

# Environmental Science Atmospheres

Accepted Manuscript

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Understanding how secondary aerosols form is extremely important as aerosols' impact on Earth's climate is one of the main sources of uncertainty for understanding global warming. The beginning stages for the formation of prenucleation complexes, that eventually lead to larger aerosols, cannot currently be investigated experimentally. We have determined the lowest Gibbs Free Energy clusters using robust computational chemistry methods, which allows for the prediction of the equilibrium concentrations of the sulfuric acid-formic acid-hydrochloric acid-ammonia-dimethylamine-water system. Comparisons to the sulfuric acid-formic acid-nitric acid-ammonia-dimethylamine-water system reveal that nitric acid drives dry nucleation better than hydrochloric acid, which is more stabilized by hydration. This indicates that different acids are able to stabilize prenucleation clusters at different stages of their growth.



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Received 00th January 20xx,  
Accepted 00th January 20xx

DOI: 10.1039/x0xx00000x

# The Driving Effects of Common Atmospheric Molecules for Formation of Clusters: The Case of Sulfuric Acid, Formic Acid, Hydrochloric Acid, Ammonia, and Dimethylamine

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**Abstract** One of the main sources of uncertainty for understanding global warming is understanding the formation of larger secondary aerosols. The beginning stages start with the formation of prenucleation complexes from precursor monomers of acids, bases, and organic molecules. The detailed interactions responsible for prenucleation and subsequent aerosol formation are difficult to decipher experimentally. We present a computational chemistry study of the interactions between three different acid molecules and two different bases. By combining a comprehensive search routine covering many thousands of configurations at the semiempirical level with high level quantum chemical calculations of approximately 1000 clusters for every possible combination of clusters containing a sulfuric acid molecule, a formic acid molecule, a hydrochloric acid molecule, an ammonia molecule, a dimethylamine molecule, and 0-3 water molecules, we have completed an exhaustive search of the DLPNO-CCSD(T)/CBS//ωB97X-D/6-31++G\*\* Gibbs free energy surface for this system. This first detailed study of HCl interacting with two other acids and two bases reveals the subtleties that exist in the formation of prenucleation complexes for this system. We find that nitric acid forms stronger interactions in dry clusters than hydrochloric acid does. Often as the clusters grow larger with hydration, the sequential energies of clusters containing hydrochloric acid become more favorable than those with nitric acid. The detailed geometries of each minimum free energy cluster are often more important than traditional acid or base strength, which makes a priori prediction of which atmospheric species will be most important for driving prenucleation growth quite difficult. The results presented in this paper add to the conclusions that hydrogen bond topology and the detailed structural interactions that are subtle interplays between enthalpy and entropy are as important as conventional ideas such as acid/base strength.

## 1. Introduction

Atmospheric aerosols play a crucial role in Earth's climate. Their ability to scatter, absorb, and emit incoming solar radiation is believed to produce net cooling of the atmosphere, but current uncertainties match or exceed the expected cooling effects.<sup>1-4</sup> The role of aerosols acting as cloud condensation nuclei (CCN) is of particular interest to better understand the climate. Secondary aerosol particles, which form in the gas phase, are estimated to account for at least half of all atmospheric new particle formation (NPF) by acting as centers for nucleation.<sup>5, 6</sup> Many studies have demonstrated the driving effect of sulfuric acid (SA) in forming prenucleation clusters, especially in the presence of other atmospheric acids and bases.<sup>7-74</sup> Hydrochloric acid (HCl), on the other hand, has not been as well

studied for its potential role in the formation of pre-nucleation clusters.<sup>75</sup> It is estimated that 85% of the HCl in the atmosphere is produced from acid displacement of sea salt aerosols.<sup>76</sup> The remaining atmospheric HCl enters as anthropogenic by-products of biomass burning,<sup>77</sup> specifically coal,<sup>78</sup> and chemical reactions driven by anthropogenic nitrogen oxide emissions.<sup>79</sup> The oxidation of HCl into more reactive chlorine species could have a significant effect on atmospheric concentrations, especially at higher altitudes.<sup>80-82</sup> The interactions of hydrochloric acid within acid-base clusters may provide important information on the outcomes and pathways of atmospheric chlorine species. Because the initial formation of neutral prenucleation clusters and their growth to small aerosols is experimentally and computationally difficult to assess, there is great uncertainty in the size and composition of acid-base clusters, leading researchers to explore a wide variety of different molecules.<sup>23, 52, 53, 63-65, 74</sup> In addition to formation of prenucleation clusters, it is thought that organic acids play a role in stabilizing smaller clusters before evaporating.<sup>74</sup> Most researchers have assumed that acids and bases will grow prenucleation clusters with equal number of acids and bases, as it is logical to assume that stabilization of the growing ensemble

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† Electronic Supplementary Information (ESI) available: Figures of minimum energy structures, sequential hydration tables, DFT and DLPNO-CCSD(T) energy tables, optimized structures, derivation of CBS extrapolation formula, and example calculations of stepwise energies. . See DOI: 10.1039/x0xx00000x

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is accounted for in a complex containing the same number of strong acids and bases.<sup>23, 52, 53, 63-65, 74</sup> However, no single experimental technique is able to analyse all of the aerosols that can be produced in the sub-nanometer to 10 nm size range,<sup>64</sup> and as the atmosphere is filled with a tremendous number of different compounds, identifying the ones most important for nucleation is akin to finding a needle in a haystack.<sup>74</sup> We have published several papers where we have examined weaker acids, as well as unequal numbers of acids and bases that form stable van der Waals complexes, and we have found that the detailed hydrogen bonding topology of particular clusters is often more important than traditional acid/base strength.<sup>55, 66, 68, 73</sup> In our previous paper, we explored the role of three acids, two bases, and 0-5 water molecules by computing all possible combinations of one sulfuric acid (SA), one formic acid (FA), one nitric acid (NA), one ammonia (A), one dimethylamine (DMA), and 0-5 waters (W).<sup>73</sup> In the present work, we continue the exploration of atmospheric aerosols by examining all possible combinations of one SA, one FA, one HCl, one A, one DMA, and up to three waters. Swapping HCl for NA adds to the insights obtained from previous work.<sup>73</sup>

## 2. Methodology

Configurational sampling was employed to generate hundreds of thousands of possible combinations of SA, FA, HCl, A, DMA, and up to three W by using the genetic-algorithm-based protocol in the OGOLEM program,<sup>83, 84</sup> followed by semi-empirical geometry optimization of the initial configurations. This protocol uses an evolutionary algorithm to rearrange the initial configurations of generated clusters until convergence upon a final structure is reached. We used a pool size of 1,000 and allowed for 100,000 global optimizations. These structures were then optimized using the GFN2-xTB semi-empirical method, resulting in a final pool of 1000 or slightly less for each simulation.<sup>85-87</sup> To complete this study we simulated 56 systems containing HCl: HCl-W<sub>0-3</sub>, HCl-A-W<sub>0-3</sub>, HCl-DMA-W<sub>0-3</sub>, SA-FA-HCl-W<sub>0-3</sub>, SA-HCl-A-W<sub>0-3</sub>, SA-HCl-DMA-W<sub>0-3</sub>, FA-HCl-A-W<sub>0-3</sub>, FA-HCl-DMA-W<sub>0-3</sub>, HCl-A-DMA-W<sub>0-3</sub>, SA-FA-HCl-A-W<sub>0-3</sub>, SA-FA-HCl-DMA-W<sub>0-3</sub>, SA-HCl-A-DMA-W<sub>0-3</sub>, FA-HCl-A-DMA-W<sub>0-3</sub>, and SA-FA-HCl-A-DMA-W<sub>0-3</sub> systems. Because of the uncertainties that arise from the use of low-level semi-empirical energies,<sup>55, 88</sup> we used a cut-off energy of 30 kcal mol<sup>-1</sup> above the GFN2-xTB global minimum, thus keeping all structures with a possibility of becoming a minimum or close to one at a better level of theory. These extensive sets of GFN2-xTB structures were optimized at the DFT level of theory using the Gaussian 16 Rev. B01 program.<sup>89</sup> The  $\omega$ B97X-D functional<sup>90, 91</sup> was used with the 6-31++G\*\* basis set<sup>92-94</sup> to obtain a final set of DFT structures. The electronic energies of all structures that were within 8 kcal mol<sup>-1</sup> of the  $\omega$ B97X-D global minimum were then recalculated using the domain-based local pair natural orbital coupled cluster (DLPNO-CCSD(T)) method<sup>95-111</sup> with singles, doubles, and semi-canonical perturbative triple excitations with three Dunning basis sets,<sup>112-115</sup> cc-pVnZ (n = D, T, Q) using ORCA 5.0.1.<sup>111</sup> The

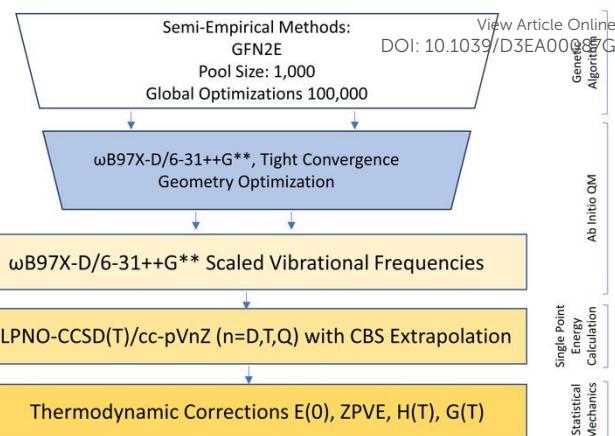


Figure 1. Computational methodology employed to obtain the minimum free energy structures and  $\Delta G^\circ$  values.

