

Electrocatalytic Properties of Quasi-2D Oxides $\text{LaSrMn}_{0.5}\text{M}_{0.5}\text{O}_4$ (M = Co, Ni, Cu, and Zn) for Hydrogen and Oxygen Evolution Reactions

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Abstract: Designing cost-effective and highly efficient electrocatalysts for water splitting is a significant challenge. We have systematically investigated a series of quasi-2D oxides, $\text{LaSrMn}_{0.5}\text{M}_{0.5}\text{O}_4$ (M = Co, Ni, Cu, Zn), to enhance the electrocatalytic properties for the two half-reactions of water-splitting, namely oxygen and hydrogen evolution reactions (OER and HER). The four materials are isostructural, as confirmed by Rietveld refinements with X-ray diffraction. The oxygen contents and metal valence states were determined by iodometric titrations and X-ray photoelectron spectroscopy. Electrical conductivity measurements in a wide range of temperatures revealed semiconducting behavior for all four materials. Electrocatalytic properties were studied for both half-reactions of water-splitting, oxygen-evolution and hydrogen-evolution reactions (OER and HER). For the four materials, the trends in both OER and HER were the same, which also matched the trend in electrical conductivities. Among them, $\text{LaSrMn}_{0.5}\text{Co}_{0.5}\text{O}_4$ showed the best bifunctional electrocatalytic activity for both OER and HER, which may be attributed to its higher electrical conductivity and favorable electron configuration.

Keywords: Solid state oxide; Quasi-2D; Hydrogen evolution reaction; Oxygen evolution reaction

1. Introduction

The development of green and sustainable energy solutions is crucial to addressing the energy challenges of the future. The persistent use of carbon-rich fossil fuels leads to carbon dioxide emissions, worsening the greenhouse effect. Additionally, the harmful gases produced during the burning of those fuels cause environmental problems. These challenges have inspired the scientific community to explore renewable and sustainable energy sources. Hydrogen is widely considered an attractive alternative to conventional fuels due to its high energy density and absence of carbon emission when used as a fuel. Among different methods of hydrogen production, the electrochemical water splitting is highly promising. It is an excellent approach for converting and storing renewable energy if cost-effective and efficient catalysts are used to facilitate large-scale hydrogen production. Electrochemical water splitting involves two half-cell reactions, the hydrogen-evolution reaction (HER) and the oxygen-evolution reaction (OER). The OER also occurs in other energy storage and conversion devices, including rechargeable metal-air batteries ($\text{M}_x\text{O}_2 \rightarrow \text{M}_x + \text{O}_2$).¹ The OER in water electrolysis involves a four-electron-transfer mechanism ($\text{H}_2\text{O} \rightarrow \text{O}_2 + 4\text{H}^+ + 4\text{e}^-$) with a more complex reaction pathway and slower kinetics compared to the HER, which follows a two-electron-transfer mechanism ($2\text{H}^+ + 2\text{e}^- \rightarrow \text{H}_2$). There are efficiency challenges associated with water electrolysis due to the sluggish kinetics of the two half-reactions, particularly the OER. This leads to substantial overpotentials beyond the theoretical thermodynamic potential of 1.23 V versus the reversible hydrogen electrode (RHE).² Although catalysts based on noble metals, such as Pt, Ru, and

Ir, often exhibit very high performance for water-splitting, their high cost and scarcity limit their wide-scale application, making it necessary to explore cost-effective and earth-abundant alternatives.^{3, 4}

Perovskite oxides have received growing interest as non-precious metal-based catalysts due to their versatility in composition, tunable electronic structures, affordability, and strong inherent catalytic performance.^{5–8} Partial replacement or doping with different metal cations leads to changes in their performance for different applications driven by the resulting modifications to their physical, chemical, and electronic properties.^{9, 10} Several perovskite oxides, such as the benchmark catalyst $\text{Ba}_{0.5}\text{Sr}_{0.5}\text{Co}_{0.8}\text{Fe}_{0.2}\text{O}_{3-\delta}$ (BSCF), have demonstrated efficient OER catalysis, gaining prominence for their OER performance in alkaline environments, on a par with IrO_2 .⁵ Studies have also explored the HER characteristics of perovskite oxides such as $\text{Pr}_{0.5}(\text{Ba}_{0.5}\text{Sr}_{0.5})_{0.5}\text{Co}_{0.8}\text{Fe}_{0.2}\text{O}_{3-\delta}$ ¹¹ and $\text{SrNb}_{0.1}\text{Co}_{0.7}\text{Fe}_{0.2}\text{O}_{3-\delta}$.¹² A variety of other structural arrangements are derived from perovskites, such as double perovskite, brownmillerite, Ruddlesden-popper, and the so-called K_2NiF_4 -type¹³ structure. The latter is the topic of the present work and has a quasi-2D structure (Figure 1) with the general formula A_2BO_4 , where A is usually a rare-earth or alkaline-earth metal, and B is frequently a transition metal. Some materials with this structure-type have been reported to show electrocatalytic activity for OER or HER, such as SrLaFeO_4 , $\text{SrLaCo}_{0.5}\text{Fe}_{0.5}\text{O}_4$, and $\text{SrLaCoO}_{4-\delta}$.¹³ Nevertheless, oxide catalysts that demonstrate efficient catalytic activity for both OER and HER are not commonly found.

In the present work, we investigated a series of quasi-2D oxides $\text{LaSrMn}_{0.5}\text{M}_{0.5}\text{O}_4$ ($\text{M} = \text{Co}, \text{Ni}, \text{Cu}, \text{Zn}$), with the aim of improving their electrocatalytic performance for HER and OER through the incorporation of different transition metals. Previous reports on these materials have been limited to the synthesis and some structural characterizations.^{14–18} The electrocatalytic properties have not been previously studied. Our comprehensive studies involved the investigation of their electrical and electrocatalytic properties, as well as detailed X-ray diffraction, X-ray photoelectron spectroscopy (XPS), and scanning electron microscopy (SEM) studies. Our findings showed semiconducting behavior in all four compounds, with electrical conductivity increasing as a function of temperature. The trend in electrical conductivity correlated with that of the electrocatalytic activity for both HER and OER. Importantly, the bifunctional electrocatalytic properties found in these materials are remarkable, particularly for $\text{LaSrMn}_{0.5}\text{Co}_{0.5}\text{O}_4$, which exhibited the highest electrical charge-transport and superior electrocatalytic performance.

2. Results and Discussion

2.1. Crystal Structure

Powder X-ray diffraction (XRD) was employed to assess the crystal structures of $\text{LaSrMn}_{0.5}\text{M}_{0.5}\text{O}_4$ ($\text{M} = \text{Co}, \text{Ni}, \text{Cu}, \text{Zn}$) using Rietveld refinements (Figure 1). Table 1 shows a representative example of refined structural parameters for $\text{LaSrMn}_{0.5}\text{Co}_{0.5}\text{O}_4$. The structural parameters for the other three materials are listed in Tables S1–S3. The data revealed that all samples consisted of a single phase, indicating the absence of impurities. The space group and structures of these materials were consistent with previous reports.^{14–18} All four materials are isostructural with a body-centered tetragonal space group $I4/mmm$. The crystal structure is shown in Figure 2. All four materials can be described by the general formula A_2BO_4 , where the A-site is occupied by a mix of La and Sr, while the B-site contains a mix of Mn with Co, Ni, Cu, or Zn. Their structures comprise BO_6 octahedra that share corners to form a 2D layer. The layers are separated from each other, and the gap between them is where the A-site cations reside. The A-site cations have 9 oxygens in their coordination environment. The morphologies of sintered pellets were probed using scanning electron microscopy (SEM), and the results are shown in Figure 3, indicating similar microstructures for the four samples. The EDS mapping showed the surface elemental compositions, as shown in Figures S1, which illustrates representative data for $\text{LaSrMn}_{0.5}\text{Co}_{0.5}\text{O}_4$, matching the expected atomic percentages.

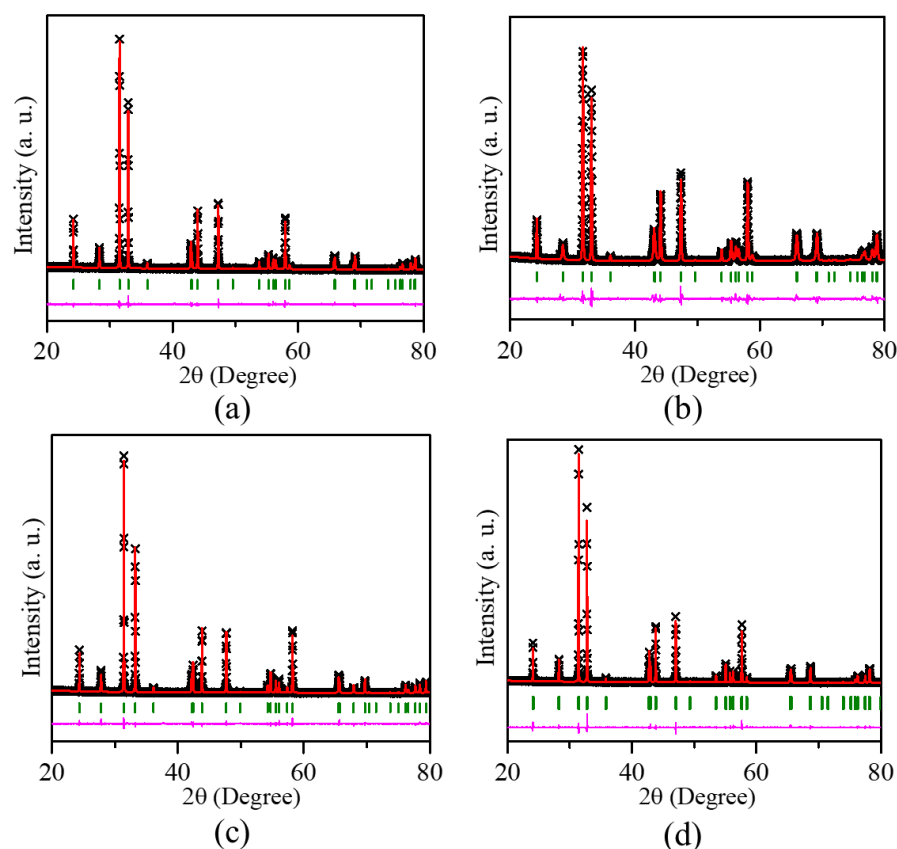


Figure 1. Rietveld refinement profiles using powder X-ray diffraction data of (a) $\text{LaSrMn}_{0.5}\text{Co}_{0.5}\text{O}_4$ (b) $\text{LaSrMn}_{0.5}\text{Ni}_{0.5}\text{O}_4$ (c) $\text{LaSrMn}_{0.5}\text{Cu}_{0.5}\text{O}_{4-\delta}$ (d) $\text{LaSrMn}_{0.5}\text{Zn}_{0.5}\text{O}_4$. The cross symbols, solid red line, olive vertical tick marks, and lower magenta line correspond to experimental data, the calculated pattern for the structural model, Bragg peak positions, and the difference plot, respectively.

