

Spatially Resolved Elastic Modulus of Magnesium Silicate Hydrate: A Cementitious Material

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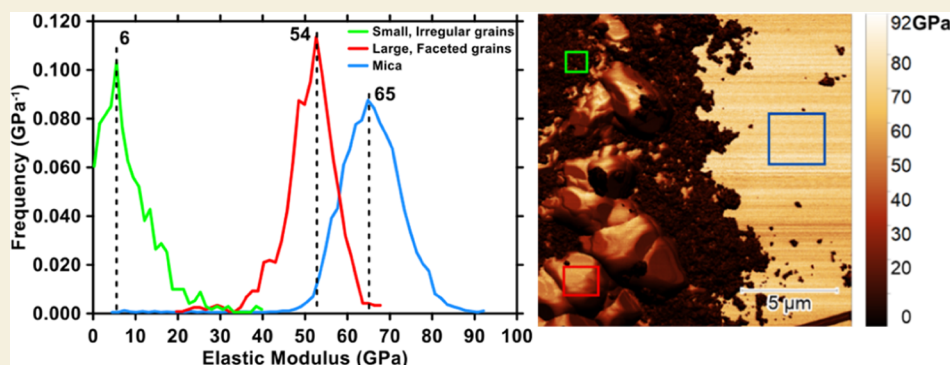
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ABSTRACT: Magnesium-based cement such as magnesium silicate hydrate (M–S–H) has drawn interest as a substitute for ordinary Portland cement. The precise determination of M–S–H's elastic modulus is essential in assessing the mechanical performance of M–S–H-based concrete. Atomic force microscopy (AFM) allows for spatially resolved quantification of nanomechanical characteristics of materials, including thin overgrowths on substrates. In this study, elastic modulus maps of M–S–H grown on single crystal mica surfaces were obtained using amplitude modulation–frequency modulation AFM. The effects of the Mg-to-Si molar ratio and morphology on the elastic modulus of M–S–H films were investigated. AFM reveals elastic moduli ranging from 5 to 15 and 40 to 64 GPa, depending on the growth reaction time and morphology. A comparison with calcium silicate hydrate found in traditional cement reveals insights into the influence of silicate polymerization on the material's stiffness. These findings demonstrate the effectiveness of AFM in quantitatively describing the mechanical characteristics of cementitious phases and suggest that M–S–H cement has the potential for use in a variety of construction applications.

KEYWORDS: mechanical properties, atomic force microscopy, amplitude modulation-frequency modulation, alternative cement

1. INTRODUCTION

Magnesium silicate hydrate (M–S–H) is a potential alternative to calcium silicate hydrate (C–S–H) composing ordinary Portland cement (OPC) traditionally used in construction.^{1–4} By substituting for C–S–H, the strain on calcium resources is reduced while taking advantage of the plentiful magnesium supply, e.g., in rocks rich in magnesium silicates.^{5–8} However, the use of M–S–H in construction applications is currently limited, in part because of a lack of knowledge regarding its cementitious qualities, particularly its mechanical properties, durability, and long-term performance.

The elastic modulus of cementitious materials describes their capacity to endure elastic deformation under applied stresses. It is crucial for evaluating the strength, resilience, and overall performance of concrete. For instance, in applications where stiffness and load-bearing capability are desired, such as in high-rise buildings or bridges, a rigid material that can effectively resist deformation and convey stress is desired.⁹ On the contrary, a low elastic modulus is characteristic of a flexible

material that can withstand loads and motions and is preferred for applications requiring compatibility with thermal expansion or fracture resistance, such as pavements or structures exposed to large temperature fluctuations.^{10,11} Therefore, the design, durability, and serviceability of concrete structures depend on an understanding of the elastic modulus.

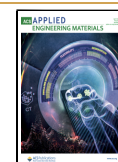
The elastic moduli of cement—typically measured for a paste, mortar, or concrete, instead of a specific hydrate in the binder such as C–S–H—have been studied using non-destructive techniques,¹² nanoindentations,¹³ finite element analysis,¹⁴ and molecular dynamics simulations.¹⁵ It has been

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shown using nanoindentation that elastic modulus increases with decreasing porosity.¹³ Nondestructive testing of concrete is increasingly acknowledged for evaluating strength, durability, and structural properties.¹² In addition, finite element analysis reveals stress accumulation and restoration patterns,¹⁴ whereas molecular dynamics simulations prove effective in predicting the elastic properties of hydrated cement paste constituents.¹⁵ The composition, water-to-cement ratio, curing conditions, and aggregate qualities affect the elastic modulus of concrete.^{16,17} For instance, it has been shown that both extensive curing and the addition of supplementary cementitious materials to OPC concrete marginally decrease workability but significantly improve mechanical properties.^{16,18} However, to the best of our knowledge, the mechanical properties, including the elastic modulus, of M–S–H remain largely unexplored.

In this study, we characterize the elastic modulus of M–S–H grown on single-crystal mica using amplitude modulation–frequency modulation atomic force microscopy (AM–FM AFM). To ascertain the nature and distribution of associated phases (if any), M–S–H was also synthesized as a bulk powder and characterized using a suite of tools including X-ray diffraction (XRD), thermogravimetric analysis (TGA), Fourier-transform infrared spectroscopy (FTIR), and scanning electron microscopy with energy-dispersive spectroscopy (SEM-EDS). This study demonstrates the unique capability of AFM in spatially resolving the mechanical properties of diverse materials, such as cement, at the nanoscale. The elastic modulus of M–S–H, measured for the first time in this study, is compared with that of C–S–H. Investigating M–S–H's mechanical properties at high spatial and temporal resolutions provides insights into the fundamental origins underlying its cementitious characteristics. Moreover, insights gained from this study can be extrapolated to lesser-explored magnesium-based cement. Our results support the viability of M–S–H as a binding material in construction and a potential alternative to conventional cement.

