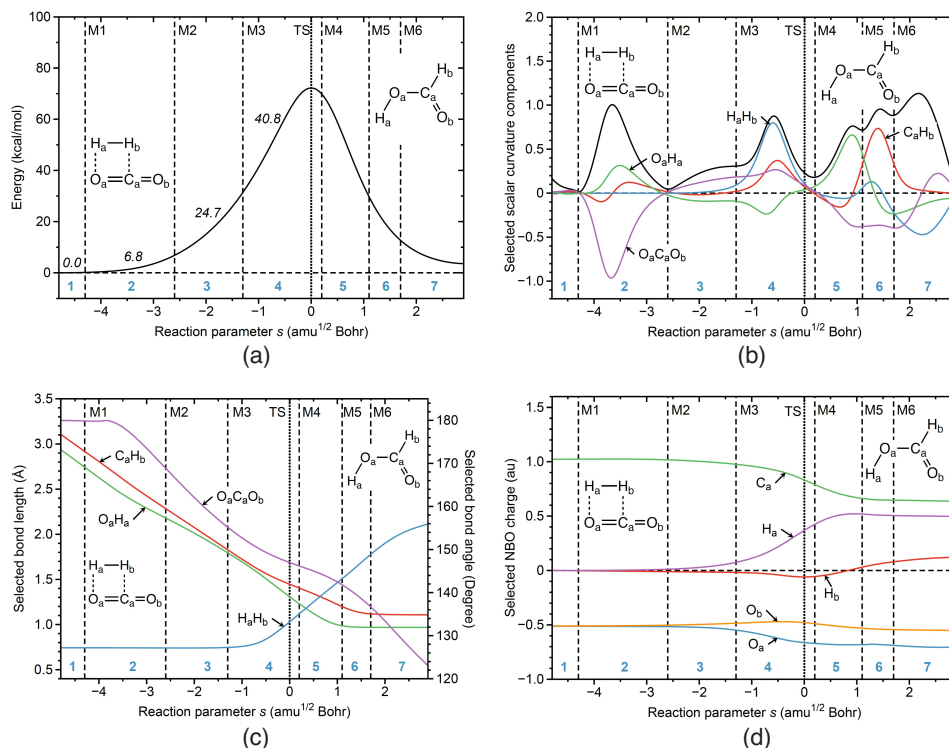


Figure 4. Properties of the gas phase reaction **R0** along the reaction path. (a) Energy profile, energy contributions of each reaction phase to the activation energy are indicated by italic numbers; (b) decomposition of the reaction curvature into selected components; (c) change of selected bond lengths along s ; (d) change of selected NBO atomic charges along s . Positions of the curvature minima are shown as dashed vertical lines and are labeled as M1, M2, and so on. The position of the TS is indicated by a dotted line. Reaction phases are indicated by blue numbers. B3LYP/6-31G(d,p)/SDD(Ir) level of theory.

The next stage of H_b transfer from Ir to C_a occurs in phase 5, being finalized far out in the exit channel in phases 8–10. Figure 5c,d confirm these processes. Largest C_aH_b bond length changes are found in phases 8 and 9 after the TS, as well as the changes in the C_a and H_b NBO charges. Figure 5b also reveals the involvement of the $\text{O}_a\text{C}_a\text{O}_b$ bond angle (purple line) in the reaction mechanism, with largest contribution between phases 4 and 9, in line with the OCO bond angle changes from 180° to 120° occurring between phases 4 and 9 (see Figure 5c) which is facilitated by the coordination of C_b to the metal atom. According to Figure 5a, the two largest energy contributions to TS energy are 4.6 kcal/mol in phase 4, caused by IrO_b and C_aH_b bond formation, together with changes of the $\text{O}_a\text{C}_a\text{O}_b$ bond angle, and 6.2 kcal/mol in phase 8, connected with IrH_b bond cleavage and IrO_b bond formation, which in summary is far less than the direct cleavage of the HH bond in reaction **R0**.

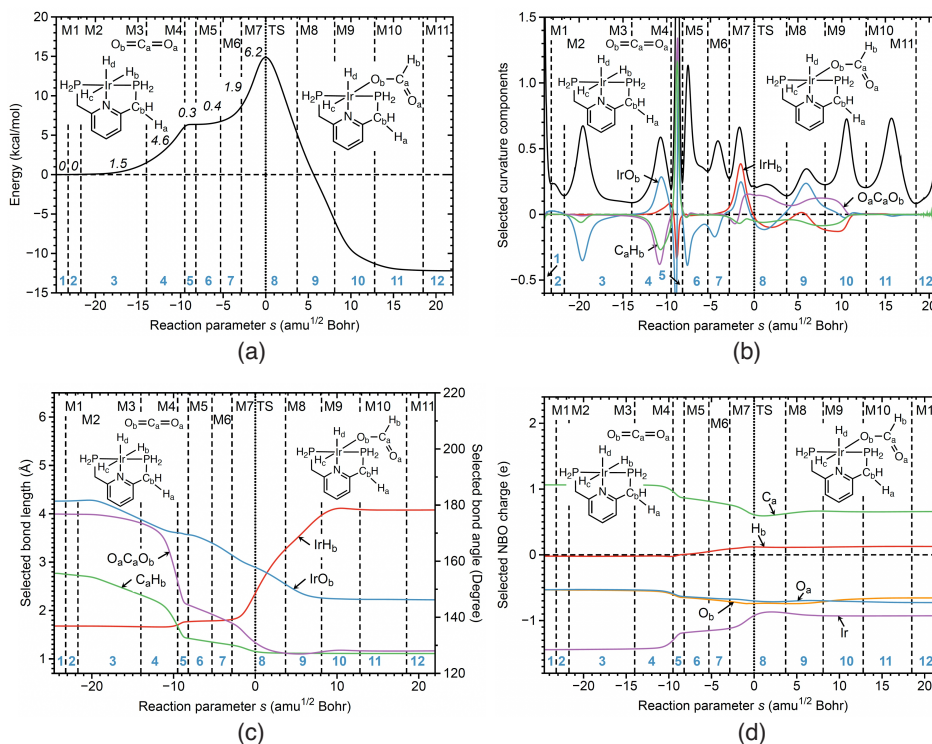


Figure 5. Properties of reaction **R1** of the catalytic cycle along the reaction path. (a) Energy profile, energy contributions of each reaction phase to the activation energy are indicated by italic numbers; (b) decomposition of the reaction curvature into selected components; (c) change of selected bond lengths along s ; (d) change of selected NBO atomic charges along s . Positions of the curvature minima are shown as dashed vertical lines and are labeled as M1, M2, and so on. The position of the TS is indicated by a dotted line. Reaction phases are indicated by blue numbers. B3LYP/6-31G(d,p)/SDD(Ir) level of theory.

Reaction R2 Reaction **R2** of the catalytic cycle is characterized by the H atom transfer from the C_b carbon atom of the catalyst to the O_a atom of CO_2 . This is a necessary step for formation of the formic acid product, which goes along with elongation of the intermediate IrO_b bond. Figure 6a–d show the corresponding URVA analysis. According to Figure 6b, C_bH_a bond cleavage (orange line) takes place in phase 5, strongly supporting, and in phase 6, strongly resisting. The C_bH_a distance starts to increase in phase 5, however it takes until the end of phase 10 to reach its final distance of 3.1 Å (see Figure 6c).

The formation of the new O_aH_a bond starts in phase 4 (red line), as depicted in Figure 6b and is finalized in phase 6, also reflected by the corresponding change of the O_aH_a distance shown in Figure 6c. What is interesting to note is that the changes of O_aH_a and C_bH_a distances complement each other. A secondary contribution to the mechanism is the change of C_aO_a from double to single bond character (purple line in Figure 6b) in phases 5 and 6. Overall, changes in the atomic charges are only marginal (see Figure 6d). According to Figure 6a, the largest contributions to the activation energy are i) O_aH_a bond formation and elongation of the IrO_b bond in phase 4 (3.8 kcal/mol) and ii) hydrogen atom transfer accompanied by the change of the CO bond from double to single bond in phase 5 (9.8 kcal/mol).