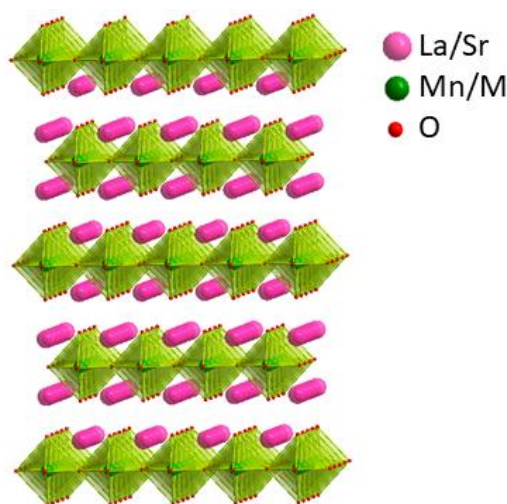


Figure 2. (a) Crystal structure of $\text{LaSrMn}_{0.5}\text{M}_{0.5}\text{O}_4$ ($\text{M} = \text{Co}, \text{Ni}, \text{Cu}, \text{Zn}$). Pink spheres denote La/Sr, small red spheres indicate oxygen, and green spheres represent Mn/M.

Table 1. Refined structural parameters for $\text{LaSrMn}_{0.5}\text{Co}_{0.5}\text{O}_4$ at room temperature using powder X-ray diffraction data. Space group: $I4/mmm$, $a = 3.8435(1) \text{ \AA}$, $c = 12.5933(2) \text{ \AA}$, $R_p = 0.0392$, $wR_p = 0.0504$.

Atom	x	y	z	Occupancy	Multiplicity	$U_{\text{iso}} (\text{\AA}^2)$
La	0	0	0.35985(8)	0.5	4	0.0192(5)
Sr	0	0	0.35985(8)	0.5	4	0.0192(5)

Mn	0	0	0	0.5	2	0.0131(9)
Co	0	0	0	0.5	2	0.0131(9)
O1	0.5	0	0	1	4	0.031(2)
O2	0	0	0.1642(6)	1	4	0.038(2)

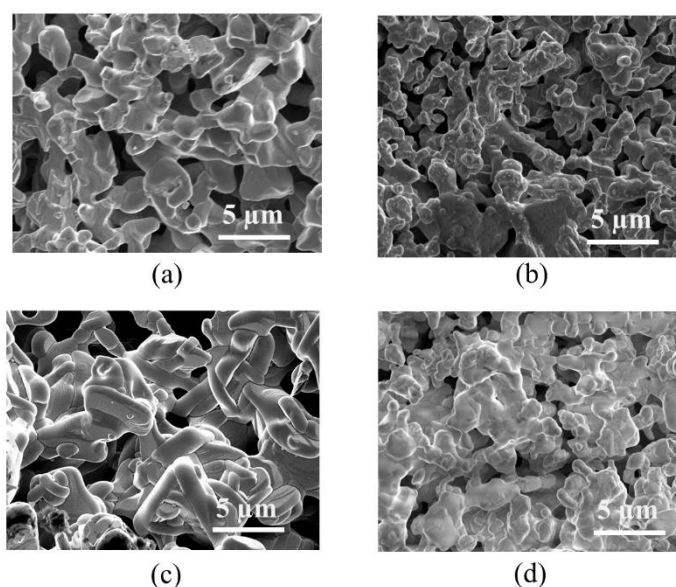


Figure 3. Scanning electron microscopy (SEM) images of (a) $\text{LaSrMn}_{0.5}\text{Co}_{0.5}\text{O}_4$, (b) $\text{LaSrMn}_{0.5}\text{Ni}_{0.5}\text{O}_4$, (c) $\text{LaSrMn}_{0.5}\text{Cu}_{0.5}\text{O}_{4-\delta}$ and (d) $\text{LaSrMn}_{0.5}\text{Zn}_{0.5}\text{O}_4$.

2.2. Oxidation State and Oxygen Content

The synthesis conditions and elevated temperatures necessary for synthesizing these compounds can result in different valence states, which can affect the oxygen contents.^{19, 20} Oxygen contents of these materials were determined by iodometric titrations, which were repeated at least three times with excellent reproducibility, giving oxygen stoichiometries of 4.0173 ± 0.0016 for $\text{LaSrMn}_{0.5}\text{Co}_{0.5}\text{O}_4$, 3.8530 ± 0.0073 for $\text{LaSrMn}_{0.5}\text{Cu}_{0.5}\text{O}_{4-\delta}$, 4.0039 ± 0.0105 for $\text{LaSrMn}_{0.5}\text{Ni}_{0.5}\text{O}_4$, and 4.0120 ± 0.0164 for $\text{LaSrMn}_{0.5}\text{Zn}_{0.5}\text{O}_4$. Therefore, iodometric titrations indicate four oxygens per formula unit for $\text{LaSrMn}_{0.5}\text{Co}_{0.5}\text{O}_4$, $\text{LaSrMn}_{0.5}\text{Ni}_{0.5}\text{O}_4$, and $\text{LaSrMn}_{0.5}\text{Zn}_{0.5}\text{O}_4$, matching the expected stoichiometries. However, $\text{LaSrMn}_{0.5}\text{Cu}_{0.5}\text{O}_{4-\delta}$ shows oxygen deficiency, with an oxygen stoichiometry of ~ 3.85 per formula unit ($\delta \approx 0.15$).

We also investigated all four materials by X-ray photoelectron spectroscopy (XPS) to determine the oxidation states of transition metals. Given the close overlap of binding energies of Mn^{3+} and Mn^{4+} peaks, we also did additional measurements of standard Mn_2O_3 and MnO_2 samples (Figure 4). The manganese spectra for three materials, $\text{LaSrMn}_{0.5}\text{Co}_{0.5}\text{O}_4$, $\text{LaSrMn}_{0.5}\text{Ni}_{0.5}\text{O}_4$, and $\text{LaSrMn}_{0.5}\text{Cu}_{0.5}\text{O}_{4-\delta}$, are very similar to each other and encompass binding energies corresponding to both trivalent and tetravalent manganese, as evident from comparison to Mn_2O_3 and MnO_2 standards in Figure 4 and previously reported manganese binding energies in other oxide materials.^{21–23} Therefore, these three materials contain manganese in a mixed $+3/+4$ state. On the other hand, the manganese spectrum of $\text{LaSrMn}_{0.5}\text{Zn}_{0.5}\text{O}_4$ shows peaks at a higher binding energy compared to the other three compounds, indicating a significantly greater tetravalent character. This observation is consistent with iodometric titration results that indicated 4 oxygens per formula unit. This oxygen content requires manganese in a tetravalent state, given the divalent state of zinc in $\text{LaSrMn}_{0.5}\text{Zn}_{0.5}\text{O}_4$.

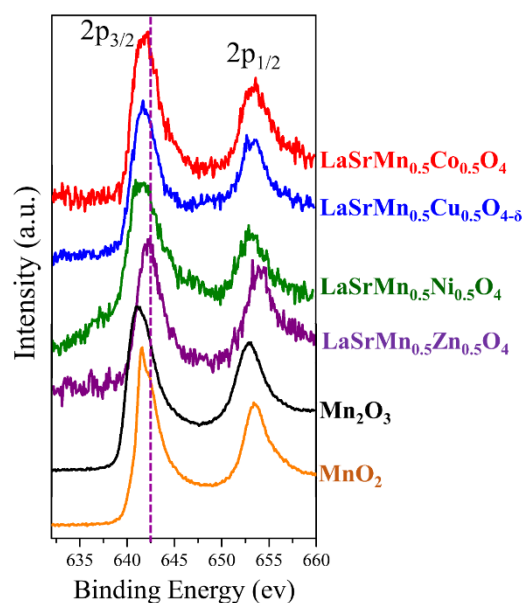


Figure 4. XPS data showing the manganese spectra for all four materials. The purple dashed line is drawn to show that the binding energy for the $2p_{3/2}$ peak of $\text{LaSrMn}_{0.5}\text{Zn}_{0.5}\text{O}_4$ is shifted to higher energy compared to those of the other three materials.

For $\text{LaSrMn}_{0.5}\text{Co}_{0.5}\text{O}_4$, the cobalt spectrum was obtained (Figure 5a), and showed binding energies consistent with a mix of +2/+3 oxidation states. Divalent cobalt is expected to give a $2p_{3/2}$ peak just below 780 eV,²⁴ while the peak for trivalent cobalt is expected to be just above 780 eV.²⁴ The Co $2p_{3/2}$ peak for our material is centered around 780 eV, with a satellite peak at about 786 eV, indicative of Co^{2+} ,²⁵ and a satellite at about 789 eV, indicative of Co^{3+} .²⁴ We ruled out the presence of tetravalent cobalt, as we did not observe any pronounced peak at ~281.5 eV.²⁴ These analyses are consistent with iodometric titration results, indicating four oxygens per formula unit for $\text{LaSrMn}_{0.5}\text{Co}_{0.5}\text{O}_4$, and the manganese XPS spectrum, indicating manganese in mixed +3/+4 state, which would require cobalt to be in mixed +2/+3 state for charge neutrality of the material.

For $\text{LaSrMn}_{0.5}\text{Ni}_{0.5}\text{O}_4$, the nickel spectrum was obtained, as shown in Figure 5b. In this spectrum, the first peak, just above 850 eV corresponds to $\text{La}^{3+} 3d_{3/2}$.²⁶ The $2p_{3/2}$ peak for Ni^{2+} is expected to appear close to 854 eV,^{26–28} but higher binding energies, such as ~855 eV, have also been reported for divalent nickel.²⁹ On the other hand, the $\text{Ni}^{3+} 2p_{3/2}$ peak is expected to be close to 856–857 eV.^{26, 27, 29} In our nickel spectra, the $2p_{3/2}$ peak is centered around ~254.6 eV, consistent with the predominant presence of Ni^{2+} , and a small high-energy shoulder at ~856.5 eV, which may indicate the presence of some Ni^{3+} .