2. MATERIALS AND METHODS

2.1. Sample Preparation

The samples were prepared by mixing stock solutions of magnesium nitrate hexahydrate ($\text{Mg}(\text{NO}_3)_2 \cdot 6\text{H}_2\text{O}$, 95% purity), sodium metasilicate pentahydrate ($\text{Na}_2\text{SiO}_3 \cdot 5\text{H}_2\text{O}$, 99% purity), and ultrapure water ($>18.2 \text{ M}\Omega\text{-cm}$) in a polypropylene centrifuge tube, following a previously published procedure.⁴ Bulk (powder) samples were prepared by adding 6 mL of each stock solution in a polypropylene centrifuge tube for a target final concentration of $[\text{Mg}] = [\text{Si}] = 100 \text{ mM}$ and a final volume of 12 mL. A multiparameter benchtop meter (ThermoFisher Scientific Orion VersaStar Pro) calibrated over $4 < \text{pH} < 12$ was used to measure the pH values of the growth solution immediately following mixing of the stock solutions (Table 1). Then, the growth solution was set aside at ambient temperature ($25 \pm 2^\circ\text{C}$) for 48 h to allow precipitates to form and settle at the bottom of the tube. At the end of the reaction period, the residual growth solution was decanted and 1 mL of ethanol was introduced into the tube to minimize further reaction. Subsequently, the tube was centrifuged twice at 2500 rpm for 20 min to separate the precipitates from the excess solution. The supernatant was decanted, and the wet precipitates were dried under vacuum conditions at 25°C . Once dried, the precipitates (wrapped in paper) were finely ground into a powder using a heavy metal chisel. A similar procedure was followed to prepare brucite ($\text{Mg}(\text{OH})_2$) powder using stock solutions of NaOH and $\text{Mg}(\text{NO}_3)_2 \cdot 6\text{H}_2\text{O}$.

To prepare M–S–H samples for elastic modulus measurements, 1 mL of each stock solution was mixed in a centrifuge tube to achieve

Table 1. Experimental Conditions for Bulk/Powder M–S–H Samples, Showing Mg/Si Ratios, Mg and Si Concentrations, Calculated pH, pH Measured Immediately after Mixing of the Stock Solutions (“Initial”), pH Measured after Reaction (“Final”), Reaction Time, and Reaction Temperature^a

[Mg]/[Si]	[Mg] (mM)	[Si] (mM)	pH (calc.)	pH (initial)	pH (final)	time (h)	temp. ($^\circ\text{C}$)
1	100	100	11.61	10.59	9.62	48.0	24.1
1	500	500	11.91	10.19	9.92	48.0	25.3
1	100	100	11.61	10.26	9.74	48.0	24.2

^apH is calculated using PHREEQC³⁹ with the Cemdata18⁴⁰ database.

Mg-to-Si molar ratios of 0.5, 1, 1.5, and 2 (Table 2). The reference material and substrate is a single-crystal disc of muscovite mica (001),

Table 2. Experimental Conditions for M–S–H Grown on Mica, Showing Mg/Si Ratios, Mg and Si Concentrations, Calculated pH, pH Measured Immediately after Mixing of the Stock Solutions (“Initial”), pH Measured after Reaction (“Final”), Reaction Time, and Reaction Temperature^a

[Mg]/[Si]	[Mg] (mM)	[Si] (mM)	pH (calc.)	pH (initial)	pH (final)	time (h)	temp. ($^\circ\text{C}$)
0.5	50	100	12.58	10.19	9.85	48.3	23.1
1.0	100	100	11.70	10.29	9.88	0.7	22.3
1.0	100	100	11.67	10.16	9.77	0.7	23.1
1.0	100	100	11.68	10.40	9.62	24.6	22.9
1.0	100	100	11.68	10.40	9.58	24.6	22.9
1.0	100	100	11.70	10.29	9.46	94.1	22.3
1.0	100	100	11.67	10.41	9.46	112.5	23.1
1.0	100	100	11.64	10.59	9.62	48.1	24.1
1.5	150	100	10.98	9.56	9.10	0.7	23.1
1.5	150	100	10.97	9.72	9.46	48.1	23.3
2.0	200	100	10.74	9.49	9.04	0.7	24.0
2.0	200	100	10.76	9.73	8.75	48.1	23.3

^apH is calculated using PHREEQC³⁹ with the Cemdata18⁴⁰ database. For a growth solution with a $[\text{Mg}]/[\text{Si}]$ ratio of 1 and a pH of 11.70, the saturation indices with respect to $\text{M}_{0.5}\text{SH}$ and $\text{M}_{1.5}\text{SH}$ are 9.61 and 8.57, respectively.

grade V1 (Ted Pella), 10 mm in diameter and 0.2 mm in thickness. The mica disc was cleaved to ensure a clean surface, and any debris was removed by brief exposure to ultrahigh purity (UHP) nitrogen gas. The cleaved mica disc is then placed immediately in a centrifuge tube containing 2 mL of growth solution. To allow M–S–H to precipitate on the mica surface, the tube was sealed and then stored at ambient temperature ($25 \pm 2^\circ\text{C}$) for up to 112 h (Table 2). The pH values of the growth solutions were measured upon mixing the stock solutions and at the end of the reaction. Following the pH measurements, the mica disc was carefully removed from the centrifuge tube. Excess liquid was removed from the edge of the mica disc by using a Kimwipe, and the disc was then briefly exposed to UHP nitrogen to ensure the removal of any remaining liquid. The samples were then stored in a lidded polyethylene container under ambient conditions until they were ready to be characterized by AFM. The elastic modulus of mica was measured using a Hysitron Ubi1 Nano-Indenter with a diamond tip.