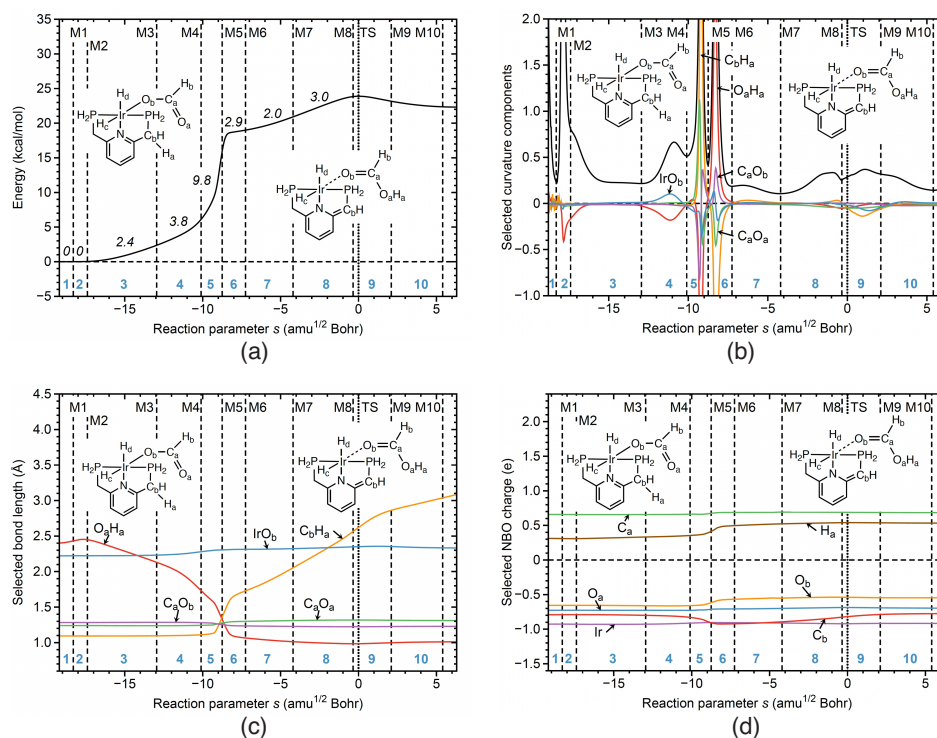


Figure 6. Properties of reaction **R2** of the catalytic cycle along the reaction path. (a) Energy profile, energy contributions of each reaction phase to the activation energy are indicated by italic numbers; (b) decomposition of the reaction curvature into selected components; (c) change of selected bond lengths along s ; (d) change of selected NBO atomic charges along s . Positions of the curvature minima are shown as dashed vertical lines and are labeled as M1, M2, and so on. The position of the TS is indicated by a dotted line. Reaction phases are indicated by blue numbers. B3LYP/6-31G(d,p)/SDD(Ir) level of theory.

Reaction R3 In reaction **R3** of the catalytic cycle, formic acid is released, accompanied by a cleavage of the intermediate IrO_b bond (already weakened in **R2**), reducing the metal coordination number from 6 to 5, which leads to a reorganization the Ir coordination sphere, in particular a change in the H_c and H_d ligand positions. The URVA results of **R3** are summarized in Figure 7a–d. The curvature profile 7b shows that IrO_b bond cleavage (blue line) starts in phase 2, with a supporting contribution. This is also reflected by a steep increase in the IrO_b distance starting in that phase, (see Figure 7c) and an increase of the Ir charge (see Figure 7d). The next stages of IrO_b bond cleavage occur in phases 3 and 4 (see IrO_b) first resisting and then supporting (phase 5, after the TS) with the final IrO_b distance of 4.1 Å reached at the end of the reaction (see Figure 7c).

According to Figure 7b, there is also a contribution from the NH_a hydrogen bond cleavage (pink line) in phase 3 (the NH_a hydrogen bond is not shown in Figure 1 for simplicity), which starts in phase 2 and is finalized in phase 6. The cleavage of the NH_a bond is also reflected in Figure 7c showing a steady increase of the NH bond distance from 1.75 Å to 3.2 Å. The change of the metal coordination sphere being related to the change of the H_c and H_d ligand positions is reflected by IrH_c (orange line) and IrH_d (red line) curvature components in phases 3 and 4. The contribution of the IrH_d component is larger, indicating a stronger electronic structure change of the metal caused by displacement of this equatorial ligand. As shown in Figure 7a, the largest energy contribution of 7.0 kcal/mol to the activation energy results from the first stage of IrO_b bond cleavage occurring in phase 2. The second stage of IrO_b bond cleavage in phase 3, together with IrH_c and IrH_d ligand

reorganization and NH_a hydrogen bond cleavage contribute 5.3 kcal/mol to the activation energy.

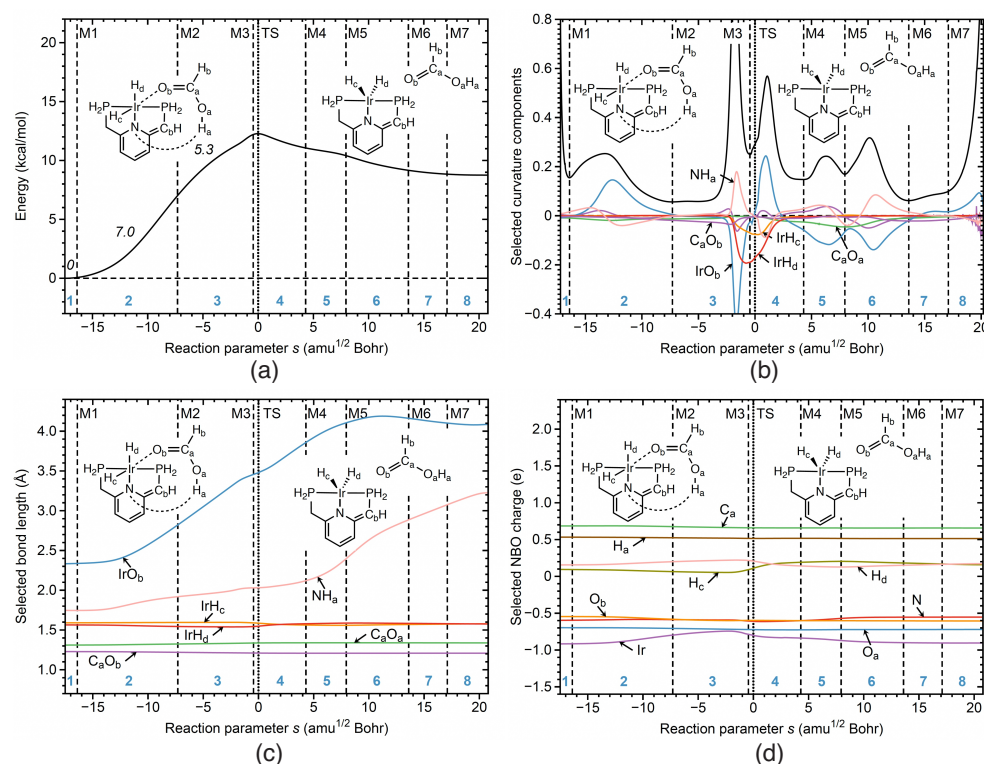


Figure 7. Properties of reaction **R3** of the catalytic cycle along the reaction path. (a) Energy profile, energy contributions of each reaction phase to the activation energy are indicated by italic numbers; (b) decomposition of the reaction curvature into selected components; (c) change of selected bond lengths along s ; (d) change of selected NBO atomic charges along s . Positions of the curvature minima are shown as dashed vertical lines and are labeled as M1, M2, and so on. The position of the TS is indicated by a dotted line. Reaction phases are indicated by blue numbers. B3LYP/6-31G(d,p)/SDD(Ir) level of theory.

Reaction R4 Figure 8a–d illustrate the change of the reaction properties along the reaction path for reaction **R4** of the catalytic cycle, which involves the addition of diatomic hydrogen to the coordination sphere of the metal via $\eta^2\text{-H}_2$ coordination; the first part of restoring the catalyst, which implies an increase of the Ir coordination number from 5 to 6. According to Figure 8b the IrH_a (green line) and IrH_b (blue line) bond components contribute to the reaction curvature in phases 2, 3, and 5. As depicted in Figure 8c IrH_a and IrH_b distances start to decrease in phase 2. Although the starting H_a distance is 4.2 Å, considerably larger than the H_b distance of 3.45 Å, both bonds are finalized simultaneously at the end of phase 5. There is also a minor contribution of the H_aH_b component (red line) to the reaction curvature in phase 5 (see Figure 8b), reflecting changes of the H_aH_b bond in this phase, where the $\eta^2\text{-H}_2$ complexation to the metal is finalized. The change of the metal coordination sphere in this reaction phase leads to an increase of negative charge on Ir atomic atom, as shown in Figure 8d. The activation energy of reaction **R4** is relatively small, (DFT result: $E^a = 1.2$ kcal/mol, Figure 8a), mainly resulting from electronic structure reorganization of the IrH_a and IrH_b bonds taking place in phase 3.