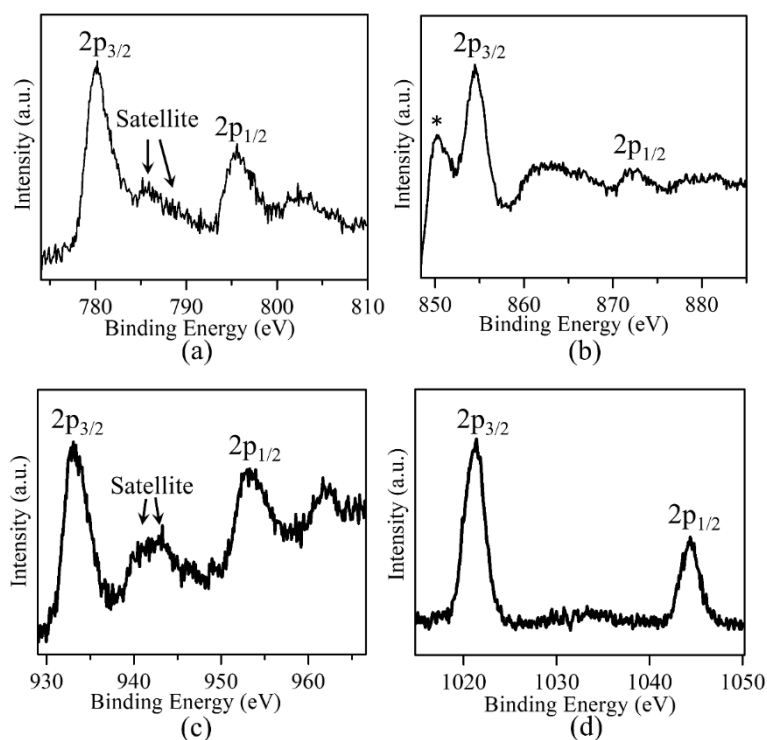


Figure 5. XPS data showing the spectra for (a) cobalt in $\text{LaSrMn}_{0.5}\text{Co}_{0.5}\text{O}_4$ (b) nickel in $\text{LaSrMn}_{0.5}\text{Ni}_{0.5}\text{O}_4$ (c) copper in $\text{LaSrMn}_{0.5}\text{Cu}_{0.5}\text{O}_{4-\delta}$, and (d) zinc in $\text{LaSrMn}_{0.5}\text{Zn}_{0.5}\text{O}_4$. The peak marked by * is the lanthanum $3d_{3/2}$ peak, which appears very close to the nickel $2p_{3/2}$.

For $\text{LaSrMn}_{0.5}\text{Cu}_{0.5}\text{O}_{4-\delta}$, the copper spectrum was obtained, as illustrated in Figure 5c. The Cu^{2+} $2p_{3/2}$ peak is expected to appear at $\sim 933.5 - 934.6$ eV,^{26, 30-32} while the peak for Cu^{1+} should be close to ~ 932.4 eV.³⁰⁻³² In our copper spectra, the $2p_{3/2}$ peak is centered around ~ 933.7 eV, with strong satellite peaks at about 7-10 eV higher than the $2p_{3/2}$ peak, which is considered the signature of Cu^{2+} .^{31, 33} These observations are consistent with the presence of copper in predominantly divalent state. They also match the observation of $\text{Mn}^{3+}/\text{Mn}^{4+}$ mixture in the manganese spectrum, and the oxygen-vacancies determined by iodometric titrations, which will be required for charge neutrality.

For $\text{LaSrMn}_{0.5}\text{Zn}_{0.5}\text{O}_4$, the zinc spectrum was obtained, as illustrated in Figure 5d. As expected, zinc is in divalent state, evident from the $2p_{3/2}$ peak at ~ 1021.3 eV.^{26, 34} As stated in the discussion of the manganese spectrum, the presence of Zn^{2+} and Mn^{4+} are consistent with iodometric titration results indicating four oxygens per formula unit.

2.3. Electrical Conductivity

We conducted a study on the electrical conductivity of all four compounds over a range of temperatures, from 25 to 800 °C (298 to 1073 K). The resistance is determined by applying Ohm's law to the current response obtained from DC measurements, and this resistance value is subsequently converted into electrical conductivity, denoted by σ (Equation 1).³⁵

$$\sigma = \frac{l}{RA} = \frac{l}{V} \cdot \frac{I}{A} \quad (1)$$

In this equation, l indicates the thickness of the measured pellet, and A represents the cross-sectional area of the pellet through which the current is applied. Table 2 lists the conductivity values at room temperature, and Figure 6a shows the plot of conductivity at different temperatures.

Table 2. Room temperature electrical conductivity and activation energies.

	Electrical conductivity (S/cm)	Activation energy (eV)
LaSrMn _{0.5} Co _{0.5} O ₄	0.0026	0.2673(2)
LaSrMn _{0.5} Ni _{0.5} O ₄	0.0014	0.2746(2)
LaSrMn _{0.5} Cu _{0.5} O _{4-δ}	8.5961×10 ⁻⁴	0.2676(3)
LaSrMn _{0.5} Zn _{0.5} O ₄	7.0755×10 ⁻⁴	0.4727(7)

As the temperature rises, the conductivities of all four materials increase, a behavior that is commonly observed in semiconducting materials. As temperature rises, charge carriers become more mobile, resulting in increased conductivity described by Equation 2.

$$\sigma = ne\mu \quad (2)$$

In this equation, σ represents electrical conductivity, n denotes the concentration of charge carriers, e stands for charge of an electron, and μ represents the mobility of charge carriers.^{36, 37} LaSrMn_{0.5}Co_{0.5}O₄ shows higher conductivity than other three materials in the whole temperature range. The electrical conductivity in oxides typically occurs by electron/hole hopping through metal–oxygen–metal (M–O–M) bonds.³⁸ For this process to happen, metals with more than one stable oxidation states, such as Mn³⁺/Mn⁴⁺ or Co³⁺/Co⁴⁺, are preferred. The transition metal 3d orbitals overlap with oxygen 2p orbitals, allowing the electron hopping through the M–O–M pathway. The activation energy (E_a) for the temperature-dependent increase in the conductivity of these materials can be calculated using the Arrhenius equation (Equation 3).^{35, 39}

$$\sigma T = \sigma^0 e^{-\frac{E_a}{kT}} \quad (3)$$

The activation energy (E_a) for the change in conductivity with temperature can be determined by analyzing the slope of the line of best-fit in the $\log \sigma T$ versus the $1000/T$ plot. In this equation, σ^0 is a pre-exponential factor, while E_a , k , and T represent activation energy, the Boltzmann constant, and absolute temperature, respectively. Figure 6b shows the Arrhenius plots for the four materials. The activation energies, determined from the slopes are presented in Table 2. It is important to emphasize that the activation energy signifies the energy barrier associated with the temperature-dependent increase in conductivity. The E_a values are consistent with the electrical conductivity trend, where LaSrMn_{0.5}Co_{0.5}O₄, which exhibits higher conductivity, also has a lower activation energy.

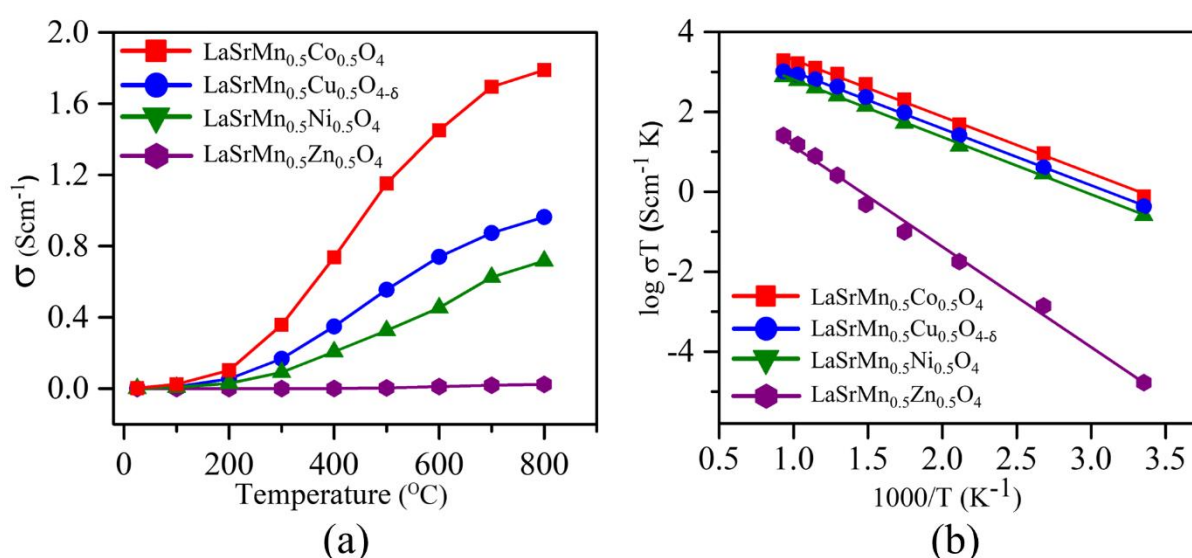


Figure 6. (a) Electrical conductivities as a function of temperature. (b) Arrhenius plots to determine the activation energies (E_a) for the temperature-activated increase in conductivity.

2.4. Hydrogen Evolution Reaction

The hydrogen evolution reaction, which is the cathodic half-reaction of water electrolysis and involves a two-electron transfer process.^{6, 40} To overcome the sluggish reaction kinetics and accelerate the rate of the HER, it is necessary to apply an overpotential. Designing efficient and stable electrocatalysts is essential to improving the reaction kinetics and minimizing the overpotential. The mechanism of the HER involves a series of steps in both acidic and alkaline media.⁴¹ Each condition has advantages and disadvantages. The efficiency of HER in acidic media is expected to be better than that of alkaline environments due to faster kinetics.⁴² The source of the hydrogen adsorbed on the catalyst in acidic media is a proton, whereas in alkaline HER, the adsorbed hydrogen comes from breaking the bond in H₂O molecule,⁴³ as shown below.