2.2. Characterization of M–S–H Powder

Geochemical modeling of the growth solution shows supersaturation with respect to M–S–H and brucite ($\text{Mg}(\text{OH})_2$) and undersaturation with respect to amorphous silica ($\text{SiO}_{2(\text{am})}$).⁴ To confirm the nature and abundance of phases that coprecipitate with M–S–H (if any), the powder samples were characterized using XRD, FTIR, TGA, and SEM-EDS. To characterize the crystallographic structure of

the powder precipitates, a D8 Advance X-ray diffractometer (Bruker) with a Cu $K\alpha$ source ($\lambda = 1.5418 \text{ \AA}$) was used. The data were collected at 2θ ranging from 5 to 80° with an increment of 0.01° and $0.5^\circ/\text{min}$ scan rate at a voltage and current of 40 kV and 40 mA , respectively. A Nicolet iSSO FTIR spectrometer (Thermo Fisher Scientific) was used to perform infrared absorbance measurements. An attenuated total reflectance (ATR) attachment and a deuterated triglycine sulfate (DTGS) KBr detector were fitted to the apparatus. The infrared spectra were obtained with the atmospheric suppression option enabled over a wavenumber range of 525 to 4000 cm^{-1} , and 32 scans were collected. TGA was performed using an SDT-Q600 simultaneous TGA/differential scanning calorimeter (DSC) (TA Instruments). Data were collected over a temperature range of 20 to 1050°C with an increment of $15^\circ\text{C}/\text{min}$ and a nitrogen gas flow rate of $20 \text{ mL}/\text{min}$. SEM-EDS was performed using a S-3000N variable pressure SEM (Hitachi) with a tungsten electron source. Before imaging, the samples were coated with lead in the presence of argon gas by using a CrC-100 sputtering system. An accelerating voltage of 20 kV , magnification of $200\times$, beam current of 62 nA , and working distance of 15 mm were used.

2.3. Measurement of Elastic Modulus at the Nanoscale

Amplitude modulation–frequency modulation (AM–FM) AFM was used to quantify the elastic modulus of the M–S–H overgrowths. In AM–FM, amplitude modulation (AM) enables the characterization of surface topography, while frequency modulation (FM) provides nanomechanical characteristics of the sample surface, including the elastic modulus.^{19,20} A high-amplitude drive is applied to the primary resonance frequency of the cantilever, while a low-amplitude drive is applied to its second resonance frequency. Shifts in the second resonance frequency are related to the sample's elastic modulus.²¹ The analysis presented herein uses a reference material with a known elastic modulus, E (herein, single-crystal mica with $E = 64 \pm 2 \text{ GPa}$), to determine the unknown elastic modulus of the sample (herein, M–S–H). The known elastic modulus of the reference material is used to obtain the best-fit parameters that describe tip–sample interactions, including tip–sample Hertzian contact mode (e.g., Hertz Punch, Hertz Cone, Hertz Sphere) and tip geometry (e.g., tip radius and half-cone angle), enabling quantitative determination of the unknown elastic modulus. Therefore, a priori knowledge of specific tip–sample interactions is not required. AFM data were collected using a silicon probe having a rectangular cantilever with a 70 nm -thick gold coating on the detector side, force constant of $3 \text{ N}/\text{m}$, nominal resonance frequency of 75 kHz , length of $225 \text{ }\mu\text{m}$, width of $28 \text{ }\mu\text{m}$, and thickness of $3 \text{ }\mu\text{m}$.

3. RESULTS AND DISCUSSION

3.1. Analysis of Associated Phases

The conditions for the bulk (powder) synthesis performed in this study are shown in Table 1. The chemical composition of M–S–H is influenced by several variables, including the reaction duration, chemical composition of precursors, temperature, Mg/Si ratio, and water activity.^{22–24} Furthermore, previous studies characterizing synthetic M–S–H reveal the presence of associated phases such as brucite and amorphous silica.^{25,26} XRD, TGA, and FTIR show that the precipitates that formed under our experimental conditions are primarily M–S–H. Specifically, XRD data for M–S–H reveal characteristic broad humps at $2\theta = 20.1, 26.7, 35.0,$ and 59.9° (Figure 1).^{25–30} FTIR analysis also shows that brucite is not present in significant quantities (estimated to be $<3\%$ by mass of the precipitates, based on the typical instrumental resolution of techniques employed herein) (Figure 2).^{26,30,31} Silicate tetrahedral polymerization in M–S–H (~ 600 – 1400 cm^{-1}) can be characterized from the relative amounts of Q^1 (one oxygen shared between two silicon atoms, forming chains), Q^2 (two shared oxygens, forming double chains), and Q^3 (three