$\omega$ B97X-D frequencies were scaled by a factor of 0.971 to account for some anharmonicity,<sup>116, 117</sup> then used to estimate the thermodynamic corrections for  $H^\circ$ ,  $S^\circ$ , and  $G^\circ$  at a 1 atm standard state and temperatures of 216.65, 273.15, and 298.15 K using the THERMO.pl script<sup>118</sup> from the National Institute of Science and Technology. The electronic energies from the three DLPNO-CCSD(T) calculations were used in a 4-5 inverse polynomial complete basis set extrapolation,<sup>119</sup> and combined with the scaled DFT thermodynamic corrections to calculate the final Gibbs free energies. The entire methodology is depicted in Figure 1. Structures that don't contain HCl were taken from our previous papers,<sup>66, 73, 88</sup> and more explanation of our methodologies can be found in our recent review.<sup>74</sup> Once all the  $G^\circ$  values were calculated for the individual clusters, the  $\Delta G^\circ$  values of binding were calculated as follows:

$$\Delta G^\circ_{\text{bind}} = G^\circ_{\text{cluster}} - \sum G^\circ_{\text{monomers}} \quad (1)$$

In addition, the sequential binding energies for addition of a new molecule to an existing cluster were calculated as:

$$\Delta G^\circ_{\text{addition}} = G^\circ_{\text{new cluster}} - (\Delta G^\circ_{\text{bind}} + G^\circ_{\text{new monomer}}) \quad (2)$$

The sequential addition energies obtained from equation 2 allow for computation of the equilibrium concentrations for every cluster, by assuming a closed system and using each  $\Delta G^\circ_{\text{addition}}$  value along with estimates of initial concentrations of the monomers at a given temperature. For these simulations, the top of the troposphere is defined as 217 K while the bottom of the troposphere is 298 K, and estimates for the monomer concentrations are taken from the literature for 298 K, and reduced by a factor of 1000 at 217 K. We used a water concentration of  $7.7 \times 10^{17} \text{ cm}^{-3}$  at 298 K and  $9.9 \times 10^{14} \text{ cm}^{-3}$  at 217 K, which corresponds to 100% humidity at the bottom and top of the troposphere.<sup>120</sup> Initial monomer concentrations from the literature<sup>22, 120-128</sup> at 298 K were reduced by three orders of magnitude at 217 K to compensate for the reduction of CCN-forming particles in the upper troposphere. This is a rough estimate based on the three orders of magnitude decrease in concentration of water, and we resort to this approximation since experimental concentrations of these monomers in the upper troposphere are difficult to measure. Combining these results with previous work allows for estimates of the most likely pathways for the formation of prenucleation complexes for these common acids and bases. The 56 HCl systems were



combined with the relevant SA-FA-NA-A-DMA-W systems previously published<sup>73</sup> to complete our analysis. Three caveats on our methodology are important to note.<sup>74</sup> First, no search routine can ensure that all of the potential energy surface (PES) for these clusters has been exhaustively explored, so we may have missed the lowest Gibbs free energy structure. Second, in the funnel methodology (Figure 1), we are computing the CCSD(T) electronic energies on the  $\omega$ B97X-D geometries, which means we are making the electron correlation corrections on the DFT PES. Third, using the DLPNO routine and scaled harmonic frequencies leads to slightly more positive values. All three of these uncertainties will produce  $\Delta G^\circ$  values that are more positive than would be obtained if (a) a lower energy cluster is obtained, or if (b) the PES is improved so that it is closer to the CCSD(T) surface than the DFT surface, and if (c) anharmonic frequencies could be obtained instead of harmonic frequencies<sup>68</sup> and CCSD(T) electronic energies could be computed directly instead of using the DLPNO routine.<sup>74</sup>

## 3. Results and Discussion

### 3.1 HCl Monomer and Dimer Systems

In the following discussion we briefly refer to the HCl monomer complexed with 1-3 waters and dimers of HCl, bound to one base, and their complexes with water, as most of these systems were previously reported and our results are consistent with published work. We have not included the two acid dimers with HCl (SA-HCl-W<sub>0-3</sub> and FA-HCl-W<sub>0-3</sub>), as previous results indicate that these clusters would not likely be atmospherically relevant due to the lack of hydrogen bonding sites and bases for deprotonation.<sup>73</sup> All our structures and energies for the monomer systems are included in the SI (Figures S1-S2 and Tables T1-T2). Experiment and theory have demonstrated that it takes at least four water molecules to dissociate HCl in the gas-phase at low temperature,<sup>129</sup> while DFT calculations predict that five water molecules are required at 298 K.<sup>130</sup> Our data for the HCl-W<sub>0-3</sub> system has very similar structures to these previous results (S1, T1).<sup>130</sup> Of all possible dimers, only the HCl-DMA and HCl-A systems have not been previously discussed. A general finding comparing SA and HCl binding to the two bases is that the HCl-base dimers have energies significantly less negative than the SA-base dimers (S3-S4, T3-T4), further affirming the significance of SA hydrogen bonding topology. Every HCl-DMA-W<sub>0-3</sub> minima undergoes proton transfer from the acid to the base, and have  $\Delta G^\circ$  values roughly 10 kcal mol<sup>-1</sup> (9.47-11.19 kcal mol<sup>-1</sup> at 217 K and 8.96-11.27 kcal mol<sup>-1</sup> at 298 K) more positive than the SA-DMA-W<sub>0-3</sub> complexes (S4, T4), resulting from DMA's well known nucleating ability in the presence of SA.<sup>10</sup> DMA and other amines have higher gas-phase basicities than ammonia,<sup>131</sup> and this explains why studies of these small complexes have more negative Gibbs free energies of binding relative to ammonia.<sup>10, 24, 132-134</sup> By contrast, HCl-A remains neutral in the dry state, but undergoes proton transfer when hydrated with one or more water molecules. The HCl-A-W<sub>0-3</sub> clusters have  $\Delta G^\circ$  values that are about 10 kcal mol<sup>-1</sup> (7.09-13.27 kcal mol<sup>-1</sup> at 217 K and 6.87-12.75 kcal mol<sup>-1</sup> at 298 K)

more positive than the SA-A-W<sub>n</sub> system (S3, T3). Regardless of atmospheric concentration, the role of acid strength is a far less influential contributor to more negative  $\Delta G^\circ$  values than hydrogen bonding topology.<sup>66</sup> Both SA and HCl are strong acids in solution, but the slightly less strong SA forms dimers with more negative energies in the gas phase, demonstrating its role as a known driver of nucleation.

### 3.2 Trimmers of one acid and two bases

Trimmers of one acid and two bases have been shown to be less atmospherically relevant than other trimers,<sup>73</sup> and this remains true for HCl-A-DMA (S5, T5). When SA is the acid, the  $\Delta G^\circ$  for the one acid-two base trimer is more negative than the HCl containing trimer. While HCl is a stronger acid than SA in the aqueous state, SA has unique bonding capabilities and can form more hydrogen bonds. SA-A-DMA has more negative energies of formation than HCl-A-DMA by roughly 11 kcal mol<sup>-1</sup> (11.29 kcal mol<sup>-1</sup> at 217 K and 10.95 kcal mol<sup>-1</sup> at 298 K) (T5). Similarly, SA-A-DMA is more negative than NA-A-DMA by 9.92 kcal mol<sup>-1</sup> at 217 K and 10.07 kcal mol<sup>-1</sup> at 298 K. NA-A-DMA and HCl-A-DMA have similar energies, bond formations, and protonation patterns. In both clusters, the acid donates its proton to the DMA, and the three molecules form a ring with the A donating a hydrogen bond to the DMA and acid. The approximately one kcal mol<sup>-1</sup> difference (1.37 kcal mol<sup>-1</sup> at 217 K and 0.88 kcal mol<sup>-1</sup> at 298 K) between HCl-A-DMA and NA-A-DMA is likely accounted for by the extra hydrogen bond in NA-A-DMA. NA-A-DMA-W<sub>0-3</sub> has more hydrogen bonds than HCl-A-DMA-W<sub>0-3</sub> in all states of hydration, yet HCl-A-DMA-W<sub>0-3</sub> has more negative hydration energies than NA-A-DMA-W<sub>0-3</sub> at all temperatures. While NA makes stronger dry pre-nucleation clusters in this system, HCl is more favorably hydrated. Additionally, the hydrogen bonds including chlorine are often longer than other interatomic distances between molecules. In HCl-A-DMA, the hydrogen bond from DMA to the chlorine ion is 1.82 Å and the London dispersion forces between the chlorine ion and the ammonium ion result in a N-H---Cl length of 2.50 Å. In the NA-A-DMA cluster, the NA-DMA hydrogen bond length is 1.54 Å and the NA-A hydrogen bond length is 2.05 Å. Shorter hydrogen bond lengths in the NA-A-DMA cluster contributes to its initial stabilization when compared to HCl-A-DMA. Even with its larger size, which increases the bond lengths within the clusters, the HCl cluster benefits more with the addition of waters, likely caused by the greater entropy associated with these structures. This is demonstrated by larger clusters containing HCl and will be further discussed in the respective sections. While these clusters are not predicted to be present in high atmospheric concentrations, they highlight the driving effects of different acids and further reinforce sulfuric acid's effectiveness when compared to other strong acids.