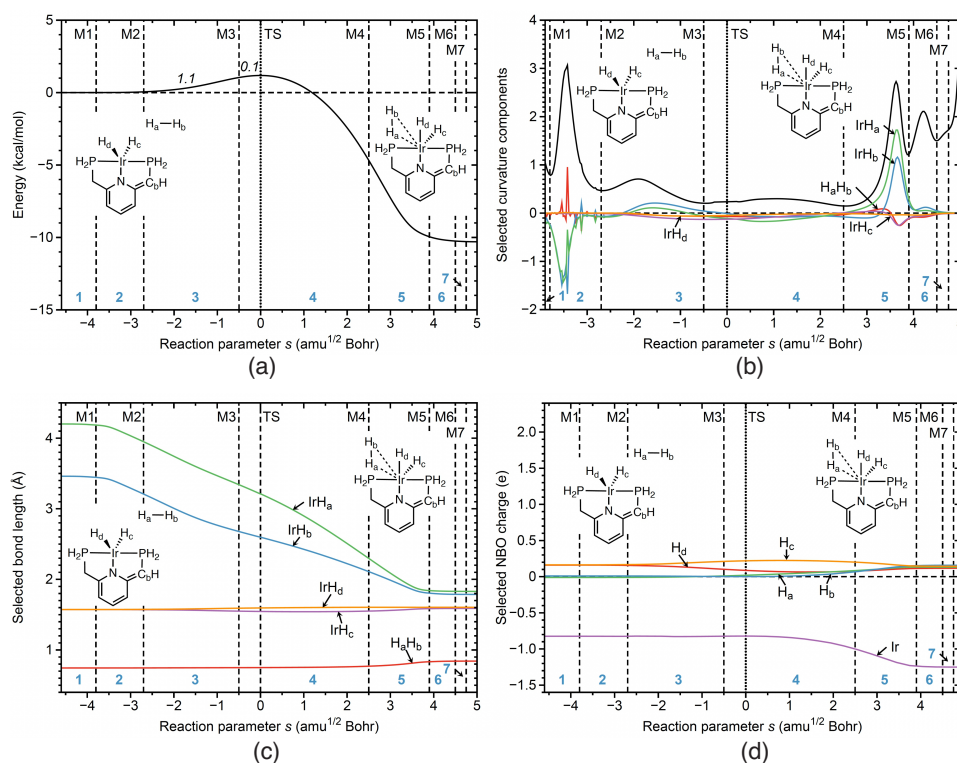


Figure 8. Properties of reaction **R4** of the catalytic cycle along the reaction path. (a) Energy profile, energy contributions of each reaction phase to the activation energy are indicated by italic numbers; (b) decomposition of the reaction curvature into selected components; (c) change of selected bond lengths along s ; (d) change of selected NBO atomic charges along s . Positions of the curvature minima are shown as dashed vertical lines and are labeled as M1, M2, and so on. The position of the TS is indicated by a dotted line. Reaction phases are indicated by blue numbers. B3LYP/6-31G(d,p)/SDD(Ir) level of theory.

Reaction R5 In reaction **R5** the initial catalyst is recovered, which implies H_aH_b bond cleavage of the $\eta^2\text{-H}_2$ coordinated H_2 ligand, IrH_a bond cleavage, migration of H_a to C_b and C_bH_a bond formation. The corresponding URVA analysis is illustrated in Figure 9a–d. Figure 9b shows that IrH_a bond cleavage (orange line) starts in phase 4 and proceeds through the following phases until phase 10. As reflected in Figure 9b, the initial IrH_a distance of 1.8 Å starts to increase in phase 4 until its final value of 3.5 Å is reached at the end of the reaction. H_aH_b bond cleavage (red line) starts at the border between phases 3 and 4, dominates phase 6 shortly before the TS with a large supporting contribution and proceeds parallel to IrH_a bond cleavage until phase 10. It is interesting to note that elongation of the H_aH_b bond starts in phase 6 (see Figure 9c) which parallels charge separation, H_a migrating to the C_b carbon atom to form a CH bond becomes more positively charged and H_b forming a stronger IrH bond, more negatively charged (see Figure 9d). According to Figure 9b, C_bH_a bond formation (green line) starts in phase 3 and is finalized in phase 7, after the TS and is revealed by a large curvature contribution.

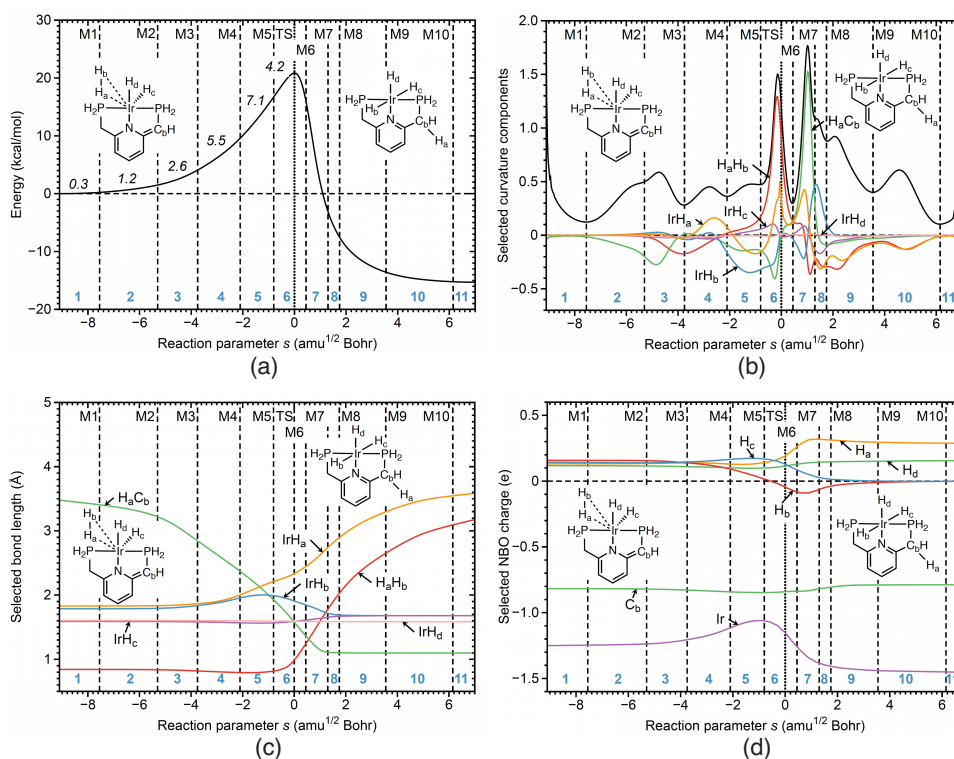


Figure 9. Properties of reaction R5 of the catalytic cycle along the reaction path. (a) Energy profile, energy contributions of each reaction phase to the activation energy are indicated by italic numbers; (b) decomposition of the reaction curvature into selected components; (c) change of selected bond lengths along s ; (d) change of selected NBO atomic charges along s . Positions of the curvature minima are shown as dashed vertical lines and are labeled as M1, M2, and so on. The position of the TS is indicated by a dotted line. Reaction phases are indicated by blue numbers. B3LYP/6-31G(d,p)/SDD(Ir) level of theory.

In phase 7 the C_bH_a bond length reached its final value, as shown in Figure 9c. According to Figure 9a, the largest energy contributions to the activation energy occur in reaction phases 4 (5.5 kcal/mol) and 5 (7.1 kcal/mol), corresponding to IrH_a bond cleavage and strengthening of the IrH_b bond (see Figure 9b). In addition, it is important to note that H_aH_b bond cleavage contributes only 4.2 kcal/mol to the activation energy in the reaction phase 6, assisted by the charge polarization of both H atoms.

3.3. Chemical Bond Analysis

The URVA analysis was complemented with LMA and QTAIM bond analyses for all stationary points of reactions R0–R5 in order to provide further insights into the mechanism and to confirm the URVA analysis. Table 3 presents bond length R , local mode force constant k^a and related bond strength order BSO and energy density H_ρ for those bonds of reactants RE, transition states TS and products PR which change their character during the reaction. In order to limit the number of entries reported in this table, we assumed that the product of the previous catalyzed reaction is the same as the reactant of the next reaction in the catalytic cycle, although the IRC calculations did finish at slightly different geometries.

Table 3. Bond length R, local mode force constant k^a , bond strength order BSO, and energy density H_p of the selected bonds for reactant RE, transition state TS, and product PR of reactions **R0–R5**. For the bond label description, see the text. B3LYP/6-31G(d,p)/SDD(Ir) level of theory.

