

PERSPECTIVE | AUGUST 15 2024

More than just smoke and mirrors: Gas-phase polaritons for optical control of chemistry

Jane C. Nelson ; Marissa L. Weichman  



J. Chem. Phys. 161, 074304 (2024)

<https://doi.org/10.1063/5.0220077>



The Journal of Chemical Physics

Special Topics Open for Submissions

[Learn More](#)

More than just smoke and mirrors: Gas-phase polaritons for optical control of chemistry

Cite as: J. Chem. Phys. 161, 074304 (2024); doi: 10.1063/5.0220077

Submitted: 22 May 2024 • Accepted: 8 July 2024 •

Published Online: 15 August 2024



View Online



Export Citation



CrossMark

Jane C. Nelson  and Marissa L. Weichman^{a)} 

AFFILIATIONS

Department of Chemistry, Princeton University, Princeton, New Jersey 08544, USA

^{a)} Author to whom correspondence should be addressed: weichman@princeton.edu

ABSTRACT

Gas-phase molecules are a promising platform to elucidate the mechanisms of action and scope of polaritons for optical control of chemistry. Polaritons arise from the strong coupling of a dipole-allowed molecular transition with the photonic mode of an optical cavity. There is mounting evidence of modified reactivity under polaritonic conditions; however, the complex condensed-phase environment of most experimental demonstrations impedes mechanistic understanding of this phenomenon. While the gas phase was the playground of early efforts in atomic cavity quantum electrodynamics, we have only recently demonstrated the formation of molecular polaritons under these conditions. Studying the reactivity of isolated gas-phase molecules under strong coupling would eliminate solvent interactions and enable quantum state resolution of reaction progress. In this Perspective, we contextualize recent gas-phase efforts in the field of polariton chemistry and offer a practical guide for experimental design moving forward.

Published under an exclusive license by AIP Publishing. <https://doi.org/10.1063/5.0220077>

I. INTRODUCTION

Since the invention of the laser, physical chemists have explored the prospect of controlling chemical reactivity with light for practical and intellectual purposes alike.^{1–4} Even with decades of progress in laser-driven chemistry, the rapid inter- and intramolecular reorganization of deposited energy in most molecular systems still presents a major obstacle.^{5,6} Laser control of chemical behavior has therefore been largely limited to simple gas-phase systems where dephasing and energy redistribution are minimized.^{2,7} Recently, evidence of altered solution-phase chemistry under strong light–matter coupling has reinvigorated the dream of optical reaction control under less-than-pristine conditions.^{8,9} While new paradigms in chemical physics have historically been developed in the gas phase before extension to the condensed phase, polariton chemistry has so far skipped this step: the first demonstrations of strong light–matter coupling were performed with atomic gases, but gaseous molecular polaritons have not been part of the conversation until our recent work.^{10,11} In this Perspective, we review these advances and provide a roadmap for future efforts.

Polaritons are hybrid light–matter states that form when the photonic mode of an optical cavity is brought into resonance with a bright transition of intracavity atoms, molecules, or material.^{9,12,13} Under strong coupling conditions, two spectral features appear in

the frequency-domain cavity transmission spectrum separated by the Rabi splitting Ω_R (Fig. 1). When an ensemble of N molecules are coupled to the same cavity mode, the Rabi splitting is observed to scale with $(N/V)^{1/2}$, where N/V is the intracavity molecular number density. These split peaks are dubbed the upper and lower polariton states and can be resolved when the light–matter coupling strength $g = \Omega_R/2$ exceeds both the molecular half-linewidth γ and the cavity half-linewidth κ . In the time domain, strong light–matter coupling indicates that the rate of exchange of photons between the molecules and cavity outcompetes all dissipative processes.

Polariton formation and the emergence of the strong coupling regime are often introduced using the language of cavity quantum electrodynamics (cQED), in which the discrete states of a quantum emitter hybridize with a quantized mode of light.¹⁴ In the simplest case, the Jaynes–Cummings (JC) model treats the interaction of a single two-level emitter with a harmonic photonic mode, considering only the first excitation manifold of the coupled light–matter system.^{15,16} When the JC Hamiltonian is diagonalized, the upper and lower polaritons emerge as eigenstates corresponding to symmetric and antisymmetric linear combinations of photonic and material excitations. The Tavis–Cummings (TC) model extends the JC model to the collective strong coupling regime by accounting for the mixing of one cavity mode with the totally symmetric collective bright state of N intracavity two-level emitters.¹⁷ Upper and lower

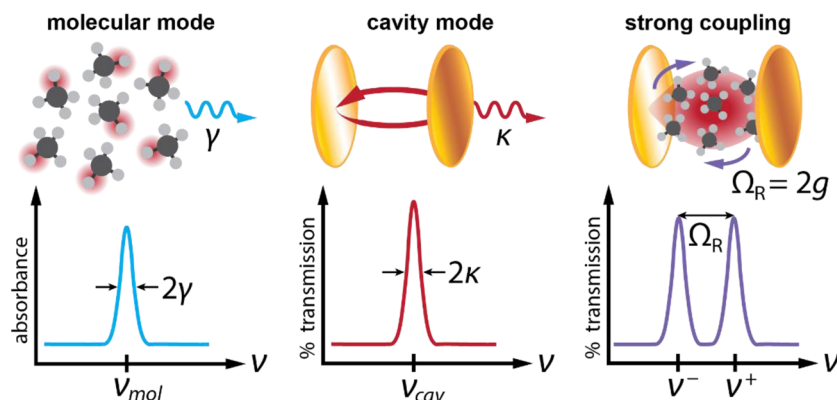


FIG. 1. Schematic of strong light–matter coupling when a molecular excitation frequency (ν_{mol}) is resonant with an optical cavity mode frequency (ν_{cav}). When the rate of photon exchange between molecules and cavity exceeds both the molecular relaxation rate, γ , and the cavity photon loss rate, κ , the light–matter system enters the strong coupling regime. In this regime, cavity transmission maxima appear at new frequencies, ν^{\pm} , separated by the Rabi frequency, Ω_R .

polariton states again emerge, separated in frequency by a collective Rabi splitting that scales with $N^{1/2}$. In addition to the two polaritonic eigenstates, $N - 1$ so-called “dark states” are left at the original energy of the uncoupled emitters. The dark states are non-totally symmetric linear combinations of emitter states that carry no photonic character and are therefore optically forbidden (so long as one neglects heterogeneity in the ensemble of emitters).¹⁸

cQED models provide a convenient conceptual description of polaritons as hybrid light–matter quasiparticles and correctly predict the Rabi splitting and its collective scaling for intracavity ensembles. However, experimental implementations of the strong light–matter coupling regime feature complexity not captured by the JC or TC models: molecules are not two-level systems and they exhibit finite excited-state lifetimes as well as energetic and orientational disorder, while optical cavities have finite lifetimes due to photonic losses and feature a geometry-dependent spectrum of longitudinal and transverse modes. In addition, most experiments use classical light sources to probe cavity transmission spectra, operating nowhere near the few-photon regime of quantum optics. As a result, many experimental groups treat cavity spectra using classical simulations, which can more easily account for the aforementioned practical factors.

Classically, the strong coupling regime arises from the dispersion of light in a cavity filled with dielectric material. In particular, we consider the self-interference of an incident electromagnetic field as it propagates through a Fabry–Pérot (FP) cavity composed of two identical mirrors spaced by length L and containing an intracavity medium with a frequency-dependent refractive index $n(\nu)$ and absorption coefficient $\alpha(\nu)$. The fractional intensity of light transmitted through such a structure is given by^{9,19,20}

$$\frac{I_T(\nu)}{I_0} = \frac{T^2 e^{-\alpha(\nu)L}}{1 + R^2 e^{-2\alpha(\nu)L} - 2R e^{-\alpha(\nu)L} \cos\left(\frac{4\pi n(\nu)L\nu}{c}\right)}, \quad (1)$$

where R and T are the reflection and transmission probabilities for each mirror with $R + T = 1$ for lossless mirrors. The widely used transfer matrix method uses the same principles of classical

wave interference as Eq. (1) to treat cavity structures containing layered media.^{9,13,21} Peaks in the cavity transmission spectrum appear through constructive interference of light traveling in the cavity at frequencies where the round-trip phase shift [$\delta\phi = 4\pi n(\nu)L\nu/c$] is equal to an integer multiple of 2π . When $n(\nu)$ is constant, cavity transmission peaks are evenly spaced by the cavity free spectral range (FSR). Near a strongly absorbing molecular transition, on the other hand, $n(\nu)$ takes on a dispersive line shape and gives rise to additional transmission fringes whose frequencies coincide with the polaritonic states predicted by cQED. Recent theoretical work has drawn an explicit connection between the cQED and classical optics results in the limit that the number of intracavity absorbers $N \rightarrow \infty$, explaining why both descriptions are successful in capturing experimental data.^{22,23}

While the community has focused on strong light–matter coupling in condensed-phase systems for the past 15 years, the original demonstrations of cQED were carried out in atomic gases. In the early 1980s, Haroche and co-workers observed the resonant enhancement of spontaneous emission under *weak* coupling of sodium Rydberg atoms to a microwave Fabry–Pérot resonator,²⁴ a manifestation of the long-predicted Purcell effect.^{25,26} Shortly thereafter, Haroche’s team reported *strong* cavity coupling in the same N -atom system,^{27–29} measuring the excited state Rydberg population in the time domain via field ionization and observing collective Rabi oscillations with frequency scaling as $N^{1/2}$. Furthermore, Haroche showed reversible spontaneous emission under strong coupling as the cavity recycles photons back to the atomic ensemble faster than excitations dissipate.²⁸ Within the same decade, Kimble and co-workers harnessed improvements in mirror coatings to strongly couple electronic transitions of sodium atoms at optical wavelengths.³⁰ Kimble’s group then advanced toward the single-atom limit, achieving the first experimental demonstration of the JC model with a cesium atom.^{31–33} Kimble and co-workers also introduced metrics for strongly coupled systems via frequency-domain cavity transmission spectroscopy, reporting both the splitting of the coupled system’s normal modes and dispersion of these modes upon detuning the cavity from resonance.³⁰ To this day, mode splitting and dispersion in cavity transmission or

reflection spectra remain the prime experimental hallmarks of strong coupling.

Polariton formation has subsequently been explored in a variety of media. Inspired by the early atomic demonstrations, solid-state physicists reported strong light–matter coupling in semiconductor quantum wells in 1992³⁴ and launched the field of inorganic exciton-polaritons.³⁵ In 1998, Lidzey and co-workers made the leap to create exciton-polaritons in organic molecular materials.³⁶ Since then, solid-state exciton-polaritons have remained an active research topic with Bose–Einstein condensation and polariton lasing being of particular interest.^{37–39} In the 2010s, Ebbesen and co-workers led a new charge to examine the chemical behavior of molecular polaritons,⁴⁰ focusing initially on electronic strong coupling (ESC) in molecular dyes and the resulting impact on photochemistry and fluorescence.^{41–43} In 2015, Ebbesen⁴⁴ and Simpkins⁴⁵ reported parallel implementations of vibrational strong coupling (VSC) by coupling infrared (IR)-active carbonyl modes of polymer films to micrometer-scale planar FP cavities. Both teams subsequently demonstrated that solution-phase molecules with strong and narrow absorption bands could be cavity-coupled in synthesis-compatible microfluidic devices.^{46,47} Ebbesen and co-workers took the next step to examine chemical reactivity under VSC, reporting slowed kinetics of a thermal liquid-phase silane deprotection reaction in a microcavity in 2016.⁴⁸ In the near-decade since, several groups have worked to understand the scope and reproducibility of both ground-state chemistry under VSC^{8,9,13,49–55} and photochemistry under ESC.^{56–60} Despite a vibrant research community and a growing body of experimental work in polariton chemistry, the field is still in search of mechanistic understanding and predictive capabilities. Solvation effects and rapid energy dissipation complicate data analysis and interpretation in liquid-phase chemical kinetics and dynamics.^{6,61} Meanwhile, theoretical efforts face the challenge of simultaneously treating the experimentally relevant collective coupling of $N > 10^6$ molecules, disorder, complex reactive surfaces, and realistic lossy multimode cavities.^{14,62,63}

Here, we do not aim to provide a comprehensive review of all prior work and proposed mechanisms for cavity-altered chemistry and refer interested readers to several recent reviews of this emerging and lively field.^{9,12–14,62,63} Instead, we lay out a case for how gas-phase molecular polariton experiments may clarify the state and scope of polariton chemistry. Gas-phase molecules feature resolved quantum states and well-characterized potential energy surfaces that may prove useful in identifying signatures and validating emerging hypotheses for cavity-mediated dynamics. For example, a leading hypothesis for chemistry under VSC is that the cavity may mediate intramolecular vibrational energy redistribution (IVR) among the ensemble of strongly coupled intracavity molecules.^{14,54,63–69} In free space, IVR can cause a selectively excited vibration to decay into the manifold of nearby states on ultrafast timescales.⁷⁰ If operative, a cavity-IVR mechanism could drain energy out of a strongly coupled unimolecular or bimolecular reaction coordinate into orthogonal molecular degrees of freedom or into the lossy photonic mode; this would serve to slow reaction rates under VSC, as has indeed been observed in seminal experiments.^{48,53} At present, just one report has provided evidence of enhanced IVR in a unimolecular isomerization under VSC.⁶⁴ Additional experiments are necessary to validate the cavity-IVR mechanism and other emerging hypotheses in a wide range of chemical systems, including bimolecular

reactions. We argue here that the gas phase is a sensible domain to conduct this work, as the historical proving ground for chemical physics which the polariton community has not returned to since the original atomic cQED experiments.