### 3.3 Trimmers of two acids and one base

The larger clusters are less studied; therefore, we report all our findings in the next several sections. In the figures, hydrogen bonds are marked in blue, which have hydrogen-bonded distances of less than 2.20 Å and hydrogen bond angles

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between 140 and 180°. Red lines are used to denote van der Waals forces where the bond angle encompassing the hydrogen is less than 140° or the hydrogen-bond distance is greater than 2.20 Å. The formation energies of all dry trimers, tetramers, and pentamers are listed in Table 1. Hydrated two acid and one base clusters allow for observation and comparison of the driving effects of ammonia and dimethylamine. The two acid and one base structures containing HCl are displayed in Figure 2 and have corresponding energetics in Table 2. The remaining two

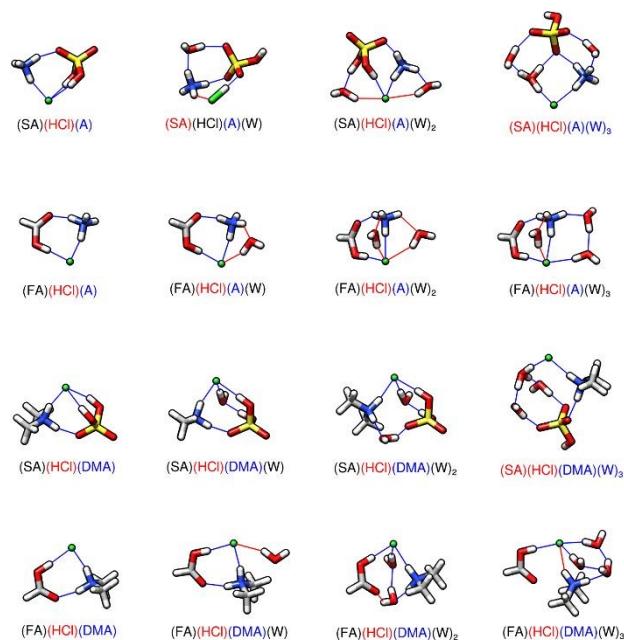
acid and one base trimers are in SI tables T6-T7 and figures S6-S7. Formation of the SA-HCl-DMA cluster from its monomers has a  $\Delta G^\circ$  that is 10.54 kcal mol<sup>-1</sup> more negative than that for formation of SA-HCl-A at 217 K, even though they both contain four hydrogen bonds and a proton transfer from the HCl to the base. This is because DMA is more favorable for pre-nucleation,<sup>10</sup> especially in dry clusters. Upon hydration, the newly formed SA-HCl-A-W consists of a neutral HCl, a

**Table 1.** DLPNO-CCSD(T)/CBS//wB97X-D/6-31++G\*\* Gibbs free energy changes (kcal mol<sup>-1</sup>) associated with the formation of all dry trimers, tetramers, and pentamer in the systems sulfuric acid-formic acid-hydrochloric acid-ammonia-dimethylamine and sulfuric acid-formic acid-nitric acid-ammonia-dimethylamine<sup>a</sup> at atmospherically relevant temperatures and 1 atm pressure. <sup>a</sup> – Reference <sup>73</sup>

Cluster	216.65 K	273.15 K	298.15 K
SA + FA + NA $\rightleftharpoons$ SA-FA-NA	-13.45	-9.50	-7.76
SA + FA + HCl $\rightleftharpoons$ SA-FA-HCl	-8.18	-4.60	-3.02
SA + FA + A $\rightleftharpoons$ SA-FA-A <sup>a</sup>	-17.13	-13.39	-11.74
SA + FA + DMA $\rightleftharpoons$ SA-FA-DMA <sup>a</sup>	-26.08	-22.19	-20.48
SA + NA + A $\rightleftharpoons$ SA-NA-A <sup>a</sup>	-15.74	-11.99	-10.40
SA + NA + DMA $\rightleftharpoons$ SA-NA-DMA <sup>a</sup>	-25.19	-21.37	-19.69
SA + HCl + A $\rightleftharpoons$ SA-HCl-A	-13.15	-9.52	-7.90
SA + HCl + DMA $\rightleftharpoons$ SA-HCl-DMA	-23.69	-20.10	-18.50
FA + NA + A $\rightleftharpoons$ FA-NA-A <sup>a</sup>	-8.42	-5.25	-3.87
FA + NA + DMA $\rightleftharpoons$ FA-NA-DMA <sup>a</sup>	-15.58	-11.46	-9.65
FA + HCl + A $\rightleftharpoons$ FA-HCl-A	-6.65	-3.04	-1.45
FA + HCl + DMA $\rightleftharpoons$ FA-HCl-DMA	-15.73	-12.09	-10.48
SA + FA + NA + A $\rightleftharpoons$ SA-FA-NA-A <sup>a</sup>	-24.77	-18.82	-16.28
SA + FA + NA + DMA $\rightleftharpoons$ SA-FA-NA-DMA <sup>a</sup>	-33.41	-27.50	-24.90
SA + FA + HCl + A $\rightleftharpoons$ SA-FA-HCl-A	-20.20	-14.83	-12.47
SA + FA + HCl + DMA $\rightleftharpoons$ SA-FA-HCl-DMA	-28.92	-23.10	-20.59
SA + FA + A + DMA $\rightleftharpoons$ SA-FA-A-DMA <sup>a</sup>	-30.63	-24.98	-22.49
SA + NA + A + DMA $\rightleftharpoons$ SA-NA-A-DMA <sup>a</sup>	-33.31	-27.23	-24.55
SA + HCl + A + DMA $\rightleftharpoons$ SA-HCl-A-DMA	-36.33	-30.87	-28.45
FA + NA + A + DMA $\rightleftharpoons$ FA-NA-A-DMA <sup>a</sup>	-19.35	-13.94	-11.57
FA + HCl + A + DMA $\rightleftharpoons$ FA-HCl-A-DMA	-17.85	-12.63	-10.32
SA + FA + NA + A + DMA $\rightleftharpoons$ SA-FA-NA-A-DMA <sup>a</sup>	-39.56	-31.51	-28.00
SA + FA + HCl + A + DMA $\rightleftharpoons$ SA-FA-HCl-A-DMA	-43.08	-35.34	-32.02

deprotonated SA, and the ammonium ion, which is the only instance of all the two acids and one base clusters where the HCl is not deprotonated (Figure 2). In nearly every case, SA-HCl-A-W<sub>0-3</sub> has more negative stepwise hydration energies than SA-HCl-DMA-W<sub>0-3</sub> (Table 2). Cluster formation and growth is more favorable in the dry state with DMA, and more favorable when hydrated with ammonia, which is why SA-HCl-A-W<sub>0-3</sub> has more negative stepwise hydration energies than SA-HCl-DMA-W<sub>0-3</sub>.<sup>135</sup> When SA is replaced with FA, the  $\Delta G^\circ$  of the trimers becomes more positive by around 6-9 kcal mol<sup>-1</sup> (6.01-6.50 kcal mol<sup>-1</sup> at 217 K and 5.51-6.45 kcal mol<sup>-1</sup> for replacing SA in SA-HCl-A-W<sub>0-3</sub> and 7.96-9.76 kcal mol<sup>-1</sup> at 217 K and 8.02-9.71 kcal mol<sup>-1</sup> at 298 K for replacing SA in SA-HCl-DMA-W<sub>0-3</sub>) (Table 2). Both dry FA-HCl-A and FA-HCl-DMA have three hydrogen bonds, with the decrease in hydrogen bonds coming from the reduced bonding capabilities of FA compared to SA. We again note that DMA yields more stable clusters in the dry state,<sup>10</sup> whereas upon hydration, ammonia enables more negative energies (Table 2). Ammonium has four hydrogen bonding sites, which allows the water molecules that are added to form more stable hydrogen bond geometries than are possible with the geometry of protonated DMA, which is restricted by only having two

possible sites for hydrogen bonding. FA-HCl-A-W and FA-HCl-DMA-W both contain 3 hydrogen bonds, but FA-HCl-A-W has two hydrogen bonds and a van der Waals interaction contributed by A, whereas FA-HCl-DMA-W has only two hydrogen bonds coming from DMA (Figure 2). The results of the



**Figure 2.** DLPNO-CCSD(T)/CBS//wB97X-D/6-31++G\*\* minimum energy clusters for the sequential hydration of hydrochloric acid with one acid and one base. The molecule labels are colored according to charge as follows: blue = +1, black = 0, red = -1. Atoms are drawn in the following colors: hydrogen – white, carbon – grey, nitrogen – blue, oxygen – red, sulfur – yellow, chlorine – green.