Bond	R (Å)			k^a (mDyn/Å)			BSO			H_p (Hr/Bohr ³)		
	RE	TS	PR	RE	TS	PR	RE	TS	PR	RE	TS	PR
R0												
C _a O _a	1.169	1.260	1.353	16.364	10.621	5.972	2.267	1.690	1.143	−0.7594	−0.6537	−0.4670
C _a O _b	1.169	1.176	1.198	16.402	14.830	13.994	2.271	2.120	2.039	−0.7597	−0.7484	−0.7315
C _a H _b	3.128	1.447	1.108	0.019	0.865	4.805	0.124	0.447	0.797	-	−0.0666	−0.2896
O _a H _a	2.959	1.308	0.969	0.017	1.686	8.005	0.119	0.560	0.947	0.0010	−0.1021	−0.5918
H _a H _b	0.743	1.030	2.120	5.898	1.210	0.469	0.854	0.501	0.364	−0.3364	−0.0861	-
R1												
C _a O _a	1.170	1.242	1.240	16.098	9.679	10.562	2.242	1.587	1.684	−0.7449	−0.6444	−0.6432
C _a O _b	1.170	1.268	1.283	16.039	8.344	8.253	2.236	1.435	1.424	−0.7436	−0.6083	−0.5787
C _a H _b	2.769	1.153	1.111	0.041	3.090	4.692	0.160	0.687	0.791	0.0014	−0.2259	−0.2828
IrO _b	4.262	2.893	2.222	0.049	0.415	1.184	0.107	0.352	0.633	-	-	−0.0038
IrH _b	1.679	2.359	4.079	1.983	0.415	-	0.576	0.212	-	−0.0574	−0.0003	-
R2												
C _a O _a	1.240	1.318	1.311	10.562	7.086	7.134	1.684	1.284	1.290	−0.6432	−0.5196	−0.5320
C _a O _b	1.283	1.224	1.229	8.253	11.827	11.316	1.424	1.818	1.765	−0.5787	−0.6731	−0.6671
C _a H _b	1.111	1.096	1.097	4.692	5.302	5.270	0.791	0.824	0.822	−0.2828	−0.3133	−0.3115
C _b H _a	1.094	2.623	3.084	5.237	0.138	0.117	0.821	0.241	0.228	−0.2896	-	-
O _a H _a	2.403	0.985	1.011	0.029	6.893	4.053	0.143	0.900	0.753	0.0020	−0.5652	−0.4934
IrO _b	2.222	2.351	2.333	1.184	0.573	0.635	0.633	0.422	0.447	−0.0038	0.0002	−0.0002
R3												
C _a O _a	1.311	1.337	1.338	7.134	6.260	6.280	1.290	1.180	1.183	−0.5320	−0.4924	−0.4907
C _a O _b	1.229	1.210	1.209	11.316	12.883	13.069	1.765	1.927	1.946	−0.6671	−0.6978	−0.6967
IrO _b	2.333	3.476	4.085	0.635	0.076	0.050	0.447	0.136	0.108	−0.0002	0.0004	-
IrH _c	1.564	1.548	1.576	3.139	3.129	2.758	0.772	0.770	0.710	−0.0981	−0.1078	−0.0949
IrH _d	1.590	1.586	1.578	2.833	2.737	2.728	0.723	0.707	0.706	−0.0874	−0.0900	−0.0939
NH _a	1.747	2.031	3.225	0.219	0.129	0.029	0.282	0.236	0.143	−0.0040	−0.0007	-
R4												
H _a H _b	0.745	0.750	0.842	5.744	5.371	1.826	0.847	0.828	0.575	−0.3330	−0.3225	−0.2020
IrH _a	4.201	3.210	1.829	0.004	0.071	0.730	0.011	0.069	0.304	-	-	-
IrH _b	3.461	2.597	1.787	0.004	0.050	0.807	0.011	0.055	0.324	0.0004	0.0001	−0.0275
IrH _c	1.576	1.545	1.589	2.758	3.318	2.783	0.710	0.799	0.715	−0.0949	−0.1091	−0.0889
IrH _d	1.578	1.595	1.604	2.728	2.760	2.762	0.706	0.711	0.711	−0.0939	−0.0859	−0.0829
R5												
H _a H _b	0.842	0.986	3.182	1.826	0.707	0.037	0.575	0.418	0.155	−0.2020	−0.0907	-
IrH _a	1.829	2.335	3.588	0.730	0.355	0.107	0.304	0.192	0.089	-	-	-
IrH _b	1.787	1.909	1.679	0.807	0.567	1.987	0.324	0.259	0.576	−0.0275	−0.0202	−0.0581
IrH _c	1.589	1.591	1.679	2.783	2.737	1.988	0.715	0.707	0.576	−0.0889	−0.0863	−0.0581
IrH _d	1.604	1.592	1.585	2.762	2.842	2.914	0.711	0.724	0.736	−0.0829	−0.0869	−0.0894
C _b H _a	3.481	1.580	1.098	0.105	0.932	5.118	0.220	0.459	0.814	-	−0.0371	−0.2792

Reaction R0 Reaction **R0** includes HH bond cleavage, formation of a new CH and OH bond, and the change of CO bond character in CO₂. There are five bonds involved in this reaction, namely C_aO_a, C_aO_b, C_aH_b, O_aH_a, and H_aH_b (see Table 3). The C_aO_a bond of CO₂ is changing its character from a double bond in RE to a single bond in PR, which is confirmed by the decreasing value of k^a (16.364, 10.621, and 5.972 mDyn/Å for RE, TS and PR, respectively) which is in line with a bond length increase (1.169, 1.260, and 1.353 Å for RE, TS and PR, respectively) and the decreasing covalent character (−0.7594, −0.6537, and −0.4670, Hr/Bohr³ for RE, TS and PR, respectively). In contrast, the C_aO_b bond of CO₂ keeps its double character during this reaction, however, its strength is slightly decreasing (16.402, 14.830, and 13.994 mDyn/Å, for RE, TS and PR, respectively), the bond length is slightly increasing (1.169, 1.176, and 1.198 Å for RE, TS and PR, respectively), and the covalent character is slightly decreasing (−0.7597, −0.7484, and −0.7315 Hr/Bohr³ for RE,

TS and PR, respectively). Two bonds are formed in this reaction, namely C_aH_b and O_aH_a , while the H_aH_b bond is cleaved. Consequently, the strength of the C_aH_b bond is increasing (0.019, 0.865, and 4.805 mDyn/Å for RE, TS and PR, respectively), its length is decreasing (3.128, 1.447, and 1.108 Å for RE, TS and PR, respectively), and its covalent character is increasing (−0.0666 and −0.2896 Hr/Bohr³, for TS and PR, respectively). Similarly, the strength of the O_aH_a bond is increasing (0.017, 1.686, and 8.005 mDyn/Å for RE, TS and PR, respectively), its length is decreasing (2.959, 1.308, and 0.969 Å for RE, TS and PR, respectively), and its covalent character is increasing (0.0010, −0.1021, and −0.5918 Hr/Bohr³ for RE, TS and PR, respectively). Finally, the H_aH_b bond is cleaved in this reaction, which is confirmed by its decreasing strength (5.898, 1.210, and 0.469 mDyn/Å for RE, TS and PR, respectively), by its increasing length (0.743 1.030, and 2.120 Å for RE, TS and PR, respectively), and by its decreasing covalent character (−0.3364 and −0.0861 Hr/Bohr³ for RE and TS, respectively). In summary, the LMA and QTAIM analyses taken at the stationary points are completely in line with the URVA analysis.

Reaction R1 Reaction **R1** involves addition of CO₂ to the metal center forming the IrO_b bond, H atom transfer from the metal to the CO₂ ligand via cleavage of the IrH_b bond, and the formation of the C_aH_b bond. Five bonds are involved in this process, namely C_aO_a , C_aO_b , C_aH_b , IrO_b, and IrH_b (see Table 3). The C_aO_a bond of CO₂ keeps more or less its double character during reaction **R1**, although its strength is somewhat decreasing as reflected by the k^a values (16.098, 9.679, and 10.562 mDyn/Å for RE, TS and PR, respectively), increasing bond length (1.170, 1.242, and 1.240 Å for RE, TS and PR, respectively) and decreasing covalent character (−0.7449, −0.6444, −0.6432 Hr/Bohr³ for RE, TS and PR, respectively). In contrast, the C_aO_b bond of CO₂ is changing its character from a double to a single bond upon CO₂ addition to the metal. The strength of the C_aO_b bond is decreasing (16.039, 8.344, and 8.253 mDyn/Å for RE, TS and PR, respectively), the bond length is increasing (1.170 1.268, and 1.283 Å for RE, TS and PR, respectively) and the covalent character is decreasing (−0.7436, −0.6083, and −0.5787 Hr/Bohr³ for RE, TS and PR, respectively). The changes of the C_aO_b bond are a consequences of the IrO_b bond formation in this reaction. The strength of the IrO_b bond is consequently increasing (k^a = 0.049, 0.415, and 1.184 mDyn/Å for RE, TS and PR, respectively) and its length is decreasing (4.262, 2.893, and 2.222 Å for RE, TS and PR, respectively). We observed only for the PR IrO_b bond a bond critical with a small covalent character (−0.0038 Hr/Bohr³). The reaction **R1** also involves a H atom transfer from the catalyst to the incoming CO₂ ligand, where the IrH_b bond of the catalyst is cleaved, and the new C_aH_b bond is formed. Consequently, the strength of the IrH_b bond is decreasing (k^a = 1.983 and 0.415 mDyn/Å for RE and TS, respectively), its length is increasing (1.679, 2.359, and 4.079 Å for RE, TS and PR, respectively), and its covalent character is decreasing (−0.0574 and −0.0003 Hr/Bohr³ for RE and TS, respectively). The formation of the C_aH_b bond is reflected by increasing bond strength (0.041, 3.090, and 4.692 mDyn/Å for RE, TS and PR, respectively), decreasing bond length (2.769 1.153, and 1.111 Å for RE, TS and PR, respectively) and increasing covalent character (0.0014, −0.2259, and −0.2828 Hr/Bohr³ for RE, TS and PR, respectively). In conclusion, we observe that the results of the LMA and QTAIM analyses of the reaction **R1** are in support of the results of the URVA analysis, similarly as for the reaction **R0**.