Volmer reaction (acidic): $\text{H}_3\text{O}^+ + \text{M} + \text{e}^- \rightleftharpoons \text{M}-\text{H}^* + \text{H}_2\text{O}$

Volmer reaction (alkaline): $\text{H}_2\text{O} + \text{M} + \text{e}^- \rightleftharpoons \text{M}-\text{H}^* + \text{OH}^-$

Heyrovsky reaction (acidic): $\text{M}-\text{H}^* + \text{H}_3\text{O}^+ + \text{e}^- \rightleftharpoons \text{M} + \text{H}_2 + \text{H}_2\text{O}$

Heyrovsky reaction (alkaline): $\text{M}-\text{H}^* + \text{H}_2\text{O} + \text{e}^- \rightleftharpoons \text{M} + \text{H}_2 + \text{OH}^-$

Tafel reaction (acidic and alkaline): $2\text{M}-\text{H}^* \rightleftharpoons 2\text{M} + \text{H}_2$

Alkaline electrolysis may be favored in industrial setting,⁴⁴ due to some limitations of acidic conditions, such as problems with the stability of metal oxides in acids.¹³ In this work, the HER experiments were performed in both alkaline (1 M KOH) and acidic (0.5 M H₂SO₄) conditions. In an acidic medium (Figure 7a), among the four material, LaSrMn_{0.5}Co_{0.5}O₄ shows the lowest overpotential of $\eta_{10} = 427$ mV at 10 mA/cm², which is a current density used by convention and corresponds to a 12.3% solar-to-hydrogen efficiency. It is followed by LaSrMn_{0.5}Cu_{0.5}O_{4-δ} showing $\eta_{10} = 535$ mV, LaSrMn_{0.5}Ni_{0.5}O₄ with $\eta_{10} = 647$ mV, and LaSrMn_{0.5}Zn_{0.5}O₄ with $\eta_{10} = 724$ mV. Therefore, the HER activities show the following order: LaSrMn_{0.5}Zn_{0.5}O₄ < LaSrMn_{0.5}Ni_{0.5}O₄ < LaSrMn_{0.5}Cu_{0.5}O_{4-δ} < LaSrMn_{0.5}Co_{0.5}O₄. The acidic HER overpotential of LaSrMn_{0.5}Co_{0.5}O₄ is not as low as those observed in some other catalysts such as MoO_{3-γ} nanofilms ($\eta_{10} = 0.201$ V),⁴⁵ WO₃ nanoplates ($\eta_{10} = 0.117$ V),⁴⁶ and LaCa₂Fe₃O₈ ($\eta_{10} = 0.400$ V).⁴⁷ However, it is lower than those reported for some other oxides in acidic medium, such as perovskite LaFeO₃ ($\eta_{10} = 0.490$ V),⁴⁸ Ruddlesden-Popper oxide Sr₂LaCoMnO₇ ($\eta_{10} = 0.612$ V),⁴⁹ bilayered brownmillerite Ca₃Fe₂MnO₈ ($\eta_{10} = 0.507$ V),²¹ and K₂NiF₄-type oxide SrLaCoO_{4-δ} ($\eta_{10} = -0.547$ V).¹³

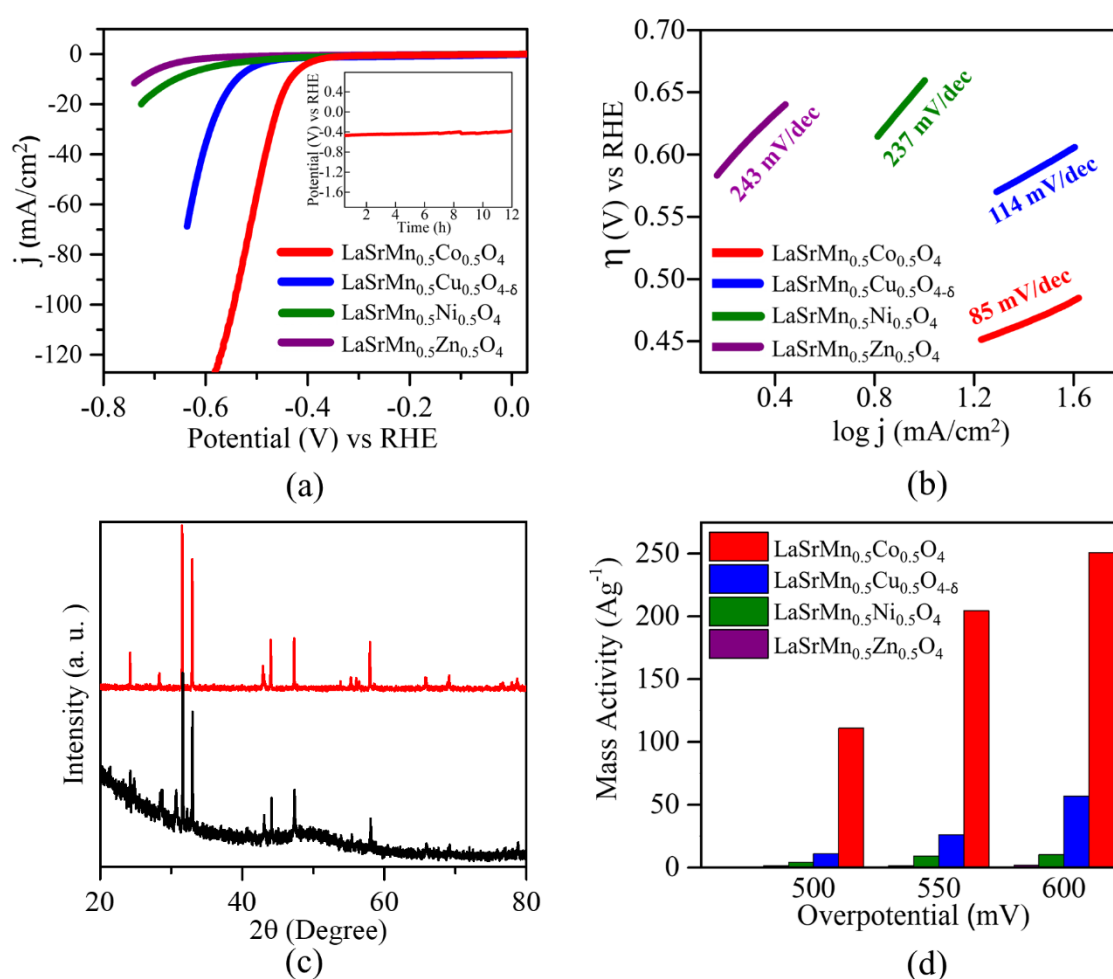


Figure 7. (a) Polarization curves for HER in 0.5 M H₂SO₄. The inset shows chronopotentiometry data for the best-performing material LaSrMn_{0.5}Co_{0.5}O₄. (b) Tafel plots and slopes. (c) X-ray diffraction data for LaSrMn_{0.5}Co_{0.5}O₄ before and after 100 cycles of HER in 0.5 M H₂SO₄. (d) HER mass activities at different overpotentials in 0.5 M H₂SO₄.

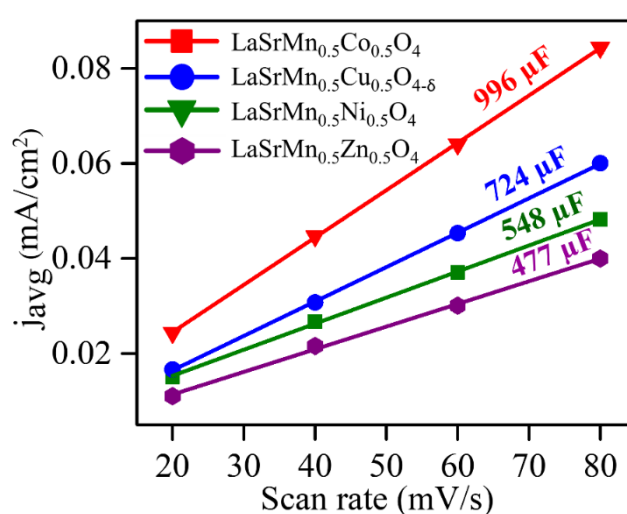


Figure 8. Plots of j_{avg} against v in 0.5 M H₂SO₄. The slopes indicate double-layer capacitance, C_{dl} .

To explore the reaction kinetics for the four materials, the Tafel equation was used (Equation 4).^{50, 51}

$$\eta = a + b \log j \quad (4)$$

Where η is the overpotential and j is the current density. The slope of η vs $\log j$ is indicative of the reaction kinetics. A smaller slope suggests a faster reaction as it indicates that a small change in overpotential is associated with a large change in current density.^{51, 52} As shown in Figure 7b, the Tafel slopes of the four materials show the same trend as their overpotentials. LaSrMn_{0.5}Co_{0.5}O₄ shows the smallest Tafel slope among the four compounds, indicating faster reaction kinetics and consistent with its greater electrocatalytic activity. Chronopotentiometry experiments performed on the most active material, LaSrMn_{0.5}Co_{0.5}O₄, exhibit a stable response for over 12 hours as shown in the inset of Figure 7a. Furthermore, X-ray diffraction data, before and after 100 cycles of HER (Figure 7c) indicates little change, confirming that the material's structural integrity remains unaltered. We also calculated the mass activity to evaluate the catalysts' inherent qualities. In Figure 7d, we have compared the mass activities of the four compounds at various overpotentials, showing that LaSrMn_{0.5}Co_{0.5}O₄ has a consistently higher mass activity.

We also determined the double-layer capacitance, C_{dl} , (Figure 8) by analyzing the average current densities (j_{avg}) at various scan rates (v), as obtained from cyclic voltammograms in a non-faradaic region (Figure S2).⁵³ The importance of C_{dl} is that it is proportional to the electrochemically active surface area.⁵⁴ The slope of the plot depicting j_{avg} against v provides the double-layer capacitance, calculated according to Equation 5.⁶

$$j_{avg} = C_{dl} \times v \quad (5)$$

Notably, the C_{dl} values align with the trend observed in electrocatalytic activity, with the most active catalyst, LaSrMn_{0.5}Co_{0.5}O₄, also displaying the highest C_{dl} value.

The HER activities of the four materials in basic condition, 1 M KOH, show a similar trend to that in acidic condition, where LaSrMn_{0.5}Co_{0.5}O₄ has the best activity, followed by LaSrMn_{0.5}Cu_{0.5}O_{4-δ}, LaSrMn_{0.5}Ni_{0.5}O₄, and LaSrMn_{0.5}Zn_{0.5}O₄ (Figure 9a). The overpotential values (η_{10}) at 10 mA/cm² are 0.482 V, 0.680 V, 0.740 V and 0.805 V, respectively. Similarly, the same trend is observed for the HER kinetics as evident from the Tafel slopes shown in Figure 9b. XRD data, before and after 100 HER cycles (Figure 9c), indicate that the structure remains unchanged, confirming the stability. In addition, the comparison of mass activity at various overpotentials in the basic medium again shows consistently higher mass activity for LaSrMn_{0.5}Co_{0.5}O₄ (Figure 9d).