In this Perspective, we report on the status and future prospects of gas-phase molecular polaritons. We lay out important experimental considerations for gas-phase polaritons and review our group's recent demonstrations in Sec. II. In Sec. III, we describe our procedure to evaluate new molecular candidates using the classical optics framework described above. We propose several next-generation species for both vibrational and electronic gas-phase strong coupling in Sec. IV, highlighting the most compelling prospects for studies of reactivity and photophysics. Finally, in Sec. V, we consider future challenges and envision how gas-phase strong coupling efforts may prove a crucial step in bridging the experiment–theory gap in cavity chemistry.

II. LAYING THE GROUNDWORK FOR GAS-PHASE MOLECULAR POLARITONS

Here, we briefly review the infrastructure our group has introduced to cavity couple gas-phase molecules.^{10,11} Reaching the strong coupling regime in diffuse gases requires some thought, given the $(N/V)^{1/2}$ dependence of the collective Rabi splitting; simply increasing the gas pressure can lead to counter-productive pressure broadening of molecular transitions.¹³ Instead, we operate in the Doppler-broadened regime at relatively low temperatures and pressures, aiming to simultaneously minimize molecular linewidths and maximize Ω_R . We use collisional cooling in a home-built cryogenic buffer gas cell (CBGC) to prepare a cold, dense intracavity sample [Fig. 2(a)]. Through cooling, we narrow Doppler line shapes and reduce the rovibrational partition function to condense population into low-lying states. To perform cavity-coupling *in situ*, we surround the CBGC with a plano-concave two-mirror FP optical cavity whose transmission we probe with a tunable narrow-band mid-infrared continuous-wave (cw) laser.

Our open centimeter-scale cavity design enables careful tunability of parameters and distinguishes our gas-phase apparatus from the micrometer-scale microfluidic cavities more commonly used for VSC.⁹ For a given experiment, we choose a cavity geometry to match the photonic mode linewidth to that of the targeted molecular transition and to ensure that the cavity mode spacing is large enough for clean, state-specific molecular coupling conditions. In practice, we have direct control over the following:

- The cavity length, L , which sets the free spectral range as $\text{FSR} = c/2nL$, where c is the speed of light and n is the intracavity refractive index.
- The reflectivity of each mirror, R , which sets the cavity finesse according to $\mathcal{F} = \pi R^{1/2}/(1 - R)$, assuming that the two mirrors are identical; practically, this parameter is used in combination with the cavity length to determine the photonic mode linewidth as $\Delta\nu = \text{FSR}/\mathcal{F}$.
- The mirror radius of curvature (ROC), which determines the spectrum of transverse Gaussian spatial modes supported by a plano-concave FP cavity.

A representative experimental cavity transmission spectrum is plotted in red in Fig. 2(b) for a near-confocal cavity with $L = 8.36$ cm,

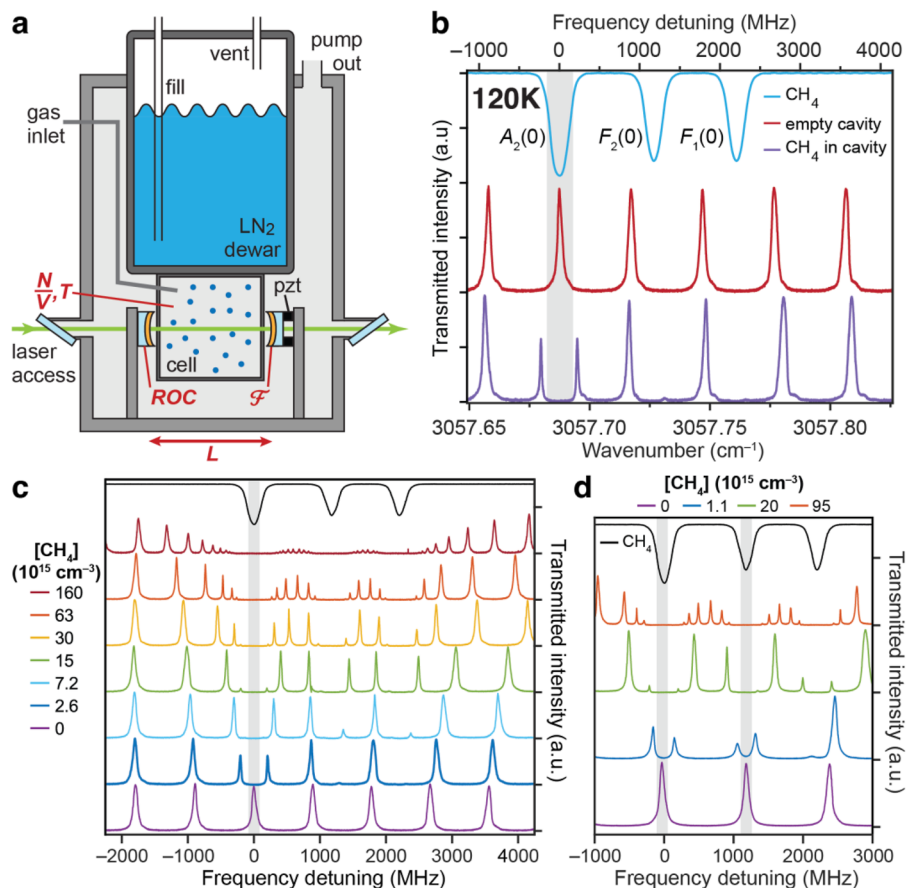


FIG. 2. (a) Intracavity cryogenic buffer gas cell (CBGC) used to achieve strong coupling with gas-phase molecules of varying number density (N/V) and temperature (T). The Fabry–Pérot (FP) cavity geometry is determined by its length (L), finesse (\mathcal{F}), and the mirror radius of curvature (ROC). (b) Transmission spectra of methane (CH_4 , blue), an empty FP cavity (red), and methane in the same cavity under strong coupling conditions for the target $\nu_3, J = 3 \rightarrow 4, A_2(0)$ transition (purple). Methane spectra are acquired at 120 K with $N/V = 1.5 \times 10^{15} \text{ cm}^{-3}$ for the extracavity sample and $N/V = 3.5 \times 10^{15} \text{ cm}^{-3}$ for the strongly coupled sample. The near-confocal cavity used here has $L = 8.36 \text{ cm}$, $\mathcal{F} = 24$, and $\text{ROC} = -8.36 \text{ cm}$. The empty cavity sustains a free spectral range of 895 MHz (0.0299 cm^{-1}) and a linewidth of 65 MHz FWHM (0.0022 cm^{-1}). A Rabi splitting of 454 MHz (0.0151 cm^{-1}) is achieved under these conditions. (c) Transmission spectra of the same cavity from (b) containing increasing intracavity number densities of methane at 120 K. (d) Transmission spectra of a near-confocal $L = 6.27 \text{ cm}$, $\mathcal{F} = 25$ cavity containing increasing intracavity methane number densities. Adjacent cavity modes are near-resonant with the $\nu_3, J = 3 \rightarrow 4 A_2(0)$ and $F_2(0)$ transitions so that strong coupling is achieved with two transitions simultaneously. (a) and (b) Adapted with permission from Wright *et al.*, *J. Am. Chem. Soc.* **145**(10), 5982–5987 (2023). Copyright 2023 American Chemical Society. (c) and (d) Adapted with permission from Wright *et al.*, *J. Chem. Phys.* **159**(16), 164202 (2023). Copyright 2023 AIP Publishing LLC.

$\text{ROC} = -8.36 \text{ cm}$, and $\mathcal{F} = 24$. We mount one cavity mirror on a piezo-electric (pzt) chip to allow for remote detuning of the cavity length. We stabilize the cavity length with active feedback on the pzt to ensure that cavity-coupling conditions remain consistent throughout an experiment.

We use rovibrational transitions of methane (CH_4) for an initial demonstration of gas-phase strong coupling.¹⁰ CH_4 is a convenient target with high symmetry, bright line strengths, and precedence in the mode-specific chemistry literature.^{4,71} We target an individual rovibrational transition in the R-branch of the ν_3 asymmetric C–H stretching band [blue trace in Fig. 2(b)]. At low temperature

and pressure, the three $A_2(0)$, $F_2(0)$, and $F_1(0)$ symmetry components of the $\nu_3, J = 3 \rightarrow 4$ rovibrational transition can be resolved at 3057.687423 , 3057.726496 , and $3057.760735 \text{ cm}^{-1}$, respectively.⁷² We record the full-width at half-maximum (FWHM) linewidth of these peaks as 180 MHz (0.006 cm^{-1}) in the CBGC, consistent with Doppler broadening at 120 K. To perform strong coupling experiments, we stabilize the cavity length so one photonic mode is resonant with a targeted molecular transition, then flow CH_4 into the cell while monitoring the cavity transmission spectrum. Split polaritonic features appear at sufficiently high molecular flow rates; a representative strong coupling experiment targeting resonance with

the $A_2(0)$ band is given in the purple trace in Fig. 2(b). All experimental traces are well reproduced by the classical cavity transmission expression in Eq. (1).

Following the initial demonstration of this gas-phase strong coupling platform, we explore the tunability of both cavity and molecular experimental parameters.¹¹ We continue to work with the rovibrational ν_3 , $J = 3 \rightarrow 4$ $A_2(0)$ band of CH_4 , first testing a wider range of molecular conditions. At molecular number densities where the collective Rabi splitting approaches the cavity FSR, we observe that nested polariton peaks emerge from the off-resonant coupling of the targeted molecular transition with adjacent cavity modes [Fig. 2(c)]. The additional peaks in the cavity transmission spectrum become even more numerous at higher flow rates where the collective Rabi splitting approaches the spacing between molecular states. These emergent features represent admixtures of multiple molecular transitions and multiple photonic modes; their light-matter composition could be quantified with Hopfield coefficient analysis.^{9,63} We also show that strong coupling of CH_4 is possible at room temperature, though the increased room temperature rovibrational partition function and Doppler broadening lower the absorption cross section of the targeted transition. We therefore need an order of magnitude higher molecular number density to reach the same room temperature Rabi splitting as in a 120 K sample, and only a small fraction of these warm molecules are in the correct lower state for resonant cavity-coupling.¹¹ Regardless, working with room temperature samples greatly relaxes experimental demands and lowers the barrier for others in the community to create and study gas-phase molecular polaritons.

We also systematically tune the cavity parameters and examine the effects on polariton formation.¹¹ Evaluating cavity transmission spectra as a function of cavity finesse, we find that polariton linewidths track with the changing cavity mode linewidth, while the Rabi splitting is unchanged, consistent with the known behavior of inhomogeneously disordered samples under strong coupling.^{13,73} Tuning the FSR via the cavity length enables further control of the photonic modes participating in strong coupling. By choosing an FSR commensurate with the spacing between rovibrational molecular peaks, we can reach a multi-state coupling condition where two transitions are simultaneously cavity-coupled [Fig. 2(d)].