trimers with HCl are energetically and structurally consistent with trimers from previous work.<sup>73</sup> FA-NA-DMA and FA-HCl-DMA have  $\Delta G^\circ$  energies of formation within one kcal mol<sup>-1</sup> (0.15 kcal mol<sup>-1</sup> at 217 K and 0.83 kcal mol<sup>-1</sup> at 298 K) (Table 1) and have very similar structures with the same number of hydrogen bonds (Figures 2 and S7). This further supports the idea that the acid identity is not always as significant in the energies of formation as hydrogen bonding topology.<sup>66</sup> SA-HCl-DMA is 1.50 kcal mol<sup>-1</sup> more positive than SA-NA-DMA at 217 K (Table 1). In SA-HCl-DMA, the hydrogen bond length from the chlorine ion to the hydrogen on DMA is 1.98 Å and the hydrogen bond lengths from the chlorine ion to the hydrogens on SA is

2.13 Å (Figure 2). When HCl is replaced with NA, tighter bonds are formed between the monomers; the DMA-NA bond length is 1.85 Å and the SA-NA bond length is 1.52 Å (Figure S7). The bond length difference between the monomers in these two clusters will have a more stabilizing effect for SA-NA-DMA, which is likely why SA-NA-DMA is lower in energy than SA-HCl-DMA. The number and length of hydrogen bonds within two acid-one base trimers is more important to the energies of formation than the identity of the acid. Elm and co-workers have used ACDC to predict that SA-NA-Base clusters such as SA-NA-DMA show high cluster formation potential.<sup>136</sup> Elm's group has used the same

**Table 2.** DLPNO-CCSD(T)/CBS//wB97X-D/6-31++G\*\* Gibbs free energy changes (kcal mol<sup>-1</sup>) associated with the formation and sequential hydration of sulfuric acid-hydrochloric acid-ammonia, formic acid-hydrochloric acid-ammonia, sulfuric acid-hydrochloric acid-dimethylamine, and formic acid-hydrochloric acid-dimethylamine trimers at atmospherically relevant temperatures and 1 atm pressure.

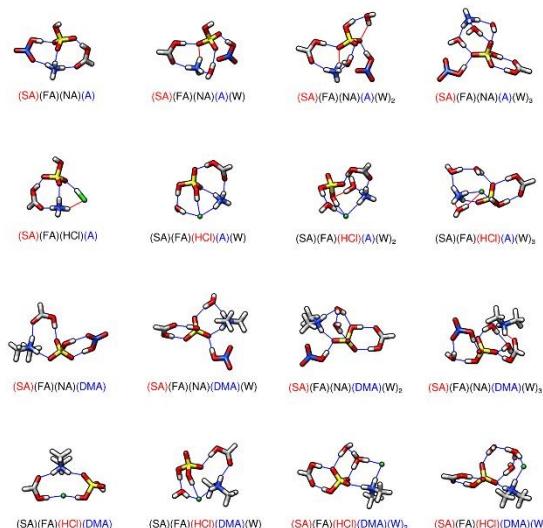
Cluster	216.65 K	273.15 K	298.15 K
SA + HCl + A $\rightleftharpoons$ SA-HCl-A	-13.15	-9.52	-7.90
SA-HCl-A + W $\rightleftharpoons$ SA-HCl-A-W	-3.92	-2.50	-1.88
SA-HCl-A-W + W $\rightleftharpoons$ SA-HCl-A-W <sub>2</sub>	-3.54	-1.39	-0.45
SA-HCl-A-W <sub>2</sub> + W $\rightleftharpoons$ SA-HCl-A-W <sub>3</sub>	-3.69	-2.06	-1.33
FA + HCl + A $\rightleftharpoons$ FA-HCl-A	-6.65	-3.04	-1.45
FA-HCl-A + W $\rightleftharpoons$ FA-HCl-A-W	-4.41	-2.80	-2.10
FA-HCl-A-W + W $\rightleftharpoons$ FA-HCl-A-W <sub>2</sub>	-3.53	-1.82	-1.17
FA-HCl-A-W <sub>2</sub> + W $\rightleftharpoons$ FA-HCl-A-W <sub>3</sub>	-3.31	-1.70	-0.88
SA + HCl + DMA $\rightleftharpoons$ SA-HCl-DMA	-23.69	-20.10	-18.50
SA-HCl-DMA + W $\rightleftharpoons$ SA-HCl-DMA-W	-3.96	-1.97	-1.09
SA-HCl-DMA-W + W $\rightleftharpoons$ SA-HCl-DMA-W <sub>2</sub>	-2.75	-0.93	-0.13
SA-HCl-DMA-W <sub>2</sub> + W $\rightleftharpoons$ SA-HCl-DMA-W <sub>3</sub>	-1.96	-0.24	0.52
FA + HCl + DMA $\rightleftharpoons$ FA-HCl-DMA	-15.73	-12.09	-10.48
FA-HCl-DMA + W $\rightleftharpoons$ FA-HCl-DMA-W	-2.68	-1.01	-0.28
FA-HCl-DMA-W + W $\rightleftharpoons$ FA-HCl-DMA-W <sub>2</sub>	-2.23	-0.16	0.75
FA-HCl-DMA-W <sub>2</sub> + W $\rightleftharpoons$ FA-HCl-DMA-W <sub>3</sub>	-2.78	-1.08	-0.33

methodology to predict that only the SA-FA-DMA clusters are significantly enhanced relative to simulations that compare SA, methanesulfonic acid, ammonia and amine bases.<sup>137</sup> Formation of the SA-FA-DMA complex has the lowest value of  $\Delta G^\circ$ , followed closely by the formation of the SA-NA-DMA cluster, and then the SA-HCl-DMA cluster (Table 1).

### 3.4 Tetramers of three acids and one base

Before discussing three acid and one base clusters, we note that we have calculated the lowest energy structures for SA-FA-NA, SA-FA-HCl, and their hydrates with one to three waters (Figure S8, Table T8). The cluster formation  $\Delta G^\circ$  values for SA-FA-NA are -13.45 kcal mol<sup>-1</sup> at 217 K and -7.76 kcal mol<sup>-1</sup> at 298 K, while those for SA-FA-HCl are about 5 kcal mol<sup>-1</sup> (5.27 kcal mol<sup>-1</sup> at 217 K and 4.74 kcal mol<sup>-1</sup> at 298 K) more positive, likely caused by the fewer hydrogen bonding

sites on HCl as compared to NA. The SA-FA-HCl-W hydrate is the most favorable, with a sequential hydration value of -4.22 kcal mol<sup>-1</sup>. As



**Figure 3.** DLPNO-CCSD(T)/CBS//wB97X-D/6-31++G\*\* minimum energy clusters for the sequential hydration of three acids with one base. The molecule labels are colored according to charge as follows: blue = +1, black = 0, red = -1. Atoms are drawn in the following colors: hydrogen – white, carbon – grey, nitrogen – blue, oxygen – red, sulfur – yellow, chlorine – green.



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the three acid trimers still have relatively positive energies of formation and hydration compared to the two acid one base systems, the structures and energies of the three acid trimers are included only in the SI. The energies of formation and sequential hydration of the three acid and one base tetramer systems containing SA, FA, NA or HCl, and A or DMA are listed in Table 3. The corresponding minimum energy structures are shown in Figure 3. The formation of the dry SA-FA-HCl-DMA cluster is 8.72-8.12 kcal mol<sup>-1</sup> lower in free energy than forming the SA-FA-HCl-A cluster at all

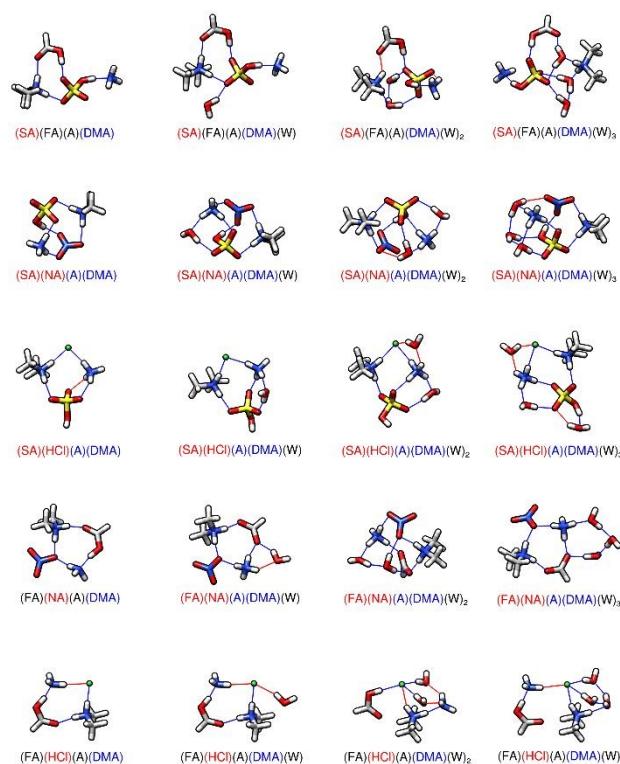
temperatures listed (Table 3). Interestingly, HCl remains neutral in dry SA-FA-HCl-A, with SA being deprotonated instead, which is unlike any of the hydrated three acid and one base clusters, where HCl is deprotonated (Figure 3). Similar to the two acid and one base clusters, HCl remains protonated only when A is present in the dry cluster. Among the dry three acid and one base tetramers, SA-FA-NA-DMA has the most negative  $\Delta G^\circ$  with -33.41 kcal mol<sup>-1</sup> at 217 K.<sup>73</sup> The second most