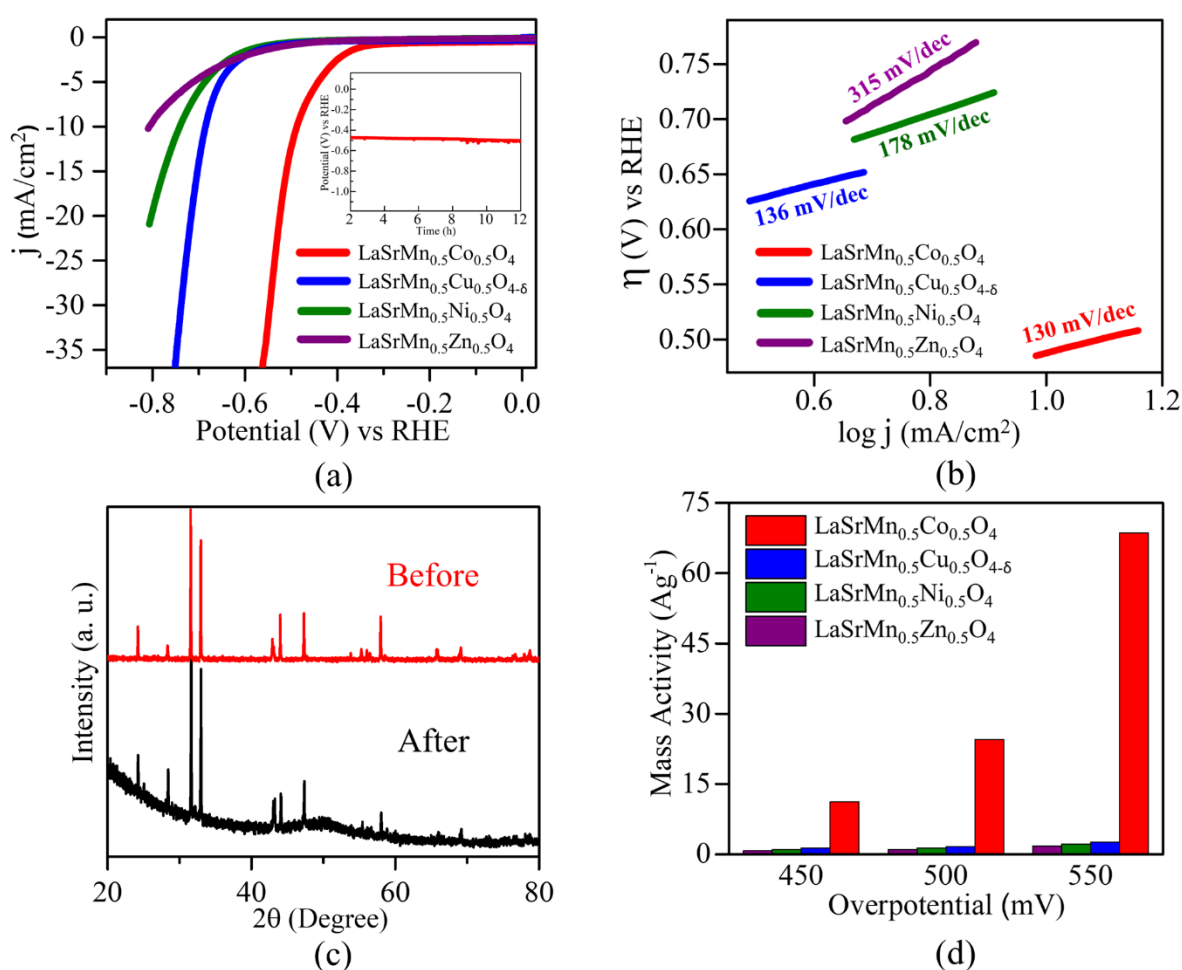


Figure 9. (a) HER polarization curves in 1 M KOH. The inset shows chronopotentiometry data for the best-performing material LaSrMn_{0.5}Co_{0.5}O₄. (b) Tafel plots and slopes. (c) X-ray diffraction data for LaSrMn_{0.5}Co_{0.5}O₄ before and after 100 cycles of HER in 1 M KOH. (d) HER mass activities at different overpotentials in 1 M KOH.

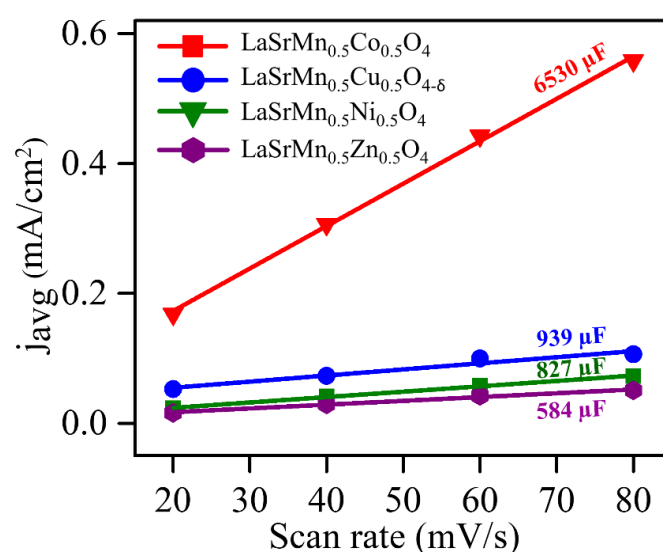


Figure 10. Plots of j_{avg} against v in 1 M KOH. The slopes indicate double-layer capacitance, C_{dl} .

We also determined the double-layer capacitance (C_{dl}) in 1 M KOH (Figure 10) using cyclic voltammograms in the non-faradaic region (Figure S2). Among the four materials,

LaSrMn_{0.5}Co_{0.5}O₄ displays the highest C_{dl} , followed by LaSrMn_{0.5}Cu_{0.5}O_{4-δ}, LaSrMn_{0.5}Ni_{0.5}O₄, and LaSrMn_{0.5}Zn_{0.5}O₄, the same trend as the electrocatalytic activity.

2.5. Oxygen Evolution Reaction

The OER electrocatalytic activities of the four materials were studied in both acidic and alkaline media. These materials showed little OER activity in acidic conditions, but the alkaline OER activity, especially for two of the materials, was substantial. Therefore, alkaline OER is discussed here. As demonstrated by the polarization curves in Figure 11a, the electrocatalytic performance for OER follows a similar trend to that observed in the HER. Among the four compounds, LaSrMn_{0.5}Ni_{0.5}O₄, and LaSrMn_{0.5}Zn_{0.5}O₄ did not produce sufficient current to reach 10 mA cm⁻² within the experimental potential range. On the other hand, LaSrMn_{0.5}Cu_{0.5}O_{4-δ} shows an overpotential of $\eta_{10} = 550$ mV, beyond the thermodynamic potential of 1.23 V at 10 mA cm⁻², while LaSrMn_{0.5}Co_{0.5}O₄ exhibits a lower overpotential of $\eta_{10} = 450$ mV. The OER overpotential of LaSrMn_{0.5}Co_{0.5}O₄ ($\eta_{10} = 450$) is not as low as those observed for some oxides such as IrO₂ (400 mV) and RuO₂ ($\eta_{10} \approx 420$ mV).⁴⁹ Nevertheless, LaSrMn_{0.5}Co_{0.5}O₄ show a better activity than several previously reported oxides, including well-known catalyst BSCF ($\eta_{10} \approx 500$ mV),⁵ La_{0.5}Sr_{0.5}Co_{0.8}Fe_{0.2}O₃ ($\eta_{10} = 600$ mV),⁵⁵ and La_{0.6}Sr_{0.4}CoO_{3-δ} ($\eta_{10} = 590$ mV).⁵⁶

The kinetics of the OER were analyzed using Tafel plots (Figure 11b), indicating a smaller Tafel slope, i.e. faster reaction kinetics, for LaSrMn_{0.5}Co_{0.5}O₄ compared to the other three materials, and the same trend as the OER activity. The stability of the best catalyst in the series, LaSrMn_{0.5}Co_{0.5}O₄, was evaluated by chronopotentiometry, showing a steady response for at least 12 hours as shown in the inset of Figure 11a. Also, X-ray diffraction data after 100 cycles of OER confirmed the material's structural integrity was preserved (Figure 11c). The mass activities for OER at various overpotentials are shown in Figure 11d, highlighting the significantly higher activity of LaSrMn_{0.5}Co_{0.5}O₄.

Some comments on the observed electrocatalytic properties are in order. The trend in electrocatalytic activity matches the trend in electrical conductivity. It is important to note that both OER and HER involve the transfer of electrons and are affected by the conductivity of catalysts. Some previous studies have shown that perovskite oxides with better electronic conductivity exhibit improved electrocatalytic performance.^{6, 36}