Extending this apparatus to strongly couple molecules other than CH_4 is straightforward. All that is required is (a) a tunable cw laser to measure cavity transmission in the targeted spectral region; (b) cavity mirrors with reflectivity in the targeted spectral region, though the gold-coated mirrors we have used so far are quite general purpose; and (c) a cavity geometry with an FSR and linewidth designed to couple cleanly to the desired molecular transition. We discuss some practicalities in screening next-generation candidates in Sec. III.

III. EVALUATING NEXT-GENERATION CANDIDATES FOR GAS-PHASE STRONG COUPLING

Here, we lay out our considerations for extending our platform to molecules of interest in polariton reactivity studies. We aim to provide a roadmap for designing gas-phase polariton experiments that target molecules with well-studied reactivity in free space, particularly those with unimolecular or IR-driven chemistry. We

must carefully consider the experimental feasibility of target systems, a step often overlooked in theoretical proposals in the literature. Specifically, we note:

1. Prospective molecules must first and foremost be amenable to strong coupling under physically realizable conditions. The collective Rabi splitting—which has to exceed the molecular and cavity linewidths to reach the strong coupling regime—is proportional to the molecular transition dipole and the square-root of the number density. A target molecule must have both sufficiently bright optical transitions and a large vapor pressure, so reasonable gas-phase number densities are experimentally accessible.
2. The molecule's fully resolved rovibrational or vibronic structure must also be considered. In the condensed phase, fine structure is washed out and an entire vibrational or electronic band can be coupled at once. In the gas phase, we can couple individual quantum states with a proper cavity design, provided that we account for the temperature- and pressure-dependent absorption line shape of the targeted transition.
3. Finally, we must choose systems with direct readouts of reaction dynamics and which are likely to feature cavity-modulated behavior. We also have to consider whether it is possible to study a reaction at the high molecular number densities required for strong coupling and evaluate possible reactant or product detection schemes.

We will discuss this last point further for various systems in Sec. IV. Addressing the first two points is more straightforward: we can evaluate prospects for strong coupling specific molecular transitions by simulating the transmission of light through a cavity containing the target species with the classical optics treatment described in Sec. I.^{10,11,19,20} The most critical components of these simulations are the molecular absorption coefficient, $\alpha(\nu)$ (cm^{-1}), and the real part of the refractive index, $n(\nu)$. $\alpha(\nu)$ is typically determined from the absorption cross section, $\sigma(\nu)$ ($\text{cm}^2/\text{molecule}$), according to $\alpha(\nu) = \sigma(\nu) \cdot N/V$ and must be evaluated in the spectral range of interest under experimentally relevant number density, Doppler broadening, and pressure broadening conditions. $n(\nu)$ can then be derived from $\alpha(\nu)$ using the Kramers-Kronig relation. Finally, we input $\alpha(\nu)$ and $n(\nu)$ into Eq. (1) along with experimentally reasonable parameters for mirror reflectivity and transmission (R, T) and cavity length (L) to simulate transmission spectra. For most simulations provided in Sec. IV, we assume a cavity geometry similar to that used in our previous work.^{10,11} Unless otherwise noted, we use a representative non-confocal $L = 5$ cm optical cavity with $R = 90\%$ mirrors. This geometry features a mode spacing of FSR = 3000 MHz (0.1 cm^{-1}), a mode linewidth of $\Delta\nu = 100$ MHz (0.003 cm^{-1}) FWHM, and a finesse of $\mathcal{F} = 30$. We confirm strong coupling prospects for each system by targeting resonance between a cavity mode and a selected molecular transition for various number densities and determine the conditions for which polaritonic splittings first appear in the transmission spectra.

It is straightforward to perform these simulations when high-resolution reference data exist for the absorption cross section $\sigma(\nu)$. In the best-case scenario, a line list is available from the HITRAN database,⁷² which can be opened directly in the PGOPHER software package.⁷⁴ The process is more involved when line lists or

high-resolution experimental reference data are not available, the target species has many isomers, or the system features intrinsic broadening. In these cases, the temperature and pressure dependence of $\sigma(\nu)$ may be less certain. When lacking a line list, we build a rovibrational or vibronic model of the system in PGOPHER using spectroscopic constants derived from Gaussian 16 calculations⁷⁵ and any available experimental spectroscopic constants. The PGOPHER documentation⁷⁴ and tutorials from Sprague⁷⁶ and Wilhelm *et al.*⁷⁷ are helpful in this process. We then use the PGOPHER spectrum fitting tools to fit the available experimental reference spectra. Once we have constructed a reasonable simulation, we set the temperature, Doppler broadening, and pressure broadening in PGOPHER to calculate $\sigma(\nu)$ under conditions relevant to our CBGC experiments and convert to the absorption coefficient $\alpha(\nu)$ for various molecular number densities.

Table I provides a summary of the gas-phase molecular transitions that we examine for strong coupling prospects in this work, which we discuss in Sec. IV. See Sec. SI of the [supplementary](#)

[material](#) for more details and Table SI for specific simulation parameters and sources of reference data.

IV. PROSPECTIVE MOLECULAR GASES FOR POLARITON CHEMISTRY

We now report specific prospects for strong cavity coupling of quantum-state-resolved transitions in molecules of interest. We target rovibrational strong coupling (RVSC) in molecules that undergo gas-phase bimolecular H-abstraction reactions and unimolecular isomerization and photodissociation. We also explore possibilities for gas-phase electronic strong coupling at ultraviolet (UV)/visible wavelengths. In our evaluation of RVSC candidates, we place a particular emphasis on systems whose behavior may inform the cavity-mediated IVR hypothesis for polariton chemistry introduced in Sec. I. We therefore target (a) reactions with established IR-driven reactivity where we can examine if intra-cavity rates and product branching are modifiable via cavity-IVR

TABLE I. Conditions for strong coupling simulations of molecular gases.

Molecule	Transition	T (K)	$2\gamma_{mol}$ [FWHM, (cm ⁻¹)] ^a	N/V (cm ⁻³) ^b	Ω_R (cm ⁻¹) ^c
CH ₄	$\nu_3, J = 3 \rightarrow 4$ 3057.6878 cm ⁻¹	295	0.009 40	3×10^{15}	0.010 87
C ₂ H ₂	$\nu_3, J = 9 \rightarrow 10$ 3304.9655 cm ⁻¹	295	0.007 99	3×10^{15}	0.011 70
CH ₃ OH	$\nu_8, J = 9 \rightarrow 10$ 1049.3578 cm ⁻¹	295 100	0.002 44 0.001 36	9×10^{15} 5×10^{15}	0.003 58 0.003 40
<i>trans</i> -HONO	$\nu_3, J = 9 \rightarrow 10$ 1271.218 cm ⁻¹	100	0.001 35	3×10^{15}	0.004 49
<i>trans</i> -C ₄ H ₆	$\nu_{11}, J = 8 \rightarrow 7$ 894.2307 cm ⁻¹ $\nu_{11}, J = 30 \rightarrow 29$ 894.2316 cm ⁻¹	100	0.001 01 ^d	1×10^{16}	0.003 78
O ₃	$\nu_3, J = 7 \rightarrow 8$ 1048.0695 cm ⁻¹	295 100	0.001 99 ^d 0.001 08	3×10^{15} 9×10^{14}	0.004 56 0.003 51
I ₂	B-X, $\nu_1 = 0 \rightarrow 29, J = 45 \rightarrow 46$ 18 599.1 cm ⁻¹ , 538 nm	295	0.028 44	5×10^{15}	0.036 47
O ₃ ^e	B-X 39 300 cm ⁻¹ , 254 nm	293	6552	3×10^{21}	7027

^aThe molecular linewidth reflects the Voigt profile of the simulated absorption coefficient and therefore represents a convolution of both Gaussian Doppler broadening and Lorentzian pressure broadening.

^bThe number density given is the smallest value simulated that reaches the strong coupling regime, as defined by the Rabi splitting exceeding both molecular transition and cavity mode FWHM linewidths.

^cClassical cavity transmission simulations were performed with a non-confocal $L = 5$ cm, $\mathcal{F} = 30$ cavity with a $\Delta\nu = 0.003\,36$ cm⁻¹ FWHM linewidth unless otherwise noted.

^dThese systems feature asymmetric molecular line shapes with distinct shoulders not captured by the molecular linewidth. In these cases, the strong coupling regime is determined when the Rabi splitting exceeds the cavity linewidth.

^eClassical cavity transmission simulations for the O₃ B-X band were performed with an $L = 254$ nm, $\mathcal{F} = 6$ cavity with a $\Delta\nu = 3267$ cm⁻¹ FWHM cavity linewidth. The absorption cross section data used in transmissions simulations are from Serdyuchenko *et al.*¹³⁶ rather than PGOPHER simulations.

and (b) molecules that natively feature IVR where we might resolve vibrational energy distributions and screen for cavity-induced changes.

A. Bimolecular reactions of small hydrocarbons

The bimolecular hydrogen (H) abstraction reactions of small organic molecules with reactive radicals are benchmark systems for laser-driven chemistry^{2,71} and feature simple, well-characterized reactive surfaces for theoretical analysis. The reactions of methane (CH₄) and acetylene (C₂H₂) with the hydroxyl radical (OH)^{78,79} and other small radicals have been the subject of decades of research.⁸⁰ We have already demonstrated RVSC of CH₄,^{10,11} here, we show that acetylene (C₂H₂) has similar prospects. We consider the rovi-

brational transitions of room-temperature C₂H₂ in the ν_3 C–H stretching band near 3 μm using line lists available in HITRAN. We predict the onset of strong coupling for $N/V = 3 \times 10^{15} \text{ cm}^{-3}$ (Fig. 3, Table I). As C₂H₂ is a commercially available gas, reaching this number density is not a practical concern. Comparable Rabi splittings are accessible for the same number densities of CH₄ and C₂H₂, though C₂H₂ features stronger polaritonic transmission features due to its narrower Doppler linewidth and therefore weaker absorption at polariton frequencies (Fig. 3). We expect cooling of C₂H₂ to reduce the number density necessary to reach strong coupling, as we saw for CH₄.¹¹

The H-abstraction reactions of strongly coupled CH₄ or C₂H₂ with OH can be monitored with pulsed laser photolysis-laser-induced fluorescence (PLP-LIF). In this approach, one would prepare OH reactants via PLP of hydrogen peroxide precursor and track the subsequent reactive loss of OH using LIF.^{78,79,81} Derived reaction rates can be compared between on- and off-resonance intracavity experiments and extracavity controls. The mechanisms of cavity-altered bimolecular reactions remain poorly understood, but the relevant experimental literature, if anything, suggests that they will feature slower kinetics under VSC of reactants.^{48,53} We therefore hypothesize that one might observe slowed reaction rates for intracavity gas-phase H-abstraction. In any event, these systems will be a natural platform to test theories such as the cavity-IVR hypothesis introduced in Sec. I, which posits that cavity-coupling can impede reaction progress by introducing a new channel to drain vibrational energy out of the reaction coordinate. Adding a detection channel to probe the distribution of post-reaction product vibrational states could explicitly illuminate intracavity vibrational energy redistribution and more directly test the cavity-IVR hypothesis. Preparation of vibrationally excited reagents is another natural direction; IR-pumping of both reactants in the CH₄ + OH system has been examined in the extracavity literature.^{82–84} Performing similar experiments under RVSC of CH₄ would provide a direct probe of cavity-mediated relaxation pathways that may funnel vibrational energy out of the pumped reaction coordinate.

The bimolecular reactions of larger organic molecules that sustain significant inter-mode vibrational couplings are also of interest for gas-phase polariton chemistry studies and could enable further insights into the cavity-IVR hypothesis. Methanol (CH₃OH) has similar reactivity to methane^{85,86} but features more extensive torsion-vibration couplings and rapid IVR.^{87–89} Similarly, extending C₂H₂ to longer acetylenic chains (e.g., 1-propyne, 1-butyne, and 1-pentyne) yields denser manifolds of vibrational states and increased coupling between modes.^{90,91} Unfortunately, we predict that it will be practically challenging to reach the strong coupling regime in these species, as the same intramolecular couplings of interest cause spectral broadening and congestion that prohibit clean, state-specific resolution of polaritonic features. For example, while transitions in the C–O stretching band of methanol near 10 μm are accessible to cavity-coupling, the intrinsic congestion of the C–H stretching region near 3 μm ⁹² will prohibit rovibrationally resolved cavity-coupling (Fig. S1). We note that state-specific resolution of coupling conditions is one of the unique strengths of working with gas-phase molecules; it may make more sense to rely on existing condensed-phase microcavity architectures when targeting broad or unresolved bands. Balancing these competing considerations is a challenge for future work, as we discuss further in Sec. IV D.