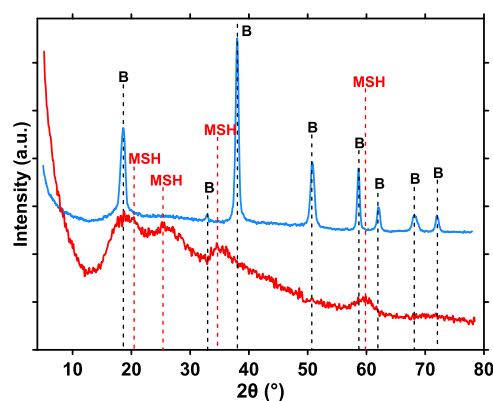


Figure 1. X-ray diffraction pattern of the precipitates grown from a solution containing either (red) $[\text{Mg}] = [\text{Si}] = 100 \text{ mM}$ or (blue) $[\text{Mg}] = [\text{OH}] = 100 \text{ mM}$ reacted for 48 h at $25 \pm 2^\circ\text{C}$, matching that of either M–S–H³⁷ or brucite.^{41,42}

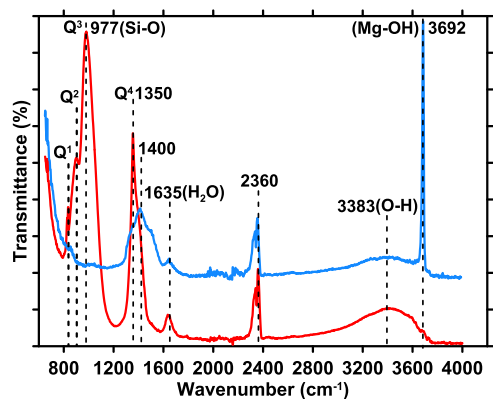


Figure 2. Infrared absorbance spectra of the precipitates grown from a solution containing either (red) $[\text{Mg}] = [\text{Si}] = 100 \text{ mM}$ or (blue) $[\text{Mg}] = [\text{OH}] = 100 \text{ mM}$ reacted for 48 h at $25 \pm 2^\circ\text{C}$. The peak at 2360 cm^{-1} is associated with atmospheric CO_2 .

shared oxygens, forming sheets) species,^{32–34} with Q^3 representing M–S–H sheets and Q^2 representing defects within these sheets.⁴ Distinct peaks between ~ 600 and 1400 cm^{-1} , specifically, the peaks at 835 cm^{-1} (Q^1), 902 cm^{-1} (Q^2), 977 cm^{-1} (Q^3), and 1350 cm^{-1} (Si-O^-) are associated with asymmetric and symmetric Si–O stretching vibrations in M–S–H (Figure 2).^{25,26,30,35,36} On the other hand, the peaks at 1635 and 3383 cm^{-1} are associated with H–O–H bending vibrations of molecularly bound H_2O and O–H stretching (Figure 2),^{4,30,37,38} whereas the sharp peak at 3692 cm^{-1} is associated with brucite.³⁷

Thermogravimetric analysis reveals weight losses in three distinct regions (Figure 3). The first region (20 – 250°C) is associated with poorly bound water, including monolayer, multilayer, and interlayer water within M–S–H.^{30,43,44} This region may also include bulk water at a high relative humidity. The second weight loss in M–S–H with a Mg/Si ratio of 1 is observed between 320 and 450°C and is primarily related to the dehydroxylation of Si–OH groups (Figure 3a).^{25,43–46} The third weight loss, occurring at 580°C , is associated with the dehydroxylation of Mg–OH groups.²⁵ Dehydroxylation of serpentine minerals has been observed between 670 and 900°C , whereas the inner OH groups linked to the magnesium in the talc structure dehydroxylates between 750 and 1000°C , in agreement with the weight losses observed for M–S–H from

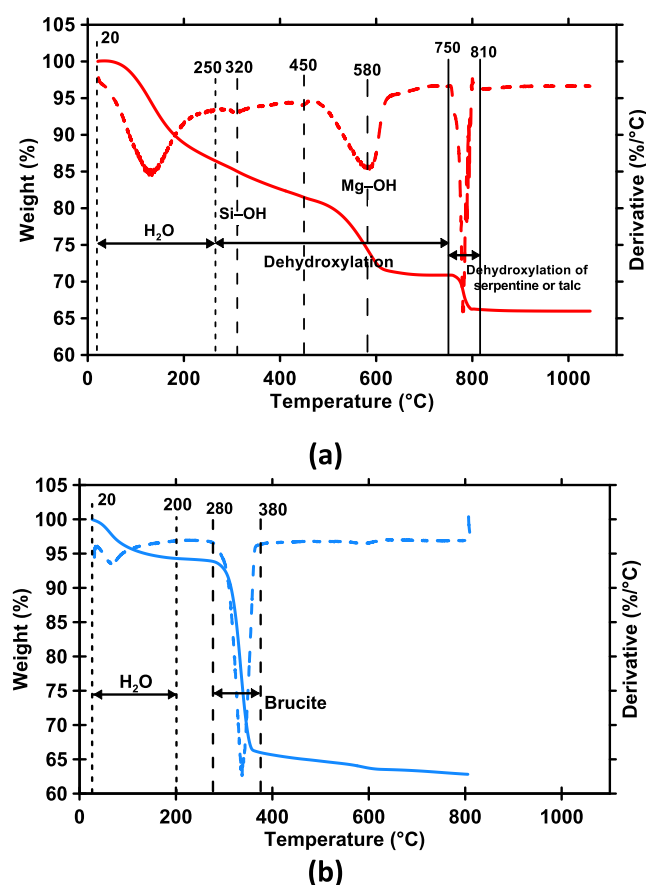


Figure 3. Decomposition curves of the precipitates grown from a solution containing either (a) $[Mg] = [Si] = 100$ mM or (b) $[Mg] = [OH] = 100$ mM for 48 h at 25 ± 2 °C. The derivative of the mass loss is shown by the dashed curves.