Reaction R2 Reaction **R2** involves H atom transfer from the catalyst to the CO₂ ligand and subsequent OH bond formation implying the cleavage of a catalyst CH bond. There are six bonds involved in this process, C_aO_a , C_aO_b , C_aH_b , C_bH_a , O_aH_a , and IrO_b (see Table 3). As the H_a atom is transferred from C_b to the CO₂ ligand, the strength of the C_bH_a bond is decreasing (k^a = 5.237, 0.138, and 0.117 mDyn/Å for RE, TS and PR, respectively), and its length is consequently increasing (1.094, 2.623, and 3.084 Å for RE, TS and PR, respectively). We observed only a bond critical point for the RE C_bH_a bond with a relatively large covalent character (−0.2896 Hr/Bohr³). The change of O_aH_a bond properties are going into the opposite direction. The O_aH_a bond strength is increasing (k^a = 0.029, 6.893, and 4.053 mDyn/Å for RE, TS and PR, respectively), the bond length is decreasing (2.403,

0.985, and 1.011 Å for RE, TS and PR, respectively), and the covalent character is increasing (0.0020, −0.5652, and −0.4934 Hr/Bohr³ for RE, TS and PR, respectively). It is interesting to note that at the TS the O_aH_a bond is stronger (6.893 mDyn/Å), shorter (0.985 Å) and more covalent (−0.5652 Hr/Bohr³) than in the PR. As the reaction proceeds, the C_aO_a bond is changing its character from double bond to a single bond, which is reflected in the RE, TS, and PR bond properties. The strength of the C_aO_a bond is decreasing (k^a = 10.562, 7.086, and 7.134 mDyn/Å for RE, TS and PR, respectively), its length is increasing (1.240, 1.318, and 1.311 Å for RE, TS and PR, respectively), and its covalent character is decreasing (−0.6432, −0.5196, and −0.5320 Hr/Bohr³ for RE, TS and PR, respectively). C_aO_b bond properties are going in the opposite direction, as the character of this bond is changing from a single to a double bond. The strength of the C_aO_b bond is increasing (k^a = 8.253, 11.827, and 11.316 mDyn/Å for RE, TS and PR, respectively), its length is decreasing (1.283, 1.224, and 1.229 Å for RE, TS and PR, respectively), and its covalent character is increasing (−0.5787, −0.6731, and −0.6671 Hr/Bohr³ for RE, TS and PR, respectively). The change of the C_aO_b bond properties, affects in turn the properties of the IrO_b bond. The IrO_b bond becomes weaker (k^a = 1.184, 0.573, and 0.635 mDyn/Å for RE, TS and PR, respectively), longer (2.222, 2.351, and 2.333 Å for RE, TS and PR, respectively), and less covalent (−0.0038, 0.0002, and −0.0002 Hr/Bohr³ for RE, TS and PR, respectively). The strength of the C_aH_b bond increases only slightly, as reflected by its k^a values (4.692, 5.302, and 5.270 mDyn/Å for RE, TS and PR, respectively), its bond length (1.111, 1.096, and 1.097 Å for RE, TS and PR, respectively), and its covalent character (−0.2828, −0.3133, and −0.3115 Hr/Bohr³ for RE, TS and PR, respectively). Again, these results are in line with the URVA analysis.

Reaction R3 Reaction **R3** involves the dissociation of the HCOOH product from the metal coordination sphere via IrO_b bond cleavage, affecting also the properties of both C_aO_a and C_aO_b bonds. In addition, this process changes the metal coordination numbers from 6 to 5, which is accompanied by the change of H_c and H_d atomic positions and the change of IrH_c and IrH_d bond properties. As such, six bonds are involved, C_aO_a, C_aO_b, IrO_b, IrH_c, IrH_d, and NH_a (see Table 3). The IrO_b bond strength is decreasing (k^a = 0.635, 0.076, and 0.050 mDyn/Å for RE, TS and PR, respectively), the bond becomes longer (2.333, 3.476, and 4.085 Å for RE, TS and PR, respectively) and less covalent (−0.0002 and 0.0004 Hr/Bohr³ for RE and TS, respectively; no bond critical point was found for PR). Cleavage of the IrO_b bond affects the properties of the C_aO_b bond, which becomes stronger (k^a = 11.316, 12.883, and 13.069 mDyn/Å for RE, TS and PR, respectively), shorter (1.229, 1.210, and 1.209 Å for RE, TS and PR, respectively), and more covalent (−0.6671, −0.6978, and −0.6967 Hr/Bohr³ for RE, TS and PR, respectively). In contrast, the C_aO_a bond becomes weaker (k^a = 7.134, 6.260, and 6.280 mDyn/Å for RE, TS and PR, respectively), longer (1.311, 1.337, and 1.338 Å for RE, TS and PR, respectively) and less covalent (−0.5320, −0.4924, and −0.4907 Hr/Bohr³ for RE, TS and PR, respectively), although the overall changes of C_aO_a bond properties are smaller than those for the C_aO_b bond. Change of the metal coordination number from 6 to 5 during this reaction implies a change of the H_c and H_d ligand positions in the metal coordination sphere, which affects the IrH_c and IrH_d bond properties. The IrH_c bond becomes weaker (3.139, 3.129, and 2.758 mDyn/Å for RE, TS and PR, respectively), longer (1.564, 1.548, and 1.576 Å for RE, TS and PR, respectively), and less covalent (−0.0981, −0.1078, and −0.0949 Hr/Bohr³ for RE, TS and PR, respectively). However, although the IrH_d bond becomes somewhat weaker (k^a = 2.833, 2.737, and 2.728 mDyn/Å for RE, TS and PR, respectively), it becomes shorter (1.590, 1.586, and 1.578 Å for RE, TS and PR, respectively), and slightly more covalent (−0.0874, −0.0900, and −0.0939 Hr/Bohr³ for RE, TS and PR, respectively). The values of the IrH_c and IrH_d PR bond properties are similar (k^a = 2.758 and 2.728 mDyn/Å, R = 1.576 and 1.578 Å, and H_p = −0.0949 and −0.0939 Hr/Bohr³, for IrH_c and IrH_d, respectively), which is consistent with their similar positions in the metal coordination sphere. There is an additional NH bond involved in reaction **R3** with hydrogen bond character N...H_aO_a, which is cleaved during the reaction. The strength of the NH_a bond continuously decreases (k^a = 0.219, 0.129, and 0.029 mDyn/Å for RE, TS and PR, respectively), the length increases (1.747, 2.031, and 3.225 Å for RE, TS and PR,

respectively), and the covalent character decreases (-0.0040 and -0.0007 Hr/Bohr³ for RE and TS R, respectively, no bond critical point was found for PR. In conclusion, the results of the LMA and QTAIM analyses are consistent with the results of the URVA analysis.

Reaction R4 Reaction **R4** involves the η^2 -H₂ coordination of H₂ to the metal center, the first step of catalyst recovery. The coordination number of the metal is changed back from 5 to 6 and two new bonds are formed between the metal and H₂, i.e., IrH_a and IrH_b, changing the properties of the H_aH_b bond from the incoming H₂ ligand. H₂ addition also changes properties of the IrH_c and IrH_d bonds caused by the change of the H_c and H_d ligand positions in the coordination sphere. Five bonds are involved in this process, H_aH_b, IrH_a, IrH_b, IrH_c and IrH_d (see Table 3). The strength of the IrH_a bond is increasing ($k^a = 0.004$, 0.071 , and 0.730 mDyn/Å for RE, TS and PR, respectively) and its length is decreasing (4.201 , 3.210 , and 1.829 Å for RE, TS and PR, respectively). No bond critical point could be found for any of the three stationary points. Similarly, the strength of the IrH_b bond is increasing as well ($k^a = 0.004$, 0.050 , and 0.807 mDyn/Å for RE, TS and PR, respectively), the length is decreasing (3.461 , 2.597 , and 1.787 mDyn/Å for RE, TS and PR, respectively), and its covalent character is increasing (0.0004 , 0.0001 , and -0.0275 Hr/Bohr³ for RE, TS and PR, respectively). The H₂ addition to the metal changes properties of the H_aH_b bond, which becomes weaker ($k^a = 5.744$, 5.371 , and 1.826 mDyn/Å for RE, TS and PR, respectively), longer (0.745 , 0.750 , and 0.842 Å for RE, TS and PR, respectively) and less covalent (-0.3330 , -0.3225 , and -0.2020 Hr/Bohr³ for RE, TS and PR, respectively). This is important for the final step of the catalytic cycle, in which the H_aH_b bond has to be cleaved. The H_c and H_d atoms are moving into new positions, affecting the IrH_c bond which becomes slightly stronger ($k^a = 2.758$, 3.318 , and 2.783 mDyn/Å for RE, TS and PR, respectively), slightly longer (1.576 , 1.545 , and 1.589 Å for RE, TS and PR, respectively), and less covalent (-0.0949 , -0.1091 , and -0.0889 Hr/Bohr³ for RE, TS and PR, respectively). Similar changes were found for the IrH_d bond. The bond strength is increasing ($k^a = 2.728$, 2.760 , and 2.762 mDyn/Å for RE, TS and PR, respectively), the bond length is increasing (1.578 , 1.595 , and 1.604 Å for RE, TS and PR, respectively), and the covalent character is decreasing (-0.0939 , -0.0859 , and -0.0829 Hr/Bohr³ for RE, TS and PR, respectively). These results fully support the URVA analysis.