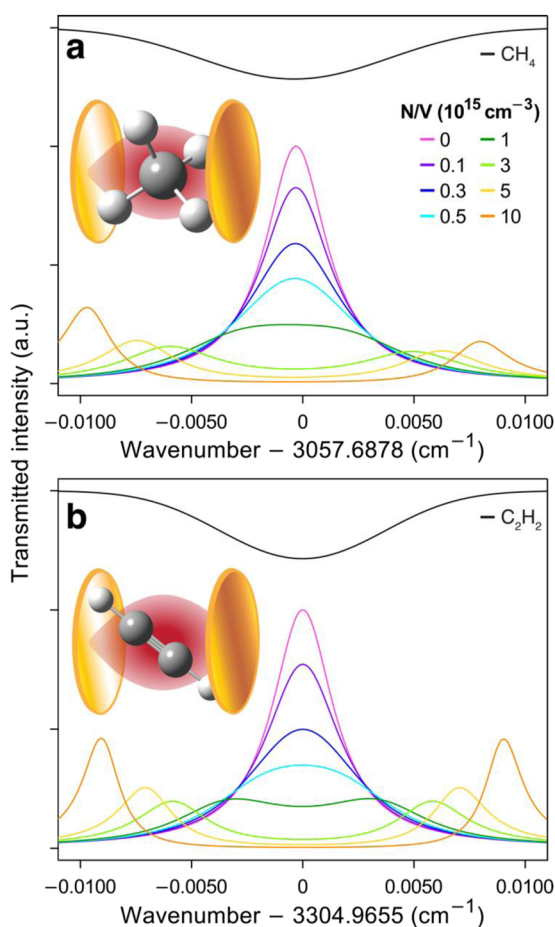


FIG. 3. Simulated classical cavity transmission spectra through an $L = 5 \text{ cm}$, $\mathcal{F} = 30$ cavity under strong coupling of individual rovibrational transitions of (a) methane (CH₄, ν_3 , $J = 3 \rightarrow 4$) and (b) acetylene (C₂H₂, ν_3 , $J = 9 \rightarrow 10$) at 295 K and varying number densities (colored traces). Simulated transmission spectra through 5 cm pathlengths of (a) CH₄ and (b) C₂H₂ at $N/V = 1 \times 10^{15} \text{ cm}^{-3}$ and 295 K are shown in black and offset for clarity. Molecular absorption cross sections are simulated using HITRAN line list parameters in PGOPHER with appropriate Doppler and pressure broadening.

B. Unimolecular isomerization: Nitrous acid and butadiene

Unimolecular isomerization reactions are compelling systems for polariton chemistry as they typically depend on IVR to funnel vibrational excitation into the isomerization coordinate. The *cis*-*trans* isomerization of nitrous acid (HONO) has been targeted by several theoretical VSC studies.^{68,93,94} IR-driven *cis*-*trans* HONO isomerization was first demonstrated in 1963⁹⁵ and has been studied since then.^{96–101} This process can be driven by pumping the $\sim 3400\text{ cm}^{-1}$ ν_1 O–H stretch of *cis*-HONO despite the reaction coordinate being more similar to the low-frequency ν_6 torsional mode.⁹⁷ Vibrational excitation of the O–H stretching coordinate must therefore redistribute into torsional motion for the reaction to progress. This IR-driven isomerization has only been observed in inert gas matrices where coupling to the matrix likely drives the redistribution of vibrational energy into the isomerization coordinate.⁹⁶ Still, it is worth considering whether cavity coupling could mediate IVR in gas-phase HONO, especially given the theoretical interest in this system.

We first evaluate strong coupling prospects for the brightest vibrational mode of the dominant *trans*-HONO isomer: the ν_3 HON bending mode (Fig. 4). We build a PGOPHER model for HONO using spectroscopic constants from Gaussian calculations and the literature^{97,99,102–105} in order to fit the available room-temperature data^{72,106} [red trace in Fig. 4(b)]. The features in the absorption cross section grow significantly narrower and brighter at 100 K [blue trace in Fig. 4(b)]. We identify the ν_3 , $J = 9 \rightarrow 10$ transition at 1271.218 cm^{-1} as an optimal rovibrational transition for cavity-coupling. We find that strong coupling of this transition is possible only with $N/V \geq 3 \times 10^{15}\text{ cm}^{-3}$ [Fig. 4(c), Table I]. Unfortunately, this number density is unlikely to be realized in the gas phase; the highest HONO number densities reported⁹⁹ do not exceed $8.4 \times 10^{13}\text{ cm}^{-3}$ and contamination with NO_x by-products is unavoidable even in dilute HONO samples.^{99,107–111} This source issue is likely prohibitive for achieving RVSC in the ν_3 band of the *trans*-HONO or any weaker vibrational transitions of either isomer. To simulate the *trans*-HONO absorption cross section accurately, one must account for the changing isomer equilibrium between room temperature and 100 K. Because *trans*-HONO is the thermodynamic ground state, we do expect a larger contribution from the *trans*-HONO population at low temperatures. We examine this further in Sec. SIII of the [supplementary material](#) but find that this consideration does not relax the inaccessibly high *trans*-HONO number density required for strong coupling.

Although RVSC of HONO is likely not realistic, there are other promising candidates for gas-phase isomerization. Cavity-altered isomerization of 1,3-butadiene (C_4H_6) has been the subject of recent theoretical work.¹¹² The planar *trans*- C_4H_6 isomer composes 95% of gaseous samples at room temperature; rotation about the central C–C bond leads to a non-planar *gauche* structure [Fig. 5(a)].¹¹³ The ultraviolet photochemistry of C_4H_6 has been studied for decades,^{114–117} and IR-driven photoisomerization should be possible, given the $\sim 2000\text{ cm}^{-1}$ torsional barrier in the electronic ground state.¹¹⁸ We evaluate prospects for cavity-coupling the ν_{11} CH_2 wagging mode of *trans*- C_4H_6 , which is the most strongly absorbing band at room temperature (Fig. 5, Table I). We build a C_4H_6 model in PGOPHER based on spectroscopic constants from calculations and the literature^{119–124} and fit it to the available experimental data^{72,106}

[red trace in Fig. 5(b)]. We neglect the temperature-dependence of the *trans*-*gauche* isomer equilibrium, which would lead to at most a 5% increase in the absorption cross section of *trans*- C_4H_6 at low temperatures. We inspect the 100 K simulated spectrum [blue trace

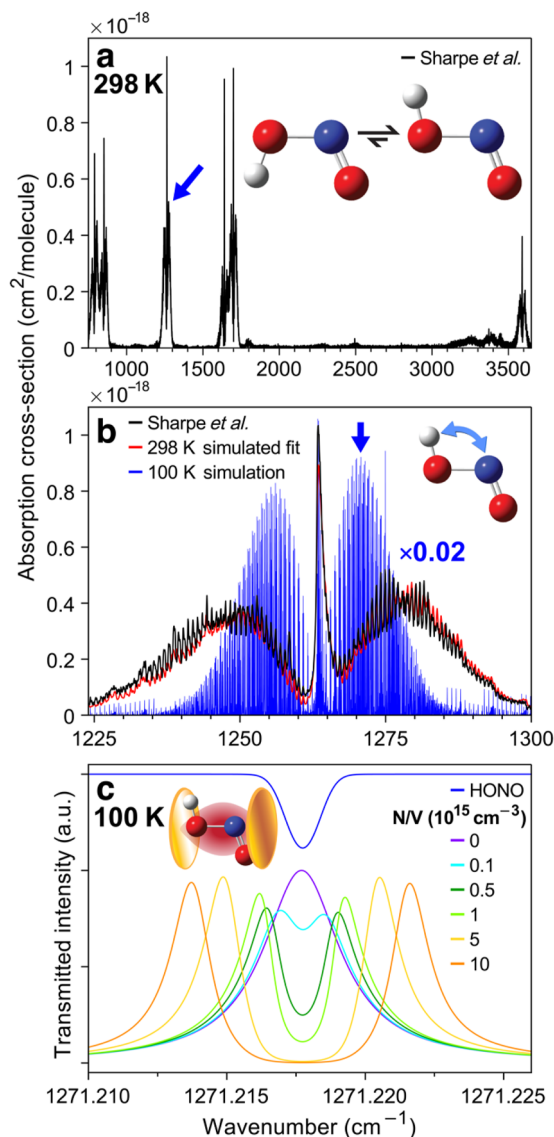


FIG. 4. (a) Broadband infrared absorption cross section of nitrous acid (HONO) at 298 K from Sharpe *et al.*¹⁰⁶ obtained via HITRAN. (b) Absorption cross section of the ν_3 HON bending mode of *trans*-HONO. The 298 K experimental absorption cross section (black) is fit with a PGOPHER simulation using spectroscopic constants from Gaussian calculations and the literature (red). The PGOPHER model is then used to calculate the low-temperature, low-pressure absorption cross section at 100 K and 0.0104 Torr ($N/V \sim 1 \times 10^{15}\text{ cm}^{-3}$), which is scaled by 0.02 for plotting (blue). (c) Classical cavity transmission simulation for coupling a ν_3 , $J = 9 \rightarrow 10$ rovibrational transition of *trans*-HONO in a $L = 5\text{ cm}$, $\mathcal{F} = 30$ cavity for varying intracavity number densities at 100 K (colored traces). A simulated trace of transmission through a 5 cm pathlength of an $N/V = 1 \times 10^{15}\text{ cm}^{-3}$, 100 K sample of *trans*-HONO is offset for clarity (blue).

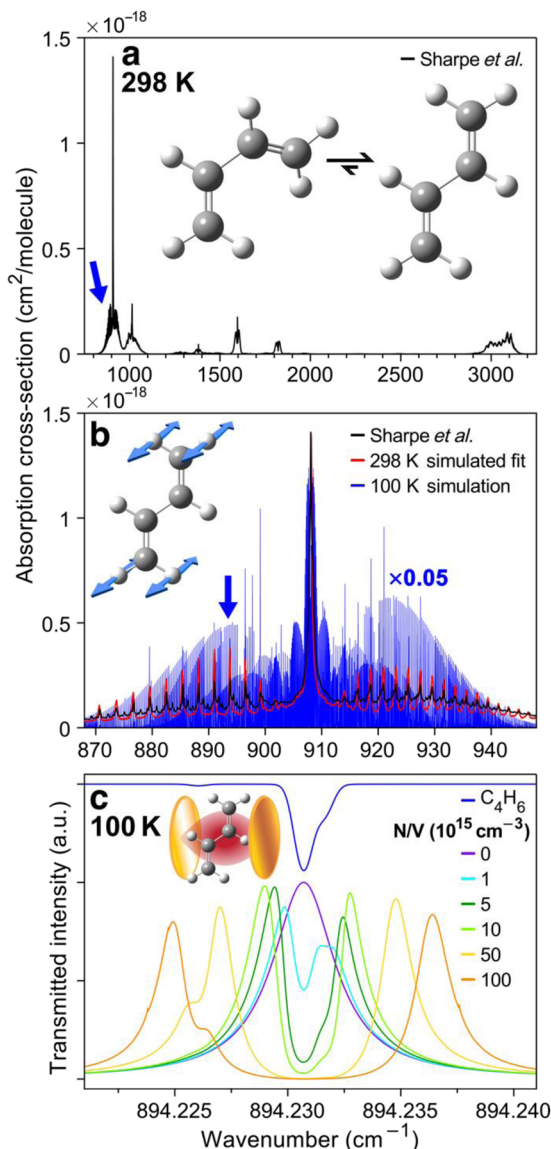


FIG. 5. (a) Broadband infrared absorption cross section of butadiene (C_4H_6) at 298 K from Sharpe *et al.*¹⁰⁶ obtained via HITRAN. (b) Absorption cross section of the ν_{11} CH_2 wagging mode of *trans*- C_4H_6 . The 298 K experimental absorption cross section (black) is fit with a PGOPHER simulation using spectroscopic constants from Gaussian calculations and the literature (red). The PGOPHER model is then used to calculate the low-temperature, low-pressure absorption cross section at 100 K and 0.0104 Torr ($N/V \sim 1 \times 10^{15} \text{ cm}^{-3}$), which is scaled by 0.05 for plotting (blue). (c) Classical cavity transmission simulation for coupling overlapping ν_{11} , $J = 8 \rightarrow 7$ and $J = 30 \rightarrow 29$ transitions of *trans*- C_4H_6 in a $L = 5 \text{ cm}$, $\mathcal{F} = 30$ cavity for varying intracavity number densities at 100 K (colored traces). A simulated trace of transmission through a 5 cm pathlength of an $N/V = 5 \times 10^{15} \text{ cm}^{-3}$, 100 K sample of *trans*- C_4H_6 is offset for clarity (blue).

in Fig. 5(b)] and select overlapping *trans*- C_4H_6 ν_{11} and $J = 8 \rightarrow 7$ and $J = 30 \rightarrow 29$ R-branch transitions at 894.2307 and 894.2316 cm^{-1} for cavity coupling. We find that the strong coupling regime is accessible for $N/V \geq 1 \times 10^{16} \text{ cm}^{-3}$ [Fig. 5(c)]. C_4H_6 is a commercially available

gas, so achieving these number densities is straightforward. Investigating the IR-driven chemistry of C_4H_6 under strong coupling could provide a platform to explore the cavity-IVR hypothesis. It may also be of interest to study IVR-mediated excited-state photochemistry of C_4H_6 under RVSC, including photoisomerization and photolysis.