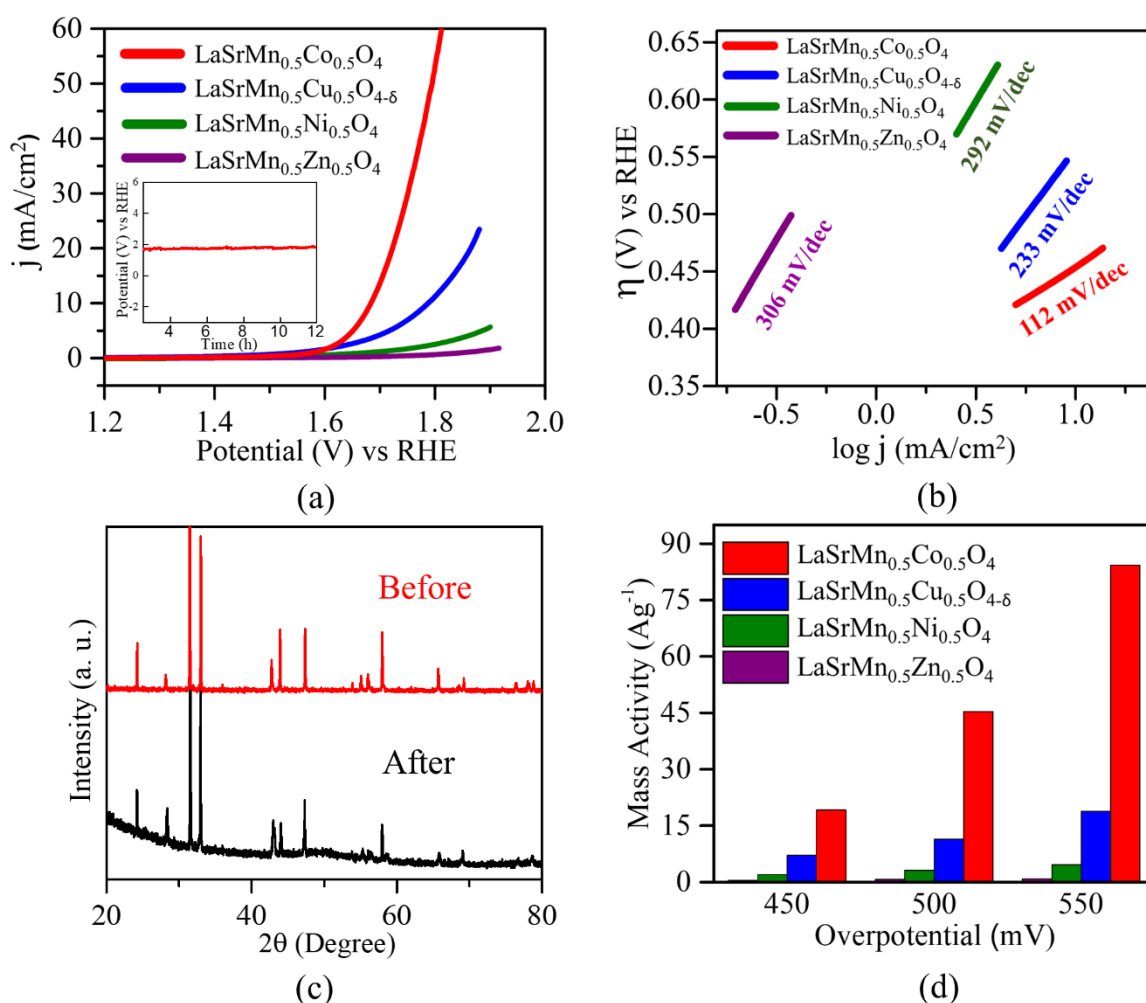


Figure 11. (a) Polarization curves for OER in 1 M KOH. The inset shows chronopotentiometry data for the best-performing material, LaSrMn_{0.5}Co_{0.5}O₄. (b) Tafel plots and slopes. (c) X-ray diffraction data for LaSrMn_{0.5}Co_{0.5}O₄ before and after 100 cycles of OER in 1 M KOH. (d) OER mass activities at different overpotentials in 1 M KOH.

There are also other parameters to consider. Some researchers have suggested that the filling of e_g level orbitals can be a descriptor for the OER activity.⁵ They proposed that high OER activity is expected for oxides, containing metal cations that have e_g occupancy close to one, due to the optimum binding with reaction intermediates. In the best catalyst in our series, LaSrMn_{0.5}Co_{0.5}O₄, a mixture of Co²⁺ and Co³⁺ oxidation states is observed. Previous studies have indicated that the intermediate spin state is prominent for Co³⁺ in some perovskite oxides.^{5,57} This intermediate spin state would give electron configuration of $t_{2g}^5 e_g^1$ for Co³⁺. Also, the mixed oxidation states of transition metals can accelerate OER activity,^{1,43} an advantage offered by the presence of cobalt. About the next best material in the series, LaSrMn_{0.5}Cu_{0.5}O_{4-δ}, it is possible that the observed activity is influenced by the presence of oxygen vacancies in this material, as confirmed by iodometric titrations and XPS, which might facilitate the adsorption of reaction intermediates.^{6,47,58} In addition, some studies on HER catalysts containing both copper and nickel have shown the better activity of copper compared to nickel. For example, a study of NiCu showed that surface Cu exhibits a d-band center closer to that of the highly active catalyst Pt.⁵⁹ The Ni site of NiCu was found to play a less substantial role in HER, primarily due to its d-band being too high, while Cu emerged as the more active catalytic site.^{59,60} Finally, the least active material in the series, LaSrMn_{0.5}Zn_{0.5}O₄, has the lowest electrical conductivity and contains zinc, which is only in a divalent state.

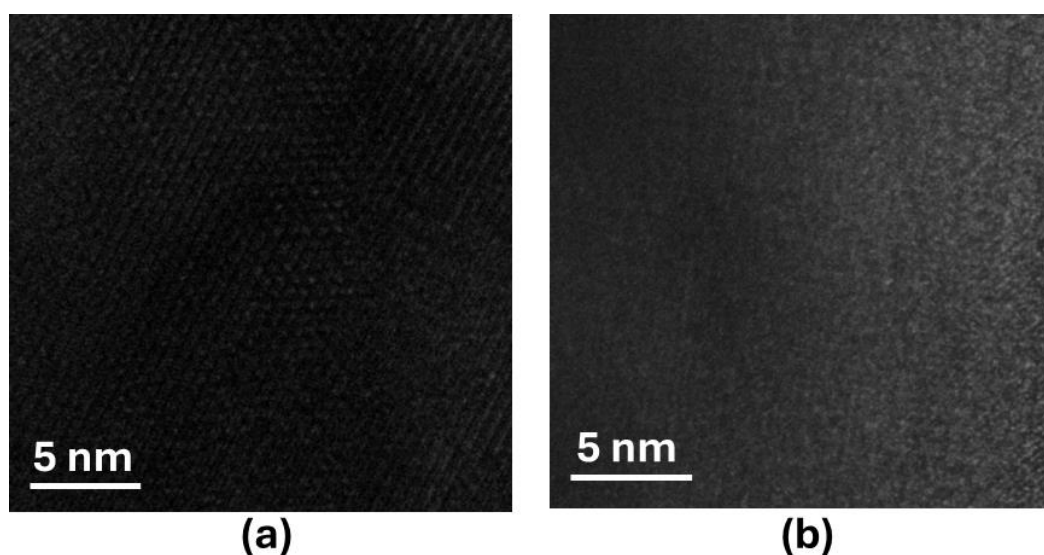


Figure 12. The TEM data: (a) before and (b) after chronopotentiometry experiment under OER conditions.

3. Experimental Methods

Synthesis and Characterization. Polycrystalline samples of $\text{LaSrMn}_{0.5}\text{Co}_{0.5}\text{O}_4$, $\text{LaSrMn}_{0.5}\text{Ni}_{0.5}\text{O}_4$, $\text{LaSrMn}_{0.5}\text{Cu}_{0.5}\text{O}_{4-\delta}$, and $\text{LaSrMn}_{0.5}\text{Zn}_{0.5}\text{O}_4$ were synthesized under argon atmosphere by solid-state synthesis method. Stoichiometric proportions of precursors La_2O_3 , SrCO_3 , MnO_4 , CoO , NiO , CuO , and ZnO were thoroughly mixed in agate mortar and pestle. For example, for the synthesis of $\text{LaSrMn}_{0.5}\text{Co}_{0.5}\text{O}_4$, a mixture of La_2O_3 (0.1500 g), SrCO_3 (0.1359 g), CoO (0.0343 g), and MnO_2 (0.0400 g) was used. The powder mixtures were pressed into pellets, and were heated at 1200 °C for 24 h followed by slow cooling. The heating and cooling rates of the furnace for all samples were set at 100 °C/h. The phase purity and structure of polycrystalline samples were determined by powder X-ray diffraction (XRD) at room temperature using Cu $K\alpha_1$ radiation ($\lambda = 1.54056 \text{ \AA}$) on an X-ray diffractometer equipped with a monochromator. Rietveld refinements were done using GSAS⁶¹ program and EXPGUI interface.⁶² The oxygen content was determined using iodometric titrations,⁶³ where excess potassium iodide (2g) and 50 mg of the sample were dissolved in 100 mL of argon-purged 1 M HCl, allowing the mixture to react overnight. Then, 5 mL of the reacted mixture, containing the generated iodine, was titrated against 0.025 M $\text{Na}_2\text{S}_2\text{O}_3$ ($\text{I}_2 + 2\text{S}_2\text{O}_3^{2-} \rightarrow 2\text{I}^- + \text{S}_4\text{O}_6^{2-}$) using starch indicator (0.6 mL), which was added near the titration endpoint. Excess KI reduced the oxide samples to form oxides with the lowest stable oxidation states of transition metals. The amount of $\text{Na}_2\text{S}_2\text{O}_3$ needed for titration indicated the amount of oxygen lost during metal ion reduction. Measurements were repeated three times for error analysis. The morphologies of the samples were studied by Scanning Electron Microscopy (SEM) on sintered pellets using a Thermo Fisher Apreo C LoVac Field Emission SEM at a magnification of 5000 \times , accompanied by an energy-dispersive X-ray spectroscopy (EDS) detector. X-ray photoelectron spectroscopy (XPS) measurements were performed using a ThermoFisher Scientific K-Alpha instrument. The instrument employed an aluminum monochromatic X-ray source with an energy of $h\nu = 1486.69 \text{ eV}$ and incorporated an electron flood gun for effective charge neutralization. Wide survey scans were carried out at a pass energy of 160 eV, while high-resolution scans were performed with a pass energy of 20 eV.

Electrical Conductivity. Electrical conductivity measurements were done at variable temperatures, 25°C to 800°C. These measurements were performed using two-probe direct current (DC) method⁶⁴ on pellets that had been sintered at 1250°C. Prior to

conducting the measurements, a layer of gold paste was applied to both sides of the pellet and then dried by heating for 3 hours at 800°C. Gold wires were attached to gold foils and used as electrodes, ensuring contact with the two gold-painted surfaces on each side of the pellet. A voltage of 0.01 V was applied for the measurements, which were conducted at approximately 100 °C intervals. Equilibrium conductivity was achieved at each measurement temperature after about 30 minutes, as indicated by a stable plateau in the DC conductivity data. Heating and cooling rates during the conductivity measurements were maintained at 3 °C/min.

Electrochemical Measurements. The catalyst ink for electrochemical measurements was prepared using 35 mg of the catalyst material, 7 mg of carbon black powder, 40 µL of Nafion D-521 solution (5% w/w in water and 1-propanol), and 7 mL of Tetrahydrofuran (THF). The mixture was ultrasonically dispersed in water for 30 minutes. The drop-casting was done by placing two coats of 10 µL of the mixture onto the surface of a glassy carbon electrode (GCE) with a diameter of 5 mm and an area of 0.196 cm², with a mass loading of 1.02 mg/cm², followed by overnight air-drying. This catalyst-loaded electrode was used as the working electrode. Electrochemical measurements were done using a standard three-electrode electrochemical cell connected to a rotating disk electrode at 1600 rpm. The OER and HER experiments were done in 1 M KOH solution by using a Hg/HgO reference electrode (1 M NaOH). The HER experiments were also conducted in 0.5 M H₂SO₄ solution using Ag/AgCl (4 M KCl) as the reference electrode. A carbon electrode was used as the counter electrode for all experiments. All potentials were iR corrected and converted to potential vs reversible hydrogen electrode (RHE) using the Nernst equation $E_{vs\ RHE} = E_{vs\ Reference\ electrode} + 0.059\ pH + E^0_{Reference\ electrode}$, where $E^0_{Ag/AgCl} = 0.197\ V$ for 4 M KCl KCl and $E^0_{Hg/HgO} = 0.098$ for 1 M NaOH.^{36, 65} For each material, the electrocatalytic measurements were repeated at least three times, using at least two different batches synthesized independently. Chronopotentiometry was utilized to investigate the catalyst stability under both HER and OER conditions, employing the same three-electrode setup and a constant current of 10 mA/cm².