Reaction R5 The last reaction of the catalytic cycle involves the dissociation of the H_aH_b bond from the H₂ ligand, the transfer of the H_a atom from the metal center to the C_b carbon atom of the catalyst accompanied by the cleavage of the IrH_a bond, and the strengthening of the two IrH_b and C_bH_a bonds, regenerating the original catalyst. Six bonds are involved in this process, H_aH_b, IrH_a, IrH_b, IrH_c, IrH_d and C_bH_a (see Table 3). The strength of the already weakened H_aH_b bond decreases further ($k^a = 1.826$, 0.707 , and 0.037 mDyn/Å for RE, TS and PR, respectively), its length increases (0.842 , 0.986 , and 3.182 Å for RE, TS and PR, respectively), and its covalent character decreases (-0.2020 and -0.0907 Hr/Bohr³ for RE and TS, respectively, no bond critical point was found for PR). The strength of the IrH_a bond is decreasing ($k^a = 0.730$, 0.355 , and 0.107 mDyn/Å for RE, TS and PR, respectively) and its length is increasing (1.829 , 2.335 , and 3.588 Å for RE, TS and PR, respectively) without bond critical points in this reaction. C_bH_a bond formation is confirmed by increasing bond strength ($k^a = 0.105$, 0.932 , and 5.118 mDyn/Å for RE, TS and PR, respectively), decreasing bond length (3.481 , 1.580 , and 1.098 Å for RE, TS and PR, respectively) and increasing covalent character (-0.0371 and -0.2792 Hr/Bohr³ for TS and PR, respectively, no bond critical point was found for RE). Strengthening of the IrH_b bond is reflected by increasing k^a values (0.807 , 0.567 , and 1.987 mDyn/Å for RE, TS and PR, respectively), decreasing bond length (1.787 , 1.909 , and 1.679 Å for RE, TS and PR, respectively), and increasing covalent character (-0.0275 , -0.0202 , and -0.0581 Hr/Bohr³ for RE, TS and PR, respectively). The strength of the IrH_c bond is decreasing ($k^a = 2.783$, 2.737 , and 1.988 mDyn/Å for RE, TS and PR, respectively), the length is increasing (1.589 , 1.591 , and 1.679 Å for RE, TS and PR, respectively), and the covalent character is decreasing (-0.0889 , -0.0863 , and -0.0581 Hr/Bohr³ for RE, TS and PR, respectively) whereas the properties of the IrH_d bond

are changing in the opposite direction. The strength of this bond increases ($k^a = 2.762, 2.842$, and 2.914 mDyn/Å for RE, TS and PR, respectively), the length decreases ($1.604, 1.592$, and 1.585 Å for RE, TS and PR, respectively), and the covalent character increases (-0.0829 , -0.0869 , and -0.0894 Hr/Bohr³ RE, TS and PR, respectively). Again, these results fully support the URVA analysis.

In conclusion, the LMA and QTAIM analyses performed at all stationary points of reactions **R0–R5** deepen the insights obtained with URVA by adding bond specific details. In the last section, overall trends of the LMA and QTAIM bond properties based on Table 3 data are presented to round up the results and discussion part.

3.4. Comparison of Relative Bond Strength Orders BSO

Local mode force constants k^a can be transformed into relative bond strength orders BSO, which assesses bond strength in a more commonly used language referring to single and multiple bond character of a bond, and in this way offering a more convenient way to compare bond strengths in a series of molecules. BSO values were obtained as described in the computational section, based on the reference molecules summarized in Table 1. Figure 10a–d show BSO values as a function of the local mode force constants k^a , for the CO, CH, HH, and OH bonds and Figure 11a–b for IrO and IrH bonds. There are two types of CO bonds, C_aO_a and C_aO_b , which are changing their bond strengths in reactions **R0–R3**. According to Figure 10a, largest CO BSO values are in the range of 2.2 indicating strong double bond character. The strongest CO bonds were found for the RE of the non-catalyzed reaction **R0** with BSO values close to the BSO of 2.271 of the CO bond in the CO₂ molecule (see Table 1). The smallest BSO values are in the range of 1.2 indicating single bond character, with the TS values in between. The weakest CO bond is found again for the PR of the non-catalyzed reaction **R0**.

Figure 10b shows the BSO values for C_aH_b and C_bH_a bonds of the catalyst, which are changing in reactions **R0–R2**, and **R5**. For example, the dissociating C_bH_a bond of reaction **R2** starts with an RE BSO value of 0.821, which is close to a single bond strength and drops down in the PR to a value of 0.228, indicating a weak CH interaction, already shown in the TS with a BSO value of 0.241. The BSO values for the H_aH_b bonds are presented in Figure 10c for reactions **R0**, **R4**, and **R5**. The HH bond of the RE for both non-catalyzed reaction **R0** and catalyzed reaction **R4** are the strongest with BSO values in the range of 0.85, which are close to the BSO value of 0.855 of the HH bond of the free H₂ molecule (see Table 1). The weakest HH interaction is found for the RE of reaction **R5** with a BSO value of 0.575, facilitating the cleavage of this bond. There is one OH formed in reactions **R0** and **R2**. The small RE BSO value of 0.119 reflects that there is only a weak interaction in the RE van der Waals complex. The BSO value increases to 0.560 at the TS and the PR BSO of 0.947 indicates OH single bond character upon OH bond formation. Figure 11a–b present plots of the BSO values for the IrO and IrH bonds. The IrO_b bonds change in reactions **R1–R3**. Corresponding BSO values are presented in Figure 11a. For example, in reaction **R1** the IrO_b bond is formed, starting with a small RE BSO value of 0.107, which increases to 0.352 for the TS and to 0.633 for the PR. It is interesting to note that the overall BSO values of all IrO bonds in this study are smaller than 1, indicating relatively weak bond strength character. Figure 11b shows the BSO values of the IrH bonds, such as IrH_a, IrH_b, IrH_c, and IrH_d, which are formed, cleaved, and changed in reactions **R1**, and **R3–R5**. For example, in reaction **R3**, upon dissociation of the HCOOH ligand, the IrH_c bond is only changing slightly its character, with BSO values of 0.772 for the RE, 0.770 for the TS, and 0.710 for the PR.

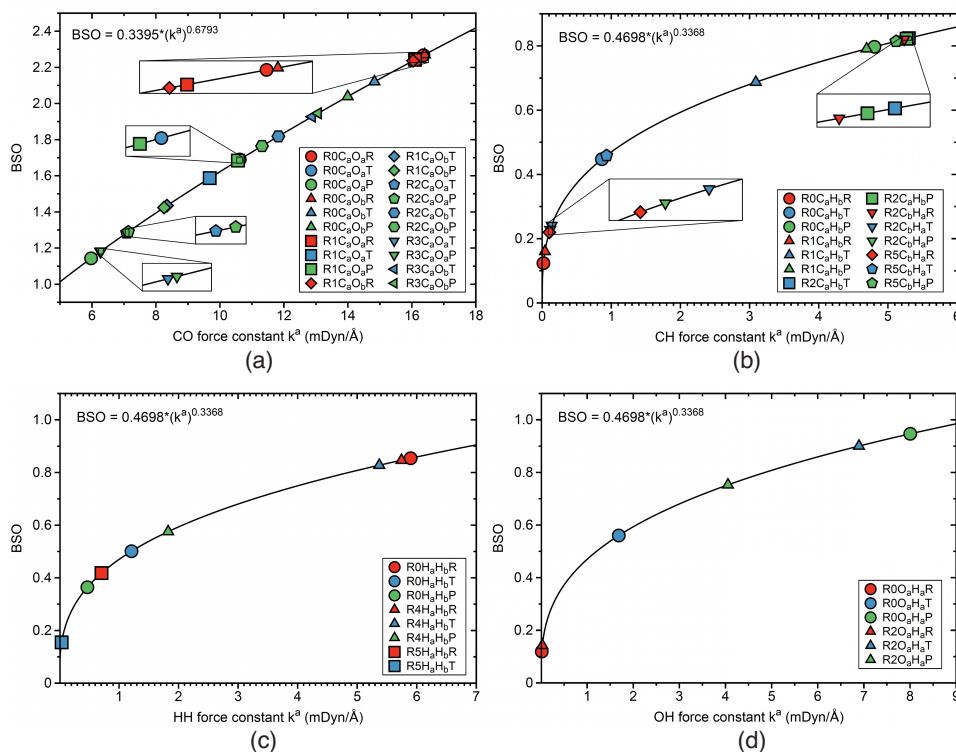


Figure 10. BSO values as a function of the local mode force constants k^a , calculated from the power relationship $BSO = A * (k^a)^B$; for details, see text and Table 1. (a) CO bonds; (b) CH bonds; (c) HH bonds; (d) OH bonds. Labels $R0C_aO_aR$, $R0C_aO_aT$, and $R0C_aO_aP$ indicate reactant RE (red color), TS (blue color), and product PR (green color) of the bond involving the C_a and O_a atoms of reaction **R0**. Similarly notation for the reactions **R1–R5**. For the bond label description see the text. B3LYP/6-31G(d,p)/SDD(Ir) level of theory.