C. Photodissociation of ozone under RVSC

We now explore cavity-coupling prospects for ozone (O_3), which features rich and well-studied photochemistry^{125–127} and is relevant to recent theoretical predictions for cavity-IVR mediated photodissociation of triatomic molecules.^{65,66} Here, we consider RVSC in ground-state O_3 ; we discuss possibilities for ESC of this system in Sec. IV D. Making use of the HITRAN line list for the infrared

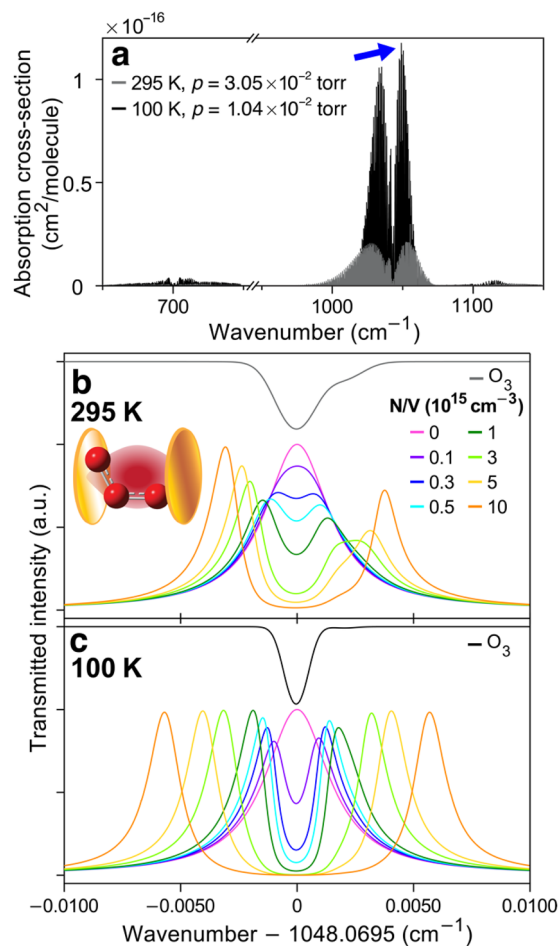


FIG. 6. (a) Broadband infrared absorption cross section of ozone (O_3) at 295 K (gray) and 100 K (black) with pressures corresponding to molecular densities of $N/V = 1 \times 10^{15} \text{ cm}^{-3}$ simulated using HITRAN line list parameters in PGOPHER. Classical cavity transmission simulations for coupling a ν_3 , $J = 7 \rightarrow 8$ rovibrational transition of O_3 in a $L = 5 \text{ cm}$, $\mathcal{F} = 30$ cavity for varying intracavity number densities at (b) 295 K and (c) 100 K (colored traces). Simulated traces of transmission through a 5 cm pathlength of (b) 295 K, $N/V = 3 \times 10^{15} \text{ cm}^{-3}$ and (c) 100 K, $N/V = 5 \times 10^{14} \text{ cm}^{-3}$ samples of O_3 are shown in gray and black, respectively, and offset for clarity.

spectrum of O_3 , we target the brightest ν_3 asymmetric stretching mode and select the $\nu_3, J = 7 \rightarrow 8$ transition at $1048.0695 \text{ cm}^{-1}$ for coupling (Fig. 6, Table I). We predict RVSC of O_3 to be possible with $N/V = 3 \times 10^{15} \text{ cm}^{-3}$ at room temperature and $N/V = 9 \times 10^{14} \text{ cm}^{-3}$ at 100 K [Figs. 6(b) and 6(c)]; these number densities are accessible with commercial ozone generators.

O_3 photodissociation under RVSC may be a useful system to test the cavity-IVR hypothesis. The theoretical work of Wang *et al.*^{65,66} predicts that generalized bent triatomic molecules resist photodissociation under strong (nearly ultrastrong) coupling to the bending vibrational mode. However, there are some important differences between the model system considered by Wang *et al.* and our proposed gas-phase strong coupling of O_3 . Here, we consider strong coupling of a single rovibrational transition of the asymmetric stretching band of O_3 . Wang *et al.*, on the other hand, do not account for rotational structure and assume that the entire vibrational oscillator strength is concentrated in a single transition. In addition, Wang *et al.* considered coupling to the bending mode of their triatomic model system, while we find that strong coupling of the bending mode of O_3 is not practical, given this band's tiny absorption cross section [Fig. 6(a)]. Studying photodissociation of O_3 under RVSC of the asymmetric stretch may still provide a useful step toward experimental tests of theory. It may also be worth exploring RVSC of heteronuclear triatomics (e.g., HOCl, OCS, or HCN) featuring brighter bending modes. Studying dissociation of any of these species following electronic or multiphoton vibrational photoexcitation under strong coupling could aid in understanding cavity mediation of vibrational energy redistribution.

D. Electronic strong coupling of molecular gases: Iodine and ozone

Electronic strong coupling of molecular gases has not yet been demonstrated but should be feasible. Here, we consider species with narrow, well-resolved vibronic transitions, so molecular linewidths are similar to centimeter-scale FP cavity linewidths. Molecular iodine (I_2) is one promising candidate. We build a model of the B–X band of I_2 near 538 nm ($18\,600 \text{ cm}^{-1}$) in PGOPHER using literature constants^{128–133} to fit room temperature experimental absorption cross section data [Fig. 7(a)].¹³⁴ Our model accounts for transitions from the $\nu_{\text{initial}} = 0, 1, 2,$ and 3 vibrational levels of the ground electronic state to the $\nu_{\text{final}} = 0–70$ vibrational levels of the excited B state in order to recover the experimental Franck–Condon envelope in the region of maximum absorption. We additionally account for the underlying I_2 hyperfine structure,^{132,135} which manifests in asymmetric line shapes under low pressure, Doppler-broadened conditions.

We confirm that this model reproduces the known state-resolved rovibronic structure of I_2 [Fig. 7(b)].¹²⁸ Finally, we simulate room temperature cavity-coupling of the B–X, $\nu_1 = 0 \rightarrow 29, J = 45 \rightarrow 46$ rovibronic transition lying near $18\,599.1 \text{ cm}^{-1}$ [Fig. 7(c), Table I]. We predict the onset of ESC for $N/V = 5 \times 10^{15} \text{ cm}^{-3}$. Solid iodine has a room-temperature vapor pressure¹³⁴ of 0.23 Torr, which corresponds to $N/V \sim 7.4 \times 10^{15} \text{ cm}^{-3}$, so the strong coupling regime should be accessible with the vapors of room-temperature or gently heated solid I_2 . Fluorescence and photodissociation yields of I_2 could be examined as a function of cavity-coupling conditions.

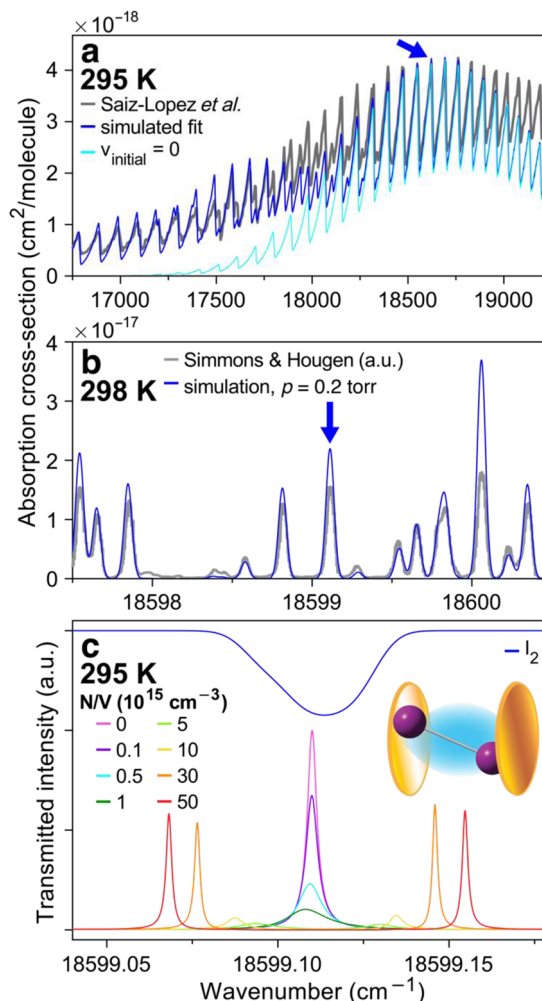


FIG. 7. (a) Absorption cross section of the B–X electronic band of iodine (I_2) at 295 K. Experimental air-broadened data from Saiz-Lopez *et al.*¹³⁴ (gray) is fit with a PGOPHER model built from literature constants (blue). Contributions from transitions originating in the $\nu_{\text{initial}} = 0$ level of the electronic ground state are shown in cyan. (b) The I_2 model (blue) reproduces rovibronic line positions of experimental absorbance data digitized from Simmons and Hougen (gray),¹²⁸ assuming a Gaussian linewidth of 0.055 cm^{-1} . The reference data from Simmons and Hougen is reported with arbitrary intensity units and strong transitions may be saturated. (c) Classical cavity transmission simulation for coupling the B–X, $\nu_1 = 0 \rightarrow 29, J = 45 \rightarrow 46$ rovibronic transition of I_2 in a $L = 5 \text{ cm}, \mathcal{F} = 30$ cavity for varying intracavity number densities at 295 K (colored traces). A simulated trace of transmission through a 5 cm pathlength of an $N/V = 1 \times 10^{15} \text{ cm}^{-3}, 295 \text{ K}$ sample of I_2 is offset for clarity (blue). The asymmetric line shape of the rovibronic transition arises from unresolved hyperfine structure.

We finally consider if ESC of a broad, unresolved electronic band is possible in the gas phase. Coupling a broad molecular transition to a single photonic mode demands a cavity with a large FSR and therefore a cavity length more typical of condensed-phase nanocavities ($L < 1 \mu\text{m}$). To give a concrete example, we consider coupling the B–X electronic band of ozone near 254 nm (Fig. S2, Table I).¹³⁶ Pumping this band leads to $O_3 \rightarrow O + O_2$

photodissociation with interesting wavelength-dependence of photoproduct electronic states,^{125–127} one could study these channels under ESC. Unfortunately, the O₃ B–X band is simply not bright enough for strong coupling, given its significant linewidth; we find that strong coupling only emerges for $N/V \geq 3 \times 10^{21} \text{ cm}^{-3}$ even in a wavelength-scale cavity with $L = 254 \text{ nm}$ [Fig. S2(b)]. For context, this number density corresponds to a physically unrealizable pressure of 122 atm at room temperature, assuming that the sample behaves as an ideal gas.

We therefore conclude that molecules with well-resolved vibronic bands and correspondingly narrower natural linewidths are better candidates for gas-phase ESC. Quantum-state resolution is a major strength of gas-phase polaritonics, and careful system selection will be central to future work.