4. Conclusions

A range of parameters can influence electrocatalytic properties, such as electrical conductivity, the presence of metals with more than one stable oxidation state, oxygen-vacancies, and electron configuration. These parameters were studied in this work through the investigation of a series of quasi-2D oxides to reveal their structural and electrocatalytic properties. X-ray diffraction, X-ray photoelectron spectroscopy, iodometric titrations, electrical conductivity studies, and OER and HER experiments were used for the investigation of the four materials LaSrMn_{0.5}M_{0.5}O₄ (M = Co, Ni, Cu, and Zn). The electrical conductivity measurements, as well as electrocatalytic HER and OER, all showed the same trend, LaSrMn_{0.5}Zn_{0.5}O₄ < LaSrMn_{0.5}Ni_{0.5}O₄ < LaSrMn_{0.5}Cu_{0.5}O_{4-δ} < LaSrMn_{0.5}Co_{0.5}O₄. The best-performing compound in this series, LaSrMn_{0.5}Co_{0.5}O₄, can act as a bifunctional catalytic material, facilitating both OER and HER. Its lower overpotential, faster kinetics, and greater mass activity were demonstrated and correlated with several descriptors that can be responsible for its higher electrocatalytic performance.

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Conflicts of Interests. Authors declare no conflict of interests.

References

1. Lee, J. G.; Hwang, J.; Hwang, H. J.; Jeon, O. S.; Jang, J.; Kwon, O.; Lee, Y.; Han, B.; Shul, Y.-G., A new family of perovskite catalysts for oxygen-evolution reaction in alkaline media: BaNiO₃ and BaNi_{0.83}O_{2.5}. *J. Am. Chem. Soc.* **2016**, *138* (10), 3541-3547.

2. Zhu, J.; Hu, L.; Zhao, P.; Lee, L. Y. S.; Wong, K.-Y., Recent advances in electrocatalytic hydrogen evolution using nanoparticles. *Chem. Rev.* **2019**, *120* (2), 851-918.
3. Lee, Y.; Suntivich, J.; May, K. J.; Perry, E. E.; Shao-Horn, Y., Synthesis and activities of rutile IrO₂ and RuO₂ nanoparticles for oxygen evolution in acid and alkaline solutions. *J. Phys. Chem. Lett.* **2012**, *3* (3), 399-404.
4. Tian, J.; Wu, W.; Tang, Z.; Wu, Y.; Burns, R.; Tichnell, B.; Liu, Z.; Chen, S., Oxygen reduction reaction and hydrogen evolution reaction catalyzed by Pd-Ru nanoparticles encapsulated in porous carbon nanosheets. *Catal.* **2018**, *8* (8), 329.
5. Suntivich, J.; May, K. J.; Gasteiger, H. A.; Goodenough, J. B.; Shao-Horn, Y., A perovskite oxide optimized for oxygen evolution catalysis from molecular orbital principles. *Science* **2011**, *334* (6061), 1383-1385.
6. Wickramaratne, K. M.; Ramezanipour, F., Impact of oxygen-vacancies on electrical conductivity and electrocatalytic activity of La_{3-x}Ca_xFe₂GaO_{9-δ} (x = 0, 2; δ = 0, 1). *Solid State Sci.* **2023**, *141*, 107208.
7. Karki, S. B.; Hona, R. K.; Yu, M.; Ramezanipour, F., Enhancement of Electrocatalytic Activity as a Function of Structural Order in Perovskite Oxides. *ACS Catal.* **2022**, *12*, 10333-10337.
8. Hona, R. K.; Karki, S. B.; Cao, T.; Mishra, R.; Sterbinsky, G. E.; Ramezanipour, F., Sustainable oxide electrocatalyst for hydrogen-and oxygen-evolution reactions. *ACS Catal.* **2021**, *11*, 14605-14614.
9. Choi, M.-J.; Kim, T. L.; Kim, J. K.; Lee, T. H.; Lee, S. A.; Kim, C.; Hong, K.; Bark, C. W.; Ko, K.-T.; Jang, H. W., Enhanced Oxygen Evolution Electrocatalysis in Strained A-Site Cation Deficient LaNiO₃ Perovskite Thin Films. *Nano Lett.* **2020**, *20* (11), 8040-8045.
10. Wang, J.; Gao, Y.; Chen, D.; Liu, J.; Zhang, Z.; Shao, Z.; Ciucci, F., Water splitting with an enhanced bifunctional double perovskite. *ACS Catal.* **2018**, *8* (1), 364-371.
11. Tian, S.; He, J.; Huang, H.; Song, T.-s.; Wu, X.; Xie, J.; Zhou, W., Perovskite-based multifunctional cathode with simultaneous supplementation of substrates and electrons for enhanced microbial electrosynthesis of organics. *ACS Appl. Mater. Interfaces* **2020**, *12* (27), 30449-30456.
12. Zhu, Y.; Zhou, W.; Zhong, Y.; Bu, Y.; Chen, X.; Zhong, Q.; Liu, M.; Shao, Z., A perovskite nanorod as bifunctional electrocatalyst for overall water splitting. *Adv. Energy Mater.* **2017**, *7* (8), 1602122.
13. Alom, M. S.; Ramezanipour, F., Layered Oxides SrLaFe_{1-x}Co_xO_{4-δ} (x=0-1) as Bifunctional Electrocatalysts for Water-Splitting. *ChemCatChem* **2021**, *13* (15), 3510-3516.
14. El Shinawi, H.; Greaves, C., Structural and Magnetic Characterisation of La_{1+x}Sr_{1-x}Co_{0.5}M_{0.5}O_{4±δ} (M= Cr, Mn). *Z. Anorg. Allg. Chem.* **2009**, *635*, 1856-1862.
15. El Shinawi, H.; Marco, J.; Berry, F.; Greaves, C., Synthesis and characterization of La_{0.8}Sr_{1.2}Co_{0.5}M_{0.5}O_{4-δ} (M= Fe, Mn). *J. Solid State Chem.* **2009**, *182* (8), 2261-2268.
16. McCabe, E.; Greaves, C., Synthesis and Structural and Magnetic Characterization of Mixed Manganese- Copper n= 1 Ruddlesden- Popper Phases. *Chem. Mater.* **2006**, *18* (24), 5774-5781.
17. Burley, J. C.; Battle, P. D.; Gaskell, P. J.; Rosseinsky, M. J., Structural and magnetic chemistry of La₂Sr₂BMnO₈ (B= Mg, Zn). *J. Solid State Chem.* **2002**, *168* (1), 202-207.
18. El Shinawi, H.; Greaves, C., Synthesis and characterization of La_{1.5+x}Sr_{0.5-x}Co_{0.5}Ni_{0.5}O_{4±δ} (x= 0, 0.2). *J. Mater. Chem.* **2010**, *20* (3), 504-511.
19. Yin, W.-J.; Weng, B.; Ge, J.; Sun, Q.; Li, Z.; Yan, Y., Oxide perovskites, double perovskites and derivatives for electrocatalysis, photocatalysis, and photovoltaics. *Energy Environ. Sci.* **2019**, *12* (2), 442-462.
20. Zhao, Y.-N.; Liu, C.; Xu, S.; Min, S.; Wang, W.; Mitsuzaki, N.; Chen, Z., A/B-Site Management Strategy to Boost Electrocatalytic Overall Water Splitting on Perovskite Oxides in an Alkaline Medium. *Inorg. Chem.* **2023**, *62* (31), 12590-12599.

21. Wickramaratne, K. M. K.; Karki, S. B.; Ramezanipour, F., Electrocatalytic Properties of Oxygen-Deficient Perovskites $\text{Ca}_3\text{Fe}_{3-x}\text{Mn}_x\text{O}_8$ ($x = 1 - 2$) for Hydrogen Evolution Reaction. *Inorg. Chem.* **2023**, *62*, 20961–20969.
22. Bhowmick, S.; Mohanta, M. K.; Qureshi, M., Transcription methodology for rationally designed morphological complex metal oxides: a versatile strategy for improved electrocatalysis. *Sustain. Energy Fuels*. **2021**, *5* (24), 6392–6405.
23. Bhowmick, S.; Dhankhar, A.; Sahu, T. K.; Jena, R.; Gogoi, D.; Peela, N. R.; Ardo, S.; Qureshi, M., Low overpotential and stable electrocatalytic oxygen evolution reaction utilizing doped perovskite oxide, $\text{La}_{0.7}\text{Sr}_{0.3}\text{MnO}_3$, modified by cobalt phosphate. *ACS Appl. Energy Mater.* **2020**, *3* (2), 1279–1285.
24. Dupin, J.-C.; Gonbeau, D.; Benqlilou, H.; Vinatier, P.; Levasseur, A., XPS Analysis of New Lithium Cobalt Oxide Thin-films Before and After Lithium Deintercalation. *Thin Solid Films* **2001**, *384*, 23–32.
25. Davison, N.; McWhinnie, W. R.; Hooper, A., X-Ray Photoelectron Spectroscopic Study of Cobalt(II) and Nickel(II) Sorbed on Hectorite and Montmorillonite. *Clays Clay Miner.* **1991**, *39*, 22–27.
26. Moulder, J. F.; Stickle, W. F.; Sobol, P. E.; Bomben, K. D.; editor Chastain, J., **1992**, *Handbook of X-ray photoelectron spectroscopy: a reference book of standard spectra for identification and interpretation of XPS data.*, Perkin-Elmer, Eden Prairie.
27. Zhao, Q.; Fang, C.; Tie, F.; Luo, W.; Peng, Y.; Huang, F.; Ku, Z.; Cheng, Y.-B., Regulating the $\text{Ni}^{3+}/\text{Ni}^{2+}$ ratio of NiOx by plasma treatment for fully vacuum-deposited perovskite solar cells. *Mater. Sci. Semicond. Process.* **2022**, *148*, 106839.
28. Fan, L.; Liu, P. F.; Yan, X.; Gu, L.; Yang, Z. Z.; Yang, H. G.; Qiu, S.; Yao, X., Atomically isolated nickel species anchored on graphitized carbon for efficient hydrogen evolution electrocatalysis. *Nat. Commun.* **2016**, *7*, 10667.
29. Cheng, M.; Fan, H.; Song, Y.; Cui, Y.; Wang, R., Interconnected hierarchical NiCo_2O_4 microspheres as high-performance electrode materials for supercapacitors. *Dalton Trans.* **2017**, *46* (28), 9201–9209.
30. Wang, P.; Liu, Z.; Han, C.; Ma, X.; Tong, Z.; Tan, B., $\text{Cu}_2\text{O}/\text{CuO}$ heterojunction formed by thermal oxidation and decorated with Pt co-catalyst as an efficient photocathode for photoelectrochemical water splitting. *J. Nanoparticle Res.* **2021**, *23* (12), 268.
31. Dan, Z.; Yang, Y.; Qin, F.; Wang, H.; Chang, H., Facile Fabrication of Cu_2O Nanobelts in Ethanol on Nanoporous Cu and Their Photodegradation of Methyl Orange. *Materials* **2018**, *11*, 446.
32. Akgul, F. A.; Akgul, G.; Yildirim, N.; Unalan, H. E.; Turan, R., Influence of thermal annealing on microstructural, morphological, optical properties and surface electronic structure of copper oxide thin films. *Mater. Chem. Phys.* **2014**, *147*, 987–995.
33. Jin, Z.; Liu, C.; Qi, K.; Cui, X., Photo-reduced Cu/CuO nanoclusters on TiO_2 nanotube arrays as highly efficient and reusable catalyst. *Sci. Rep.* **2017**, *7*, 39695.
34. Xu, D.; Fan, D.; Shen, W., Catalyst-free direct vapor-phase growth of $\text{Zn}_{1-x}\text{Cu}_x\text{O}$ micro-cross structures and their optical properties. *Nanoscale Res. Lett.* **2013**, *8*, 1–9.
35. Andoulsi, R.; Horchani-Naifer, K.; Ferid, M., Electrical conductivity of $\text{La}_{1-x}\text{Ca}_x\text{FeO}_{3-\delta}$ solid solutions. *Ceram. Int.* **2013**, *39* (6), 6527–6531.
36. Karki, S. B.; Ramezanipour, F., Pseudocapacitive Energy Storage and Electrocatalytic Hydrogen-Evolution Activity of Defect-Ordered Perovskites $\text{Sr}_x\text{Ca}_{3-x}\text{GaMn}_2\text{O}_8$ ($x = 0$ and 1). *ACS Appl. Energy Mater.* **2020**, *3* (11), 10983–10992.
37. Asenath-Smith, E.; Lokuhewa, I. N.; Mixture, S. T.; Edwards, D. D., p-Type thermoelectric properties of the oxygen-deficient perovskite $\text{Ca}_2\text{Fe}_2\text{O}_5$ in the brownmillerite structure. *J. Solid State Chem.* **2010**, *183* (7), 1670–1677.
38. Kozhevnikov, V.; Leonidov, I.; Mitberg, E.; Patrakeeve, M.; Petrov, A.; Poeppelmeier, K., Conductivity and carrier traps in $\text{La}_{1-x}\text{Sr}_x\text{Co}_{1-z}\text{Mn}_z\text{O}_{3-\delta}$ ($x = 0.3$; $z = 0$ and 0.25). *J. Solid State Chem.* **2003**, *172* (2), 296–304.