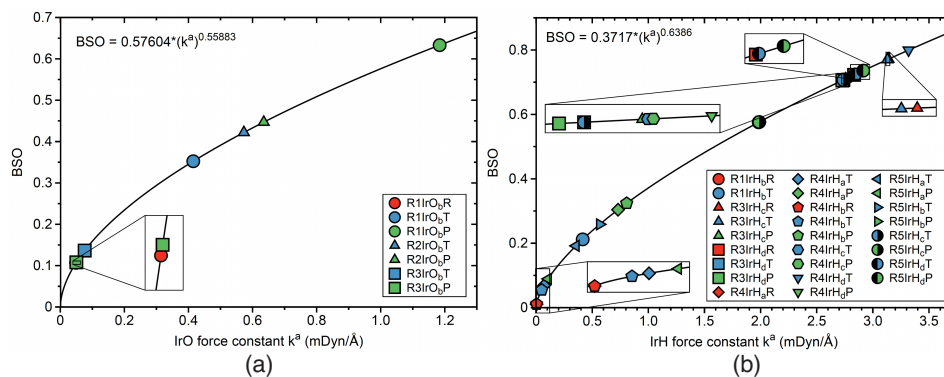


Figure 11. BSO values as a function of the local mode force constants k^a , calculated from the power relationship $BSO = A * (k^a)^B$; for details, see text and Table 1. (a) IrO bonds; (b) IrH bonds. Labels $R1IrO_bR$, $R1IrO_bT$, and $R1IrO_bP$ indicate reactant RE (red color), TS (blue color), and product PR (green color) of the bond involving the Ir and O_b atoms of thereaction **R1**. Similarly notation for reactions **R2–R5**. For the bond label description see the text. B3LYP/6-31G(d,p)/SDD(Ir) level of theory.

3.5. Correlations between Bond Properties and Local Mode Force Constants

Figures 12 and 13 present the correlation between bond strength expressed by the local mode force constants and bond length, which according to the Badger rule [139,170] leads to a power relationship. Figure 12a shows the correlation for CO bonds of reactions R0–R3.

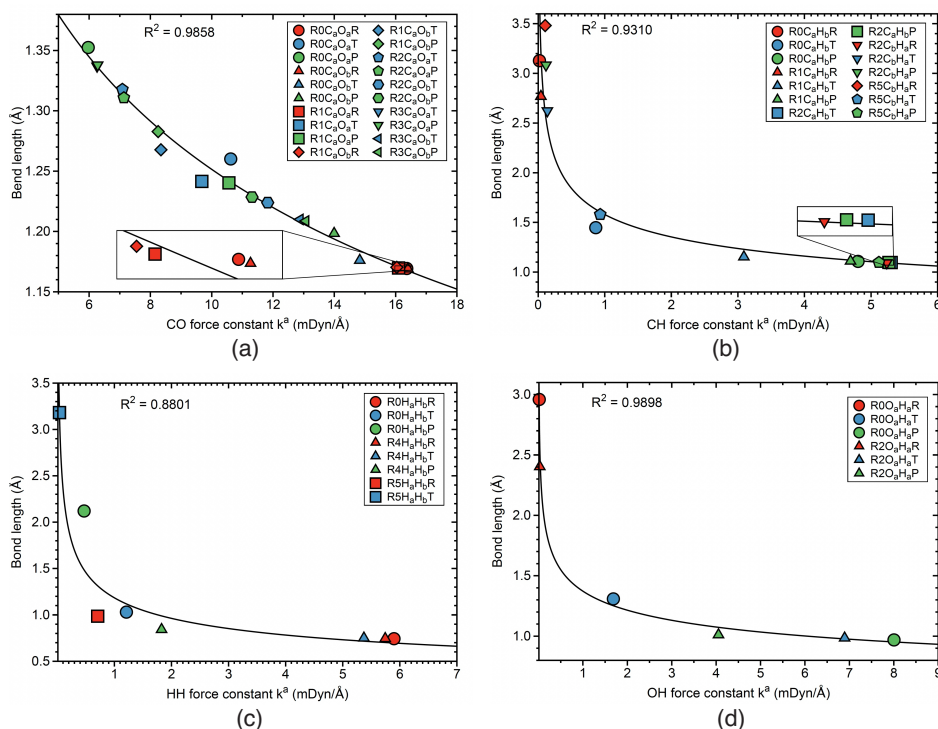


Figure 12. Correlation between bond lengths and local mode force constants. (a) CO bonds; (b) CH bonds; (c) HH bonds; (d) OH bonds. Labels such as R0C_aO_aR, R0C_aO_aT, and R0C_aO_aP indicate reactant RE (red color), TS (blue color), and product PR (green color) C_aO_a bond in reaction R0. Similarly notation for reactions R1–R5. For the bond label description see the text. B3LYP/6-31G(d,p)/SDD(Ir) level of theory.

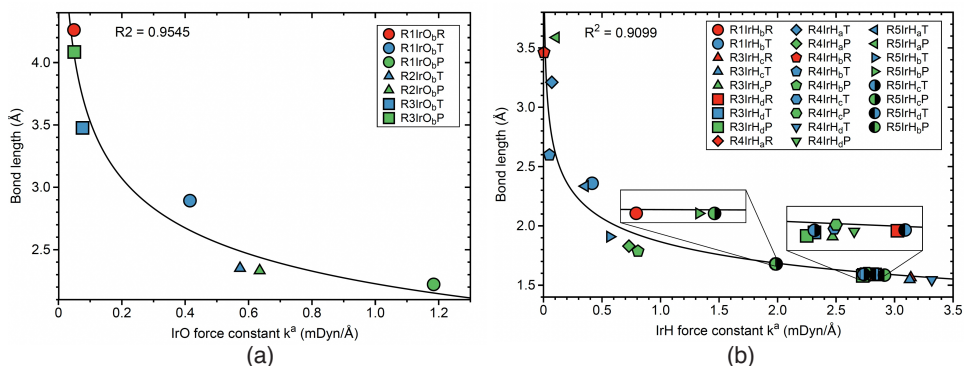


Figure 13. Correlation between bond length and the local mode force constants. (a) IrO bond; (b) IrH bonds. Labels such as R1IrO_bR, R1IrO_bT, and R1IrO_bO indicate reactant RE (red color), TS (blue color), and product PR (green color) IrO_b bond in reaction R1. Similarly notation for reactions R2–R5. For the bond label description see the text. B3LYP/6-31G(d,p)/SDD(Ir) level of theory.

According to Figure 12a, stronger CO bonds are connected to shorter CO bonds. The R² value of 0.9858 confirms a significant correlation between these two properties. larger

divergencies are observed for the C_aO_a bond at TS of reaction **R0**, and C_aO_a and C_aO_b bonds at the TSs of reaction **R1**. Figure 12b shows a similar correlation between bond length and bond strength for the CH bonds in reactions **R0–R2**, and **R5**. The R^2 factor of this correlation has a somewhat lower value of 0.9310, corresponding to a good correlation. The divergency occurs for the C_aH_b bond at the TS of reaction **R0**. The correlation between bond strength and bond length for the HH bonds are presented in Figure 12c, showing the results for the H_aH_b bonds in reactions **R0**, **R4**, and **R5**. This correlation has an R^2 value of 0.8801, with outliers for PRs of reactions **R0** and **R4** and the RE of reaction **R5**. Figure 12d shows a significant correlation between bond strength and bond length for the O_aH_a bonds of reactions **R0** and **R2**. The value of the R^2 factor of this correlation is 0.9898 with no substantial outlier. The bond strength bond versus length correlation for the IrO_b bonds in reactions **R1–R3** is presented in Figure 13a. The R^2 factor of this correlation has a moderate value 0.9545, with the largest outlier for the TSs of reactions **R1** and **R2**. Figure 13b shows the correlation between bond strength and bond length for the IrH bonds of reactions **R1**, and **R3–R5**. The R^2 value of 0.9099 indicates a moderate correlation, smaller than that found for the IrO bonds. Largest outliers are found for the IrH_b TS bond of reaction **R1**, the IrH_a and IrH_b TS bonds of reactions **R5**, and for the PR IrH_a and IrH_c bonds of reaction **R4**. In summary, we observed a satisfactory correlation between bond strength and bond length, in line with the popular rule that the stronger bond is also the shorter bond. However, this is not always the case as demonstrated by the increasing number of exceptions from this simple rule, reported in the literature [87,88,90,171,172]. The outliers in the correlations presented in this work are mostly related to the TS stationary points. At the TS, in particular for bond breaking/forming processes, significant electronic structure rearrangements take place, e.g., leading to a stage between two different hybridization forms, which can cause an unusual bonding situation that cannot be captured by such an empirical correlation.

Energy density versus Local mode force constant As outlined above, the energy density H_ρ is convenient measure of the covalent character of a chemical bond or weak chemical interaction based on the electron density. Overall, we observed in our study that the stronger bonds have more covalent character, expressed by a more negative value of H_ρ . Figures 14 and 15 present the correlations between H_ρ and k^a based on a linear relationship. Figure 14a shows the correlation for the C_aO_a and C_aO_b bonds of reactions **R0–R3**. The R^2 factor has a value of 0.9411, with a number of small outliers mostly for TSs. The bond strength- energy density correlation for the C_aH_b and C_bH_a bonds of reactions **R0–R2**, is presented in Figure 14b. The R^2 value of 0.9758 reflects a good correlation, with the largest outlier for the TS C_aH_b bond of reaction **R1**. Figure 14c shows the same correlation for the H_aH_b bonds of reactions **R0**, **R4**, and **R5**, with a R^2 value of 0.9438. The largest outlier was found for the PR H_aH_b bond of reaction **R4**. The correlation between energy density and bond strength for the O_aH_a bonds of reactions **R0** and **R2** is shown in Figure 14d. The R^2 factor for this reaction, has a value of 0.9256 reflecting a moderate correlation with the largest outlier for the PR O_aH_a bond of reaction **R2**. Figure 15a presents the correlation between H_ρ and k^a for the IrO_b bonds in reactions **R1–R3**. As obvious from the R^2 factor of 0.7816, there is only some trend between these two properties. One also has to consider the scarcity of the data points. It is interesting to note that the PR IrO_b bond of reaction **R2** has an H_ρ close to zero, indicating a more electrostatic character of this bond. A similar situation occurs for the TS IrO_b bond of reaction **R3**. A better correlation between energy density and bond strength is observed for the IrH bonds of reactions **R1** and **R3–R5** shown in Figure 15b. The R^2 factor of that correlation has a value of 0.9829, with some outlier such as the TS IrH_b bond of the reaction **R1**, or the TS IrH_c bond of reaction **R3**.