V. FUTURE CHALLENGES AND CONCLUSIONS

As polariton chemistry continues into its second decade, gas-phase experiments may provide a bridge between condensed-phase demonstrations and theoretical understanding. We have argued here that many exciting proposals for cavity-altered reaction dynamics are well-suited for study in isolated gas-phase molecules. There are practical experimental benefits: generating polaritons in open plano-concave Fabry–Pérot cavities enables unique flexibility and control of both molecular and photonic degrees of freedom. We conclude this Perspective by highlighting some of these opportunities, as well as future challenges.

Orthogonal optical access. The relatively long centimeter-scale pathlengths of gas-phase FP cavities provide the opportunity for optical access orthogonal to the cavity axis, which is considerably harder to implement in microcavities. This access will allow for initiation and interrogation of intracavity processes without confounding optical filtering effects. For instance, the hypothetical intracavity PLP-LIF experiments discussed in Sec. IV A could be implemented with PLP (pump) and LIF (probe) beams aligned along the cavity axis and fluorescence detected orthogonally. One could also consider orthogonally aligned pump (or probe) beams, though one must take care to ensure that the molecules addressed along the orthogonal axis belong to the cavity-coupled ensemble. Designing the cavity length and beam diameter to be comparable in size (e.g., a few mm) would enable the orthogonal beam to cover the entire cavity mode volume. Such a beam would inevitably sample many uncoupled molecules, so modulation and lock-in schemes would likely be necessary to correlate off-axis signals with those from a beam aligned through the cavity. Relatedly, the polarizations of on- and off-cavity axis beams must be chosen carefully to address molecules of the same orientational class. In any event, this additional orthogonal spatial degree of freedom provides a wide space to explore in future work.

Photonic mode engineering. Engineering the photonic mode structure of gas-phase FP cavities is another untapped degree of freedom. Cavity-altered chemistry in planar microcavities appears to emerge only when a cavity mode with $k = 0$ in-plane momentum is strongly coupled to a transition of reactant molecules.^{53,63,137,138} The significance of this normal-incidence resonance condition is unknown and remains an open area of experimental and theoretical

work. It is unclear the extent to which the chemical and photophysical behavior of polaritonic systems arises from the continuous dispersion of modes in a tilt-tuned planar FP microcavity. In contrast, our plano-concave FP cavities support *discrete* Gaussian transverse spatial modes, whose spectra can be tuned by changing the ratio between the mirror radius of curvature and the cavity length.²⁰ The ability to manipulate the photonic density of states sets this platform apart from microfluidic cavities and may prove useful to test connections between mode dispersion and cavity-altered chemistry.

Homogeneous and inhomogeneous broadening. There are also open questions surrounding the significance of inhomogeneous broadening and molecular disorder to polariton formation, photophysics, and chemistry.^{8,13,73,139–142} Cavity transmission spectra and polariton dynamics are sensitive to whether the intracavity molecules chiefly exhibit homogeneous or inhomogeneous broadening. The tunability of gas-phase molecular line shapes is a powerful control knob to enable further investigation into this area: one can explore the interplay between inhomogeneous and homogeneous regimes by tuning the relative contributions of Gaussian Doppler broadening and Lorentzian pressure or transit-time broadening. One could then systematically probe the role of inhomogeneity on polariton dynamics, including the chemical processes discussed in Sec. IV, and nonlinear spectroscopy (*vide infra*). In addition, future gas-phase work might test Herrera and Owrutsky's prediction that strong coupling is accessible in the pressure-broadened regime for any gas whose absorption cross section exceeds its pressure broadening coefficient.¹³

Nonlinear spectroscopy. The gas phase is also a promising arena to solidify the understanding of the nonlinear spectroscopy of polaritons. Dunkelberger,^{52,143} Xiong,^{55,144} and Kubarych¹⁴⁵ and co-workers have made major strides in studying the ultrafast optical response of solution-phase vibrational polaritons—though not without some healthy debate.^{145,146} In parallel, a growing body of work from Scholes, Zanni, and others examines the nonlinear response of systems under electronic strong coupling.^{147–149} Nonlinear spectroscopy of gas-phase polaritons would provide complementary information to these condensed-phase experiments with well-resolved spectral features, non-overlapping excited state absorption signals, and the possibility of directly addressing the “dark state” reservoir off-axis. Altogether, many outstanding questions related to the intrinsic properties and optical response of vibrational and electronic polaritons can be directly probed in the gas phase and in turn inform spectroscopic measures of polariton reactivity.

We must note that there are significant challenges to working with polaritons in the gas phase as well. Absolute Rabi splittings are four to five orders of magnitude smaller for dilute gas samples than those typical in solution-phase microcavities. Centimeter-scale cavities also contain a large number of cavity-coupled molecules (e.g., $N \sim 10^{12}$ for $N/V = 3.5 \times 10^{15} \text{ cm}^{-3}$), so the effective coupling-per-molecule is orders of magnitude smaller than that of most microcavity implementations. The coupling-per-molecule could be increased by working with smaller mode volumes, more akin to geometries used in few-atom cQED experiments.^{30,31} Future work should examine whether moving in this direction is necessary to reveal cavity modification of gas-phase chemistry.

Finally, we note that spectroscopy of molecular gases in optical cavities is by no means a new development: the cavity-enhanced spectroscopy (CES) community has long used cavities to enhance light-matter interactions for more sensitive molecular absorption spectroscopy.¹⁵⁰ A deeper connection to the frameworks developed by the CES community would be of great benefit to researchers in the polariton field. One distinguishing factor of the CES and polaritonic regimes is that in the latter, a photon emitted by a molecule into the cavity mode is more likely to be reabsorbed by the molecular ensemble than to escape the cavity for detection. Experiments that systematically probe chemical reactivity and photophysics at the threshold of strong coupling are of utmost importance to clarify when polaritonic phenomena exceed the grasp of the classical optics language of CES. Gas-phase molecules again make a compelling platform to explore this threshold.

In summary, gas-phase molecular strong coupling may provide a powerful means to develop mechanistic understanding of polariton chemistry, harnessing pristine cavity-molecule coupling conditions and state-resolved detection schemes. Drawing inspiration from atomic cavity quantum electrodynamics, cavity-enhanced spectroscopy, and mode-specific chemistry, future work in the gas phase can build toward the dream of optical control of chemistry.

SUPPLEMENTARY MATERIAL

See the [supplementary material](#) for additional details and strong coupling simulations. In Sec. SI, we define absorption line shape broadening and tabulate specific broadening coefficients for all molecular systems considered in this work (Table SI). In Sec. SII, Figs. S1 and S2 depict absorption cross sections and strong coupling prospects for RVSC of CH₃OH and ESC of O₃. In Sec. SIII, we explicitly consider the temperature dependence of the HONO *trans-cis* isomer ratio.

ACKNOWLEDGMENTS

Marissa L. Weichman acknowledges support from the NSF CAREER under Award No. CHE-2238865 and startup funds provided by Princeton University. Jane C. Nelson acknowledges support from a National Defense Science and Engineering Graduate fellowship. The authors acknowledge helpful conversations with David Chandler and Jeff Owrutsky.

AUTHOR DECLARATIONS

Conflict of Interest

The authors have no conflicts to disclose.

Author Contributions

Jane C. Nelson: Conceptualization (equal); Data curation (lead); Formal analysis (lead); Writing – original draft (lead). **Marissa L. Weichman:** Conceptualization (equal); Funding acquisition (lead); Project administration (lead); Supervision (lead); Validation (lead); Writing – review & editing (lead).

DATA AVAILABILITY

The data that support the findings of this study are available from the corresponding author upon reasonable request.

REFERENCES

- 1 B. Kohler, J. L. Krause, F. Raksi, K. R. Wilson, V. V. Yakovlev, R. M. Whitnell, and Y. Yan, *Acc. Chem. Res.* **28**(3), 133–140 (1995).
- 2 R. N. Zare, *Science* **279**(5358), 1875–1879 (1998).
- 3 H. Rabitz, R. de Vivie-Riedle, M. Motzkus, and K. Kompa, *Science* **288**(5467), 824–828 (2000).
- 4 F. F. Crim, *Proc. Natl. Acad. Sci. U.S.A.* **105**(35), 12647–12648 (2008).
- 5 T. Brixner and G. Gerber, *ChemPhysChem* **4**(5), 418–438 (2003).
- 6 C. G. Elles and F. F. Crim, *Annu. Rev. Phys. Chem.* **57**(1), 273–302 (2006).
- 7 S. Karmakar and S. Keshavamurthy, *Phys. Chem. Chem. Phys.* **22**(20), 11139–11173 (2020).
- 8 B. S. Simpkins, A. D. Dunkelberger, and J. C. Owrutsky, *J. Phys. Chem. C* **125**(35), 19081–19087 (2021).
- 9 B. S. Simpkins, A. D. Dunkelberger, and I. Vurgaftman, *Chem. Rev.* **123**(8), 5020–5048 (2023).
- 10 A. D. Wright, J. C. Nelson, and M. L. Weichman, *J. Am. Chem. Soc.* **145**(10), 5982–5987 (2023).
- 11 A. D. Wright, J. C. Nelson, and M. L. Weichman, *J. Chem. Phys.* **159**(16), 164202 (2023).
- 12 R. F. Ribeiro, L. A. Martínez-Martínez, M. Du, J. Campos-Gonzalez-Angulo, and J. Yuen-Zhou, *Chem. Sci.* **9**(30), 6325–6339 (2018).
- 13 F. Herrera and J. Owrutsky, *J. Chem. Phys.* **152**(10), 100902 (2020).
- 14 A. Mandal, M. A. D. Taylor, B. M. Weight, E. R. Koessler, X. Li, and P. Huo, *Chem. Rev.* **123**(16), 9786–9879 (2023).
- 15 E. T. Jaynes and F. W. Cummings, *Proc. IEEE* **51**(1), 89–109 (1963).
- 16 B. W. Shore and P. L. Knight, *J. Mod. Opt.* **40**(7), 1195–1238 (1993).
- 17 M. Tavis and F. W. Cummings, *Phys. Rev.* **170**(2), 379–384 (1968).
- 18 T. Khazanov, S. Gunasekaran, A. George, R. Lomlu, S. Mukherjee, and A. J. Musser, *Chem. Phys. Rev.* **4**(4), 041305 (2023).
- 19 Y. Zhu, D. J. Gauthier, S. E. Morin, Q. Wu, H. J. Carmichael, and T. W. Mossberg, *Phys. Rev. Lett.* **64**(21), 2499–2502 (1990).
- 20 W. G. Nagourney, *Quantum Electronics for Atomic Physics* (Oxford University Press, Oxford, 2010).
- 21 H. A. MacLeod, in *Thin-Film Optical Filters*, 4th ed. (CRC Press, Boca Raton, FL, 2010).
- 22 M. A. Zeb, P. G. Kirton, and J. Keeling, *ACS Photonics* **5**(1), 249–257 (2018).
- 23 J. Yuen-Zhou and A. Koner, *J. Chem. Phys.* **160**(15), 154107 (2024).
- 24 P. Goy, J. M. Raimond, M. Gross, and S. Haroche, *Phys. Rev. Lett.* **50**(24), 1903–1906 (1983).
- 25 E. M. Purcell, H. C. Torrey, and R. V. Pound, *Phys. Rev.* **69**, 681 (1946).
- 26 D. Kleppner, *Phys. Rev. Lett.* **47**(4), 233–236 (1981).
- 27 Y. Kaluzny, P. Goy, M. Gross, J. M. Raimond, and S. Haroche, *Phys. Rev. Lett.* **51**(13), 1175–1178 (1983).
- 28 S. Haroche and D. Kleppner, *Phys. Today* **42**(1), 24–30 (1989).
- 29 S. Haroche, *Rev. Mod. Phys.* **85**(3), 1083–1102 (2013).
- 30 M. G. Raizen, R. J. Thompson, R. J. Brecha, H. J. Kimble, and H. J. Carmichael, *Phys. Rev. Lett.* **63**(3), 240–243 (1989).
- 31 R. J. Thompson, G. Rempe, and H. J. Kimble, *Phys. Rev. Lett.* **68**(8), 1132–1135 (1992).
- 32 G. Rempe, R. J. Thompson, and H. J. Kimble, *Phys. Scr.* **T51**, 67–77 (1994).
- 33 H. J. Kimble, *Phys. Scr.* **T76**(1), 127 (1998).
- 34 C. Weisbuch, M. Nishioka, A. Ishikawa, and Y. Arakawa, *Phys. Rev. Lett.* **69**(23), 3314–3317 (1992).
- 35 M. S. Skolnick, T. A. Fisher, and D. M. Whittaker, *Semicond. Sci. Technol.* **13**(7), 645–669 (1998).
- 36 D. G. Lidzey, D. D. C. Bradley, M. S. Skolnick, T. Virgili, S. Walker, and D. M. Whittaker, *Nature* **395**(6697), 53–55 (1998).