450 to 1000 °C.⁴⁷ In contrast, the brucite sample shows a single distinct sharp peak in this region at around 320 °C, which is primarily attributed to the release of hydroxylated water (Figure 3b).⁴⁶

SEM-EDS analysis (Figure 4) shows that the Mg-to-Si molar ratios of the precipitates range from 0.64 to 1.02 with an average of 0.85 (Table 3), indicating the predominance of M–S–H, since previous studies have shown that the Mg/Si ratio of M–S–H precipitates closely matches that of the growth solution.^{29,36} Taken together, the combination of analytical techniques employed herein provides evidence for the

Table 3. Results of EDS Analysis of Precipitates Grown from a Solution Containing $[Mg] = [Si] = 100$ mM, Showing the Molar % for Selected Points and Areas ($\sim 510 \mu\text{m} \times 680 \mu\text{m}$) in SEM Images

	N	O	Na	Mg	Si	Mg/Si
point 1	1.63	67.26	5.19	13.10	12.82	1.02
point 2	1.35	67.29	3.42	13.82	14.12	0.98
point 3	0.00	54.06	2.93	16.68	26.32	0.64
point 4	0.00	63.32	4.44	14.59	17.65	0.83
point 5	3.23	59.76	7.13	13.17	16.71	0.79
area 1	0.00	63.97	7.23	13.40	15.39	0.87
area 2	0.00	63.00	8.00	12.00	14.00	0.86

persistence of primarily M–S–H under our experimental conditions.

3.2. Influence of Morphology and Reaction Time on the Elastic Modulus of M–S–H

The conditions for the synthesis of M–S–H overgrowths on mica are shown in Table 2. The discrepancy between the calculated and initial pH is likely caused by instantaneous (within seconds) precipitation of M–S–H.⁴

A representative topography (amplitude modulation) image and its corresponding frequency distribution show M–S–H overgrowths that are up to $\sim 2 \mu\text{m}$ thick on the mica (001) surface (Figure 5a). A uniform film is not obtained; instead, two types of morphologies are observed: (1) large and faceted and (2) small and irregular. The corresponding elastic modulus map and its frequency distribution are shown in Figure 5b. For the sake of clarity, frequency distributions for subset areas of the elastic modulus map are also shown (right, Figure 5b). The results reveal significant variations in elastic modulus depending on morphology; i.e., large, faceted grains feature an average elastic modulus of 54.0 ± 14.9 GPa, whereas small, irregular grains have an average elastic modulus of 6.0 ± 7.0 GPa (Figure 4b). As high as 75 GPa is obtained for faceted grains, indicating increased stiffness and resistance to deformation. These observations highlight the potential role of morphological control in fine-tuning the elastic modulus of M–S–H cements. For comparison, the elastic modulus of brucite has been previously reported as 11 GPa based on neutron diffraction data⁴⁸ and 15 GPa based on Brillouin scattering.⁴⁹ The elastic modulus of amorphous alkali–silica reaction gels has been measured to be between 24.9 and 34.0 GPa.⁵⁰ Using nanoscale three-point bending tests with an atomic force microscope, amorphous SiO_2 nanowires have been found to

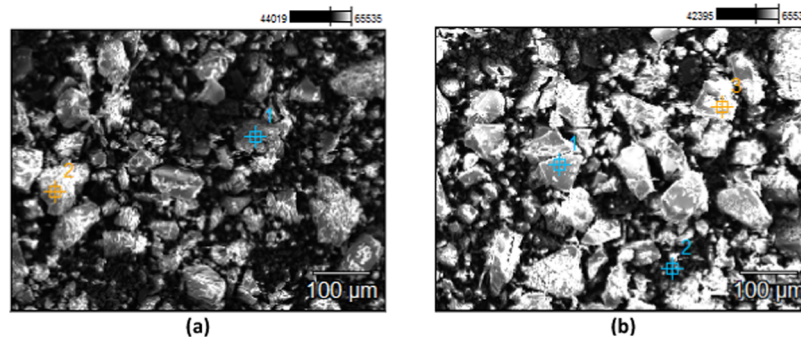


Figure 4. Scanning electron micrographs of the M–S–H precipitates grown from a solution containing $[Mg] = [Si] = 100$ mM for 48 h at 25 ± 2 °C, showing corresponding micrographs where the chemical compositions for (a) Point 1, Point 2, and Area 1 and (b) Point 3, Point 4, Point 5, and Area 2 in Table 3 are collected.