**Table 3.** DLPNO-CCSD(T)/CBS//ωB97X-D/6-31++G\*\* Gibbs free energy changes (kcal mol<sup>-1</sup>) associated with the formation and sequential hydration of sulfuric acid-formic acid-hydrochloric acid-dimethylamine, sulfuric acid-formic acid-hydrochloric acid-ammonia, sulfuric acid-formic acid-nitric acid-dimethylamine, and sulfuric acid-formic acid-nitric acid-ammonia tetramers at atmospherically relevant temperatures and 1 atm pressure. <sup>a</sup> – Reference <sup>73</sup>

Cluster	216.65 K	273.15 K	298.15 K
SA + FA + NA + A ⇌ SA-FA-NA-A <sup>a</sup>	-24.77	-18.82	-16.28
SA-FA-NA-A + W ⇌ SA-FA-NA-A-W <sup>a</sup>	-3.77	-2.38	-1.71
SA-FA-NA-A-W + W ⇌ SA-FA-NA-A-W <sub>2</sub> <sup>a</sup>	-2.81	-1.14	-0.40
SA-FA-NA-A-W <sub>2</sub> + W ⇌ SA-FA-NA-A-W <sub>3</sub> <sup>a</sup>	-0.98	0.86	1.68
SA + FA + HCl + A ⇌ SA-FA-HCl-A	-20.20	-14.83	-12.47
SA-FA-HCl-A + W ⇌ SA-FA-HCl-A-W	-4.65	-2.48	-1.51
SA-FA-HCl-A-W + W ⇌ SA-FA-HCl-A-W <sub>2</sub>	-2.88	-1.14	-0.38
SA-FA-HCl-A-W <sub>2</sub> + W ⇌ SA-FA-HCl-A-W <sub>3</sub>	-3.05	-1.03	-0.36
SA + FA + NA + DMA ⇌ SA-FA-NA-DMA <sup>a</sup>	-33.41	-27.50	-24.90
SA-FA-NA-DMA + W ⇌ SA-FA-NA-DMA-W <sup>a</sup>	-2.39	-0.48	0.36
SA-FA-NA-DMA-W + W ⇌ SA-FA-NA-DMA-W <sub>2</sub> <sup>a</sup>	-1.81	-0.01	0.78
SA-FA-NA-DMA-W <sub>2</sub> + W ⇌ SA-FA-NA-DMA-W <sub>3</sub> <sup>a</sup>	-0.87	0.82	1.57
SA + FA + HCl + DMA ⇌ SA-FA-HCl-DMA	-28.92	-23.10	-20.59
SA-FA-HCl-DMA + W ⇌ SA-FA-HCl-DMA-W	-4.93	-3.06	-2.17
SA-FA-HCl-DMA-W + W ⇌ SA-FA-HCl-DMA-W <sub>2</sub>	-2.69	-0.85	-0.03
SA-FA-HCl-DMA-W <sub>2</sub> + W ⇌ SA-FA-HCl-DMA-W <sub>3</sub>	-2.04	-0.34	0.42

negative is SA-FA-HCl-DMA, which has a  $\Delta G^\circ$  of -28.92 kcal mol<sup>-1</sup> at 216 K. The larger size of the chloride anion results in longer hydrogen bond distances with SA (1.88 Å) and FA (2.00 Å) while for the SA-FA-NA-DMA cluster, NA forms shorter and stronger hydrogen bonds with SA (1.58 & 1.85 Å). When examining the tetramers with A instead of DMA, the SA-FA-NA-A cluster is also more stable than the SA-FA-HCl-A cluster by 4.57 kcal mol<sup>-1</sup> at 216 K (Table 3). In both clusters, SA undergoes a proton exchange with A, leaving the other acid, NA or HCl, neutral. SA-FA-NA-A forms five hydrogen bonds where SA-FA-HCl-A only forms four hydrogen bonds and a van der Waals interaction. When these clusters are hydrated, the sequential energies of hydration are more negative for the clusters containing HCl than those containing NA. For the clusters containing HCl, hydronium ions are present in the DMA clusters with two and three waters and in the ammonia cluster with three waters. The hydronium ions contribute three hydrogen bonds in all cases, providing additional stability to the hydrated clusters. The clusters containing NA instead of HCl do not contain any hydronium ions and have a smaller number of hydrogen bonds than the HCl clusters with hydronium ions. Regardless of the base present, NA forms more stable complexes than HCl in three acid and one base dry tetramers. However, hydration is more favorable for three acid and one base tetramers containing HCl. Thus, it may be that NA is more important for the first step of prenucleation and HCl is more important for growth and hydration of prenucleation clusters.



**Figure 4.** DLPNO-CCSD(T)/CBS//ωB97X-D/6-31++G\*\* minimum energy clusters for the sequential hydration of two acids with two bases. The molecule labels are colored according to charge as follows: blue = +1, black = 0, red = -1. Atoms are drawn in the following colors: hydrogen – white, carbon – grey, nitrogen – blue, oxygen – red, sulfur – yellow, chlorine – green.

**Table 4.** DLPNO-CCSD(T)/CBS//ωB97X-D/6-31++G\*\* Gibbs free energy changes (kcal mol<sup>-1</sup>) associated with the formation and sequential hydration of sulfuric acid-hydrochloric acid-ammonia-dimethylamine, formic acid-hydrochloric acid-ammonia-dimethylamine, sulfuric acid-formic acid-ammonia-dimethylamine, sulfuric acid-hydrochloric acid-ammonia-dimethylamine, and formic acid-nitric acid-ammonia-dimethylamine tetramers at atmospherically relevant temperatures and 1 atm pressure. <sup>a</sup> – Reference <sup>73</sup>

Cluster	216.65 K	273.15 K	298.15 K
SA + FA + A + DMA ⇌ SA-FA-A-DMA <sup>a</sup>	-30.63	-24.98	-22.49
SA-FA-A-DMA + W ⇌ SA-FA-A-DMA-W <sup>a</sup>	-2.25	-0.50	0.27
SA-FA-A-DMA-W + W ⇌ SA-FA-A-DMA-W <sub>2</sub> <sup>a</sup>	-1.07	1.10	1.99
SA-FA-A-DMA-W <sub>2</sub> + W ⇌ SA-FA-A-DMA-W <sub>3</sub> <sup>a</sup>	-2.05	-0.48	0.28
SA + NA + A + DMA ⇌ SA-NA-A-DMA <sup>a</sup>	-33.31	-27.23	-24.55
SA-NA-A-DMA + W ⇌ SA-NA-A-DMA-W <sup>a</sup>	-6.79	-4.84	-3.99
SA-NA-A-DMA-W + W ⇌ SA-NA-A-DMA-W <sub>2</sub> <sup>a</sup>	-1.60	0.48	1.40
SA-NA-A-DMA-W <sub>2</sub> + W ⇌ SA-NA-A-DMA-W <sub>3</sub> <sup>a</sup>	-1.62	0.22	1.04
SA + HCl + A + DMA ⇌ SA-HCl-A-DMA	-36.33	-30.87	-28.45
SA-HCl-A-DMA + W ⇌ SA-HCl-A-DMA-W	-4.26	-2.45	-1.66
SA-HCl-A-DMA-W + W ⇌ SA-HCl-A-DMA-W <sub>2</sub>	-3.18	-1.56	-0.84
SA-HCl-A-DMA-W <sub>2</sub> + W ⇌ SA-HCl-A-DMA-W <sub>3</sub>	-2.14	-0.36	0.42
FA + NA + A + DMA ⇌ FA-NA-A-DMA <sup>a</sup>	-19.35	-13.94	-11.57
FA-NA-A-DMA + W ⇌ FA-NA-A-DMA-W <sup>a</sup>	-2.16	-0.32	0.50
FA-NA-A-DMA-W + W ⇌ FA-NA-A-DMA-W <sub>2</sub> <sup>a</sup>	-1.82	0.38	1.36
FA-NA-A-DMA-W <sub>2</sub> + W ⇌ FA-NA-A-DMA-W <sub>3</sub> <sup>a</sup>	-2.73	-1.23	-0.66
FA + HCl + A + DMA ⇌ FA-HCl-A-DMA	-17.85	-12.63	-10.32
FA-HCl-A-DMA + W ⇌ FA-HCl-A-DMA-W	-2.50	-0.79	-0.04
FA-HCl-A-DMA-W + W ⇌ FA-HCl-A-DMA-W <sub>2</sub>	-2.16	-0.10	0.81
FA-HCl-A-DMA-W <sub>2</sub> + W ⇌ FA-HCl-A-DMA-W <sub>3</sub>	-2.15	-0.74	-0.12