- ³⁷S. Kéna-Cohen and S. R. Forrester, *Nat. Photonics* **4**(6), 371–375 (2010).
- ³⁸H. Deng, H. Haug, and Y. Yamamoto, *Rev. Mod. Phys.* **82**(2), 1489–1537 (2010).
- ³⁹J. Keeling and S. Kéna-Cohen, *Annu. Rev. Phys. Chem.* **71**(1), 435–459 (2020).
- ⁴⁰T. W. Ebbesen, *Acc. Chem. Res.* **49**(11), 2403–2412 (2016).
- ⁴¹T. Schwartz, J. A. Hutchison, C. Genet, and T. W. Ebbesen, *Phys. Rev. Lett.* **106**(19), 196405 (2011).
- ⁴²J. A. Hutchison, T. Schwartz, C. Genet, E. Devaux, and T. W. Ebbesen, *Angew. Chem., Int. Ed.* **51**(7), 1592–1596 (2012).
- ⁴³S. Wang, T. Chervy, J. George, J. A. Hutchison, C. Genet, and T. W. Ebbesen, *J. Phys. Chem. Lett.* **5**(8), 1433–1439 (2014).
- ⁴⁴A. Shalabney, J. George, J. Hutchison, G. Pupillo, C. Genet, and T. W. Ebbesen, *Nat. Commun.* **6**(1), 5981 (2015).
- ⁴⁵J. P. Long and B. S. Simpkins, *ACS Photonics* **2**(1), 130–136 (2015).
- ⁴⁶J. George, A. Shalabney, J. A. Hutchison, C. Genet, and T. W. Ebbesen, *J. Phys. Chem. Lett.* **6**(6), 1027–1031 (2015).
- ⁴⁷B. S. Simpkins, K. P. Fears, W. J. Dressick, B. T. Spann, A. D. Dunkelberger, and J. C. Owrutsky, *ACS Photonics* **2**(10), 1460–1467 (2015).
- ⁴⁸A. Thomas, J. George, A. Shalabney, M. Dryzhakov, S. J. Varma, J. Moran, T. Chervy, X. Zhong, E. Devaux, C. Genet, J. A. Hutchison, and T. W. Ebbesen, *Angew. Chem., Int. Ed.* **55**(38), 11462–11466 (2016).
- ⁴⁹G. D. Wiesehan and W. Xiong, *J. Chem. Phys.* **155**(24), 241103 (2021).
- ⁵⁰M. V. Imperatore, J. B. Asbury, and N. C. Giebink, *J. Chem. Phys.* **154**(19), 191103 (2021).
- ⁵¹K. Nagarajan, A. Thomas, and T. W. Ebbesen, *J. Am. Chem. Soc.* **143**(41), 16877–16889 (2021).
- ⁵²A. D. Dunkelberger, B. S. Simpkins, I. Vurgaftman, and J. C. Owrutsky, *Annu. Rev. Phys. Chem.* **73**(1), 429–451 (2022).
- ⁵³W. Ahn, J. F. Triana, F. Recabal, F. Herrera, and B. S. Simpkins, *Science* **380**(6650), 1165–1168 (2023).
- ⁵⁴A. P. Fidler, L. Chen, A. M. McKillop, and M. L. Weichman, *J. Chem. Phys.* **159**(16), 164302 (2023).
- ⁵⁵W. Xiong, *Acc. Chem. Res.* **56**(7), 776–786 (2023).
- ⁵⁶R. Bhuyan, J. Mony, O. Kotov, G. W. Castellanos, J. Gómez Rivas, T. O. Shegai, and K. Börjesson, *Chem. Rev.* **123**(18), 10877–10919 (2023).
- ⁵⁷E. Eizner, L. A. Martínez-Martínez, J. Yuen-Zhou, and S. Kéna-Cohen, *Sci. Adv.* **5**(12), eaax4482 (2019).
- ⁵⁸H. Zeng, J. B. Pérez-Sánchez, C. T. Eckdahl, P. Liu, W. J. Chang, E. A. Weiss, J. A. Kalow, J. Yuen-Zhou, and N. P. Stern, *J. Am. Chem. Soc.* **145**(36), 19655–19661 (2023).
- ⁵⁹I. Lee, S. R. Melton, D. Xu, and M. Delor, *J. Am. Chem. Soc.* **146**(14), 9544–9553 (2024).
- ⁶⁰P. A. Thomas, W. J. Tan, V. G. Kravets, A. N. Grigorenko, and W. L. Barnes, *Adv. Mater.* **36**(7), 2309393 (2024).
- ⁶¹A. J. Orr-Ewing, *Chem. Soc. Rev.* **46**(24), 7597–7614 (2017).
- ⁶²D. S. Wang and S. F. Yelin, *ACS Photonics* **8**(10), 2818–2826 (2021).
- ⁶³J. A. Campos-Gonzalez-Angulo, Y. R. Poh, M. Du, and J. Yuen-Zhou, *J. Chem. Phys.* **158**(23), 230901 (2023).
- ⁶⁴T.-T. Chen, M. Du, Z. Yang, J. Yuen-Zhou, and W. Xiong, *Science* **378**(6621), 790–794 (2022).
- ⁶⁵D. S. Wang, T. Neuman, S. F. Yelin, and J. Flick, *J. Phys. Chem. Lett.* **13**(15), 3317–3324 (2022).
- ⁶⁶D. S. Wang, J. Flick, and S. F. Yelin, *J. Chem. Phys.* **157**(22), 224304 (2022).
- ⁶⁷C. Schäfer, J. Flick, E. Ronca, P. Narang, and A. Rubio, *Nat. Commun.* **13**(1), 7817 (2022).
- ⁶⁸J. Sun and O. Vendrell, *J. Phys. Chem. Lett.* **13**(20), 4441–4446 (2022).
- ⁶⁹E. W. Fischer, J. Anders, and P. Saalfrank, *J. Chem. Phys.* **156**(15), 154305 (2022).
- ⁷⁰D. J. Nesbitt and R. W. Field, *J. Phys. Chem.* **100**(31), 12735–12756 (1996).
- ⁷¹K. Liu, *Annu. Rev. Phys. Chem.* **67**(1), 91–111 (2016).
- ⁷²I. E. Gordon, L. S. Rothman, R. J. Hargreaves, R. Hashemi, E. V. Karlovets, F. M. Skinner, E. K. Conway, C. Hill, R. V. Kochanov, Y. Tan, P. Wcislo, A. A. Finenko, K. Nelson, P. F. Bernath, M. Birk, V. Boudon, A. Campargue, K. V. Chance, A. Coustenis, B. J. Drouin, J. –M. Flaud, R. R. Gamache, J. T. Hodges, D. Jacquemart, E. J. Mlawer, A. V. Nikitin, V. I. Perevalov, M. Rotger, J. Tennyson, G. C. Toon, H. Tran, V. G. Tyuterev, E. M. Adkins, A. Baker, A. Barbe, E. Canè, A. G. Császár, A. Dudaryonok, O. Egorov, A. J. Fleisher, H. Fleurbaey, A. Foltynowicz, T. Furtenbacher, J. J. Harrison, J. –M. Hartmann, V. –M. Horneman, X. Huang, T. Karman, J. Karns, S. Kass, I. Kleiner, V. Kofman, F. KwabiaTchana, N. N. Lavrentieva, T. J. Lee, D. A. Long, A. A. Lukashovskaya, O. M. Lyulin, V. Y. Makhnev, W. Matt, S. T. Massie, M. Melosso, S. N. Mikhailenko, D. Mondelain, H. S. P. Müller, O. V. Naumenko, A. Perrin, O. L. Polyansky, E. Raddaoui, P. L. Raston, Z. D. Reed, M. Rey, C. Richard, R. Tóbiás, I. Sadiek, D. W. Schwenke, E. Starikova, K. Sung, F. Tamassia, S. A. Tashkun, J. Vander Auwera, I. A. Vasilenko, A. A. Vigin, G. L. Villanueva, B. Vispoel, G. Wagner, A. Yachmenev, and S. N. Yurchenko, *J. Quant. Spectrosc. Radiat. Transfer* **277**, 107949 (2022).
- ⁷³R. Houdré, R. P. Stanley, and M. Ilegems, *Phys. Rev. A* **53**(4), 2711–2715 (1996).
- ⁷⁴C. M. Western, *J. Quant. Spectrosc. Radiat. Transfer* **186**, 221–242 (2017).
- ⁷⁵M. J. Frisch, G. W. Trucks, H. B. Schlegel, G. E. Scuseria, M. A. Robb, J. R. Cheeseman, G. Scalmani, V. Barone, G. A. Petersson, H. Nakatsuji, X. Li, M. Caricato, A. V. Marenich, J. Bloino, B. G. Janesko, R. Gomperts, B. Mennucci, H. P. Hratchian, J. V. Ortiz, A. F. Izmaylov, J. L. Sonnenberg, D. Williams-Young, F. Ding, F. Lipparini, F. Egidi, J. Goings, B. Peng, A. Petrone, T. Henderson, D. Ranasinghe, V. G. Zakrzewski, J. Gao, N. Rega, G. Zheng, W. Liang, M. Hada, M. Ehara, K. Toyota, R. Fukuda, J. Hasegawa, M. Ishida, T. Nakajima, Y. Honda, O. Kitao, H. Nakai, T. Vreven, K. Throssell, J. A. Montgomery, Jr., J. E. Peralta, F. Ogliaro, M. J. Bearpark, J. J. Heyd, E. N. Brothers, K. N. Kudin, V. N. Staroverov, T. A. Keith, R. Kobayashi, J. Normand, K. Raghavachari, A. P. Rendell, J. C. Burant, S. S. Iyengar, J. Tomasi, M. Cossi, J. M. Millam, M. Klene, C. Adamo, R. Cammi, J. W. Ochterski, R. L. Martin, K. Morokuma, O. Farkas, J. B. Foresman, and D. J. Fox, *Gaussian 16*, Revision C.01, 2016.
- ⁷⁶M. K. Sprague, Ph.D. thesis, California Institute of Technology, Pasadena, CA, 2012.
- ⁷⁷M. J. Wilhelm, T. J. Johnson, and T. L. Myers, *AIP Adv.* **13**(5), 055133 (2023).
- ⁷⁸K. W. McKee, M. A. Blitz, P. A. Cleary, D. R. Glowacki, M. J. Pilling, P. W. Seakins, and L. Wang, *J. Phys. Chem. A* **111**(19), 4043–4055 (2007).
- ⁷⁹A. Bonard, V. Daële, J.-L. Delfau, and C. Vovelle, *J. Phys. Chem. A* **106**(17), 4384–4389 (2002).
- ⁸⁰J. A. Manion, R. E. Huie, R. D. Levin, D. R. Burgess, Jr., V. L. Orkin, W. Tsang, W. S. McGivern, J. W. Hudgens, V. D. Knyazev, D. B. Atkinson, E. Chai, A. M. Tereza, C.-Y. Lin, T. C. Allison, W. G. Mallard, F. Westley, J. T. Herron, R. F. Hampson, and D. H. Frizzell, NIST Chemical Kinetics Database, NIST Standard Reference Database 17, Version 7.0 (Web Version), Release 1.6. 8, Data version 2015.09, National Institute of Standards and Technology, 2015.
- ⁸¹M. G. Bryukov, V. D. Knyazev, S. M. Lomnicki, C. A. McFerrin, and B. Dellinger, *J. Phys. Chem. A* **108**(47), 10464–10472 (2004).
- ⁸²K. Yamasaki, A. Watanabe, T. Kakuda, N. Ichikawa, and I. Tokue, *J. Phys. Chem. A* **103**(4), 451–459 (1999).
- ⁸³M. D. Wheeler, M. Tsiouris, M. I. Lester, and G. Lendvay, *J. Chem. Phys.* **112**(15), 6590–6602 (2000).
- ⁸⁴M. Tsiouris, M. D. Wheeler, and M. I. Lester, *J. Chem. Phys.* **114**(1), 187–197 (2001).
- ⁸⁵T. J. Dillon, D. Hölscher, V. Sivakumaran, A. Horowitz, and J. N. Crowley, *Phys. Chem. Chem. Phys.* **7**(2), 349–355 (2005).
- ⁸⁶R. J. Shannon, M. A. Blitz, A. Goddard, and D. E. Heard, *Nat. Chem.* **5**(9), 745–749 (2013).
- ⁸⁷O. V. Boyarkin, L. Lubich, R. D. F. Settle, D. S. Perry, and T. R. Rizzo, *J. Chem. Phys.* **107**(20), 8409–8422 (1997).
- ⁸⁸L.-H. Xu, X. Wang, T. J. Cronin, D. S. Perry, G. T. Fraser, and A. S. Pine, *J. Mol. Spectrosc.* **185**(1), 158–172 (1997).
- ⁸⁹X. Wang and D. S. Perry, *J. Chem. Phys.* **109**(24), 10795–10805 (1998).
- ⁹⁰A. McIlroy and D. J. Nesbitt, *J. Chem. Phys.* **91**(1), 104–113 (1989).
- ⁹¹A. McIlroy and D. J. Nesbitt, *J. Chem. Phys.* **92**(4), 2229–2243 (1990).
- ⁹²J. J. Harrison, N. D. C. Allen, and P. F. Bernath, *J. Quant. Spectrosc. Radiat. Transfer* **113**(17), 2189–2196 (2012).
- ⁹³M. Du, R. F. Ribeiro, and J. Yuen-Zhou, *Chem* **5**(5), 1167–1181 (2019).
- ⁹⁴J. Sun and O. Vendrell, *J. Phys. Chem. Lett.* **14**(38), 8397–8404 (2023).
- ⁹⁵R. T. Hall and G. C. Pimentel, *J. Chem. Phys.* **38**(8), 1889–1897 (1963).