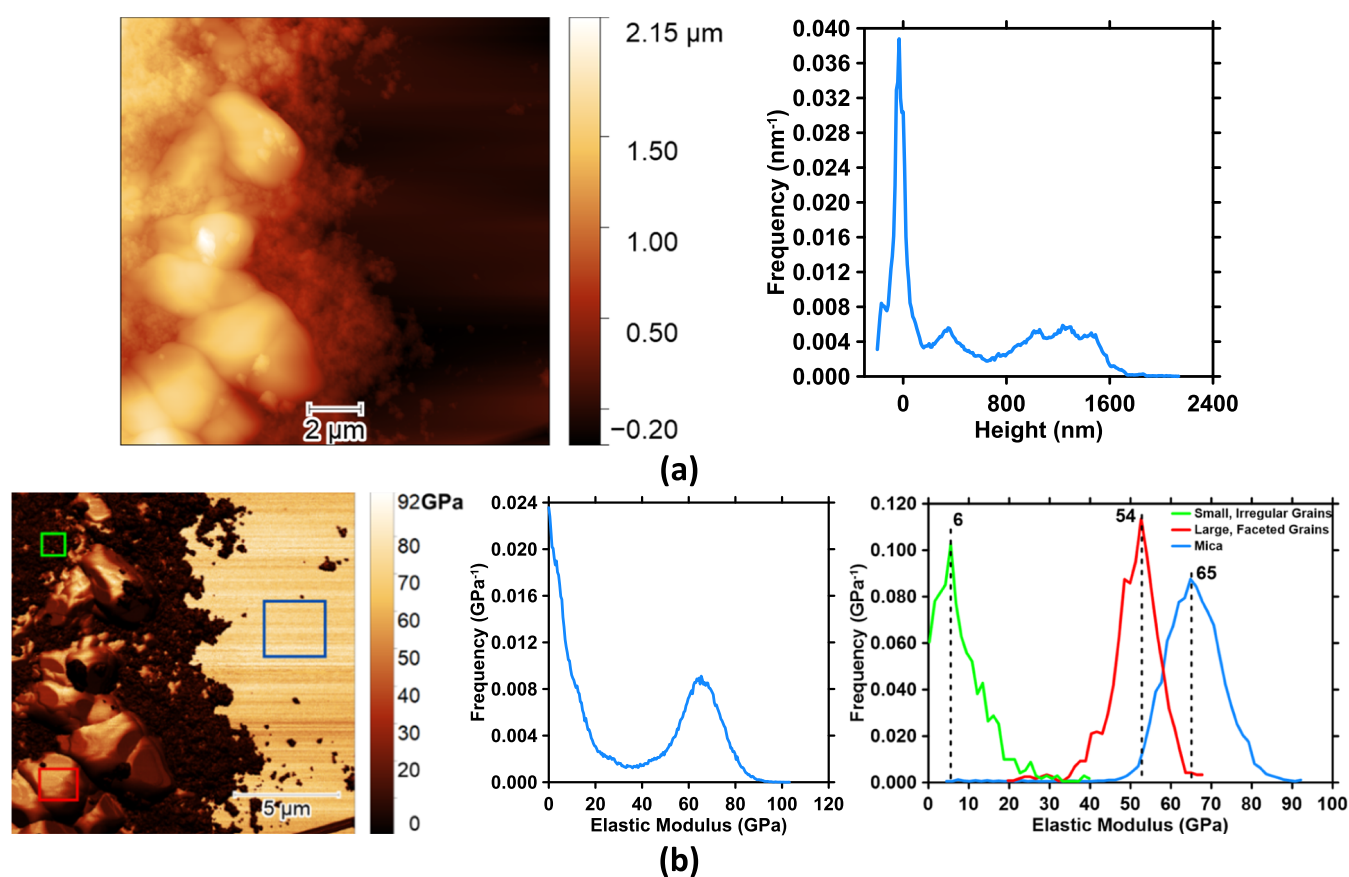


Figure 5. Representative atomic force microscopy images of (a, left) surface topography and (a, right) corresponding frequency distribution of height with an RMS (roughness) of 545.3 nm, and (b) elastic modulus and (b, middle and right) corresponding frequency distribution of elastic modulus of mica (001) reacted with a supersaturated solution of M–S–H with a final concentration of $[Mg] = [Si] = 100$ mM for 48 h at 24.1 °C. In panel (b, middle), the frequency distribution for the entire area of panel (b, left) is shown. In panel (b, right), the frequency distributions for subset areas in panel (b, left) are shown.

have an elastic modulus of 76.6 ± 7.2 GPa.⁵¹ Additionally, the elastic modulus of silica glass, as determined using molecular dynamics simulation, is 72.56 GPa.⁵²

Representative topography and elastic modulus data for relatively short (40 min) and long (48 h) experimental durations are shown in Figure 6. Higher spatial resolution images are listed in Figure 7. Specifically, a shorter reaction time resulted in the formation of monodisperse grains that are small and irregular, whereas a longer reaction time resulted in the occurrence of both morphology types, as shown in Figure 5. This suggests that the extended reaction time allowed for the development of precipitates with greater structural diversity. In agreement with the findings in Figure 5, early age precipitates exhibited relatively low elastic modulus values of 10.0 ± 7.0 GPa (Figure 6a). Late-age precipitates featured a significantly higher elastic modulus value of approximately 57.2 ± 14.9 GPa (Figure 6b,c). Elastic modulus measurements on 24 h samples (see Table 2) show a similar bimodal distribution of elastic modulus that depends on the precipitate morphology.

The tip–sample contact area, which may change as a result of surface roughness, may affect elastic modulus measurements.⁵³ Nonetheless, it is evident in Figure 6a that precipitate roughness does not result in significant differences in the elastic modulus (see Figure 7 for a detailed characterization of the morphology). On the other hand, changes in the contact area can partly or wholly explain the varying elastic modulus of faceted grains depending on the facet orientation (Figure 5b).