In summary, overall, we observed a good correlation between energy density H_ρ and bond strength expressed by the local mode force constant k^a , which indicates that the strength of the bonds investigated in this work is predominantly based on their covalent character. However, a caveat is appropriate. H_ρ reflects bonding at just one point of the electron density, namely the bond critical point, whereas the k^a as a second order property picks up the electronic environment [91]. As such it is more sensitive in the case of complex

bonding situation as, e.g., found for the IrO bonds showing only a trend but not a significant correlation between these two properties.

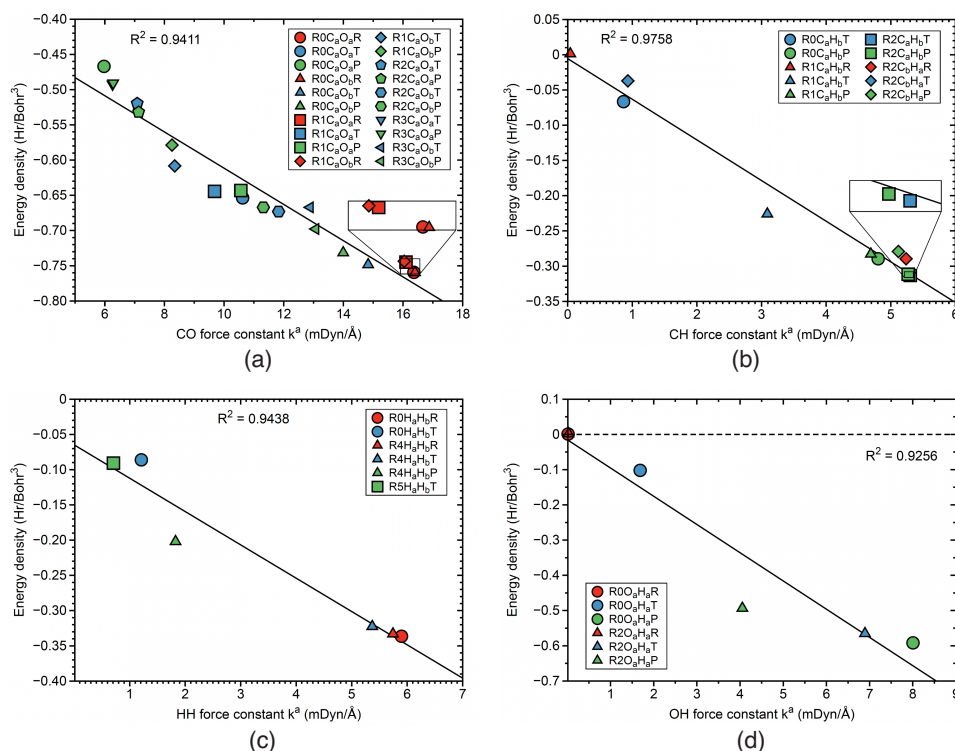


Figure 14. Correlation between energy density H_p and local mode force constant k^a . (a) CO bonds; (b) CH bonds; (c) HH bonds; (d) OH bonds. Labels such as $R0C_aO_aR$, $R0C_aO_aT$, and $R0C_aO_aP$ indicate reactant RE (red color), TS (blue color), and product PR (green color) C_aO_a in reaction **R0**. Similar notation for reactions **R1–R5**. For the bond label description see the text. B3LYP/6-31G(d,p)/SDD(Ir) level of theory.

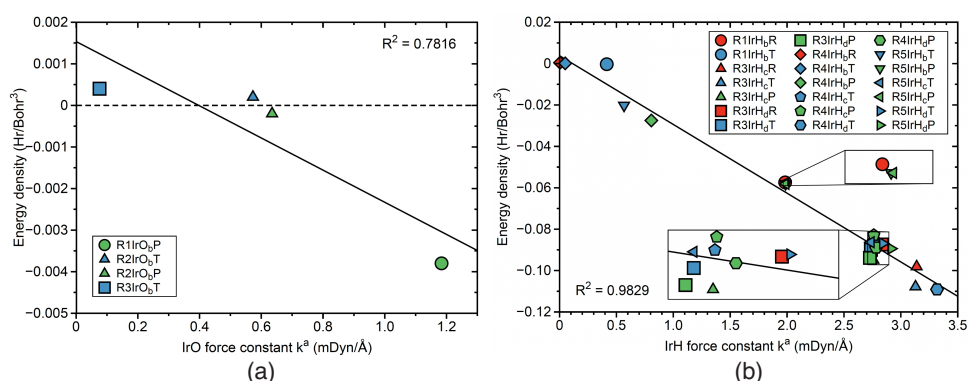


Figure 15. Correlation between energy density H_p and local mode force constant k^a . (a) IrO bonds; (b) IrH bonds. label such as $R1IrO_bR$, $R1IrO_bT$, and $R1IrO_bP$ indicate reactant RE (red color), TS (blue color), and product PR (green color) IrO_b bond in **R1**. Similar notation for reactions **R2–R5**. For the bond label description see the text. B3LYP/6-31G(d,p)/SDD(Ir) level of theory.

4. Conclusions and Outlook

We investigated in this study the catalytic cycle (reactions **R1–R5**) of carbon dioxide hydrogenation catalyzed with an iridium pincer model complex, which reduces the unfavorable activation energy of 76 kcal/mol for the non-catalyzed gas phase reaction **R0**

by more than 100%. As a computational tool we used URVA to explore the mechanistic features of each reaction step in all detail, shedding new light on which are the most energy consuming events in the non-catalyzed reaction, and how are they avoided in the catalyzed reaction. URVA was complemented by LMA and QTAIM analyses at all stationary points of each reaction. We used the DFT level of theory for URVA, LMA, and QTAIM, whereas for the energetics a dual approach (i.e., single point DLPNO-CCSD(T) energy calculations for all stationary points based on the DFT geometries) was applied. The most important findings of our study showing for the first time how the catalyst effectively manages to keep energy demand at a balance are summarized in the following list:

- According to the URVA analysis, the most energy consuming process of the non-catalyzed reaction **R0** is the cleavage of the HH bond of the H₂ reactant, which takes place before TS, and as such contributes to the activation energy with 40.8 kcal/mol.
- In order to avoid direct HH bond cleavage, the catalyst divides the one-step non-catalytic reaction into a multistep catalytic cycle; CO₂ addition to the catalyst, H atom transfer between catalyst and CO₂ ligand, release of HCOOH product, addition of H₂, reorganization of the coordination sphere of the catalyst to achieve its original form. Each of these decisive steps could be clearly identified in the characteristic curvature profiles for the first time.
- The expensive cleavage of the HH bond in the non-catalyzed reaction, is replaced in the catalyzed reaction by H transfer (reaction **R2**) requiring the cleavage of an IrH bond with a significantly smaller contribution of 9.8 kcal/mol to the activation energy, which is revealed by both energy and curvature profiles.
- The dissociation of the final product from the catalyst (reaction **R3**) is characterized by the cleavage of an IrO bond and an intermediate NH hydrogen bond, and according to our URVA analysis, both contribute to the activation energy with a moderate amount of 12.3 kcal/mol.
- As unravelled by the URVA curvature profiles most of the events related to the reorganization of the catalyst to restore its original form (reaction **R5**) occur after the TS, i.e., they do not contribute to the activation energy.

In summary, the URVA analysis complemented with the LMA and QTAIM methods provided new comprehensive mechanistic details of all reaction steps forming the catalytic cycle of the hydrogenation of carbon dioxide, which hopefully will stimulate and inspire the community working on this important topic. Experiments suggest that the catalytic activity of iridium pincer complexes for the hydrogenation of carbon dioxide can be considerably increased under basic reaction conditions [43,44]. A follow-up URVA study is planned to explore this.

Supplementary Materials: The following supporting information can be downloaded at: <https://www.mdpi.com/article/10.3390/inorganics10120234/s1>, containing reaction movies **R0–R5** and the Cartesian coordinates of the stationary point, i.e, the reactant complex, TS and product complex for all six reactions **R0–R5**.

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Abbreviations

The following abbreviations are used in this manuscript:

URVA	unified reaction valley approach
DFT	density functional theory
DLPNO-CCSD(T)	domain-based local pair natural orbital of coupled cluster single and double with perturbative triple excitations
PNP	2,6-bis(di-isopropylphosphinomethyl)pyridine
TOF	turnover frequency
TON	turnover number
LMA	Local Modes Analysis
PES	potential energy surface
RC	reaction complex
QTAIM	quantum theory of atoms-in-molecules
IRC	intrinsic reaction coordinate
NBO	natural bond orbital
BSO	bond strength order
Hr	hartree
RE	reactant
TS	transition state
PR	product

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