- ⁹⁶V. Botan, Ph.D. thesis, University of Zurich, Zurich, 2008.
- ⁹⁷V. Sironneau, J.-M. Flaud, J. Orphal, I. Kleiner, and P. Chelin, *J. Mol. Spectrosc.* **259**(2), 100–104 (2010).
- ⁹⁸R. Armante, A. Perrin, F. Kwabia Tchana, and L. Manceron, *Mol. Phys.* **120**(15–16), e1951860 (2022).
- ⁹⁹W. S. Barney, L. M. Wingen, M. J. Lakin, T. Brauers, J. Stutz, and B. J. Finlayson-Pitts, *J. Phys. Chem. A* **104**(8), 1692–1699 (2000).
- ¹⁰⁰A. Bongartz, J. Kames, F. Welter, and U. Schurath, *J. Phys. Chem.* **95**(3), 1076–1082 (1991).
- ¹⁰¹V. Botan and P. Hamm, *J. Chem. Phys.* **129**(16), 114510 (2008).
- ¹⁰²R. H. Kagann and A. G. Maki, *J. Quant. Spectrosc. Radiat. Transfer* **30**(1), 37–44 (1983).
- ¹⁰³C. M. Deeley and I. M. Mills, *Mol. Phys.* **54**(1), 23–32 (1985).
- ¹⁰⁴A. G. Maki, *J. Mol. Spectrosc.* **127**(1), 104–111 (1988).
- ¹⁰⁵J. M. Guilmot, M. Godefroid, and M. Herman, *J. Mol. Spectrosc.* **160**(2), 387–400 (1993).
- ¹⁰⁶S. W. Sharpe, T. J. Johnson, R. L. Sams, P. M. Chu, G. C. Rhoderick, and P. A. Johnson, *Appl. Spectrosc.* **58**(12), 1452–1461 (2004).
- ¹⁰⁷M. Taira and Y. Kanda, *Anal. Chem.* **62**(6), 630–633 (1990).
- ¹⁰⁸A. Febo, C. Perrino, M. Gherardi, and R. Sparapani, *Environ. Sci. Technol.* **29**(9), 2390–2395 (1995).
- ¹⁰⁹B. J. Finlayson-Pitts, L. M. Wingen, A. L. Sumner, D. Syomin, and K. A. Ramazan, *Phys. Chem. Chem. Phys.* **5**(2), 223–242 (2003).
- ¹¹⁰M. Lao, L. R. Crilley, L. Salehpoor, T. C. Furlani, I. Bourgeois, J. A. Neuman, A. W. Rollins, P. R. Veres, R. A. Washenfelder, C. C. Womack, C. J. Young, and T. C. VandenBoer, *Atmos. Meas. Tech.* **13**(11), 5873–5890 (2020).
- ¹¹¹G. Villena and J. Kleffmann, *Atmos. Meas. Tech.* **15**(3), 627–637 (2022).
- ¹¹²M. Severi and F. Zerbetto, *J. Phys. Chem. Lett.* **14**(40), 9145–9149 (2023).
- ¹¹³J. H. Baraban, M. Martin-Drumel, P. B. Changala, S. Eibenberger, M. Nava, D. Patterson, J. F. Stanton, G. B. Ellison, and M. C. McCarthy, *Angew. Chem., Int. Ed.* **57**(7), 1821–1825 (2018).
- ¹¹⁴I. Haller and R. Srinivasan, *J. Chem. Phys.* **40**(7), 1992–1997 (1964).
- ¹¹⁵W. Fuß, W. E. Schmid, and S. A. Trushin, *Chem. Phys. Lett.* **342**(1–2), 91–98 (2001).
- ¹¹⁶J. C. Robinson, S. A. Harris, W. Sun, N. E. Sveum, and D. M. Neumark, *J. Am. Chem. Soc.* **124**(34), 10211–10224 (2002).
- ¹¹⁷A. Makida, H. Igarashi, T. Fujiwara, T. Sekikawa, Y. Harabuchi, and T. Taketsugu, *J. Phys. Chem. Lett.* **5**(10), 1760–1765 (2014).
- ¹¹⁸R. Engeln, D. Consalvo, and J. Reuss, *Chem. Phys.* **160**(3), 427–433 (1992).
- ¹¹⁹N. C. Craig, J. L. Davis, K. A. Hanson, M. C. Moore, K. J. Weidenbaum, and M. Lock, *J. Mol. Struct.* **695–696**, 59–69 (2004).
- ¹²⁰M. Halonen, L. Halonen, and D. J. Nesbitt, *J. Phys. Chem. A* **108**(16), 3367–3372 (2004).
- ¹²¹N. C. Craig, P. Groner, and D. C. McKean, *J. Phys. Chem. A* **110**(23), 7461–7469 (2006).
- ¹²²N. C. Craig and R. L. Sams, *J. Phys. Chem. A* **112**(49), 12637–12646 (2008).
- ¹²³S. V. Krasnoshchekov, N. C. Craig, P. Boopalachandran, J. Laane, and N. F. Stepanov, *J. Phys. Chem. A* **119**(43), 10706–10723 (2015).
- ¹²⁴M.-A. Martin-Drumel, J. P. Porterfield, M. Goubet, P. Asselin, R. Georges, P. Soulard, M. Nava, P. B. Changala, B. Billingham, O. Pirali, M. C. McCarthy, and J. H. Baraban, *J. Phys. Chem. A* **124**(12), 2427–2435 (2020).
- ¹²⁵P. F. Zittel and D. D. Little, *J. Chem. Phys.* **72**(11), 5900–5905 (1980).
- ¹²⁶R. P. Wayne, *Atmos. Environ.* **21**(8), 1683–1694 (1987).
- ¹²⁷Y. Matsumi and M. Kawasaki, *Chem. Rev.* **103**(12), 4767–4782 (2003).
- ¹²⁸J. D. Simmons and J. T. Hougen, *J. Res. Natl. Bur. Stan. Sect. A* **81A**(1), 25 (1977).
- ¹²⁹S. Gerstenkorn and P. Luc, *Atlas du spectre d'absorption de la molécule d'iode: 14 800–20 000 cm⁻¹* (Editions du Centre National de le Recherche Scientifique, Paris, 1978).
- ¹³⁰J. Tellinghuisen, *J. Quant. Spectrosc. Radiat. Transfer* **19**(2), 149–161 (1978).
- ¹³¹S. Gerstenkorn and P. Luc, *J. Phys. France* **46**(6), 867–881 (1985).
- ¹³²B. Bodermann, H. Knöckel, and E. Tiemann, *Eur. Phys. J. D* **19**(1), 31–44 (2002).
- ¹³³J. Tellinghuisen, *J. Chem. Phys.* **134**(8), 084301 (2011).
- ¹³⁴A. Saiz-Lopez, R. W. Saunders, D. M. Joseph, S. H. Ashworth, and J. M. C. Plane, *Atmos. Chem. Phys.* **4**(5), 1443–1450 (2004).
- ¹³⁵J. Ye, L. Robertsson, S. Picard, L.-S. Ma, and J. L. Hall, *IEEE Trans. Instrum. Meas.* **48**(2), 544 (1999).
- ¹³⁶A. Serdyuchenko, V. Gorshchev, M. Weber, W. Chehade, and J. P. Burrows, *Atmos. Meas. Tech.* **7**(2), 625–636 (2014).
- ¹³⁷A. Thomas, L. Lethuillier-Karl, K. Nagarajan, R. M. A. Vergauwe, J. George, T. Cheryy, A. Shalabney, E. Devaux, C. Genet, J. Moran, and T. W. Ebbesen, *Science* **363**(6427), 615–619 (2019).
- ¹³⁸S. Montillo Vega, W. Ying, and P. Huo, *ChemRxiv:7t4c* (2024).
- ¹³⁹G. D. Scholes, C. A. DelPo, and B. Kudisch, *J. Phys. Chem. Lett.* **11**(15), 6389–6395 (2020).
- ¹⁴⁰A. George, T. Geraghty, Z. Kelsey, S. Mukherjee, G. Davidova, W. Kim, and A. J. Musser, *Adv. Opt. Mater.* **12**(11), 2302387 (2024).
- ¹⁴¹T. Liu, G. Yin, and W. Xiong, *ChemRxiv:zfb2* (2024).
- ¹⁴²A. Dutta, V. Tiainen, L. Duarte, N. Markesevic, I. Sokolovskii, D. Morozov, H. A. Qureshi, S. Pikker, G. Groenhof, and J. J. Toppari, *ChemRxiv:6vohv.v2* (2024).
- ¹⁴³A. D. Dunkelberger, B. T. Spann, K. P. Fears, B. S. Simpkins, and J. C. Owrutsky, *Nat. Commun.* **7**(1), 13504 (2016).
- ¹⁴⁴B. Xiang, R. F. Ribeiro, A. D. Dunkelberger, J. Wang, Y. Li, B. S. Simpkins, J. C. Owrutsky, J. Yuen-Zhou, and W. Xiong, *Proc. Natl. Acad. Sci. U.S.A.* **115**(19), 4845–4850 (2018).
- ¹⁴⁵R. Duan, J. N. Mastron, Y. Song, and K. J. Kubarych, *J. Phys. Chem. Lett.* **12**(46), 11406–11414 (2021).
- ¹⁴⁶B. S. Simpkins, Z. Yang, A. D. Dunkelberger, I. Vurgafman, J. C. Owrutsky, and W. Xiong, *J. Phys. Chem. Lett.* **14**(4), 983–988 (2023).
- ¹⁴⁷T. Schwartz, J. A. Hutchison, J. Léonard, C. Genet, S. Haacke, and T. W. Ebbesen, *ChemPhysChem* **14**(1), 125–131 (2013).
- ¹⁴⁸C. A. DelPo, B. Kudisch, K. H. Park, S.-U.-Z. Khan, F. Fassioli, D. Fausti, B. P. Rand, and G. D. Scholes, *J. Phys. Chem. Lett.* **11**(7), 2667–2674 (2020).
- ¹⁴⁹M. Son, Z. T. Armstrong, R. T. Allen, A. Dhavamani, M. S. Arnold, and M. T. Zanni, *Nat. Commun.* **13**(1), 7305 (2022).
- ¹⁵⁰G. Gagliardi and H. Looock, *Cavity-Enhanced Spectroscopy and Sensing*, 1st ed. (Springer, Berlin, Heidelberg, 2014).