Despite these uncertainties, the observation of a bimodal distribution (and the respective average values) of the elastic modulus is unchanged. Furthermore, the discrepancy between local tip–sample contact areas increases with increasing indentation depth; this source of error is minimized through the use of small indentation depths in AM–FM.⁵⁴

3.3. Effect of Silicate Polymerization on Elastic Modulus: Comparison with C–S–H and across Mg-to-Si Ratios

The elastic modulus of M–S–H is also evaluated as a function of solution $[Mg]/[Si]$ molar ratios: 0.5, 1.0, 1.5, and 2.0. Notably, it has been shown previously that the precipitate Mg/Si ratio reasonably matches the $[Mg]/[Si]$ of the initial growth solution.⁴⁴ The elastic modulus of faceted grains is invariant with the Mg/Si ratio and ranges from ~ 40 to 64 GPa across different samples and areas within samples (Table 2, Figure 8). On the other hand, small, irregular grains whose elastic modulus ranged from ~ 5 to 15 GPa showed a slight decrease with the Mg/Si ratio (Figure 8). The lower elastic modulus of small, irregular grains can be attributed to their lower degree of crystallinity⁵⁵ compared to the faceted grains.⁵⁶

C–S–H has an elastic modulus that ranges between 20 and 40 GPa, as measured using nanoindentation and viscoelastic modulus mapping.⁵⁷ This range lies between the average values for faceted and irregular M–S–H precipitates (Figure 8). This suggests that the higher elastic modulus in faceted M–S–H grains arises from their sheet-like structure (i.e., dominance of Q^3 sites) as in phyllosilicates^{58–61} and that the lower elastic

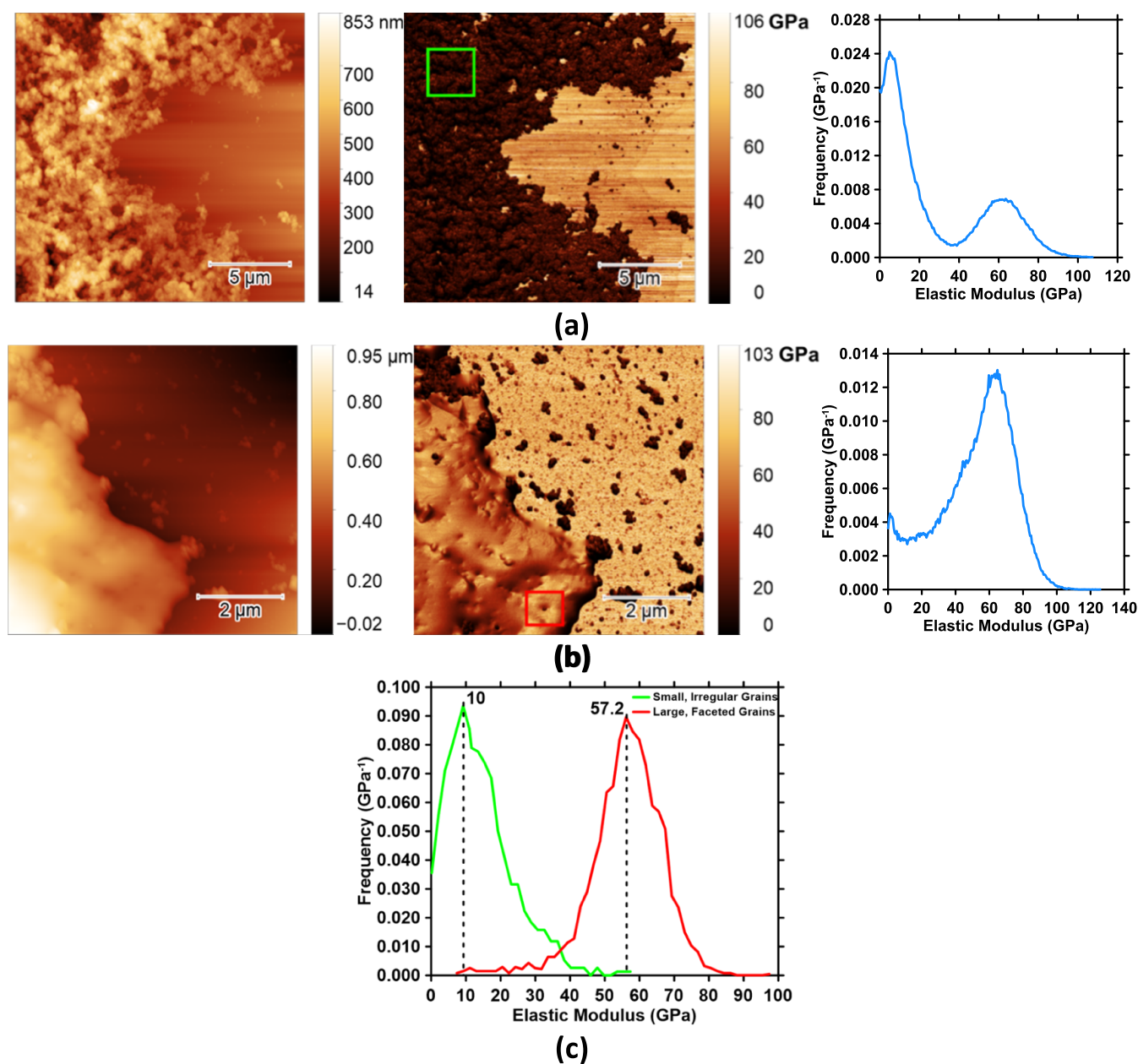


Figure 6. Representative atomic force microscopy maps of (left) topography and (middle) corresponding elastic modulus, and (right) frequency distribution of elastic modulus of mica (001) reacted with a supersaturated solution of M–S–H with a final concentration of $[Mg] = [Si] = 100$ mM for (a) 40 min at 23.1 °C with an RMS (roughness) value of 100.8 nm or (b) 48 h at 24.1 °C with an RMS (roughness) value of 197.4 nm. In panel (c), the elastic modulus distributions for small, irregular grains and large, faceted grains marked by boxes in panels (a) and (b) are shown.