39. Sudha, L.; Sukumar, R.; Uma Rao, K., Evaluation of activation energy (E_a) profiles of nanostructured alumina polycarbonate composite insulation materials. *Int. J. Mater. Mech. Manuf.* **2014**, *2* (1), 96-100.
40. Alom, M. S.; Kananke-Gamage, C. C.; Ramezanipour, F., Perovskite Oxides as Electrocatalysts for Hydrogen Evolution Reaction. *ACS omega* **2022**, *7* (9), 7444-7451.
41. Si, C.; Zhang, W.; Lu, Q.; Guo, E.; Yang, Z.; Chen, J.; He, X.; Luo, J., Recent Advances in Perovskite Catalysts for Efficient Overall Water Splitting. *Catalysts* **2022**, *12* (6), 601.
42. Strmcnik, D.; Uchimura, M.; Wang, C.; Subbaraman, R.; Danilovic, N.; Van Der Vliet, D.; Paulikas, A. P.; Stamenkovic, V. R.; Markovic, N. M., Improving the hydrogen oxidation reaction rate by promotion of hydroxyl adsorption. *Nat. Chem.* **2013**, *5* (4), 300-306.
43. Khan, R.; Mehran, M. T.; Naqvi, S. R.; Khoja, A. H.; Mahmood, K.; Shahzad, F.; Hussain, S., Role of perovskites as a bi-functional catalyst for electrochemical water splitting: A review. *Int. J. Energy Res.* **2020**, *44* (12), 9714-9747.
44. Wang, S.; Lu, A.; Zhong, C.-J., Hydrogen production from water electrolysis: role of catalysts. *Nano Conver.* **2021**, *8*, 1-23.
45. Wang, H.; Zhou, H.; Zhang, W.; Yao, S., Urea-assisted synthesis of amorphous molybdenum sulfide on P-doped carbon nanotubes for enhanced hydrogen evolution. *J. Mater. Sci. Mater.* **2018**, *53*, 8951-8962.
46. Han, H.; Nayak, A. K.; Choi, H.; Ali, G.; Kwon, J.; Choi, S.; Paik, U.; Song, T., Partial dehydration in hydrated tungsten oxide nanoplates leads to excellent and robust bifunctional oxygen reduction and hydrogen evolution reactions in acidic media. *ACS Sustain. Chem. Eng.* **2020**, *8* (25), 9507-9518.
47. Karki, S. B.; Andriotis, A. N.; Menon, M.; Ramezanipour, F., Bifunctional Water-Splitting Electrocatalysis Achieved by Defect Order in $\text{LaA}_2\text{Fe}_3\text{O}_8$ ($A = \text{Ca}, \text{Sr}$). *ACS Appl. Energy Mater.* **2021**, *4* (11), 12063-12066.
48. Galal, A.; Hassan, H. K.; Atta, N. F.; Jacob, T., An Efficient and Durable Electrocatalyst for Hydrogen Production Based on Earth-Abundant Oxide-Graphene Composite. *ChemistrySelect* **2017**, *2* (31), 10261-10270.
49. Kananke-Gamage, C. C.; Ramezanipour, F., Variation of the electrocatalytic activity of isostructural oxides $\text{Sr}_2\text{LaFeMnO}_7$ and $\text{Sr}_2\text{LaCoMnO}_7$ for hydrogen and oxygen-evolution reactions. *Dalton Trans.* **2021**, *50* (40), 14196-14206.
50. Murthy, A. P.; Theerthagiri, J.; Madhavan, J., Insights on Tafel constant in the analysis of hydrogen evolution reaction. *J. Phys. Chem. C* **2018**, *122* (42), 23943-23949.
51. Shinagawa, T.; Garcia-Esparza, A. T.; Takanabe, K., Insight on Tafel slopes from a microkinetic analysis of aqueous electrocatalysis for energy conversion. *Sci. Rep.* **2015**, *5* (1), 1-21.
52. Ling, C.; Ouyang, Y.; Shi, L.; Yuan, S.; Chen, Q.; Wang, J., Template-grown MoS_2 nanowires catalyze the hydrogen evolution reaction: ultralow kinetic barriers with high active site density. *ACS Catal.* **2017**, *7* (8), 5097-5102.
53. Hona, R. K.; Ramezanipour, F., Remarkable Oxygen-Evolution Activity of a Perovskite Oxide from the $\text{Ca}_{2-x}\text{Sr}_x\text{Fe}_2\text{O}_{6-\delta}$ Series. *Angew Chem.* **2019**, *131* (7), 2082-2085.
54. Pan, Y.; Chen, Y.; Li, X.; Liu, Y.; Liu, C., Nanostructured nickel sulfides: phase evolution, characterization and electrocatalytic properties for the hydrogen evolution reaction. *RSC Adv.* **2015**, *5* (127), 104740-104749.
55. Park, H. W.; Lee, D. U.; Park, M. G.; Ahmed, R.; Seo, M. H.; Nazar, L. F.; Chen, Z., Perovskite-nitrogen-doped carbon nanotube composite as bifunctional catalysts for rechargeable lithium-air batteries. *ChemSusChem* **2015**, *8* (6), 1058-1065.
56. Oh, M. Y.; Jeon, J. S.; Lee, J. J.; Kim, P.; Nahm, K. S., The bifunctional electrocatalytic activity of perovskite $\text{La}_{0.6}\text{Sr}_{0.4}\text{CoO}_{3-\delta}$ for oxygen reduction and evolution reactions. *RSC Adv.* **2015**, *5* (25), 19190-19198.

57. Moritomo, Y.; Higashi, K.; Matsuda, K.; Nakamura, A., Spin-state transition in layered perovskite cobalt oxides: $\text{La}_{2-x}\text{Sr}_x\text{CoO}_4$ ($0.4 \leq x \leq 1.0$). *Phys. Rev. B.* **1997**, *55* (22), R14725.
58. Alom, M. S.; Ramezanipour, F., Vacancy effect on the electrocatalytic activity of $\text{LaMn}_{1/2}\text{Co}_{1/2}\text{O}_{3-\delta}$ for hydrogen and oxygen evolution reactions. *Chem. Commun.* **2023**, *59*, 5870-5873.
59. Wei, C.; Sun, Y.; Scherer, G. n. G.; Fisher, A. C.; Sherburne, M.; Ager, J. W.; Xu, Z. J., Surface composition dependent ligand effect in tuning the activity of nickel–copper bimetallic electrocatalysts toward hydrogen evolution in alkaline. *J. Am. Chem. Soc.* **2020**, *142* (17), 7765-7775.
60. Rajput, A.; Kundu, A.; Chakraborty, B., Recent Progress on Copper-Based Electrode Materials for Overall Water-Splitting. *ChemElectroChem* **2021**, *8* (10), 1698-1722.
61. Larson, A. C.; Von Dreele, R. B., *General Structure Analysis System (GSAS)*, Los Alamos National Laboratory Report LAUR **2004**, 86-748.
62. Toby, B. H., EXPGUI, a graphical user interface for GSAS. *J. Appl. Crystallogr.* **2001**, *34* (2), 210-213.
63. Hona, R. K.; Karki, S. B.; Ramezanipour, F., Oxide Electrocatalysts Based on Earth-Abundant Metals for Both Hydrogen-and Oxygen-Evolution Reactions. *ACS Sustain. Chem. Eng.* **2020**, *8*, 11549-11557.
64. Li, Q.; Thangadurai, V., A comparative 2 and 4-probe DC and 2-probe AC electrical conductivity of novel co-doped $\text{Ce}_{0.9-x}\text{RE}_x\text{Mo}_{0.1}\text{O}_{2.1-0.5x}$ (RE= Y, Sm, Gd; $x= 0.2, 0.3$). *J. Mater. Chem.* **2010**, *20* (37), 7970-7983.
65. Niu, S.; Li, S.; Du, Y.; Han, X.; Xu, P., How to reliably report the overpotential of an electrocatalyst. *ACS Energy Lett.* **2020**, *5* (4), 1083-1087.