### 3.5 Tetramers of two acids and two bases

The two acid and two base tetramers are shown in Figure 4, and the corresponding energies are listed in Table 4. It has been previously noted that a tetra-ionic cluster provides additional stability in the form of proton transfers,<sup>73</sup> and that is also the case for the SA-HCl-A-DMA tetramer. The tetramer SA-HCl-A-DMA is 18.13–18.48 kcal mol<sup>-1</sup> lower in  $\Delta G^\circ$  than FA-HCl-A-DMA at the temperatures listed. This energy difference is likely caused by SA-HCl-A-DMA being tetra-ionic while FA-HCl-A-DMA is di-ionic, with FA and A remaining neutral. SA-HCl-A-DMA has four hydrogen bonds, while FA-HCl-A-DMA has three hydrogen bonds. Both structures have a ring-like geometry with acids and bases opposite each other, which differs from the structures of di-ionic SA-FA-A-DMA and tetra-ionic SA-NA-A-DMA. SA-FA-A-DMA is 5.70 kcal mol<sup>-1</sup> more positive in Gibbs free energy than SA-HCl-A-DMA at 217 K and SA-NA-A-DMA is 3.02 kcal mol<sup>-1</sup> higher in  $\Delta G^\circ$ . SA-FA-A-DMA only has four hydrogen bonds and is a di-ionic cluster. When comparing FA-HCl-A-DMA and FA-NA-A-DMA in the dry state, the tetramer with NA has a lower  $\Delta G^\circ$  than the tetramer with HCl. This is different from the tetramers of SA-HCl-A-DMA and SA-NA-A-DMA, where the tetramer with HCl is more stable in the dry state. Both FA-HCl-A-DMA and FA-NA-A-DMA are di-ionic clusters with four

hydrogen bonds. The 1.50 kcal mol<sup>-1</sup> difference between the dry clusters at 217 K is likely a result of subtle hydrogen bond strength differences. When the tetramers are hydrated, their relative stabilities change due to the geometric stabilization of the added water molecule. As discussed by Bready et al.,<sup>73</sup> SA-NA-A-DMA-W has a remarkably large first sequential hydration energy of -6.79 kcal mol<sup>-1</sup> at 217 K and exhibits special stability. The tetra-ionic cluster is stabilized by the addition of the water molecule allowing for 7 hydrogen bonds. When SA-HCl-A-DMA is hydrated with one water, only 6 hydrogen bonds are formed for a first sequential hydration energy of -4.26 kcal mol<sup>-1</sup> at 217 K. SA-HCl-A-DMA-W only has three hydrogen bonds that encompass SA whereas SA-NA-A-DMA-W utilizes 4 hydrogen bonding sites on SA. Upon hydration with one water, both SA-FA-A-DMA-W and FA-NA-A-DMA-W have 5 hydrogen bonds and FA-HCl-A-DMA-W has 3 hydrogen bonds. As the number of hydrogen bonds decrease in the hydrated tetramers, the stability of the clusters decreases as well. For the second and third sequential hydrations, SA-HCl-A-DMA-W<sub>2-3</sub> have more negative sequential hydration values than SA-NA-A-DMA-W<sub>2-3</sub>, which is like the results in section 3.3 where two acid-one base clusters with HCl have more negative hydration energies than when NA replaces HCl. These results support the statement that NA may be better for pre-

**Table 5.** DLPNO-CCSD(T)/CBS//ωB97X-D/6-31++G\*\* Gibbs free energy changes (kcal mol<sup>-1</sup>) associated with the formation and sequential hydration of sulfuric acid-formic acid hydrochloric acid-ammonia-dimethyl amine and sulfuric acid-formic acid-nitric acid-ammonia-dimethylamine pentamers at atmospherically relevant temperatures and 1 atm pressure. <sup>a</sup> – Reference <sup>73</sup>

Cluster	216.65 K	273.15 K	298.15 K
SA + FA + NA + A + DMA ⇌ SA-FA-NA-A-DMA <sup>a</sup>	-39.56	-31.53	-28.00
SA-FA-NA-A-DMA + W ⇌ SA-FA-NA-A-DMA-W <sup>a</sup>	-4.13	3.09	4.03
SA-FA-NA-A-DMA-W + W ⇌ SA-FA-NA-A-DMA-W <sub>2</sub> <sup>a</sup>	-4.07	-1.54	-0.64
SA-FA-NA-A-DMA-W <sub>2</sub> + W ⇌ SA-FA-NA-A-DMA-W <sub>3</sub> <sup>a</sup>	1.16	3.69	4.58
SA + FA + HCl + A + DMA ⇌ SA-FA-HCl-A-DMA	-43.08	-35.34	-32.02
SA-FA-HCl-A-DMA + W ⇌ SA-FA-HCl-A-DMA-W	-4.35	-2.67	-1.82
SA-FA-HCl-A-DMA-W + W ⇌ SA-FA-HCl-A-DMA-W <sub>2</sub>	-2.85	-0.85	0.04
SA-FA-HCl-A-DMA-W <sub>2</sub> + W ⇌ (SA-FA-HCl-A-DMA-W <sub>3</sub> )	0.00	2.07	2.86

nucleation cluster formation while HCl may be better for the growth and hydration of a prenucleation cluster.

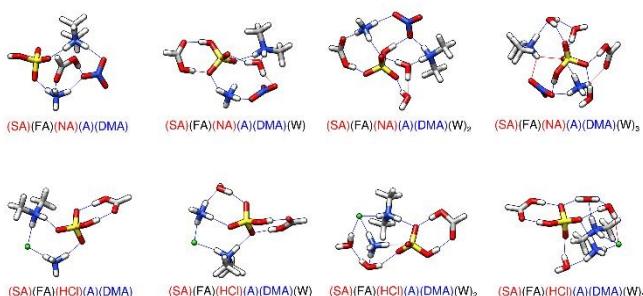
### 3.6 Pentamer of three acids and two bases

The system containing all three acids and both bases with zero to three waters is pictured in Figure 5 with the respective energies listed in Table 5. In the dry state, SA-FA-HCl-A-DMA has 6 hydrogen bonds and an SA-FA dimer, noted for its strength and stability by Harold et al.<sup>66</sup> SA-FA-HCl-A-DMA is 3.52 kcal mol<sup>-1</sup> lower in Gibbs free energy than SA-FA-NA-A-DMA at 217 K (Table 5), possibly resulting in part from SA-FA-NA-A-DMA lacking the SA-FA dimer. Both pentamers are tetra-ionic with FA remaining neutral. Upon hydration with one water, SA-FA-HCl-A-DMA is slightly more stabilized than SA-FA-NA-A-DMA. This is likely due to SA-FA-HCl-A-DMA gaining two hydrogen bonds, whereas SA-FA-NA-A-DMA only gains one hydrogen bond and a van der Waals interaction. When the second water molecule is added, SA-FA-NA-A-DMA gains three hydrogen bonds and has a sequential hydration energy 1.22 kcal mol<sup>-1</sup> lower than SA-FA-HCl-A-DMA at 217 K, which only gains two new hydrogen bonds. The third water destabilizes both clusters, as SA-FA-NA-A-DMA gains three weak London dispersion forces and SA-FA-HCl-A-DMA gains one hydrogen bond and one London dispersion forces, so that entropy is decreased more than enthalpy is increased. In all structures dry and

hydrated, SA-FA-HCl-A-DMA-W<sub>0-3</sub> has an SA-FA dimer within its geometry, likely helping stabilize it. The pentamer containing HCl has a more consistent geometry than SA-FA-NA-A-DMA-W<sub>0-3</sub> throughout hydration and has more stabilizing first and third hydration energies than the NA pentamer and less stabilizing second hydration energies (Table 5). The large size of the chlorine ion leads to its Cl---H distances being longer as the system size increases, compared to a traditional hydrogen bond between NA and another monomer.

### 3.7 Equilibrium concentrations and pathways of formation

To calculate the equilibrium concentrations of the cluster, equilibrium constants were first calculated using the  $\Delta G^\circ$  values of the cluster at 217 K and 298 K. For both of these temperatures, a system of equations adapted from Odbadrakh et al.<sup>54</sup> was used to calculate the equilibrium concentrations assuming a closed system of these three acids, two bases, and three waters. We note that the calculations account only for thermodynamics and do not account for the kinetics of the reaction. Therefore, the results of the calculations would likely change if additional kinetic atmospheric conditions like evaporation were included in the calculations. Initial starting concentrations of the monomers were  $5 \times 10^7 \text{ cm}^{-3}$  for SA,  $2 \times 10^{11} \text{ cm}^{-3}$  for FA,  $1 \times 10^9 \text{ cm}^{-3}$  for HCl,  $2 \times 10^{11} \text{ cm}^{-3}$  for A, and  $2 \times 10^9 \text{ cm}^{-3}$  for DMA at 298 K. These concentrations were chosen because they are atmospherically relevant over inland and urban areas.<sup>120, 121, 122, 123, 124</sup> We chose a water concentration of  $7.7 \times 10^{17} \text{ cm}^{-3}$  at 298 K and  $9.9 \times 10^{14} \text{ cm}^{-3}$  at 217 K, which corresponds to 100% humidity at the bottom and top of the troposphere.<sup>120</sup> Initial monomer concentrations from the literature<sup>22, 120-128</sup> at 298 K were reduced by three orders of magnitude at 217 K to compensate for the reduction of CCN-forming particles in the upper troposphere. This rough estimate is based on the three orders of magnitude decrease in water concentration, and we used this approximation since experimental concentrations of these monomers in the upper troposphere are difficult to measure. HCl is the product of the oxidation of reactive chlorine species from sea spray and biomass burning and can have concentrations largely affected by the proximity to a coast as well as the concentration of other atmospheric constituents.<sup>127, 128</sup> Ozone concentrations, gas phase



**Figure 5.** DLPNO-CCSD(T)/CBS//ωB97X-D/6-31++G\*\* minimum energy clusters for the sequential hydration of three acids with two bases. The molecule labels are colored according to charge as follows: blue = +1, black = 0, red = -1. Atoms are drawn in the following colors: hydrogen – white, carbon – grey, nitrogen – blue, oxygen – red, sulfur – yellow, chlorine – green.