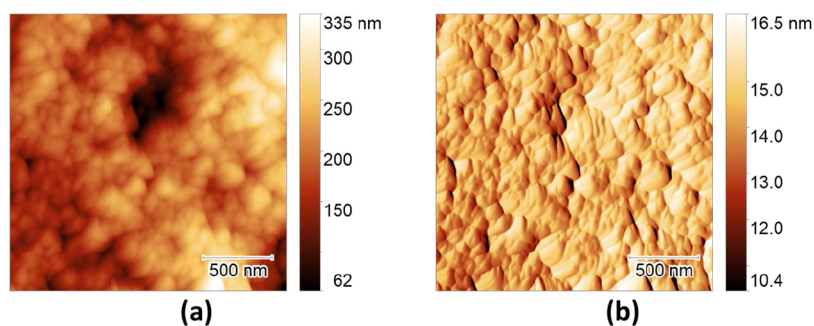


Figure 7. Representative atomic force microscopy (a) height and (b) amplitude error images of mica (001) reacted with a supersaturated solution of M–S–H with a final concentration of $[Mg] = 50$ mM and $[Si] = 100$ mM for 48.3 h, showing detail of M–S–H precipitates similar to those shown in Figure 6a.

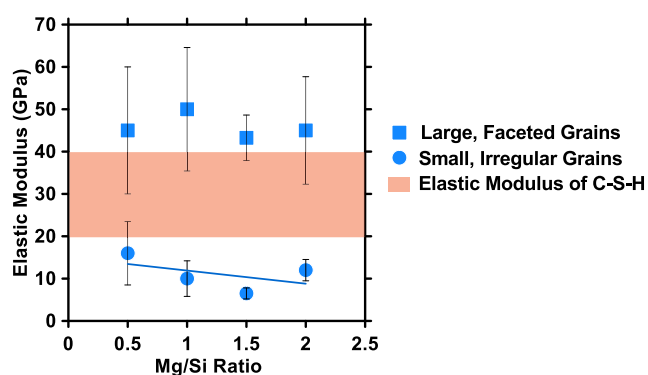


Figure 8. Elastic modulus of M–S–H as a function of Mg/Si ratio, given for large, faceted grains and small, irregular grains. Error bars represent the standard deviation for two repeat measurements on different areas within the same sample (i.e., 48-h samples in Table 2). The elastic modulus range for C–S–H is also shown for comparison. The RMS (roughness) values are 183.2 and 104.3 nm for Mg/Si = 0.5, 545.3 and 197.4 nm for Mg/Si = 1, 421.4 and 218.8 nm for Mg/Si = 1.5, and 89.59 and 131.7 nm for Mg/Si = 2. Therefore, the average RMS is 236.5 nm, and the standard deviation of the RMS is 162.2 nm.

modulus of C–S–H compared to M–S–H can be explained by C–S–H's less polymerized chain-like silicate structure.⁶²

In C–S–H, it has been shown that as the Ca/Si ratio increases, the length of the silicate chains decreases.^{63,64} It has also been shown that Ca/Si ratios influence the relative distributions of Q^2 and Q^1 species; specifically, an increase in the Ca/Si ratio results in the relative reduction of Q^2 species,⁵⁹ suggesting a decreased degree of polymerization of silica tetrahedra.^{58,65} We have previously shown that the degree of polymerization of M–S–H decreases with increasing Mg/Si ratio.⁴ Therefore, we propose that the small, irregular M–S–H precipitates (e.g., Figure 6a)⁵⁵ may have varying degrees of polymerization that are controlled by the Mg/Si ratio. This implies that the observed reduction in elastic modulus in poorly crystalline M–S–H is analogous to that observed in C–S–H.⁵⁹ Significantly, in addition to the Ca/Si ratio, particle packing, porosity, and interplanar distance influence the mechanical characteristics of C–S–H.⁵⁹ We postulate that similar operative mechanisms can be applied to M–S–H.

4. CONCLUSIONS

This study presents the first experimental measurements of the elastic modulus of the cementitious material magnesium silicate hydrate. Detailed characterization of precipitates shows that M–S–H is the primary phase that forms under the experimental conditions. Topographical imaging using AFM revealed two distinct morphological features: small, irregular grains and large, faceted grains. Simultaneous elastic modulus mapping indicates that these features have corresponding elastic modulus values of 5 to 15 GPa and 40 to 64 GPa. The elastic modulus of poorly crystalline M–S–H decreased with increasing Mg/Si ratio possibly because of a corresponding decreased degree of polymerization analogous to the behavior observed in C–S–H. This could also explain the higher elastic modulus of large, faceted M–S–H compared to that of C–S–H. The results of this study have important implications in the synthesis of cementitious materials for construction applications. Taken together, this study demonstrates that the deliberate control of the morphology and

chemical composition of M–S–H may be a viable pathway for tailoring cement properties to meet specific performance requirements.

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Notes

The authors declare no competing financial interest.

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