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ammonia concentrations, and atmospheric acidity all play a role in the relative concentration of HCl in different locations. HCl concentrations of  $1 \times 10^9 \text{ cm}^{-3}$  have been reported in coastal urban locations and mountain-top locations.<sup>123, 126</sup> To account for the decrease in monomer concentrations at the top of the troposphere,

the initial concentrations of monomers have been decreased by three orders of magnitude to  $5 \times 10^4 \text{ cm}^{-3}$  for SA,<sup>124</sup>  $1 \times 10^8 \text{ cm}^{-3}$  for FA,<sup>125</sup>  $1 \times 10^6 \text{ cm}^{-3}$  for HCl,  $2 \times 10^8 \text{ cm}^{-3}$  for A, and  $2 \times 10^6 \text{ cm}^{-3}$  for DMA at 217 K. The calculated equilibrium concentrations of atmospherically relevant clusters are shown in Table 6, where we have defined

**Table 6.** Equilibrium concentrations of clusters that form at more than  $1 \text{ cm}^{-3}$  at 217 and 298 K. Initial concentrations of the monomers at 217 K were: SA =  $5.00 \times 10^4$ , FA =  $2.00 \times 10^8$ , HCl =  $1.00 \times 10^6$ , A =  $2.00 \times 10^8$ , DMA =  $2.00 \times 10^6$ . Initial concentrations of the monomers at 298 K were: SA =  $5.00 \times 10^7$ , FA =  $2.00 \times 10^{11}$ , HCl =  $1.00 \times 10^9$ , A =  $2.00 \times 10^{11}$ , DMA =  $2.00 \times 10^9$ . The water concentrations were  $7.7 \times 10^{17} \text{ cm}^{-3}$  at 298 K and  $9.9 \times 10^{14} \text{ cm}^{-3}$  at 217 K, which corresponds to 100% humidity at the bottom and top of the troposphere.<sup>120</sup>

Cluster	216.65 K	298.15 K
SA	$2.84 \times 10^1$	$1.48 \times 10^7$
FA	$2.00 \times 10^8$	$1.99 \times 10^{11}$
HCl	$1.00 \times 10^6$	$9.99 \times 10^8$
A	$2.00 \times 10^8$	$2.00 \times 10^{11}$
DMA	$1.95 \times 10^6$	$1.98 \times 10^9$
SA-FA	$3.77 \times 10^{-2}$	$1.03 \times 10^3$
SA-A	$8.35 \times 10^{-3}$	$5.64 \times 10^2$
SA-DMA	$2.42 \times 10^4$	$1.14 \times 10^7$
FA-A	$4.88 \times 10^{-4}$	$2.90 \times 10^1$
FA-DMA	$6.05 \times 10^{-1}$	$5.00 \times 10^2$
HCl-A	$2.26 \times 10^{-4}$	3.37
HCl-DMA	$4.39 \times 10^{-3}$	4.24
SA-FA-DMA	$1.96 \times 10^3$	$9.87 \times 10^3$
SA-HCl-DMA	$3.81 \times 10^{-2}$	1.75
SA-A-DMA	$2.41 \times 10^{-3}$	1.05
SA-HCl-A-DMA	1.27	$2.80 \times 10^{-1}$
SA-W1	$2.39 \times 10^1$	$1.50 \times 10^7$
SA-W2	3.28	$2.90 \times 10^6$
SA-W3	$7.19 \times 10^{-2}$	$1.45 \times 10^5$
FA-W1	$1.90 \times 10^5$	$1.13 \times 10^9$

FA-W2	$1.06 \times 10^3$	$1.99 \times 10^7$
FA-W3	$1.77 \times 10^{-1}$	$5.22 \times 10^4$
HCl-W1	$1.24 \times 10^1$	$7.01 \times 10^5$
HCl-W2	$3.79 \times 10^{-4}$	$3.39 \times 10^2$
A-W1	$6.57 \times 10^3$	$3.79 \times 10^8$
A-W2	$5.61 \times 10^{-2}$	$4.52 \times 10^4$
A-W3	$1.88 \times 10^{-5}$	$1.03 \times 10^2$
DMA-W1	$1.86 \times 10^2$	$5.01 \times 10^6$
DMA-W2	$1.04 \times 10^{-2}$	$3.76 \times 10^3$
SA-FA-W1	$1.12 \times 10^{-2}$	$3.79 \times 10^2$
SA-FA-W2	$4.47 \times 10^{-4}$	$2.71 \times 10^1$
SA-A-W1	$6.27 \times 10^{-4}$	$8.20 \times 10^1$
SA-A-W2	$2.28 \times 10^{-3}$	$1.12 \times 10^2$
SA-A-W3	$8.32 \times 10^{-5}$	$1.21 \times 10^1$
SA-DMA-W1	$2.29 \times 10^4$	$5.42 \times 10^6$
SA-DMA-W2	$7.10 \times 10^2$	$3.06 \times 10^5$
SA-DMA-W3	$7.73 \times 10^1$	$4.30 \times 10^4$
FA-A-W1	$2.63 \times 10^{-3}$	$4.56 \times 10^1$
FA-DMA-W1	$3.53 \times 10^{-3}$	$1.12 \times 10^1$
HCl-DMA-W1	$2.48 \times 10^{-2}$	$2.01 \times 10^1$
HCl-DMA-W2	$5.95 \times 10^{-3}$	5.60
HCl-DMA-W3	$7.62 \times 10^{-4}$	$5.52 \times 10^{-1}$
SA-FA-DMA-W1	$1.25 \times 10^2$	$1.19 \times 10^3$
SA-FA-DMA-W2	$3.24 \times 10^{-1}$	8.87



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**Table 7.** Optimal pathways for growing the dry pentamer at 1 atm pressure according to equilibrium concentrations calculated using initial concentrations of  $[SA]_0 = 5.00 \times 10^7$ ,  $[FA]_0 = 2.00 \times 10^{11}$ ,  $[HCl]_0 = 1.00 \times 10^9$ ,  $[A]_0 = 2.00 \times 10^{11}$ , and  $[DMA]_0 = 2.00 \times 10^9$  at 298 K. Concentrations were decreased by 3 orders of magnitude for 217 K. The water concentrations were  $7.7 \times 10^{17} \text{ cm}^{-3}$  at 298 K and  $9.9 \times 10^{14} \text{ cm}^{-3}$  at 217 K, which corresponds to 100% humidity at the bottom and top of the troposphere.<sup>120</sup>

216.65 K		298.15 K
Optimal Pathway	$SA + DMA \rightleftharpoons SA\text{-}DMA$	$SA + DMA \rightleftharpoons SA\text{-}DMA$
	$+ FA \rightleftharpoons SA\text{-}FA\text{-}DMA$	$+ HCl \rightleftharpoons SA\text{-}HCl\text{-}DMA$
	$+ A \rightleftharpoons SA\text{-}FA\text{-}A\text{-}DMA$	$+ A \rightleftharpoons SA\text{-}HCl\text{-}A\text{-}DMA$
	$+ HCl \rightleftharpoons SA\text{-}FA\text{-}HCl\text{-}A\text{-}DMA$	$+ FA \rightleftharpoons SA\text{-}FA\text{-}HCl\text{-}A\text{-}DMA$
2 <sup>nd</sup> Best Pathway	$SA + DMA \rightleftharpoons SA\text{-}DMA$	$SA + DMA \rightleftharpoons SA\text{-}DMA$
	$+ HCl \rightleftharpoons SA\text{-}HCl\text{-}DMA$	$+ FA \rightleftharpoons SA\text{-}FA\text{-}DMA$
	$+ A \rightleftharpoons SA\text{-}HCl\text{-}A\text{-}DMA$	$+ A \rightleftharpoons SA\text{-}FA\text{-}A\text{-}DMA$
	$+ FA \rightleftharpoons SA\text{-}FA\text{-}HCl\text{-}A\text{-}DMA$	$+ HCl \rightleftharpoons SA\text{-}FA\text{-}HCl\text{-}A\text{-}DMA$
3 <sup>rd</sup> Best Pathway	$SA + FA \rightleftharpoons SA\text{-}FA$	$SA + FA \rightleftharpoons SA\text{-}FA$
	$+ DMA \rightleftharpoons SA\text{-}FA\text{-}DMA$	$+ DMA \rightleftharpoons SA\text{-}FA\text{-}DMA$
	$+ A \rightleftharpoons SA\text{-}FA\text{-}A\text{-}DMA$	$+ A \rightleftharpoons SA\text{-}FA\text{-}A\text{-}DMA$
	$+ HCl \rightleftharpoons SA\text{-}FA\text{-}HCl\text{-}A\text{-}DMA$	$+ HCl \rightleftharpoons SA\text{-}FA\text{-}HCl\text{-}A\text{-}DMA$

atmospherically relevant as equilibrium concentrations above one per  $\text{cm}^{-3}$ . Atmospherically relevant dimers are comprised of mainly acid-base pairings and SA-FA. SA-DMA is the only dimer relevant at both 217 K and 298 K, likely a result of the role of SA and DMA as drivers of nucleation. The concentration of HCl-DMA is less than for FA-DMA and SA-DMA. HCl is not a known driver of nucleation and lacks the hydrogen bonding ability of SA and FA. Upon the addition of one and two waters to HCl-DMA, its concentration at 217 K and 298 K increases. The hydration of HCl-DMA provides an increase in hydrogen bonding sites and a more stable cluster. Two acid and one base clusters are the most atmospherically relevant of the trimers, particularly SA-HCl-DMA and SA-FA-DMA. SA and FA form a particularly strong complex, which has been previously studied by Harold et al.,<sup>66</sup> and has a higher equilibrium concentration than most other dimers. When the magnitude of concentrations for all monomers is decreased by three orders, corresponding to the top of the troposphere at 217 K, clusters with an SA-FA complex have lower concentrations. This is a result of the driving force of SA and the importance of hydrogen bonding topology. The dry tetramer SA-HCl-A-DMA is barely atmospherically relevant at 217 K, but is not at 298 K. The pentamer is not atmospherically relevant at any temperature in our closed system simulation of these three acids, two bases, and three water molecules. To gain insight into prenucleation cluster growth, the ideal pathways of formation for the pentamer are derived from Table 6 by using an atmospherically relevant dimer

cluster and then adding monomers sequentially until the pentamer is formed. These pathways are shown in Table 7. A high starting concentration of SA-DMA begins the optimal growth pathway and second-best pathway for the dry pentamer at 298 K and

**Table 8.** Optimal pathways for growing the SA-FA-DMA-W cluster at 1 atm pressure according to equilibrium concentrations calculated using initial concentrations of  $[SA]_0 = 5.00 \times 10^7$ ,  $[FA]_0 = 2.00 \times 10^{11}$ ,  $[HCl]_0 = 1.00 \times 10^9$ ,  $[A]_0 = 2.00 \times 10^{11}$ , and  $[DMA]_0 = 2.00 \times 10^9$  at 298 K. Concentrations were decreased by 3 orders of magnitude for 217 K. The water concentrations were  $7.7 \times 10^{17} \text{ cm}^{-3}$  at 298 K and  $9.9 \times 10^{14} \text{ cm}^{-3}$  at 217 K, which corresponds to 100% humidity at the bottom and top of the troposphere.<sup>120</sup>

216.65 K		298.15 K
Optimal Pathway	$SA + DMA \rightleftharpoons SA\text{-}DMA$	$SA + W \rightleftharpoons SA\text{-}W$
	$+ W \rightleftharpoons SA\text{-}DMA\text{-}W$	$+ DMA \rightleftharpoons SA\text{-}DMA\text{-}W$
	$+ FA \rightleftharpoons SA\text{-}FA\text{-}DMA\text{-}W$	$+ FA \rightleftharpoons SA\text{-}FA\text{-}DMA\text{-}W$
	$SA + W \rightleftharpoons SA\text{-}W$	$SA + DMA \rightleftharpoons SA\text{-}DMA$
2 <sup>nd</sup> Best Pathway	$+ DMA \rightleftharpoons SA\text{-}DMA\text{-}W$	$+ W \rightleftharpoons SA\text{-}DMA\text{-}W$
	$+ FA \rightleftharpoons SA\text{-}FA\text{-}DMA\text{-}W$	$+ FA \rightleftharpoons SA\text{-}FA\text{-}DMA\text{-}W$
	$DMA + W \rightleftharpoons DMA\text{-}W$	$DMA + W \rightleftharpoons DMA\text{-}W$
	$+ SA \rightleftharpoons SA\text{-}DMA\text{-}W$	$+ SA \rightleftharpoons SA\text{-}DMA\text{-}W$
3 <sup>rd</sup> Best Pathway	$+ FA \rightleftharpoons SA\text{-}FA\text{-}DMA\text{-}W$	$+ FA \rightleftharpoons SA\text{-}FA\text{-}DMA\text{-}W$

217 K. Starting clusters with SA and DMA will result in good cluster growth because both are known drivers of pre-nucleation cluster formation. In the optimal pathway, HCl is added last at 217 K and is the first addition to the SA-DMA dimer at 298 K (Table 7). The only atmospherically relevant hydrated trimers are SA-FA-DMA-W<sub>1-2</sub> and the optimal pathways of formation for SA-FA-DMA-W<sub>1</sub> at 217 K and 298 K are shown in Table 8. When compared to the number of relevant hydrated clusters for the SA-FA-NA-A-DMA system, the SA-FA-HCl-A-DMA system is less atmospherically relevant. Clusters containing HCl often have more negative energies of hydration than clusters containing NA, but HCl clusters are less favourable for the first step in prenucleation than SA or NA. The low concentrations of this pentamer system may point to HCl being a poor driver of nucleation.

## Conclusions

This is the fourth study on formation of prenucleation complexes of three different acid molecules and two different bases,<sup>73, 136, 137</sup> and adds to the insights obtained from previous work. Accurate  $\Delta G^\circ$  values for the formation of every possible cluster that can be formed from a sulfuric acid molecule, a formic acid molecule, a hydrochloric acid molecule, an ammonia molecule, a dimethylamine molecule, and 0-3 water



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molecules were determined from a comprehensive search of the GFN2-xTB and  $\omega$ B97X-D potential energy surfaces combined with DLPNO-CCSD(T)/CBS electronic energy calculations on the DFT geometries. This first detailed study of HCl interacting with two other acids and two bases reveals the subtleties in play when HCl is considered as a potential actor in the formation of prenucleation clusters. Because of the larger size of chlorine, many of the clusters that include HCl form longer distances between neighbouring molecules and have weaker interactions in the dry state. By comparison, nitric acid forms stronger interactions in small, dry, clusters than hydrochloric acid. However, as the clusters grow larger with hydration, the interactions with HCl often become stronger than those with  $\text{HNO}_3$ . In many of the clusters explored in this paper, the SA-FA dimer stabilizes the overall clusters. Many subtleties are at play in the beginning stages of pre-nucleation and the importance of different factors changes with the system being investigated. Hydrogen bonding topology, acid/base strength, and complex structural interactions all play an important role in structure energetics and atmospheric relevance. Sometimes detailed hydrogen bonding topology is as important as conventional notions like acid/base strength, which makes a priori prediction of which atmospheric species will be most important for driving prenucleation growth quite difficult. Complexes with DMA and various acids form stronger dry complexes than does ammonia, yet the sequential hydration energies favor ammonia since the ammonium cation can form up to four hydrogen bonds while protonated DMA can only form two. For the three acid-one base systems, NA forms a stronger dry cluster than HCl, while stepwise hydration switches to favor the HCl clusters (Table 3). For the two acid-two base clusters containing SA, HCl is a stronger nucleator than NA. The first hydration favors the formation of the unusually strong SA-NA-A-DMA-W cluster, yet successive hydrations with a second and third water revert to the clusters with HCl (Table 4). While HCl forms a stronger two acid-two base system than NA when SA is the second acid, NA forms a stronger two acid-two base system when FA is the second acid (Table 1). As shown in Figure 5 and Table 5, the SA-FA-HCl-A-DMA-W<sub>0.1</sub> structures that are stabilized by the SA-FA dimer have more negative Gibbs free energies of formation than the comparable SA-FA-NA-A-DMA-W<sub>0.1</sub> structures. Yet as an additional water is added, the stepwise hydration energies favor the structures with NA instead of HCl. Taken as a whole, the results presented in this paper add to the conclusions that hydrogen bond topology and the detailed structural interactions that are subtle interplays between enthalpy and entropy can be as important as conventional ideas such as acid/base strength.<sup>66</sup> There is much future work that must be completed to gain a better understanding of the beginning stages of prenucleation.

## Author Contributions

The manuscript was written through contributions of all authors. All authors have given approval to the final version of the manuscript.

## Conflicts of interest

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DOI: 10.1039/D3EA00087G

There are no conflicts to declare.

## Acknowledgements

Funding for this work was provided by grants CHE-16626238, CHE-1903871, and CHE-2018427 from the National Science Foundation (GCS), the Arnold and Mabel Beckman Foundation Beckman Scholar Award (CJB), and the Barry M. Goldwater Scholarship (CJB). High-performance computing resources were provided by the Research Corporation for Science Advancement (27446) and the MERCURY Consortium ([www.mercuryconsortium.org](http://www.mercuryconsortium.org)).<sup>138, 139</sup> Molecular graphics and analyses performed with UCSF Chimera, developed by the Resource for Biocomputing, Visualization, and Informatics at the University of California, San Francisco, with support from NIH P41-GM103311.

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View Article Online  
DOI: 10.1039/D3EA00087G

